

Article

# Metathetical Redox Reaction of (Diacetoxyiodo)arenes and Iodoarenes

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**Abstract:** The oxidation of iodoarenes is central to the field of hypervalent iodine chemistry. It was found that the metathetical redox reaction between (diacetoxyiodo)arenes and iodoarenes is possible in the presence of a catalytic amount of Lewis acid. This discovery opens a new strategy to access (diacetoxyiodo)arenes. A computational study is provided to rationalize the results observed.

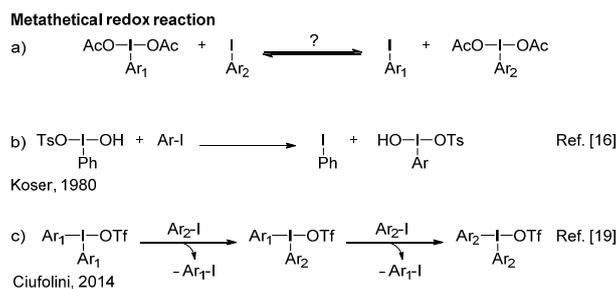
**Keywords:** hypervalent iodine; (diacetoxyiodo)arenes; metathesis

## 1. Introduction

In the recent years, the development of hypervalent iodine-mediated synthetic transformations has been receiving growing attention [1–5]. This is not surprising considering the importance of improving oxidative reactions and reducing their environmental impact. Hypervalent iodine reagents are a great alternative to toxic heavy metals often used to effect similar transformations [6–10]. Among the plethora of iodine(III) compounds [11], it is undeniable that the (diacetoxyiodo)arenes remain the most common and used ones [12]. They are usually stable solids and can serve as precursors to numerous other iodanes, such as other (diacyloxyiodo)arenes and [hydroxy(tosyloxy)iodo]arenes [13,14]. Access to (diacetoxyiodo)arenes is, of course, usually accomplished by the oxidation of the corresponding iodoarene substrate. The traditional method involves peracetic acid oxidation in acetic acid [15]. One unexploited strategy to access these specific reagents is the metathetical redox reaction between (diacetoxyiodo)arenes and related iodoarenes (Scheme 1a).

While it is widely accepted that the ligands on iodane species are typically easily exchanged, the interconversion of oxidation states between  $\lambda^3$ -iodane and iodoarenes is much rarer. The first example of such a metathetical redox reaction was reported by Koser *et al.*, using [hydroxy(tosyloxy)-iodo]benzene (HTIB) as the oxidant, enabling transfer to a variety of iodoarenes (Scheme 1b) [16]. Two applications of this initial method were subsequently reported, including one example using [bis(trifluoroacetoxy)iodo]benzene (BTI) [17,18]. More recently, Ciufolini *et al.* have reported the iodonium salt metathesis reaction with iodoarenes (Scheme 1c) [19].

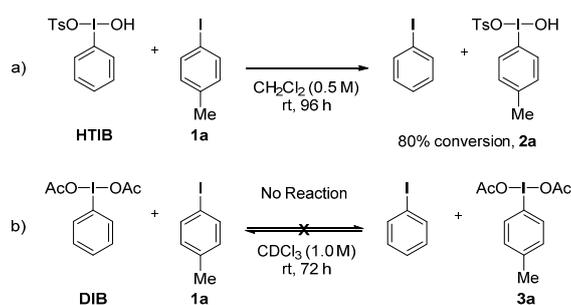
To the best of our knowledge, the metathetical redox reaction between (diacetoxyiodo)arenes and iodoarenes has never been reported. Expanding the scope of iodane metathesis to these commonly used reagents would be of great interest as an alternative synthetic method. Herein we report a protocol to achieve such a metathesis of (diacetoxyiodo)benzene and iodoarenes. We also provide a computational study to determine the origin of the observed reactivity.



**Scheme 1.** Concept and examples of metathetical redox reactions of iodanes and iodoarenes.

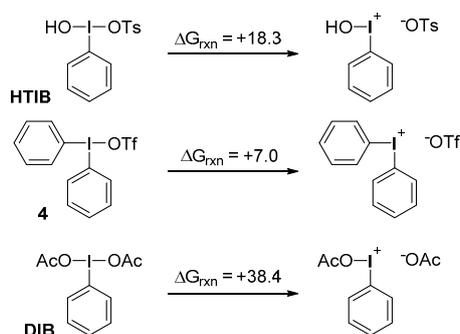
## 2. Results and Discussion

The reactivity of HTIB and (diacetoxyiodo)benzene (DIB) toward iodoarenes was evaluated, using conditions analogous to Koser's successful iodane metathesis with HTIB [16]. We submitted *p*-methyliodobenzene and *p*-bromiodobenzene to iodane metathesis with HTIB (Scheme 2a) and DIB at room temperature (Scheme 2b). In accordance with the results reported by Koser, equilibrium is achieved for the reaction with HTIB within 96 h. In contrast, no metathesis was observed after 72 h with DIB.



**Scheme 2.** Evaluation of the reactivity of DIB toward the metathetical redox process.

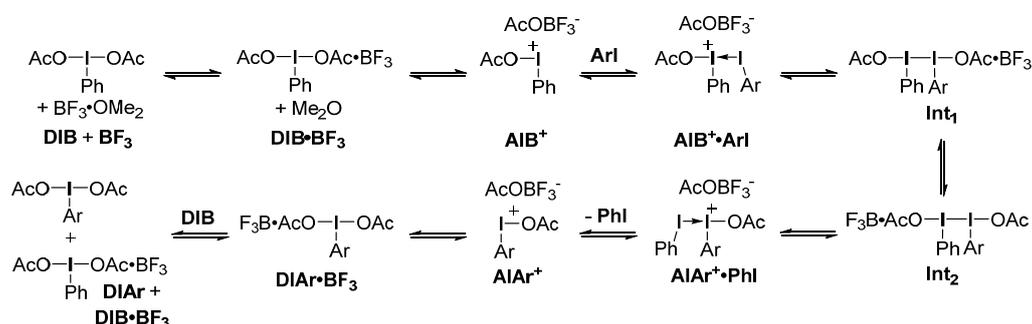
These results demonstrate the very different reactivity profiles of DIB and HTIB. The origin of this reactivity dichotomy was investigated. In a recent computational study, we have suggested that reaction with HTIB would involve, at room temperature, iodonium intermediates, through a ligand dissociation mechanism [20]. Due to the large difference in effective electronegativity between the hydroxy and tosyloxy groups on HTIB, ionization in solution is achievable at room temperature [21,22]. In contrast, (diacetoxyiodo)arenes have identical ligands involved in the three-center four-electron bonding, and thus dissociation is expected to be difficult [23]. We have evaluated the energetic of dissociation for HTIB, DIB, and diphenyliodonium salt **4** (see the computational details in Section 4), the results are illustrated in Scheme 3.



**Scheme 3.** Energetics of ligand dissociation on DIB, **4**, and HTIB (energies reported in kcal/mol).

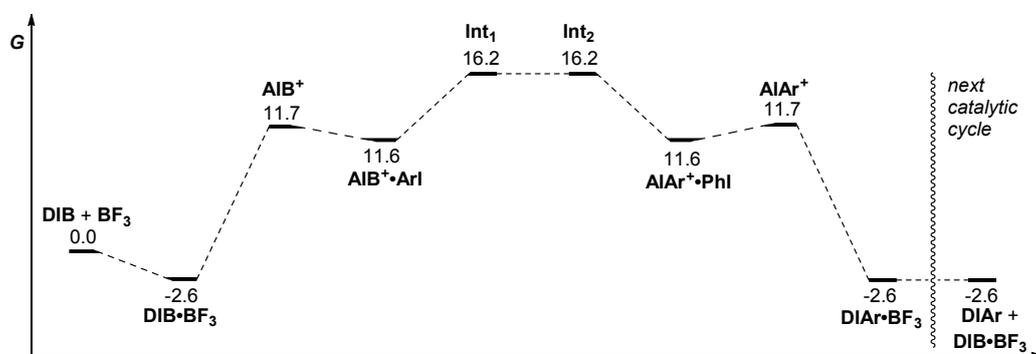


DIB. The transition states leading to the different intermediates described were not studied, as it was assumed that these processes of association and dissociation do not require large activation energies.



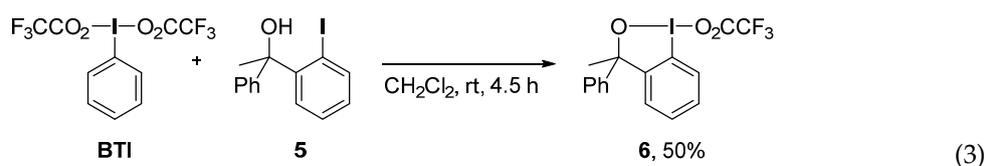
**Scheme 5.** Reaction mechanism evaluated using DFT calculations (Ar = Ph for simplicity).

The energetic properties of all these intermediates were calculated and are illustrated in an energy diagram (Scheme 6). As described in Scheme 6, formation of intermediate **AIB**<sup>+</sup>, following complexation by  $\text{BF}_3 \cdot \text{OME}_2$ , was found to be an endergonic but feasible process at room temperature. At this point, the iodonium intermediate can act as a Lewis acid toward an iodoarene (in this case iodobenzene) to form adduct **AIB**<sup>+</sup>·**ArI**. Association of the  $\text{AcOBF}_3^-$  counterion led to **Int**<sub>1</sub>. By internal Lewis acid migration, **Int**<sub>2</sub> would be accessed. This intermediate can then lead to **AIAr**<sup>+</sup>·**PhI** adduct, which can liberate iodobenzene and finally provide, by transfer of  $\text{BF}_3$  to another DIB molecule, the newly formed (diacetoxyiodo)arene (**DIAr**). The calculations thus suggest an ionic mechanism relying on the asymmetry or activation of the iodane moiety to initiate the metathetical redox process, through a key iodonium intermediate.



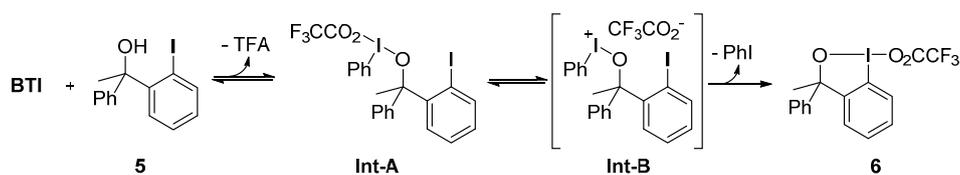
**Scheme 6.** Energy diagram of the proposed mechanism (energies reported in kcal/mol).

Passage through a non-symmetric iodane intermediate could explain the metathesis reported by Koser *et al.*, in which iodoarene **5** is converted to its corresponding cyclic iodane **6** by the action of BTI, without the need of a catalyst (Equation (3)) [17].



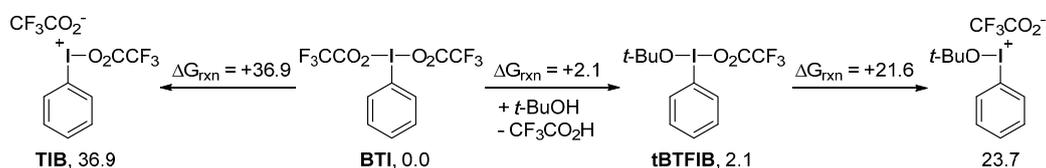
In this reaction, rapid ligand exchange would lead to intermediate **Int-A** (Scheme 7). Due to the large difference in effective electronegativity between the alkoxy and trifluoroacetoxy groups, dissociation of the latter to furnish the iodonium intermediate **Int-B** is conceivable. Furthermore,

the fact that **Int-B** binds together both I(I) and I(III) partners might facilitate the metathetical redox process.



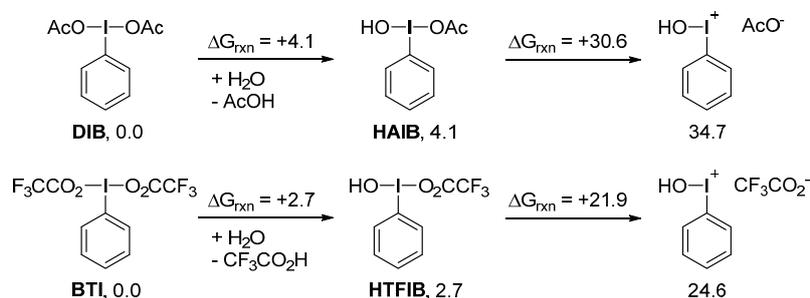
**Scheme 7.** Proposed intermediate to rationalize reactivity of BTI toward 5.

Calculations using *t*-BuOH as a model alcohol support this proposal, as described by the energetic values illustrated in Scheme 8. While dissociation of a trifluoroacetate group on the symmetric BTI was calculated to be too costly to occur at room temperature. In contrast, due to the lower effective electronegativity of the *t*-butoxy compared to the trifluoroacetoxy group, the dissociation on the non-symmetric tBTfIB is facilitated.



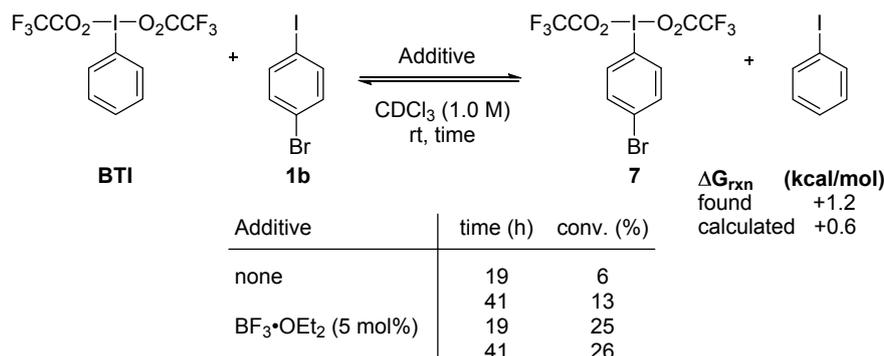
**Scheme 8.** Alcohol-promoted formation of an iodonium intermediate.

This alternative activation mechanism raises an important point concerning the presence of water in these reactions. As a protic solvent similar to an alcohol, water might act as an activator to promote acyloxy group dissociation. The energetic values of water exchange and possible acyloxy dissociation were computed for both DIB and BTI; the results are illustrated in Scheme 9. In the case of DIB, even the intermediate species [hydroxy(acetyloxy)iodo]benzene (HAIB) would not permit dissociation of the acetyloxy group to form the iodonium intermediate at room temperature. In contrast, the calculations suggest that dissociation of the trifluoroacetoxy group on the analogous [hydroxy(trifluoroacetoxy)iodo]benzene (HTfIB) intermediate might be possible at room temperature. In essence, even if BTI is found to be more stable than HTfIB, the latter would possess reactivity similar to Koser's reagent (HTIB). The main difference between DIB, BTI, and HTIB can simply be explained by the better nucleofugality of  $\text{TsO}^-$  and  $\text{CF}_3\text{CO}_2^-$  vs.  $\text{AcO}^-$ .



**Scheme 9.** Water-promoted formation of iodonium intermediates.

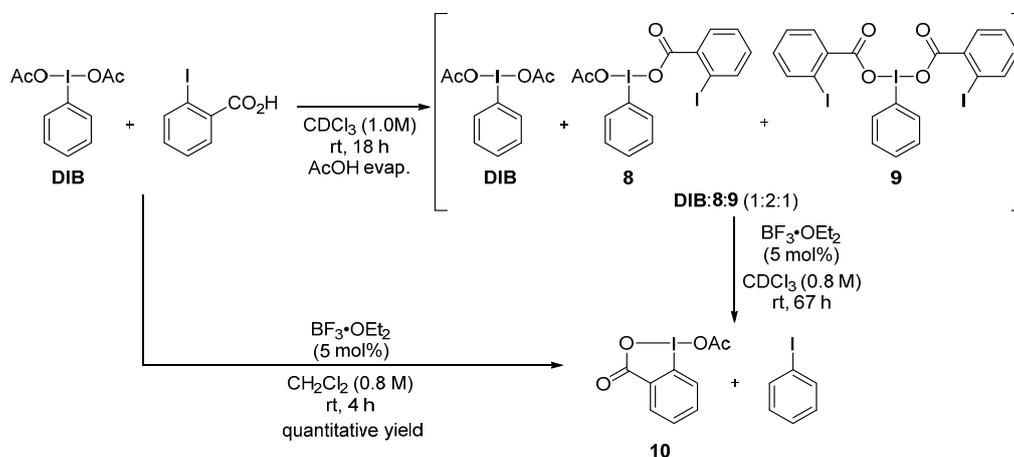
These calculations support the non-reactivity observed for DIB even in non-anhydrous solvents. In contrast, we could expect reactivity from BTI even at room temperature under non-anhydrous conditions. To test this hypothesis, BTI was let to react with *p*-bromoiodobenzene at room temperature without strict exclusion of water. For comparison, a similar experiment involving 5 mol % of  $\text{BF}_3 \cdot \text{OEt}_2$  was also performed. The results are illustrated in Scheme 10.



**Scheme 10.** Evaluation of BTI toward uncatalyzed and catalyzed iodane metathesis.

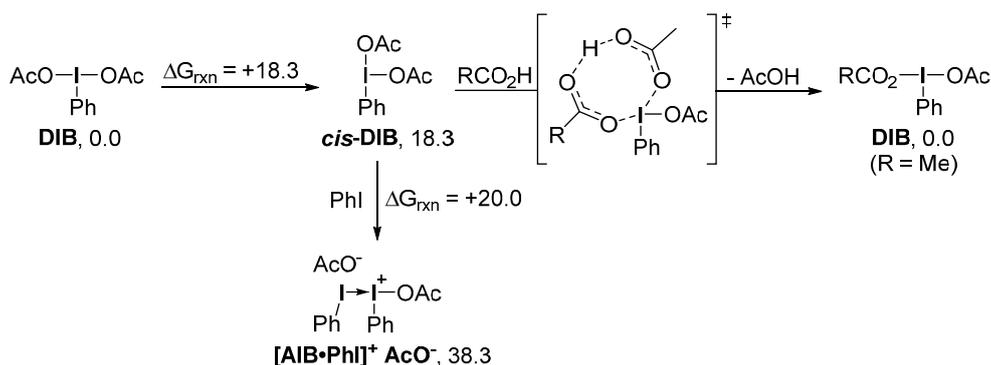
The experimental outcome of the iodane metathesis reactions with BTI is in agreement with the calculation results. Metathesis occurs slowly in the absence of a Lewis acid; equilibrium could not be achieved, even after 41 h. As with DIB, the Lewis acid has a drastic effect on the rate of metathesis, as equilibrium is attained within 19 h. Again, the calculated free energy of reaction is in good agreement with the experimentally observed conversion (Scheme 10).

As the iodane metathesis reactions are under thermodynamic control, we selected a substrate that would favor a complete metathetical redox process. We performed the iodane metathesis using DIB and *o*-iodobenzoic acid (*o*-I-BzOH) as the iodoarene (Scheme 11). The thermodynamic drive (*i.e.*, entropic gain) of the reaction was predicted to lead to a complete conversion ( $\Delta G_{\text{rxn}} = -11.8$  kcal/mol). Additionally, *o*-I-BzOH enables the evaluation of the proximity effect of the iodine(I) and iodine(III) partners on the rate of the metathetical redox reaction. Indeed, *o*-I-BzOH can serve as an acyloxy ligand. We thus performed AcO/*o*-I-BzO exchange on DIB in the absence of a Lewis acid. Treatment of an equimolar mixture of DIB and *o*-I-BzOH in chloroform, followed by AcOH removal through hexane azeotrope, resulted in a clean mixture of DIB, **8**, and **9** in a 1:2:1 ratio, indicative of an almost statistical equilibrium. It is important to note that no iodobenzene was observed at this point. The proximity effect of the I(I) and I(III) partners did not facilitate in any way the metathetical redox process in the absence of a Lewis acid. Treatment of this statistical mixture of [bis(acyloxy)iodo]benzenes with a catalytic quantity of BF<sub>3</sub>·OEt<sub>2</sub> resulted in the smooth and clean conversion to the single cyclic iodane **10** and iodobenzene. The same efficient metathetical redox process could be done in one-pot by treating an equimolar quantity of DIB and *o*-I-BzOH with 5 mol % of BF<sub>3</sub>·OEt<sub>2</sub> for 4 h in dichloromethane. Compound **10** could be obtained in essentially quantitative yield by simple removal of PhI by trituration with hexane.



**Scheme 11.** Ligand exchange between DIB and *o*-I-BzOH followed by iodane metathesis.

There is a need to reconcile the fact that ligand exchange on DIB is facile at room temperature without the need for any activation, while the metathetical redox process is not possible. The nature of the exchange processes needs to be considered. While the calculations suggest a mandatory dissociative process through an iodonium intermediate to achieve metathesis, an associative mechanism could still be considered for ligand exchange (Scheme 12). It was calculated that *cis*-DIB, a T-shaped isomer of DIB in which both acetoxy ligands are *cis* to each other, would be energetically accessible at room temperature. This intermediate neutral iodane would react readily with a carboxylic acid (AcOH was used for simplicity), through proton transfer, to afford in return DIB in its most stable conformation, bearing its two most electron withdrawing groups in *trans* relation. Addition of iodobenzene on *cis*-DIB cannot occur readily, as it is not a strong enough donor to stabilize expulsion of the acetoxy group.



**Scheme 12.** Plausible mechanism for ligand exchange (energies reported in kcal/mol).

In summary, we have demonstrated that, with the simple action of a Lewis acid, it is possible to achieve the metathetical redox reaction of (diacetoxyiodo)arenes and iodoarenes. Both the experimental results and computational insights draw a clearer picture of what is needed for these metathetical redox reactions to proceed. The results presented herein open the way for further exploration.

### 3. Experimental Section

#### 3.1. General Information

All glassware was stored in the oven and/or was flame dried prior to use under an inert atmosphere of gas.  $\text{CDCl}_3$  was dried over anhydrous  $\text{K}_2\text{CO}_3$  but not thoroughly dried. Analytical thin-layer chromatography (TLC) was performed on precoated, glass-backed silica gel (Merck 60 F<sub>254</sub>). Visualization of the developed chromatogram was performed by UV absorbance, aqueous cerium molybdate, ethanolic phosphomolybdic acid, iodine, or aqueous potassium permanganate. Nuclear magnetic resonance spectra ( $^1\text{H}$ ,  $^{13}\text{C}$ , DEPT) were recorded either on an Avance III HD 300 (Bruker, Billerica, MA, USA) or Mercury+ 400 (Agilent Technology, Santa Clara, CA, USA) spectrometers. Chemical shifts for  $^1\text{H}$ -NMR spectra are recorded in parts per million from tetramethylsilane with the solvent resonance as the internal standard. Data are reported as follows: chemical shift, multiplicity (s = singlet, d = doublet, t = triplet, q = quartet, qn = quintet, sext = sextuplet, m = multiplet and br = broad), coupling constant in Hz, integration. Chemical shifts for  $^{13}\text{C}$ -NMR spectra are recorded in parts per million from tetramethylsilane with the solvent resonance as the internal standard. All spectra were obtained with complete proton decoupling. When ambiguous, proton and carbon assignments were established using COSY, NOESY, HMQC and DEPT experiments. High resolution mass spectra were recorded on a Maxis ESI-Q-Tof (Bruker, Billerica, MA, USA) at the Université de Sherbrooke.

### 3.2. Metathetical Redox Reaction of HTIB and *p*-Methyliodobenzene

To a round-bottom flask were added [hydroxy(tosyloxy)iodo]benzene (HTIB) (333.4 mg, 0.850 mmol), *p*-methyliodobenzene (185.3 mg, 0.850 mmol) and dichloromethane (1.7 mL, 0.5 M). The white suspension was stirred at room temperature. Aliquots (5–10  $\mu$ L) were taken from clear dichloromethane supernatant after allowing the suspension to settle and diluted in HPLC-grade hexanes (500–750  $\mu$ L) at various moments. Conversions were determined by GC-MS by integration of signals of iodobenzene and *p*-methyliodobenzene.

### 3.3. Lewis-Acid-Catalyzed Metathetical Redox Reaction of DIB and *p*-Bromiodobenzene

To a flame-dried vial equipped with a septum, under an argon atmosphere, were added DIB (372.1 mg, 1.155 mmol), *p*-bromiodobenzene (325.6 mg, 1.155 mmol) and CDCl<sub>3</sub> (442  $\mu$ L). Diluted BF<sub>3</sub>·Et<sub>2</sub>O (713  $\mu$ L of a freshly made solution of BF<sub>3</sub>·Et<sub>2</sub>O [100  $\mu$ L BF<sub>3</sub>·Et<sub>2</sub>O in 10 mL CDCl<sub>3</sub>], 0.058 mmol) was added for a total volume of 1.16 mL of CDCl<sub>3</sub> (1 M). The solution was stirred at room temperature under inert atmosphere. Conversions were determined by integration of well-defined <sup>1</sup>H-NMR signals (5 s relaxation time) of the two (diacetoxyiodo)arenes present in solution. The free energies of reaction (Equation (1) and Scheme 10) were determined with the measured conversions (see Supplementary Materials for more details).

### 3.4. One-Pot Metathetical Redox Reaction of DIB and *o*-Iodobenzoic Acid

To a flame-dried vial, under argon, were added *o*-iodobenzoic acid (247.7 mg, 1.00 mmol), DIB (322.1 mg, 1.00 mmol) and a solution of BF<sub>3</sub>·Et<sub>2</sub>O (1.22 mL, [100  $\mu$ L BF<sub>3</sub>·Et<sub>2</sub>O in 10 mL CH<sub>2</sub>Cl<sub>2</sub>], 0.012 mmol). The final reaction concentration is 0.8 M. The solution was allowed to stir at room temperature for 4 h. The solvent was evaporated and the crude residue washed twice with hexane. The crude product was dried under vacuum to afford compound **10** in quantitative yield. The spectral data is consistent with the reported one in the literature [26].

## 4. Computational Details

The geometry optimizations were done using the Gaussian 09 software package [27] with the M06-2X [28] density functional in combination with the 6-31+G(d,p) basis set [29,30] for all atoms except iodine, for which LANL2DZdp + ECP was used [31,32]. The structures were optimized with a solvation model (SMD) for chloroform [33]. Unless otherwise stated, a fine grid density was used for numerical integration in the calculations. Conformational searches were done on all the species described throughout the mechanistic pathways. Harmonic vibrational frequencies were computed for all optimized structures to verify that they were either minima or transition states, possessing zero or one imaginary frequency, respectively. All the free energies are reported in kcal/mol and incorporate unscaled free energy corrections based on the vibrational analyses and temperature of 298 K.

**Supplementary Materials:** Supplementary materials can be accessed at: <http://www.mdpi.com/1420-3049/20/12/19874/s1>. Cartesian coordinates, electronic and zero-point vibrational energies.

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**Author Contributions:** A.J.-D.L. and C.Y.L. conceived and designed the experiments; A.J.-D.L. performed the experiments; C.Y.L. performed the computational calculations; C.Y.L. wrote the paper; A.J.-D.L. helped with the correction of the paper.

**Conflicts of Interest:** The authors declare no conflict of interest.

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**Sample Availability:** Not available.



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