

Supporting Information

Ti group metallocene-catalyzed synthesis of 1-hexene dimers and tetramers

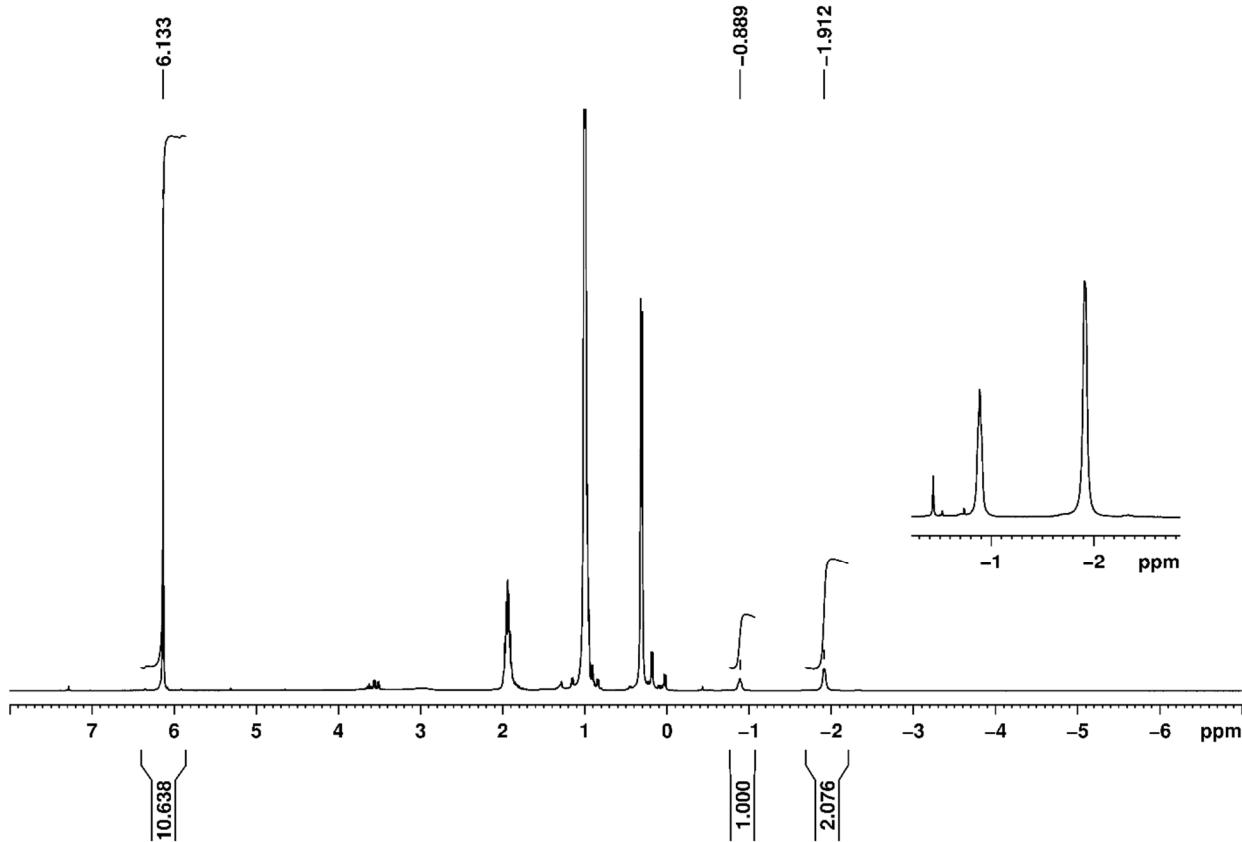
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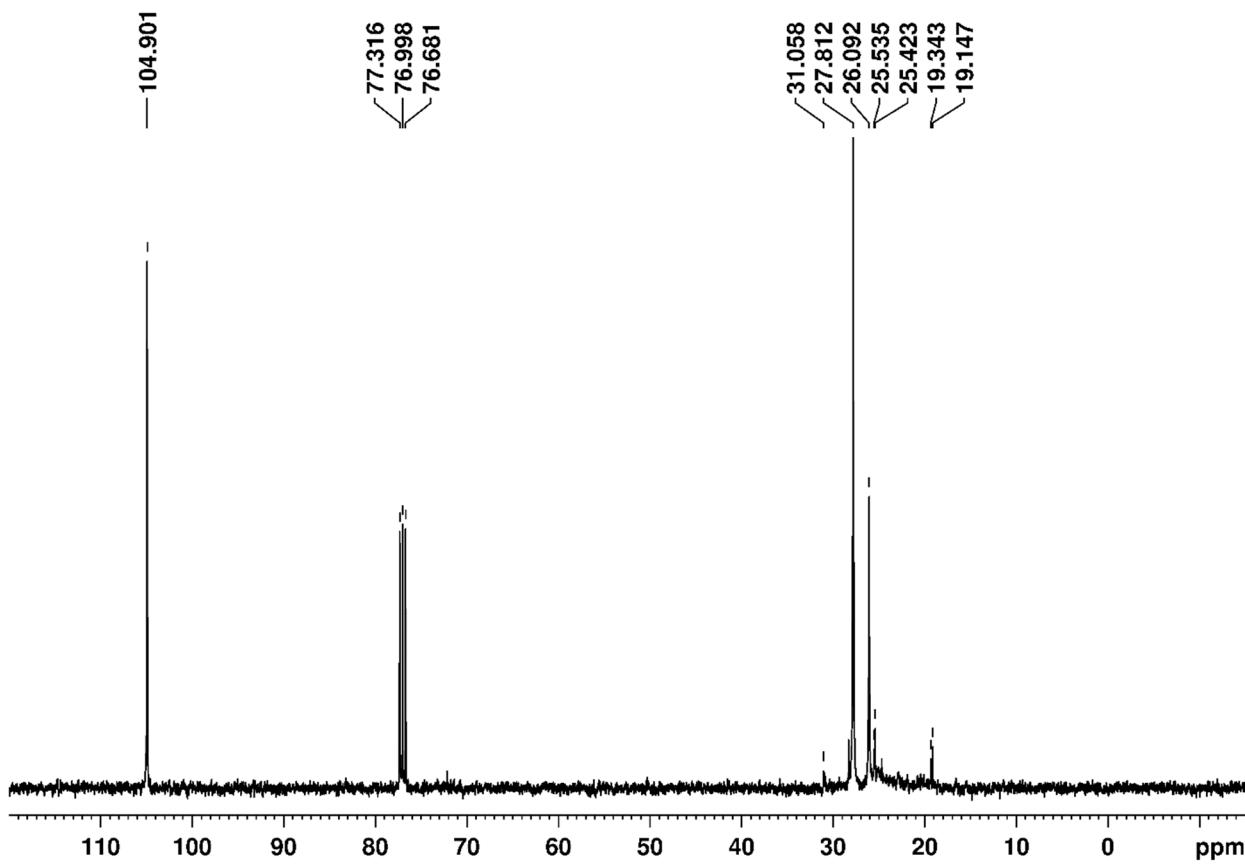


Figure S3. ¹H NMR of system Cp₂ZrCl₂-HAIBu₄-MMAO-12-1-hexene (1:4:8:0.1) in CDCl₃ (T = 299 K).

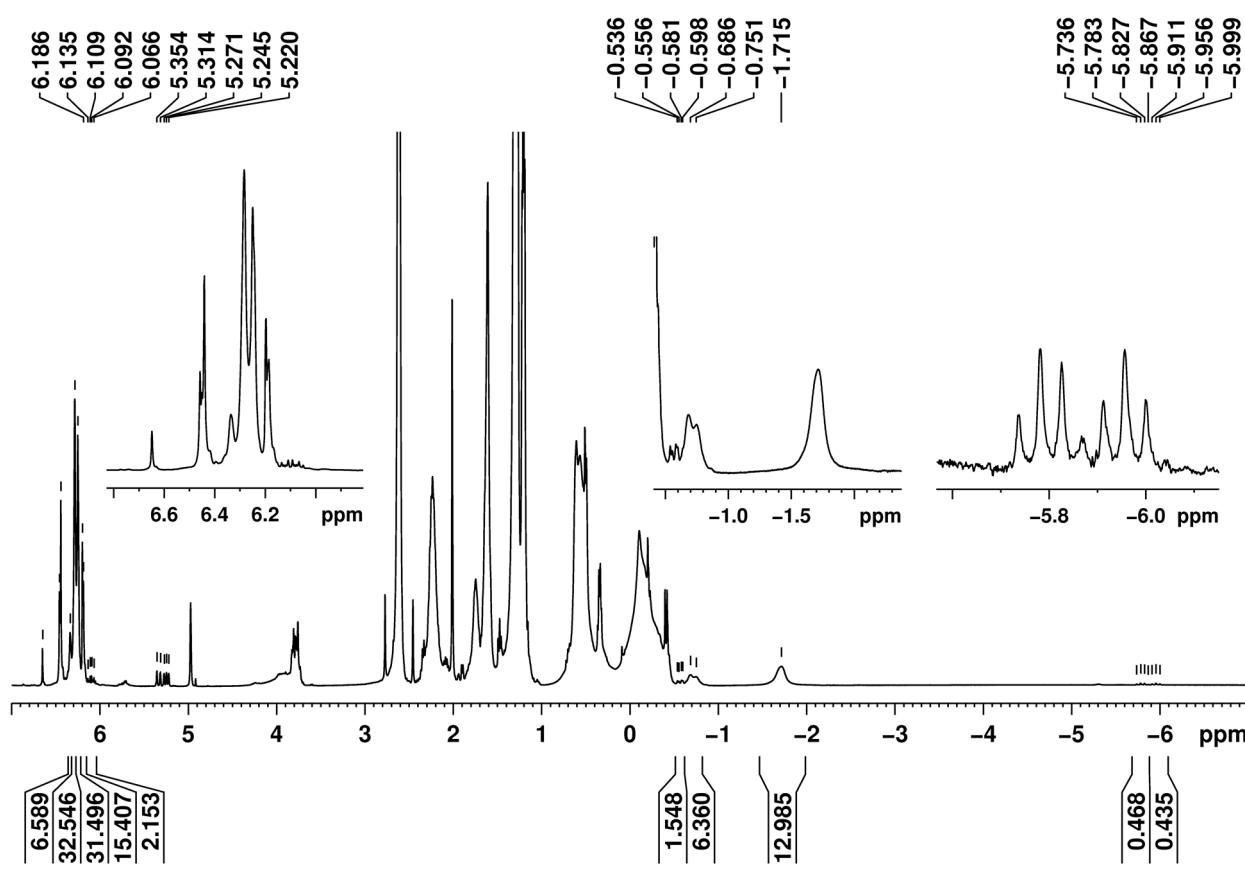


Figure S4. ^{13}C NMR of system $\text{Cp}_2\text{ZrCl}_2\text{--HAlBu}_2\text{--MMAO-12--1-hexene}$ (1:4:8:0.1) in CDCl_3 ($T=299\text{ K}$).

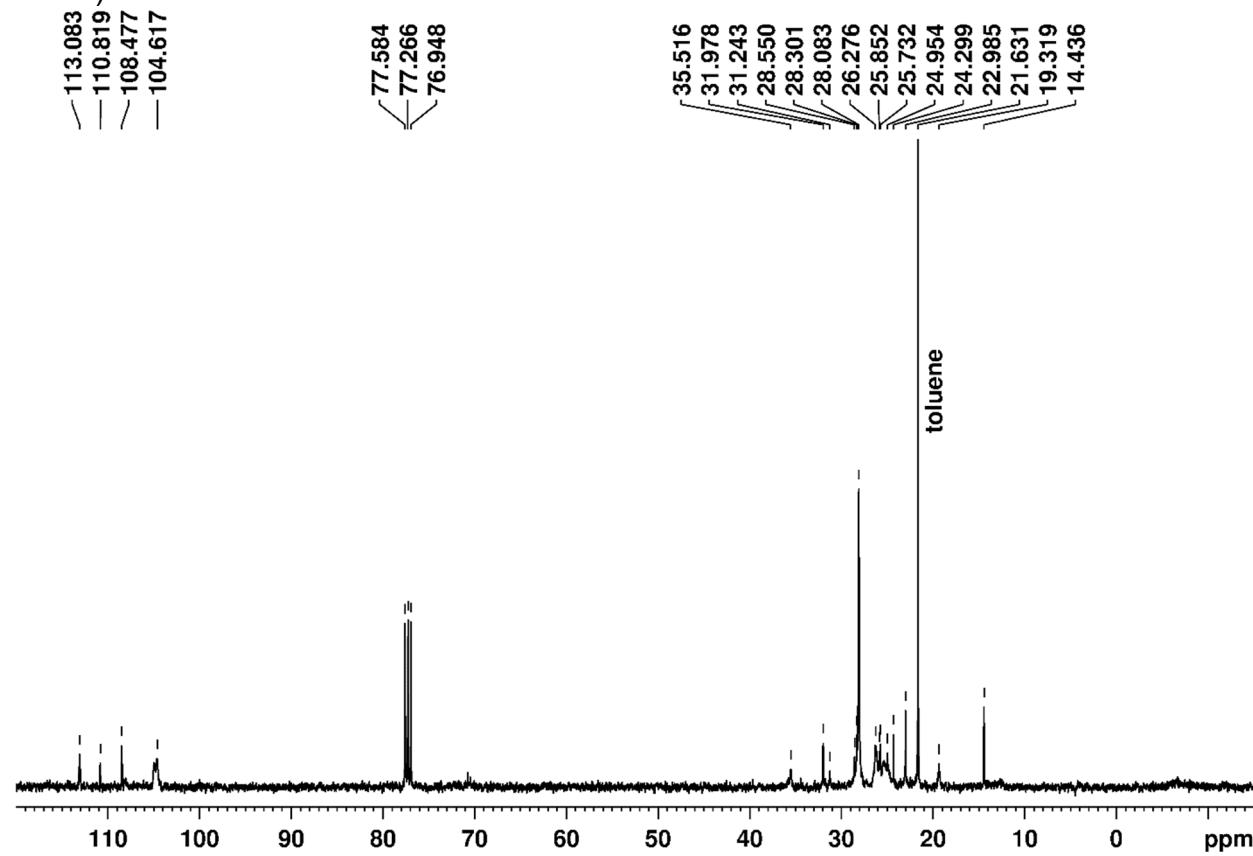


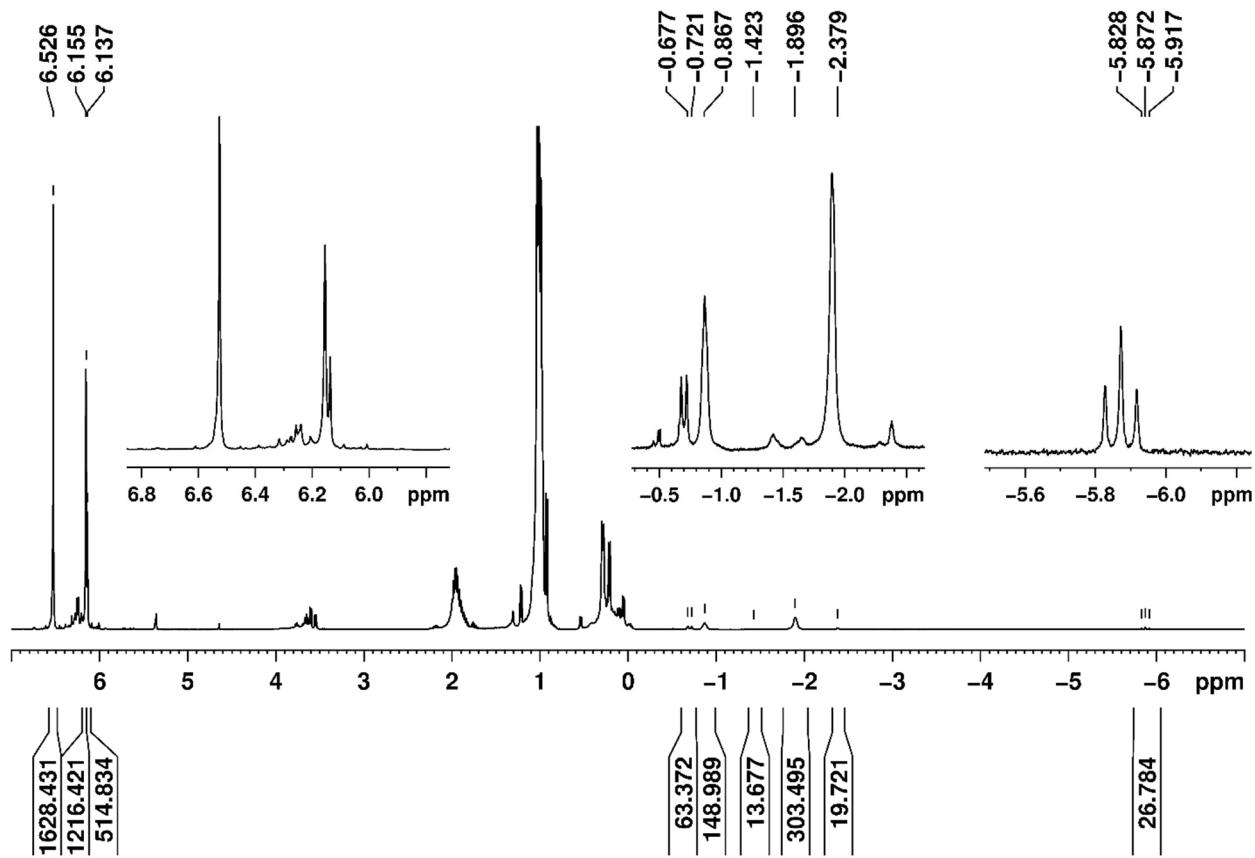
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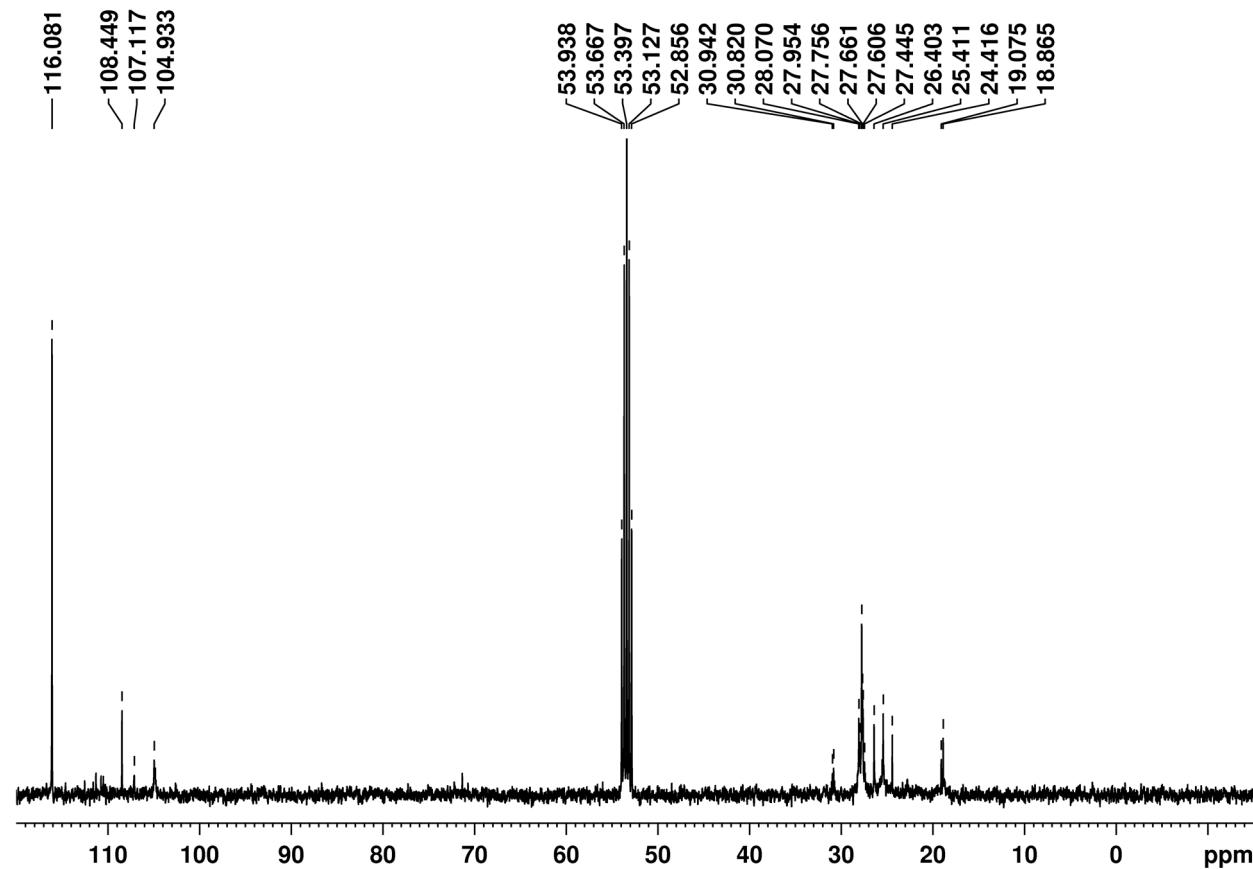
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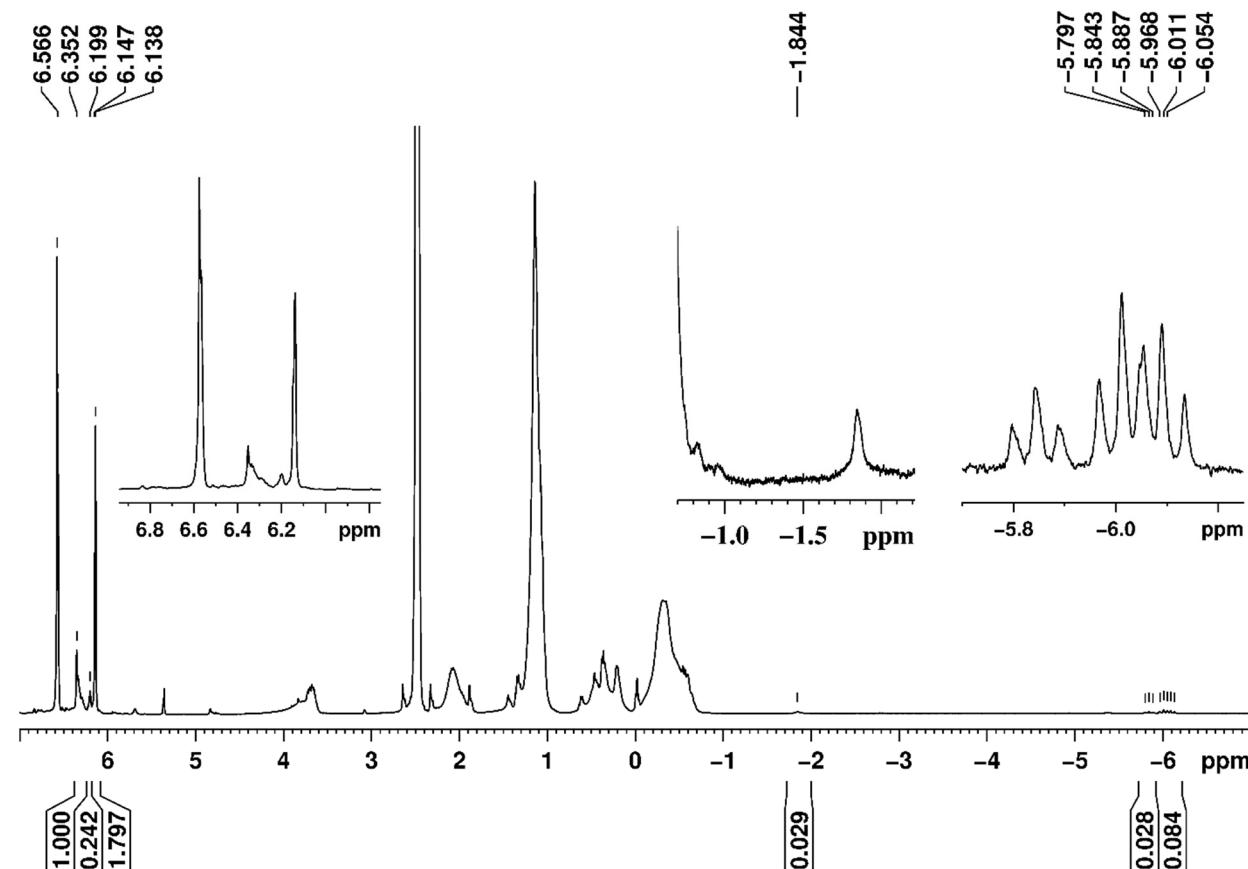
Figure S7. ^1H NMR of system $\text{Cp}_2\text{ZrCl}_2\text{--HAlBu}_2\text{--MMAO-12}$ (1:2:5) in CD_2Cl_2 ($T = 298.6\text{ K}$).

Figure S8. ^{13}C NMR of system $\text{Cp}_2\text{ZrCl}_2\text{--HAlBu}^{\text{i}}_2\text{--MMAO-12}$ (1:2:5) in CD_2Cl_2 ($T = 299.3\text{ K}$).

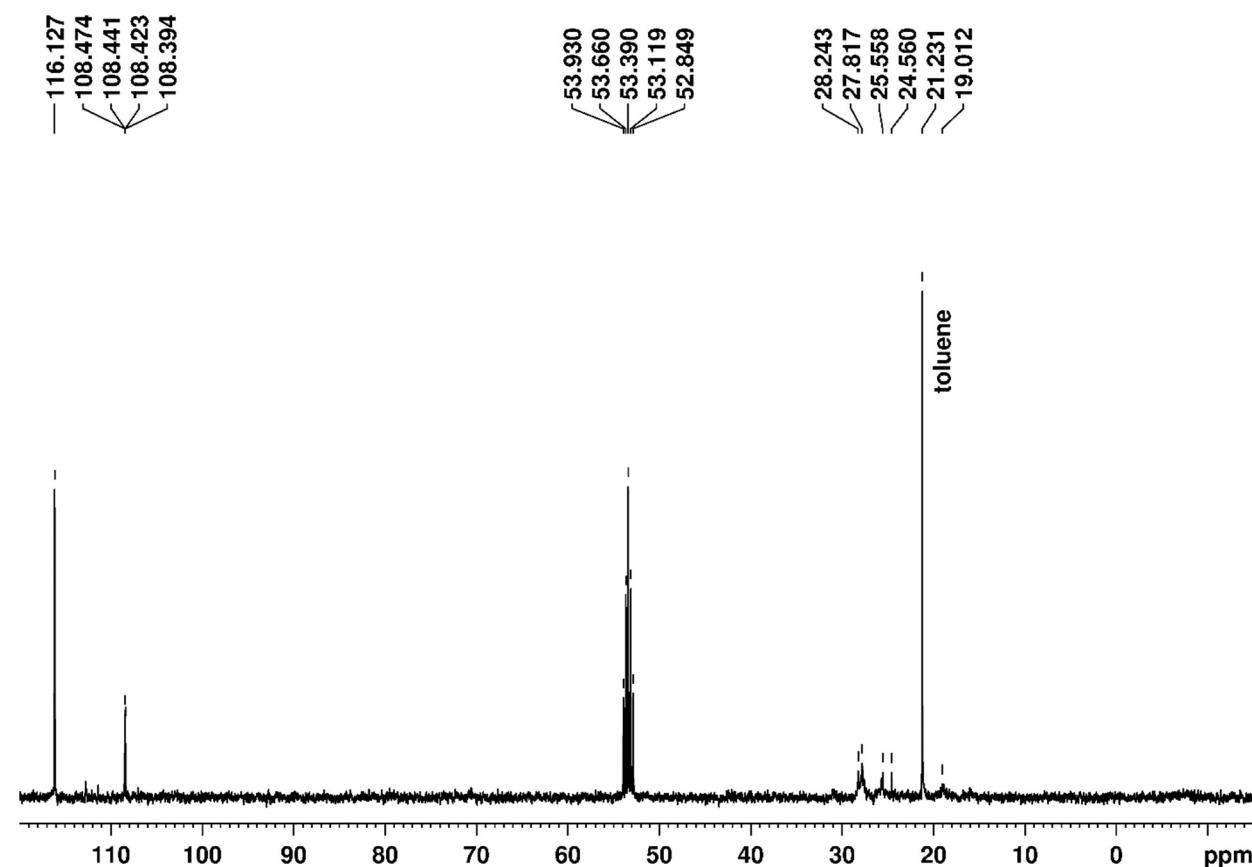


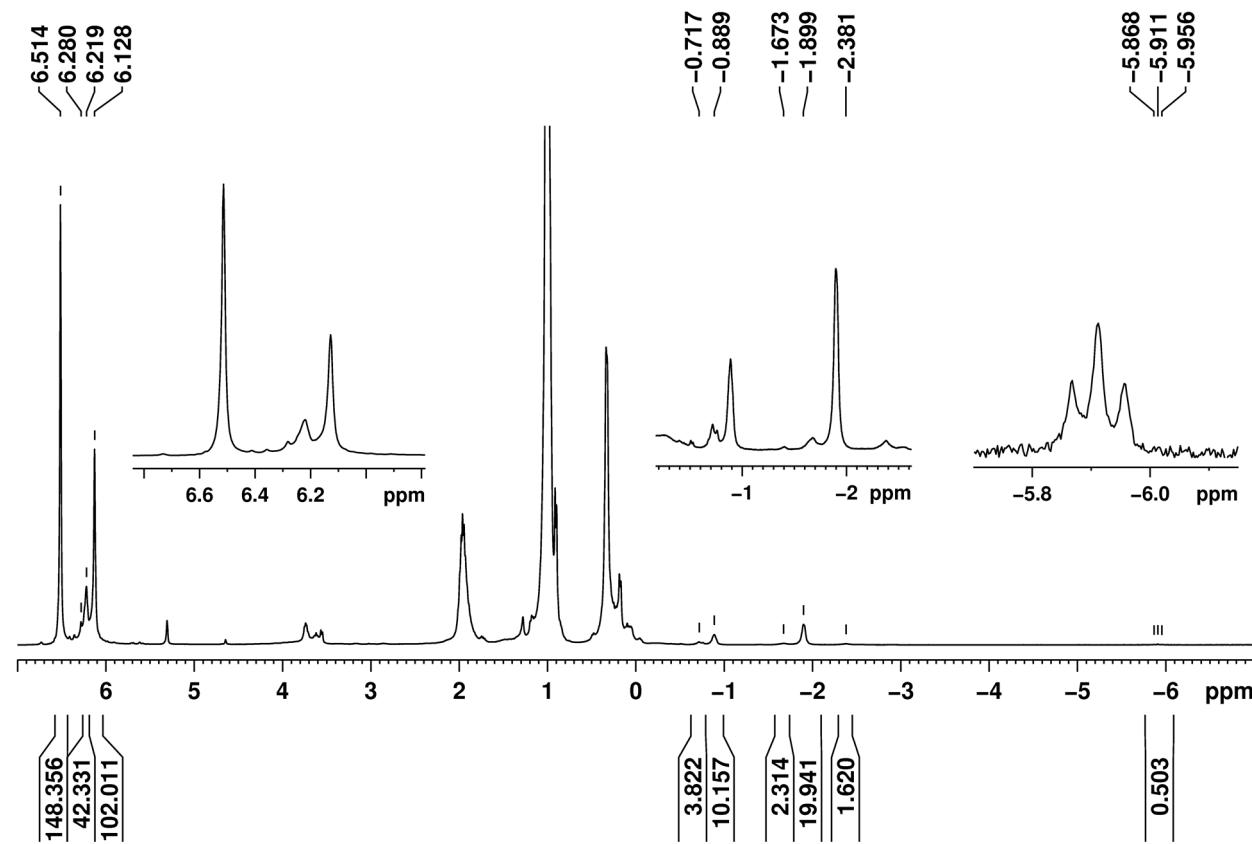
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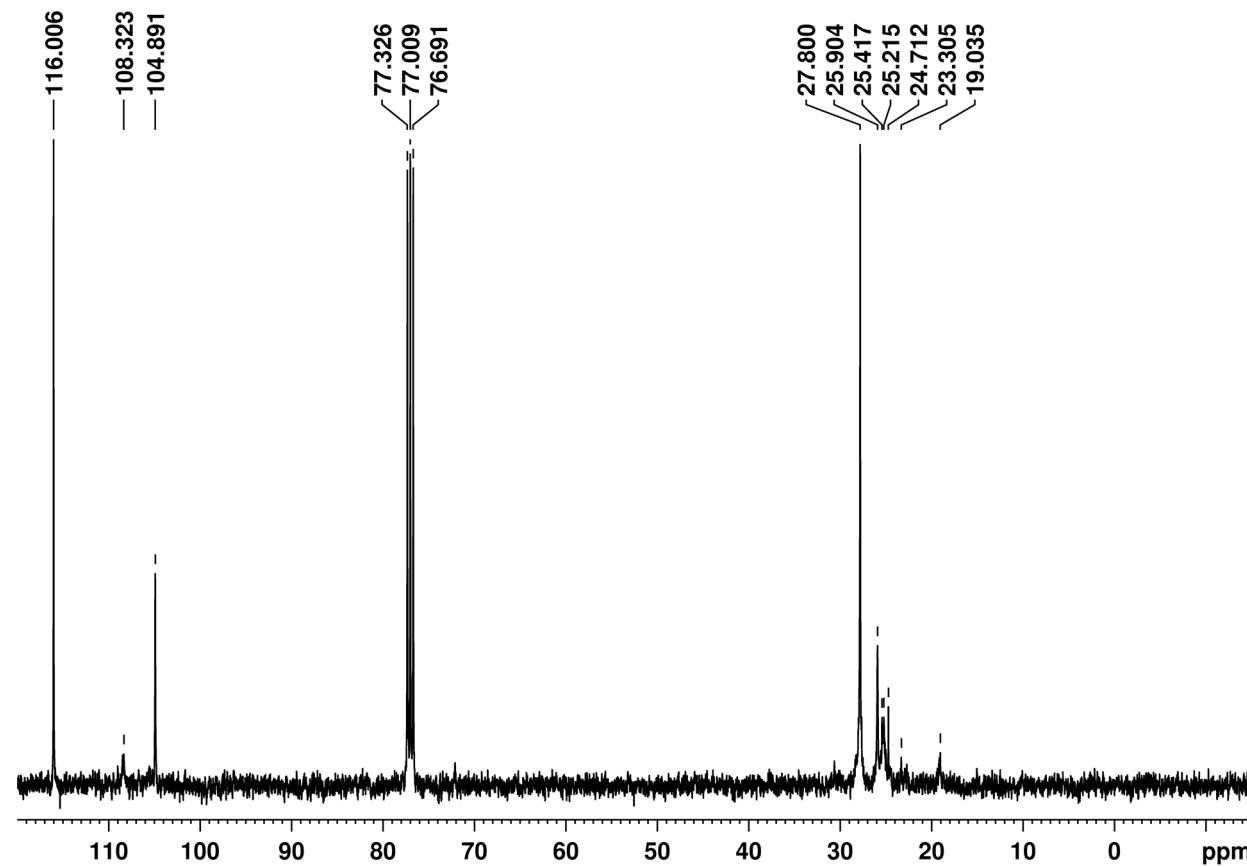
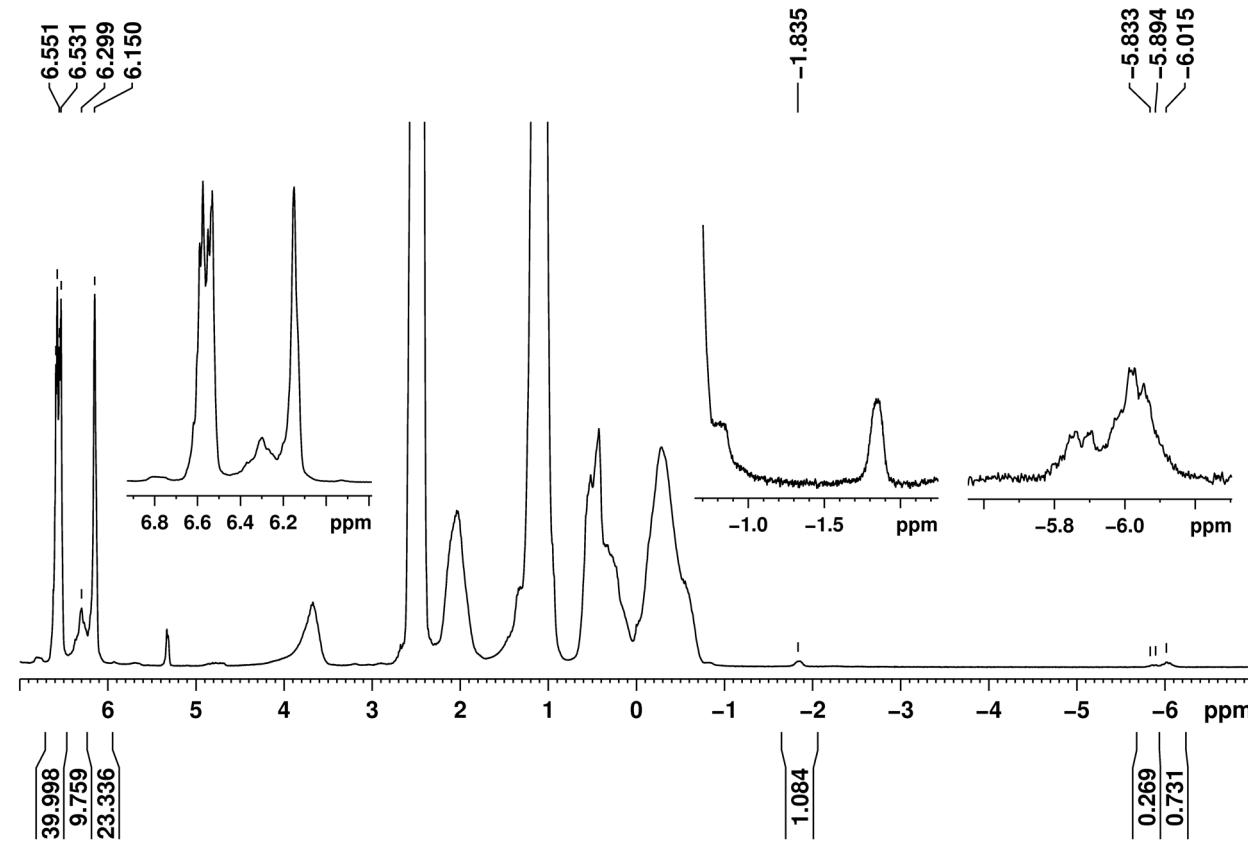
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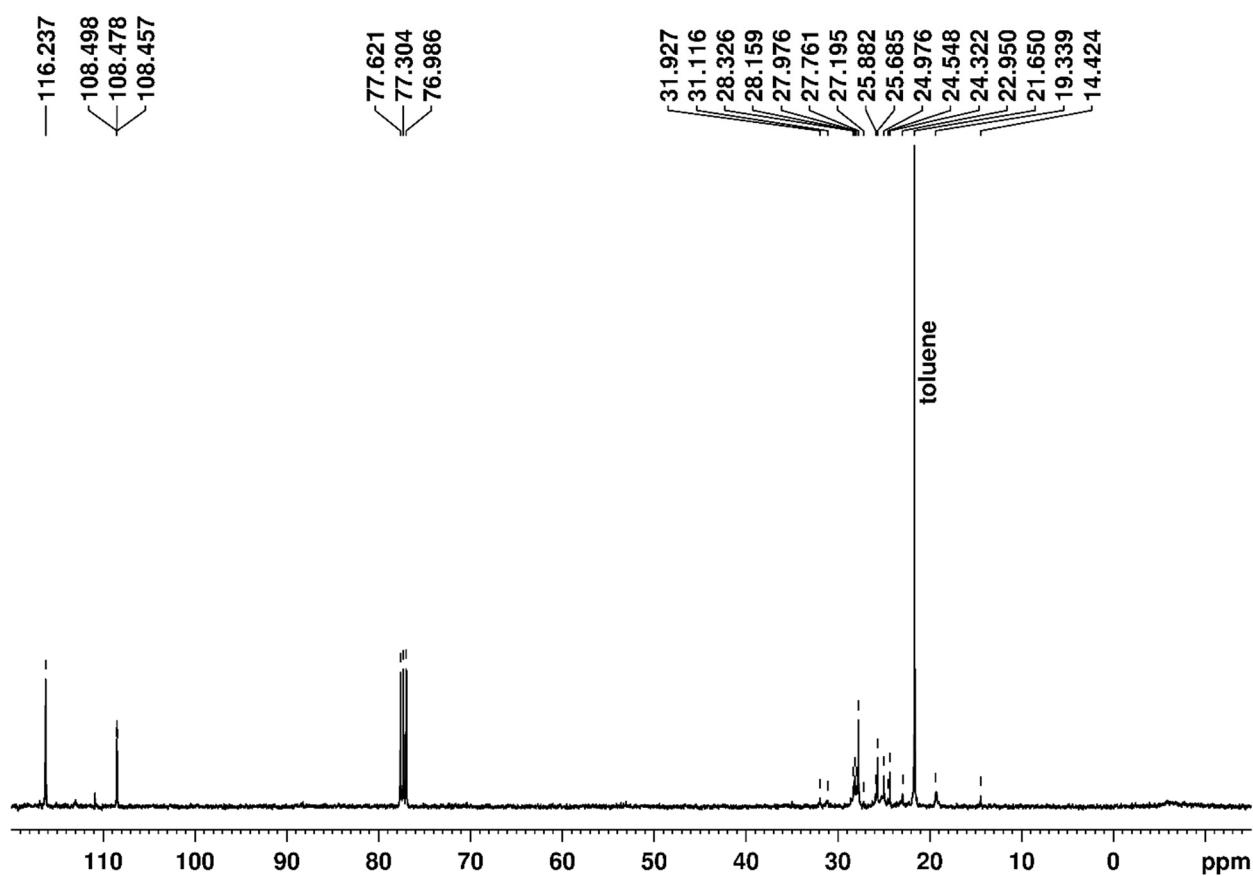
Figure S12. ^{13}C NMR of system $[\text{Cp}_2\text{ZrH}_2]_2\text{--CIAIBu}_2\text{--MMAO-12}$ (1:2:11) in CDCl_3 ($T = 299.2\text{ K}$).

Figure S13. COSY HH of system $[\text{Cp}_2\text{ZrH}_2]_2\text{--ClAlBu}^{\text{i}}_2\text{--MMAO-12}$ (1:2:11) in CDCl_3 ($T = 298.7\text{ K}$).

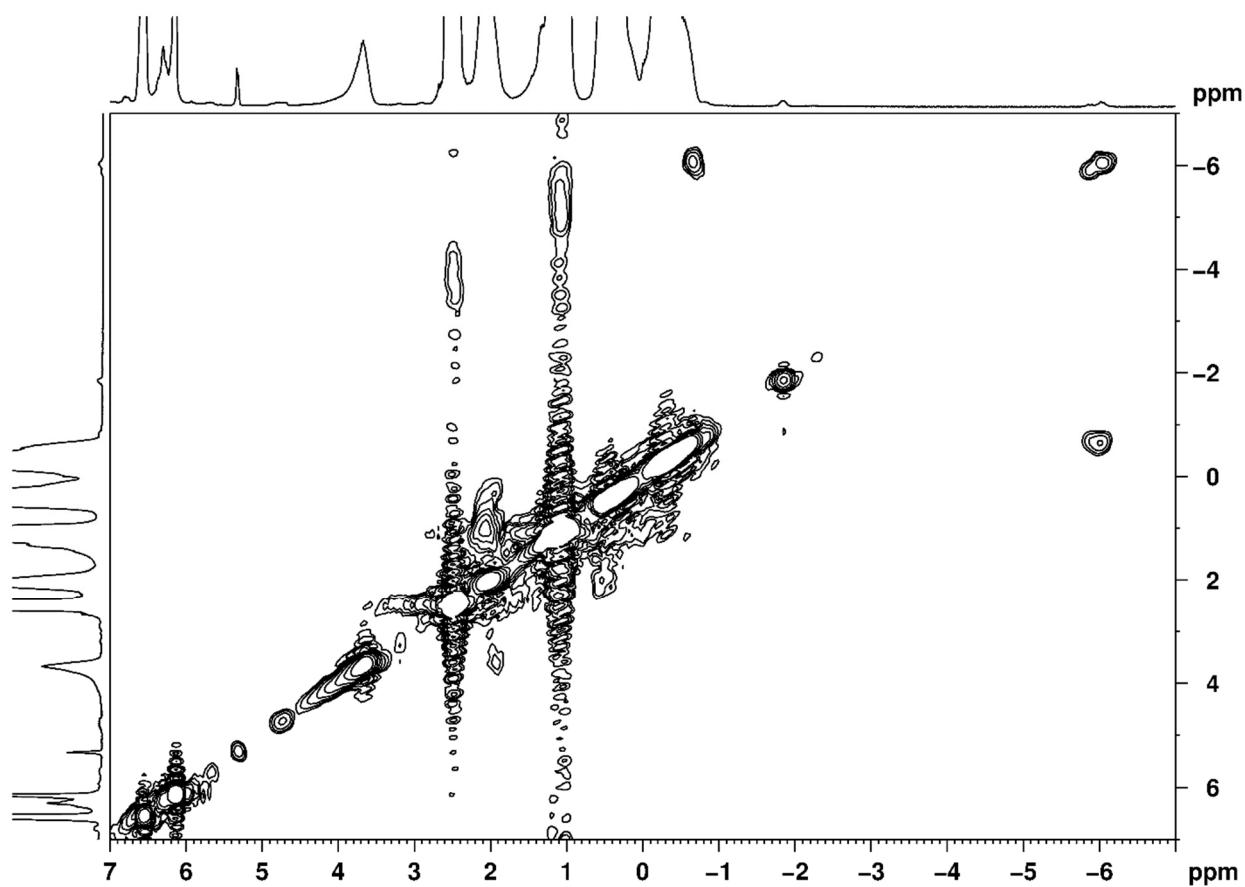


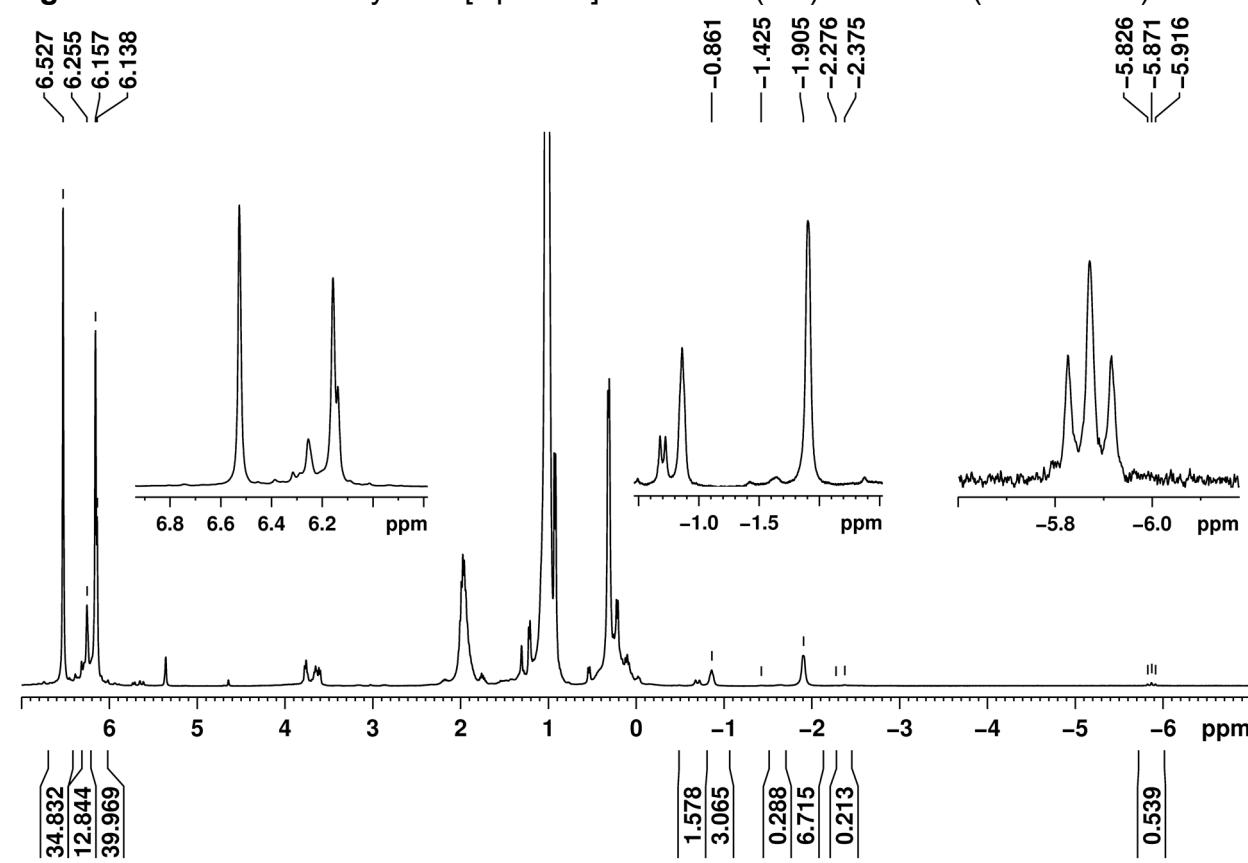
Figure S14. ^1H NMR of system $[\text{Cp}_2\text{ZrH}_2]_2\text{--ClAlBu}^\text{i}_2$ (1:2) in CD_2Cl_2 ($T = 299.1\text{ K}$).

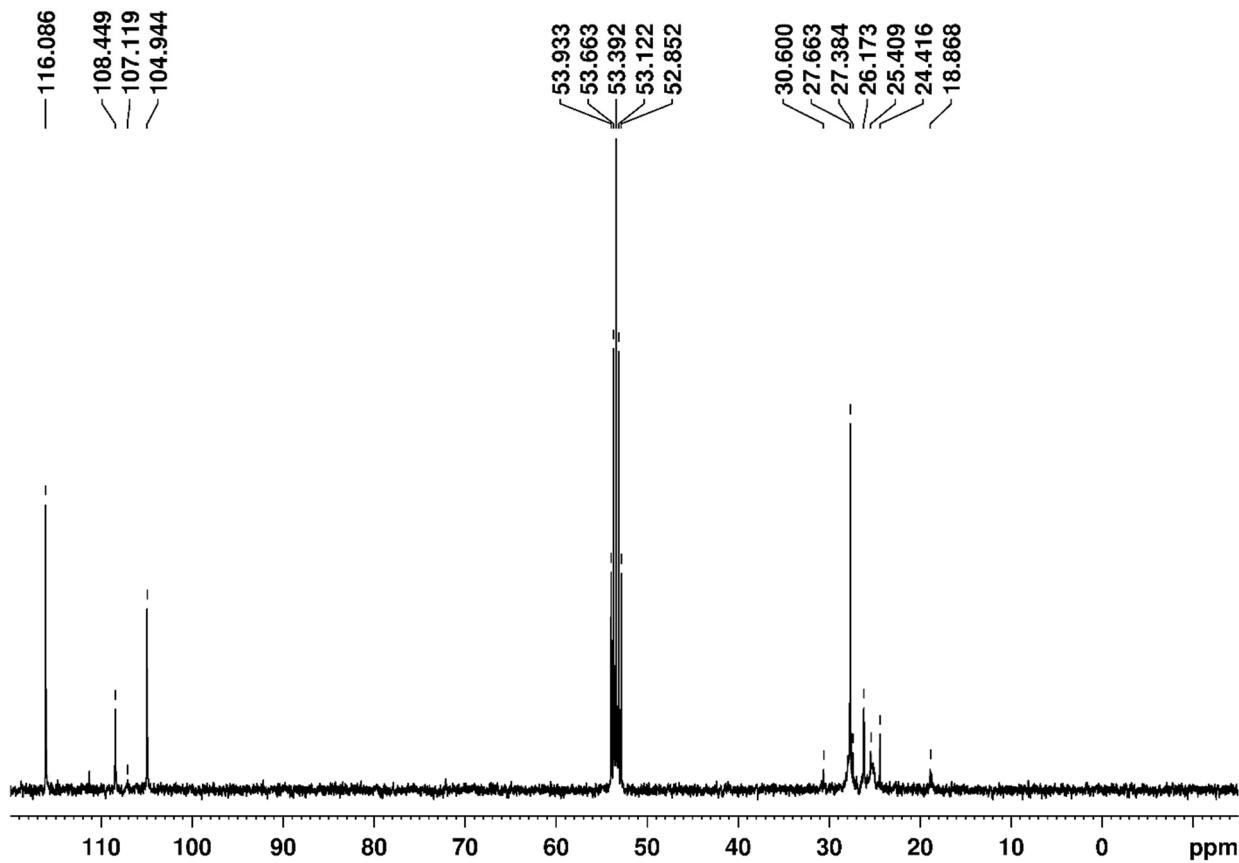
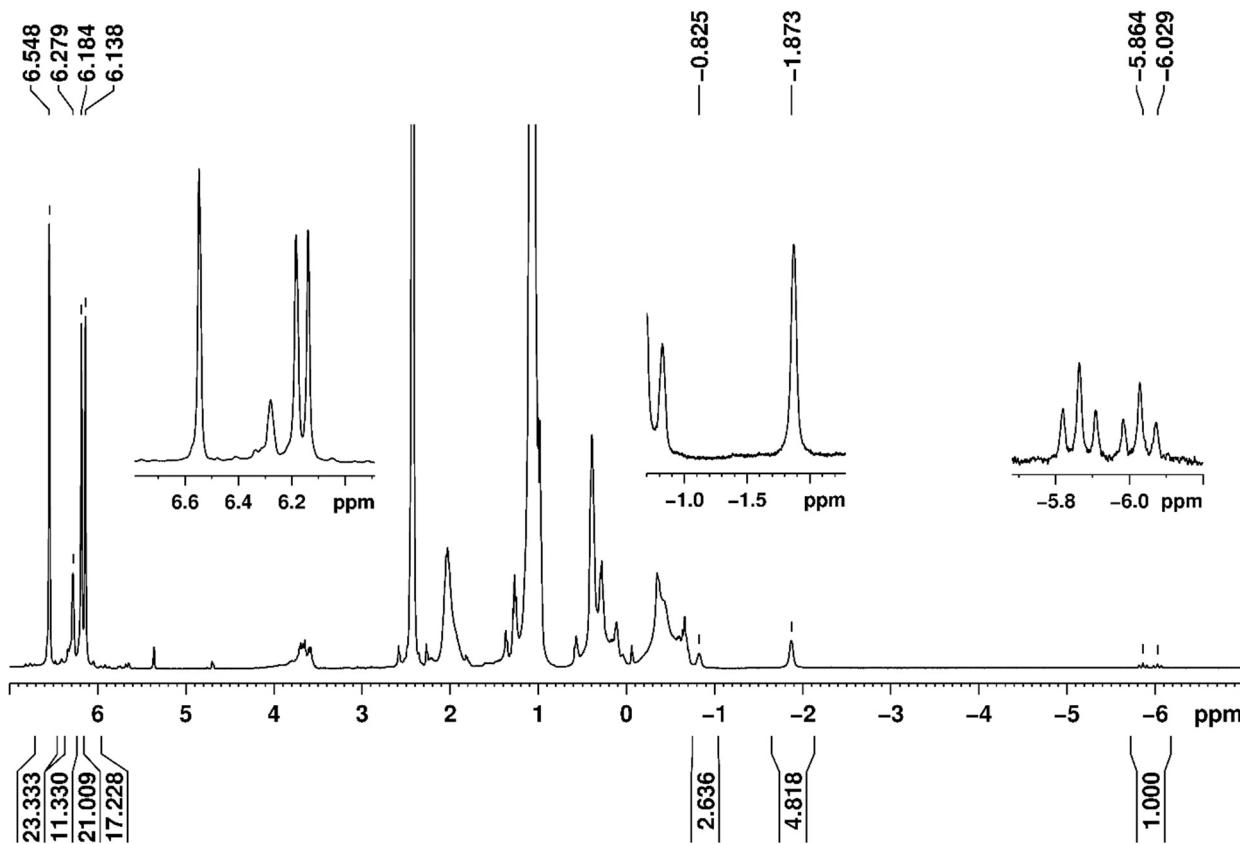
Figure S15. ^{13}C NMR of system $[\text{Cp}_2\text{ZrH}_2]_2\text{--CIAIBu}^{\ddagger}_2$ (1:2) in CD_2Cl_2 ($T = 299.7\text{ K}$).**Figure S16.** ^1H NMR of system $[\text{Cp}_2\text{ZrH}_2]_2\text{--CIAIBu}^{\ddagger}_2\text{--MMAO-12}$ (1:2:5) in CD_2Cl_2 ($T = 299.1\text{ K}$).

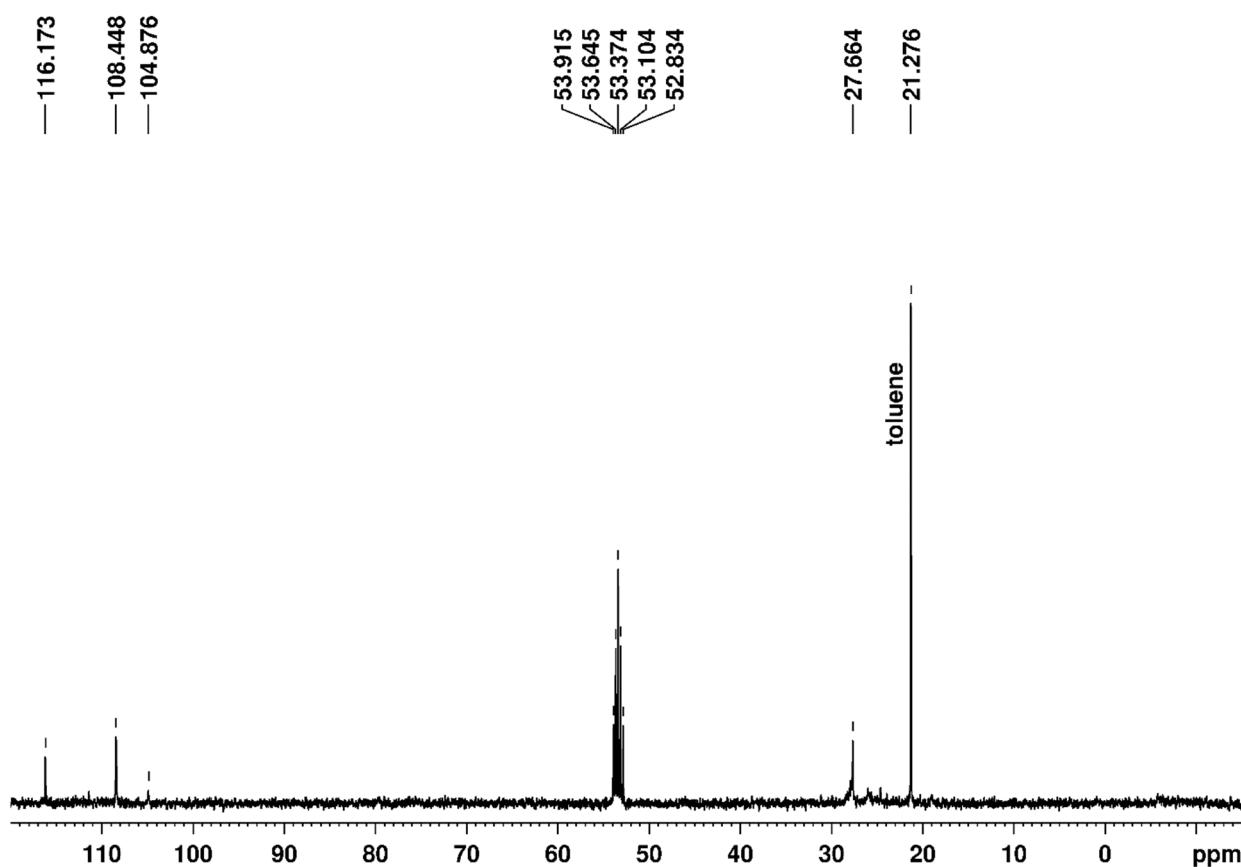
Figure S17. ^{13}C NMR of system $[\text{Cp}_2\text{ZrH}_2]_2\text{--CIAIBu}_2\text{--MMAO-12}$ (1:2:5) in CD_2Cl_2 ($T = 299.5\text{ K}$).

Figure S18. COSY HH NMR of system $[\text{Cp}_2\text{ZrH}_2]_2\text{--ClAlBu}_2\text{--MMAO-12}$ (1:2:5) in CD_2Cl_2 ($T = 298.7$ K).

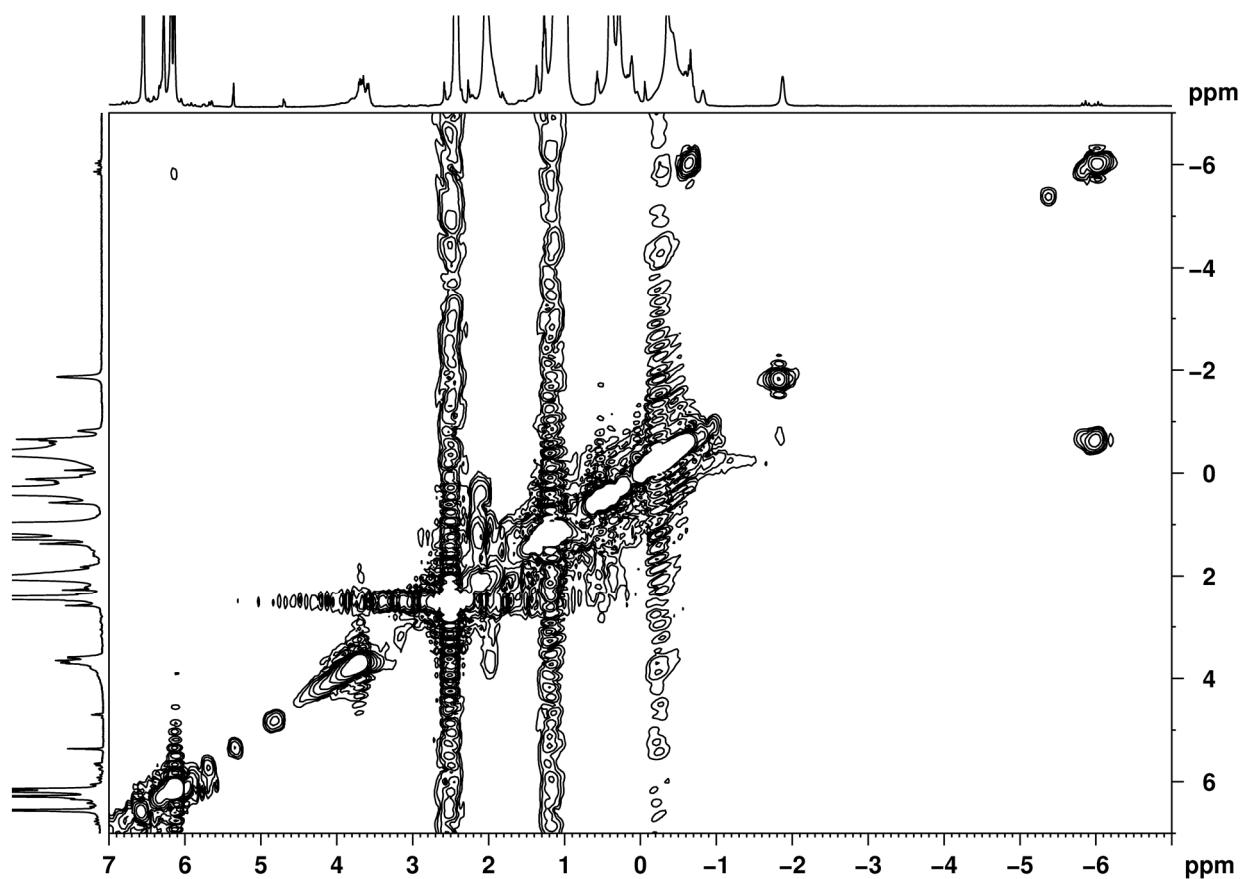


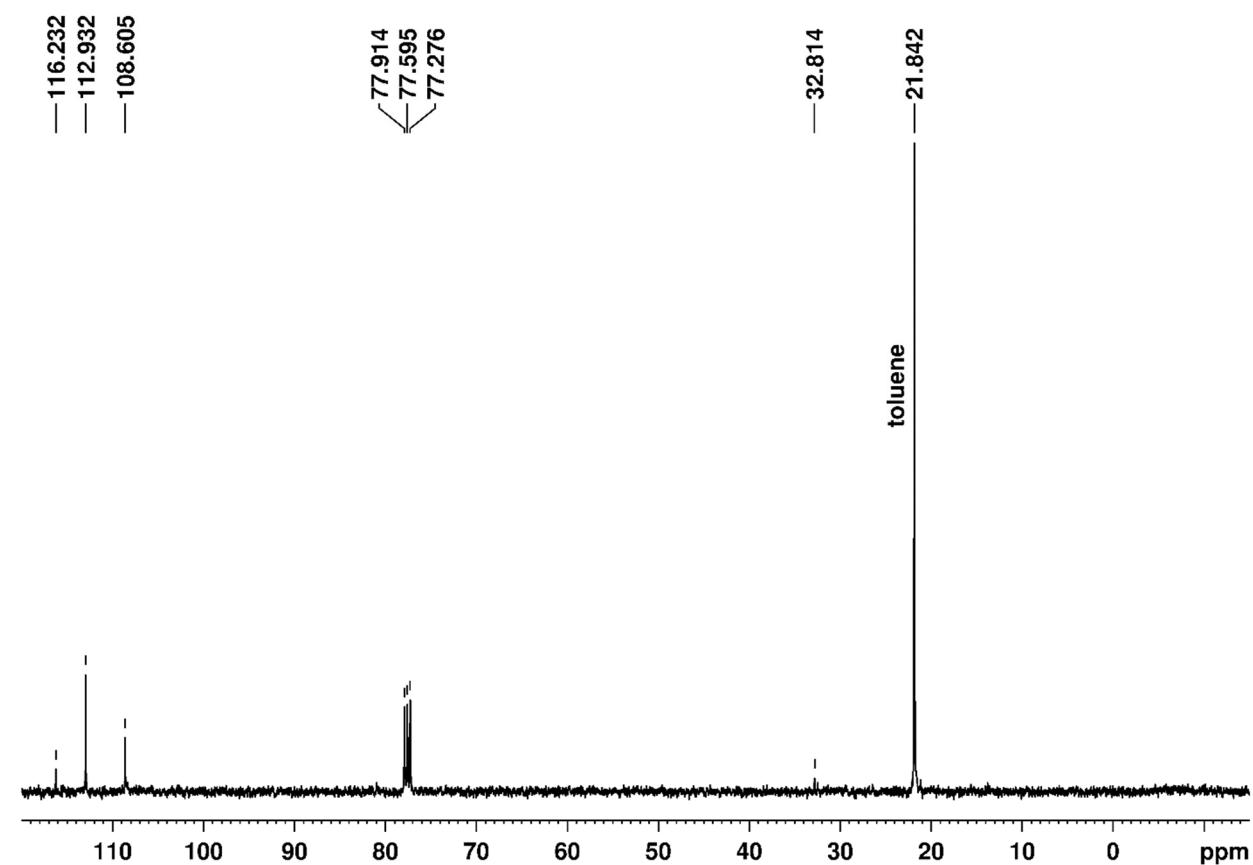
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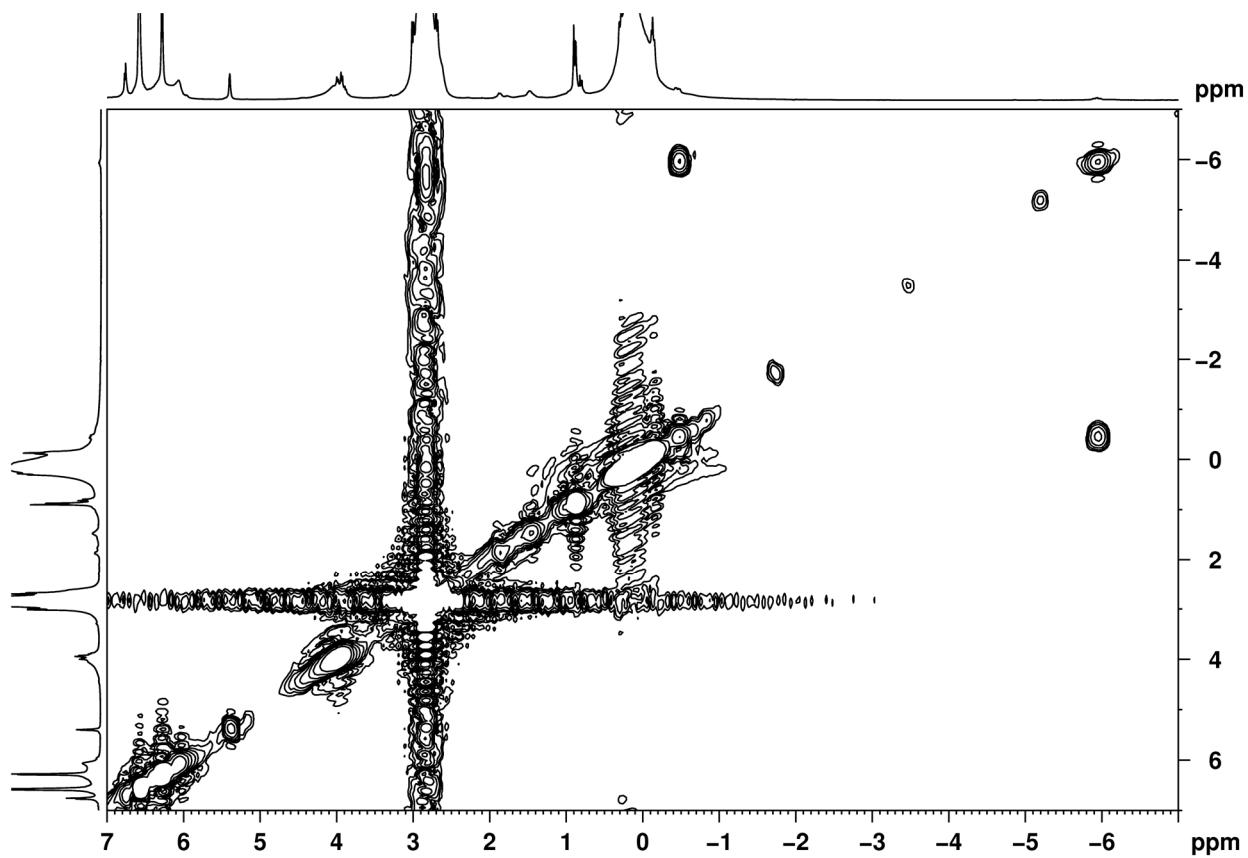
Figure S20. COSY HH NMR of system $[\text{Cp}_2\text{ZrH}_2]_2\text{--MMAO-12}$ (1:12) in CDCl_3 ($T = 296 \text{ K}$).

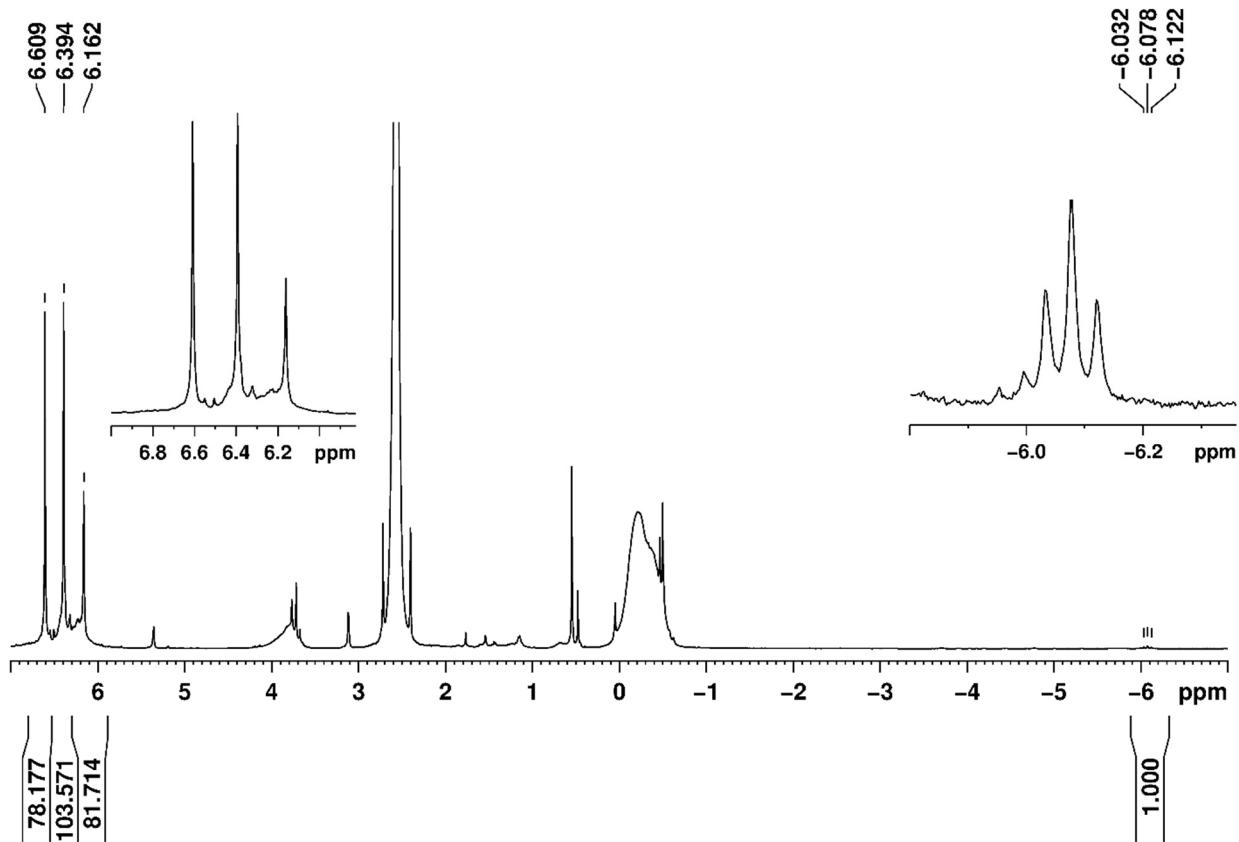
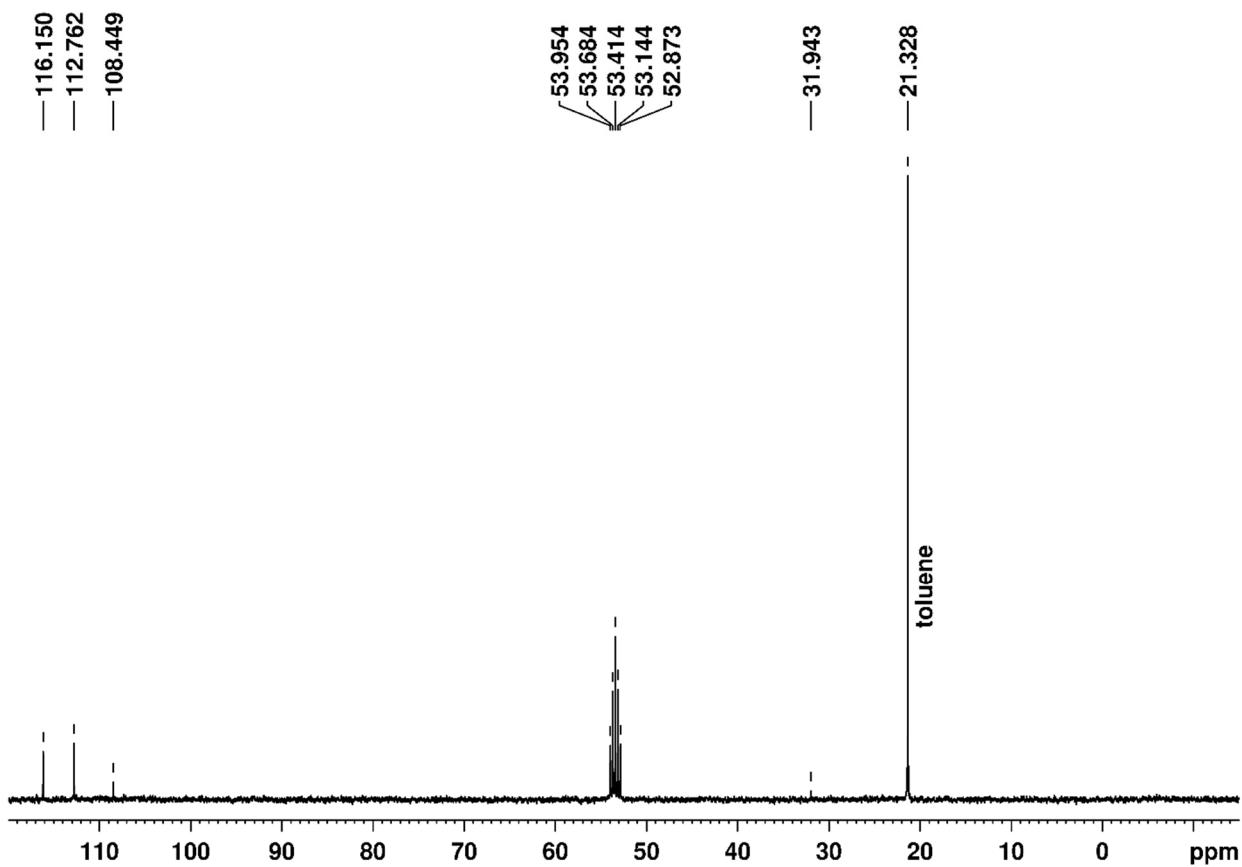
Figure S21. ^1H NMR of system $[\text{Cp}_2\text{ZrH}_2]_2$ – MMAO-12 (1:7) in CD_2Cl_2 ($T = 296.3\text{ K}$).**Figure S22.** ^{13}C NMR of system $[\text{Cp}_2\text{ZrH}_2]_2$ – MMAO-12 (1:7) in CD_2Cl_2 ($T = 296.6\text{ K}$).

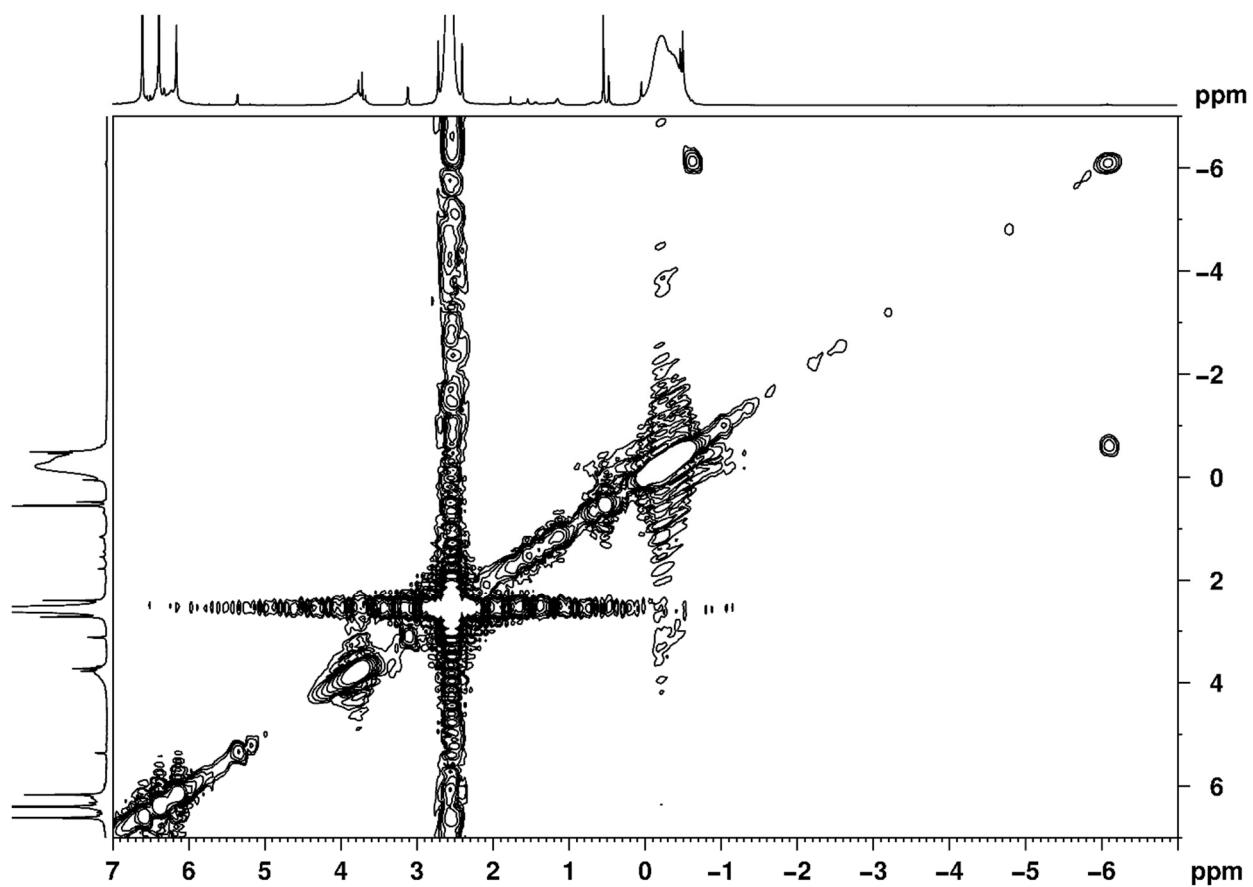
Figure S23. COSY HH of system $[\text{Cp}_2\text{ZrH}_2]_2$ – MMAO-12 (1:7) in CD_2Cl_2 ($T = 296 \text{ K}$)

Figure S24. ^1H NMR of dimer **5** obtained in the system $\text{Cp}_2\text{ZrCl}_2\text{--HAlBu}^{\text{i}}_2\text{--MMAO-12}$, $[\text{Zr}]:[\text{Al}]:[\text{MMAO-12}]:[1\text{-alkene}] = 1:4:8:1$ in CDCl_3 ($T = 298.7\text{ K}$), 20 min.

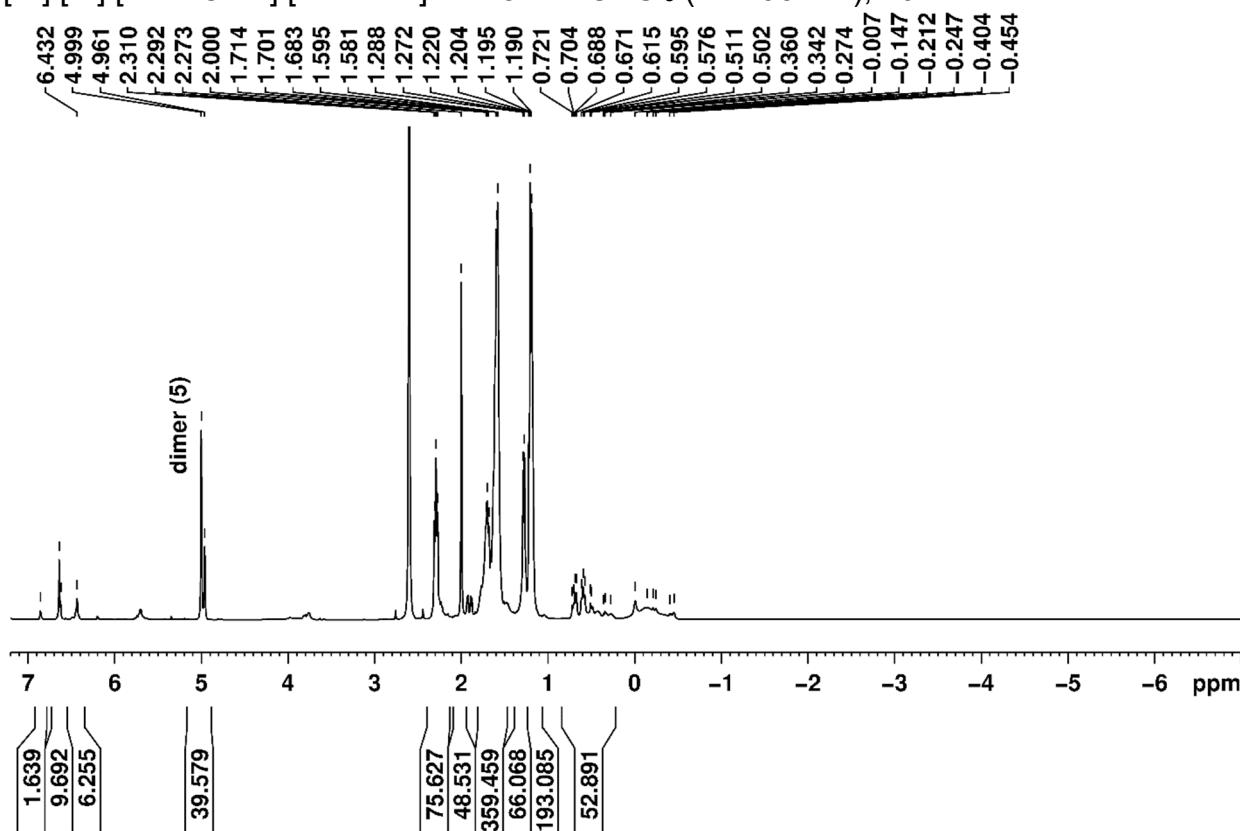


Figure S25. ^{13}C NMR of dimer **5** obtained in the system $\text{Cp}_2\text{ZrCl}_2\text{--HAlBu}^{\text{i}}_2\text{--MMAO-12}$, $[\text{Zr}]:[\text{Al}]:[\text{MMAO-12}]:[1\text{-alkene}] = 1:4:8:1$ in CDCl_3 ($T = 298.7\text{ K}$), 20 min.

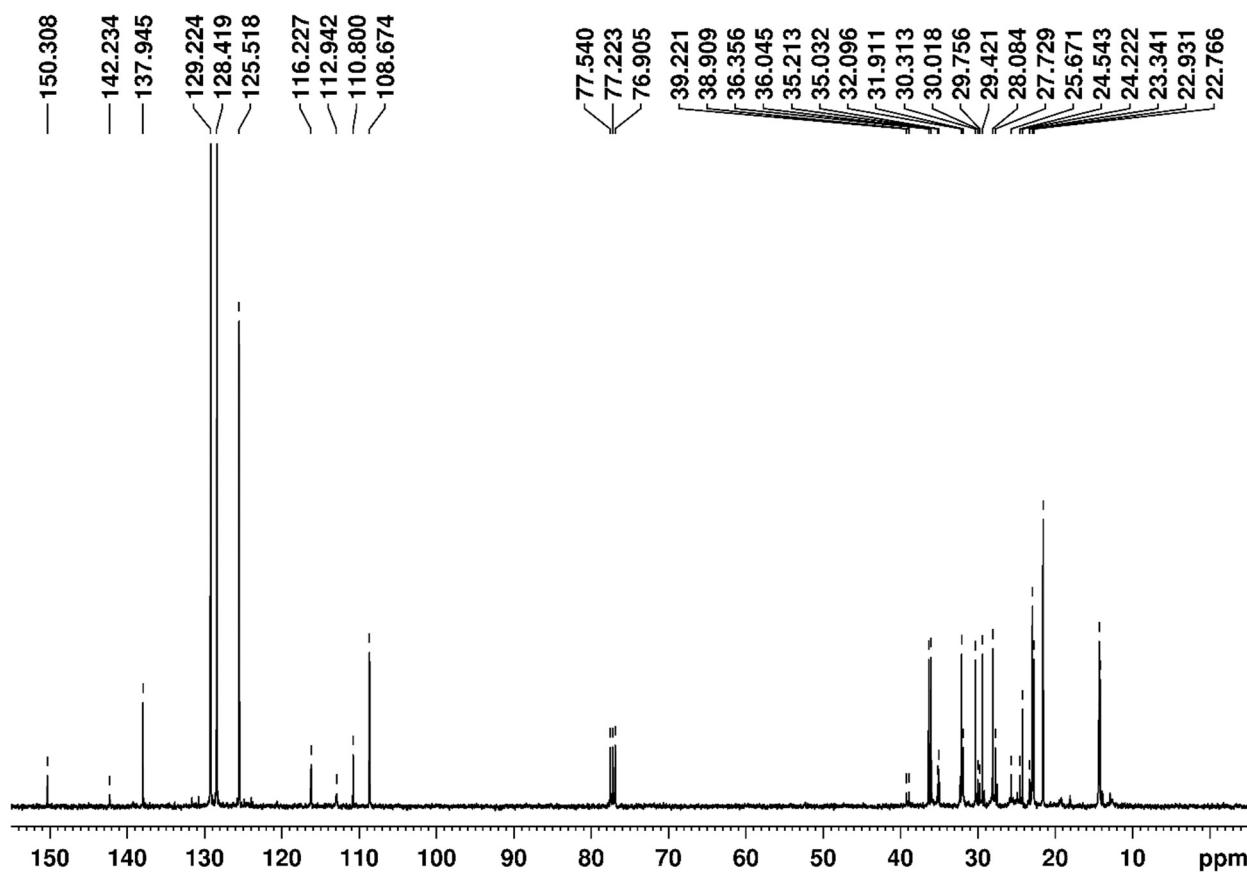


Figure S26. ^1H NMR of tetramers **7** obtained in the system $\text{Cp}_2\text{ZrCl}_2\text{--HAlBu}^{\text{i}2}\text{--(Ph}_3\text{C)[B(C}_6\text{F}_5\text{)}_4\text{]}$, $[\text{Zr}]:[\text{Al}]:[\text{B}]:[1\text{-alkene}] = 4:16:1:1000$, CHCl_3 , 40°C , 16 h.

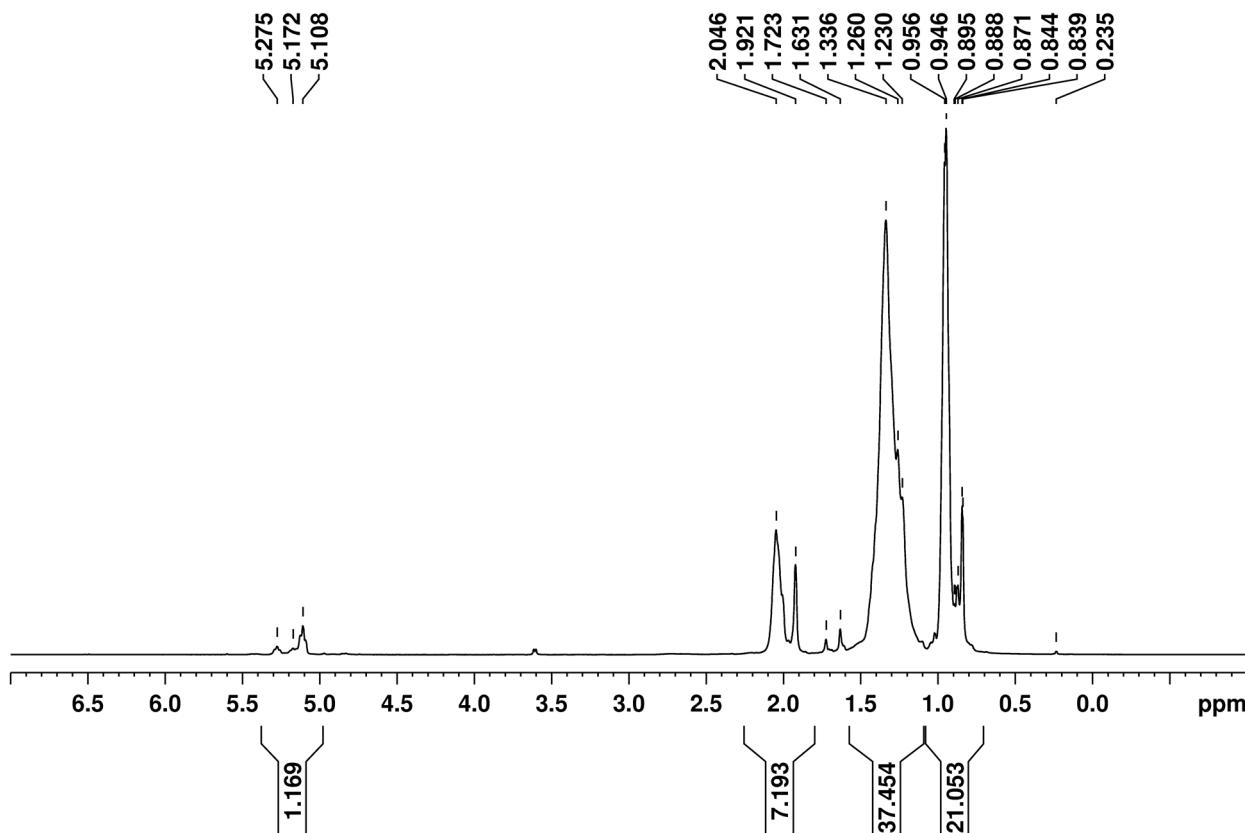
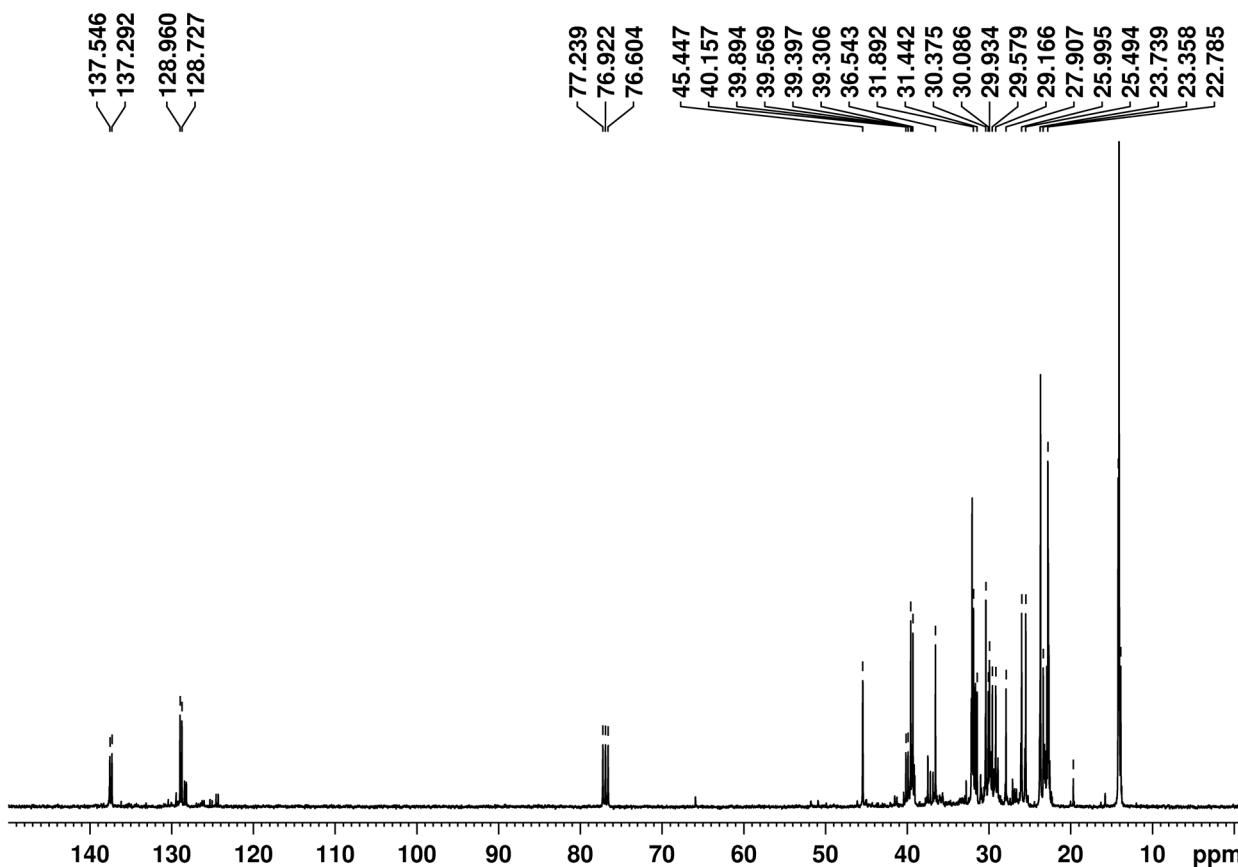


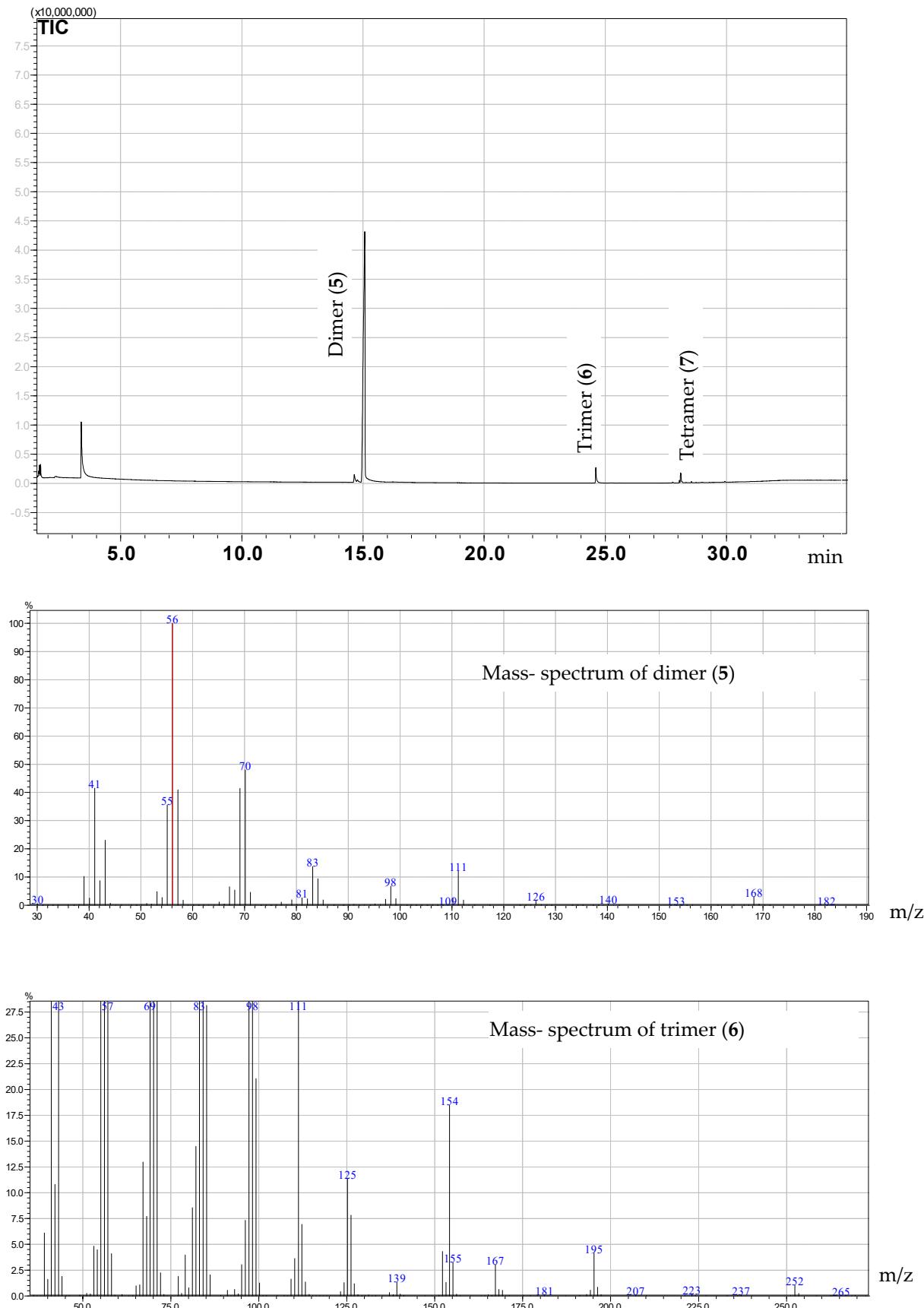
Figure S27. ^{13}C NMR of tetramers **7** obtained in the system $\text{Cp}_2\text{ZrCl}_2\text{--HAlBu}^{\text{i}2}\text{--(Ph}_3\text{C)[B(C}_6\text{F}_5\text{)}_4\text{]}$, $[\text{Zr}]:[\text{Al}]:[\text{B}]:[1\text{-alkene}] = 4:16:1:1000$, CHCl_3 , 40°C , 16 h.



GC-MS analysis of products

Before each series of mass spectral analysis, calibration was performed using 1-hexene – dimer – tetramer mixtures with various mass fractions to determine response factors (RF). Response factors of dimers or tetramers were calculated as $\text{RF}(\text{dimer or tetramer}) = \text{Slope(1-hexene)}/\text{Slope(dimer or tetramer)}$, where Slope(1-hexene) was found from the dependence Peak area (1-hexene) – Concentration (1-hexene), and $\text{Slope(dimer or tetramer)}$ from the dependence Peak area (dimer or tetramer) – Concentration (dimer or tetramer). 1-Hexene was used as a standard with $\text{RF}=1$. RFs were found for dimer and tetramer as 1.1 and 1.2, correspondingly. Thus, product composition was determined via peak areas multiplied by response factors.

Figure S28. Example of GC-MS analysis of dimers and oligomers obtained in the system $[\text{Cp}_2\text{ZrH}_2]_2\text{--CIAIBu}^{\text{i}}_2\text{--MMAO-12}$, $[\text{Zr}]:[\text{Al}]:[\text{MMAO-12}]:[1\text{-alkene}] = 1:3:30:1000$, CH_2Cl_2 , 40°C , 16 h.



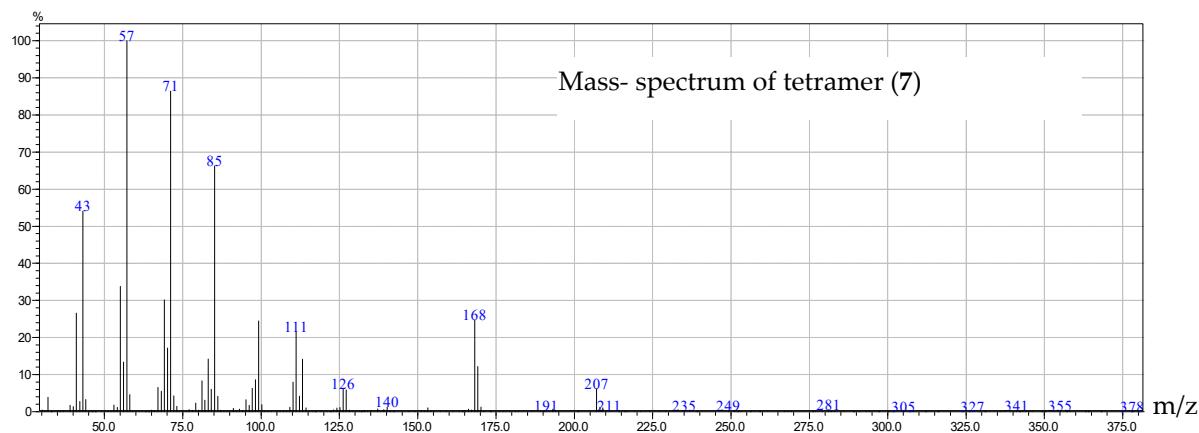
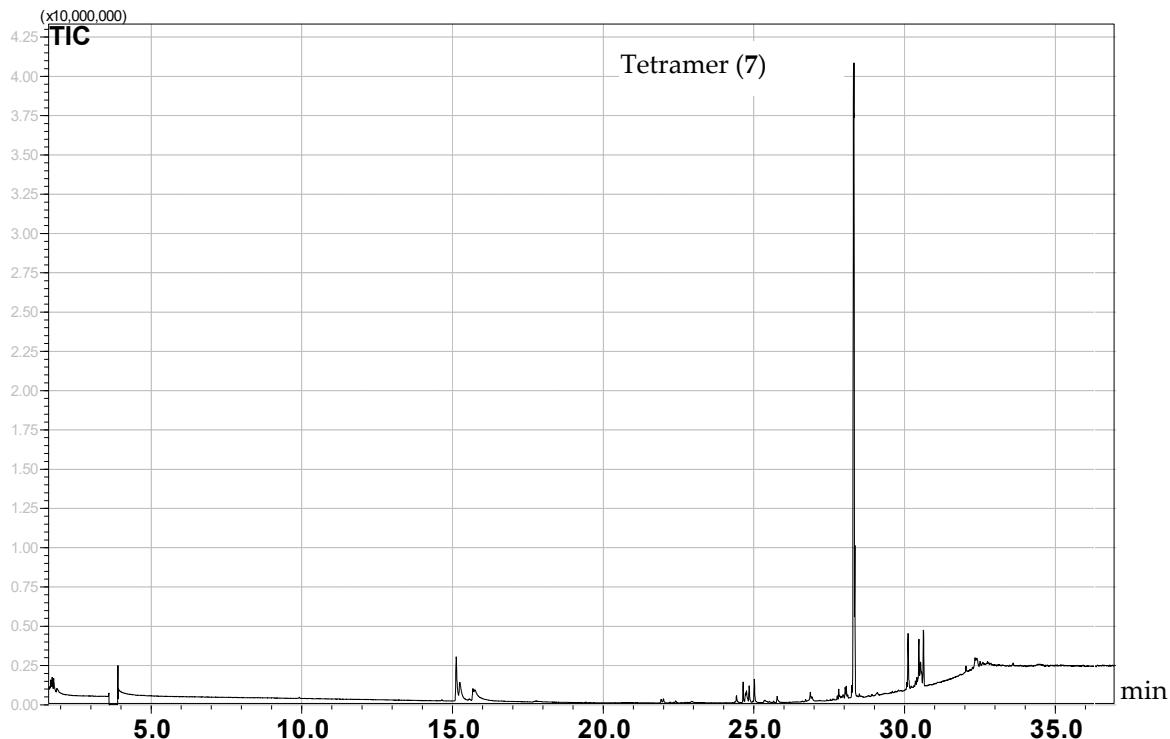


Figure S29. Example of GC-MS analysis of dimers and oligomers obtained in the system Cp_2ZrCl_2 – HAIBu_2 – $(\text{Ph}_3\text{C})[\text{B}(\text{C}_6\text{F}_5)_4]$, $[\text{Zr}]:[\text{Al}]:[\text{B}]:[1\text{-alkene}] = 4:16:1:1000$, CHCl_3 , 40°C , 16 h.



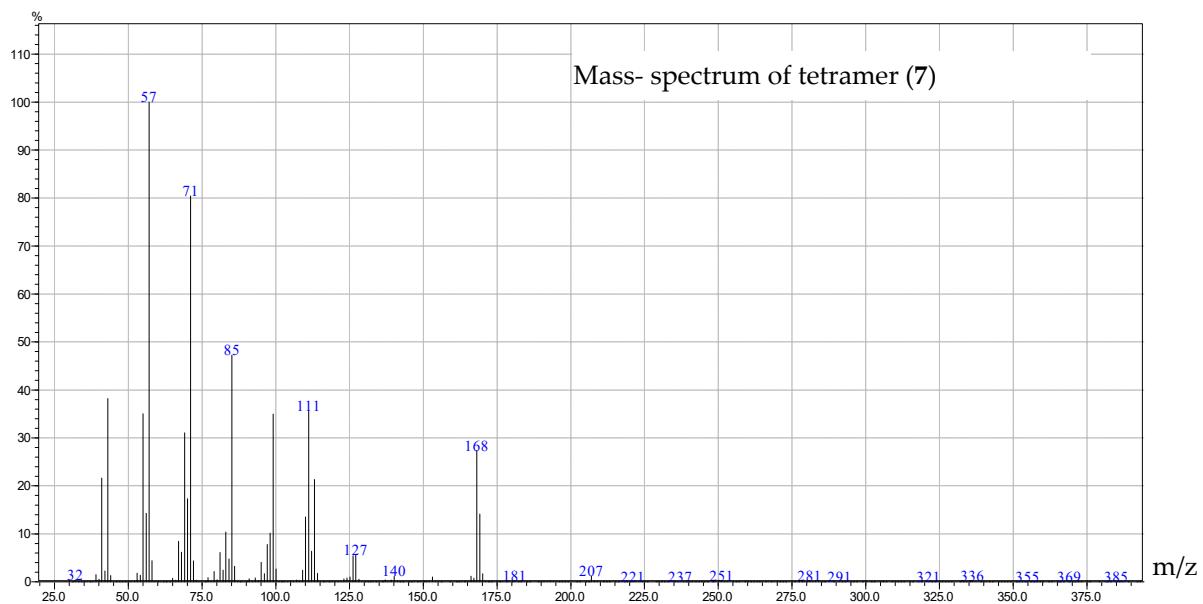


Table S1. The energy at 0K, the ZPVE correction, the enthalpy, the Gibbs free energy in gas, the temperature multiplied by the entropy for complexes **9a-e**, calculated at PBE/3 ζ level (298.15 K).

Complex	E, Hartree	ZPVE, Hartree	H, kcal/mol	G, kcal/mol	T*S, cal/mol
9a	-9177.483313	0.479046	-5758635.6	-5758702.0	66417.6
9b	-9177.470210	0.478892	-5758627.4	-5758692.4	65081.6
9c	-9177.480291	0.479422	-5758633.2	-5758698.7	65459.5
9d	-9177.461167	0.480010	-5758621.2	-5758685.2	63942.4
9e	-10038.215279	0.607638	-6298665.2	-6298747.2	81996.9

Table S2. Relative thermodynamic parameters of isomeric complexes **9a-d**, calculated at PBE0/def2TZVP level (298.15 K).

Complex	ΔE , Hartree	ΔE_{ZPVE} , Hartree	ΔH , kcal/mol	ΔG , kcal/mol	$T\Delta S$, cal/mol
9a	0.001420	0.000052	0.0	0.0	2362.8
9b	0.011620	0.010971	6.8	8.4	796.1
9c	0.000000	0.000000	0.0	1.2	1132.1
9d	0.020137	0.020704	12.7	15.1	0.0

Table S3. Relative thermodynamic parameters of isomeric complexes **9a-d**, calculated at PBE0-GD3/def2TZVP level (298.15 K).

Complex	ΔE , Hartree	ΔE_{ZPVE} , Hartree	ΔH , kcal/mol	ΔG , kcal/mol	$T\Delta S$, cal/mol
9a	0.001174	0.000050	0.0	0.0	941.6
9b	0.010265	0.009999	6.3	6.5	686.6
9c	0.000000	0.000000	0.1	0.2	898.6
9d	0.020161	0.020487	12.8	13.7	0.0

Table S4. Relative thermodynamic parameters isomeric complexes **9a-d** in CHCl₃ solution, calculated at PBE0/def2TZVP level (298.15 K).

Complex	ΔE , Hartree	ΔE_{ZPVE} , Hartree	ΔH , kcal/mol	ΔG , kcal/mol	$T\Delta S$, cal/mol
9a	0.000000	0.000000	0.0	0.0	2322.0
9b	0.004409	0.005690	3.6	4.8	1062.3
9c	0.002469	0.002932	2.1	2.3	2049.2
9d	0.021519	0.023216	14.3	16.6	0.0

Table S5. Relative thermodynamic parameters of isomeric complexes **9a-d** in CH₂Cl₂ solution, calculated at PBE0/def2TZVP level (298.15 K).

Complex	ΔE , Hartree	ΔE_{ZPVE} , Hartree	ΔH , kcal/mol	ΔG , kcal/mol	$T\Delta S$, cal/mol
9a	0.000000	0.000000	0.0	0.0	2337.8
9b	0.002884	0.004539	2.8	4.2	906.4
9c	0.003194	0.003504	2.5	2.6	2158.6

9d	0.022043	0.023679	14.6	17.0	0.0
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Atomic coordinates of the complexes **9a-e**

9a

C	3.579096680	-1.182145120	2.772116080
C	2.692045280	-0.086220410	2.960824760
C	3.172921060	1.011578820	2.194632780
C	4.590397310	-0.764442430	1.865318010
C	4.338653910	0.593690410	1.508025440
H	3.513442520	-2.147786050	3.258561900
H	1.811129990	-0.092078430	3.595230600
H	2.712778870	1.990377220	2.127821760
H	5.432051580	-1.365961360	1.539545320
H	4.950478160	1.204947060	0.851304440
Zr	2.462548190	-0.792445320	0.540541630
C	4.139309800	-1.673368520	-1.095318320
C	3.068528890	-2.607475830	-1.100619470
C	1.930701140	-1.951407000	-1.649648020
C	3.662198400	-0.442176090	-1.635319280
C	2.298908610	-0.626303840	-1.991142190
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H	3.116113600	-3.642189800	-0.782739350
H	0.943935190	-2.385835160	-1.769440610
H	4.257567880	0.450215750	-1.802894530
H	1.631910700	0.114671980	-2.416832500
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C	-0.701726880	4.133904760	1.812510580
C	-1.670094420	2.057116370	1.707607220
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H	-2.360003790	1.250456950	1.483885180
H	-2.717353220	3.852942090	0.874816760
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C	2.214337580	-2.070553090	-1.602420490
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C	2.024020600	4.302268290	-0.279339080
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9e

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