

Article

Biochar-Supported BiO_x for Effective Electrosynthesis of Formic Acid from Carbon Dioxide Reduction

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Abstract: The electrochemical reduction of carbon dioxide (CO₂) to value-added chemicals and fuels has attracted worldwide interest for its potential to address various contemporary global issues such as CO₂-related climate change, the earth's carbon deficit and the energy crisis. In the development of this technology, many efforts have been focused on the design of inexpensive, eco-friendly and effective catalysts. In this work, a bismuth (Bi)-based material was simply synthesized via a scalable method and fully characterized by physical, chemical and electrochemical techniques. The catalyst material consisted of Bi/Bi₂O₃ nanoparticles and a biochar prevented from the pyrolysis of brewed coffee waste. It was observed that the surface of the biochar was thoroughly decorated with nanoparticles. Due to its uniform surface, the biochar–BiO_x electrode demonstrated good selectivity for CO₂ reduction, showing a faradaic efficiency of more than 90% for CO and HCOOH formation in a wide potential range. Particularly, the selectivity for HCOOH reached more than 80% from −0.9 V to −1.3 V vs the reversible hydrogen electrode and peaks at 87%. Besides the selectivity, the production rate of HCOOH also achieved significant values with a maximum of 59.6 mg cm^{−2} h^{−1}, implying a good application potential for biochar–BiO_x material in the conversion of CO₂ to HCOOH.

Keywords: carbon dioxide reduction; electrochemistry; catalyst; bismuth; formic acid; production rate



Citation: Zeng, J.; Jagdale, P.; Lourenço, M.A.O.; Farkhondehfal, M.A.; Sassone, D.; Bartoli, M.; Pirri, C.F. Biochar-Supported BiO_x for Effective Electrosynthesis of Formic Acid from Carbon Dioxide Reduction. *Crystals* **2021**, *11*, 363. <https://doi.org/10.3390/cryst11040363>

Academic Editor: Monica Distaso

Received: 16 March 2021

Accepted: 26 March 2021

Published: 30 March 2021

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1. Introduction

Carbon dioxide (CO₂) is a well-known greenhouse gas, and its accumulation in the atmosphere can trap infrared radiation, thereby disturbing the energy balance on the earth's surface. In the last decade, CO₂ emissions and atmospheric CO₂ concentrations have become the most important topics in politics and scientific debate because unprecedented wildfires and superstorms have made global warming impossible to ignore. Using CO₂ as feedstock to produce value-added chemicals and fuels is considered to be a promising approach for reducing atmospheric CO₂ and consequently mitigating climate change. Many technologies have been exploited for CO₂ valorization, including thermocatalysis, photocatalysis, electrocatalysis and bio-synthesis [1–3]. Among these methods, the electrochemical conversion of CO₂ is closest to commercialization with start-ups and well-established companies such as Carbon Recycling International, Mitsui Chemicals, Dioxide Materials, Carbon Electrocatalytic Recycling Toronto and Opus-12 [4]. Electrochemical conversion presents several advantages: (1) possible use of green electrolytes and electricity without contributing to new CO₂ emissions [5]; (2) Tunable products and production rates by using different catalysts and test conditions [6]; and (3) development

of electrolyzers and processes for CO₂ conversion based on existing technology such as water electrolyzers, polymer electrolyte membrane fuel cells and solid-oxide fuel cells [7–9]. Despite the above-mentioned merits, the electrochemical CO₂ reduction reaction (CO₂RR) involves several proton-assisted, multiple electron-transfer processes with similar standard potentials [10,11]. Moreover, a H₂ evolution reaction (HER) occurs in aqueous solutions. While it is a challenge to control the selectivity of the CO₂RR thermodynamically, it is even more challenging to form chemical bonds for the complex and energetic molecule products kinetically due to the high energy barrier for the intermediate formation [12,13]. Consequently, very high overpotentials are usually needed to drive the CO₂RR. In order to reduce the overpotential, a well-designed catalyst is of vital importance to enable the activation of the CO₂ molecules, the formation of the intermediates and the production of a specific product.

Among many products, formic acid (HCOOH) is an important raw material for industries such as leather, rubber, medicine and fiber. It can also be used as a fuel in direct formic acid fuel cells and as an excellent carrier of H₂ storage [14,15]. According to a techno-economic analysis by Verma et al., HCOOH is the most economically viable and atom-economic target besides carbon monoxide (CO) [16]. Many metals such as Pb, Hg, In, Sn and Cd are selective for HCOOH formation according to the pioneer work of Hori et al. [17]. Among them, Sn-based catalysts have become the most interesting due to the high selectivity and non-noble, eco-friendly and low-cost characteristics [18–22]. In recent years, many studies also explored other types of catalysts for the CO₂RR to HCOOH. Pd appears to be an appealing catalyst for HCOOH formation with high activity and good stability at extremely low overpotentials [23]. Copper sulfide (Cu_xS) is also intensively studied and it can selectively convert CO₂ to HCOOH with good performance [24–27]. Bismuth (Bi)-based materials have also been studied [28–34] and shown to be highly selective and active for HCOOH production. Because of their low cost and low toxicity, Bi-based materials are becoming as important as Sn-based ones and are used in HCOOH production at high rates [35].

In this work, we present a biochar-supported BiO_x and applied it to a CO₂RR for the first time. The biochar was produced by the valorization of brewed coffee waste by pyrolysis [36,37]. This biochar has numerous channels and pores and consequently a high specific surface area, making it an interesting carbon support for bismuth-based nanoparticles. The composite was simply synthesized using a scalable route through a solid-state reaction method. The resultant material was thoroughly characterized by FESEM–EDS and XRD analyses to study its surface and structural properties and by operando TGA–IR to study the physicochemical properties. Finally, the performance of the biochar–BiO_x catalyst for CO₂RR to HCOOH was studied by different electrochemical techniques.

2. Experimental Section

2.1. Materials

Bismuth(III) nitrate pentahydrate (Bi(NO₃)₃·5H₂O, 99%), Potassium bicarbonate (KHCO₃, 99.7%), Nafion[®] 117 solution (5 wt.%) and isopropanol were purchased from Sigma–Aldrich (Merck). All chemicals were used as received.

2.2. Synthesis

The coffee waste was collected from Bar Katia (Turin, Italy) supplied by Vergnano. First, the coffee waste was dried at 105 °C for 72 h to eliminate any moisture. Then, 100 g of the dried waste was pyrolyzed using a tubular furnace (Carbolite TZF 12/65/550) under a nitrogen flow (40 mL min^{−1}). The heating rate was 15° min^{−1} until a temperature of 800 °C was reached. The sample was kept at 800 °C for 30 min and cooled down to room temperature. The obtained biochar was collected.

The Bi(NO₃)₃·5H₂O was used as precursor salt and used biochar as support. In a typical synthesis procedure, 3.5 g of Bi(NO₃)₃·5H₂O salt was initially added slowly to 35 mL (1.5 M) of a nitric acid (HNO₃) solution with continuous stirring (2500 rpm) till the

solution became clear. Then 1.0 g of pyrolyzed biochar was added slowly and continuously into the solution with constant stirring for more than 30 min. To achieve homogeneity in the composite solution, the mixture was sonicated for 20 min (amplitude 40%, Sonic Vibracel sonicator). The composite mixture was then heated in a furnace at 150 °C for 2 h to completely evaporate the water content. The obtained solid composite material was then calcinated at 400 °C in a N₂ atmosphere for 2 h. After calcination, the sample was allowed to cool naturally inside the furnace. The final free flowing catalyst was then ready for the fabrication of the electrodes.

2.3. Physical and Chemical Characterizations

The morphology and elemental composition of the composite powder was investigated by field-emission scanning electron microscopy (FE-SEM, ZEISS-MERLIN) equipped with high-resolution energy-dispersive (ENERGY 450) X-ray spectroscopy (EDS, OXFORD INCA). For FESEM-EDS analysis, the composite sample was coated with a 5 nm thin layer of platinum to avoid the charging effect.

The crystalline phases of the catalysts were examined by X-ray diffraction (XRD). Measurements were performed using a PANalytical X'Pert Pro instrument (Cu-K α radiation, 40 kV and 30 mA) equipped with an X'Celerator detector.

Operando TGA-IR was made on a Thermogravimetric Analyzer (TGA) NETZSCH TG209 F1Libra instrument coupling line to a Fourier transform infrared (FTIR) Bruker Tensor II spectrophotometer equipped with DTGS detector and a Bruker heated gas chamber analysis accessory. Approximately 3 mg of sample in alumina pans were heated from 30 to 900 °C, using a heating ramp of 20 °C min⁻¹, under air (flow rate = 20 mL min⁻¹). Nitrogen flow was used as protection gas (flow rate = 20 mL min⁻¹). The FTIR analysis was collected in the absorbance mode in the 650–4400 cm⁻¹ range. The gas development kinetics were followed at the adsorption maximum for different molecules: CO, 2179 cm⁻¹ [38]; CO₂, 2349 cm⁻¹ [39]; and H₂O, 3735 cm⁻¹ [40].

2.4. Preparation of the Electrodes

Electrodes were made by preparing a slurry containing 3.45 mg of the synthesized catalyst, 30 μ L of a Nafion[®] 117 solution, and 90 μ L of isopropanol through sonication for 30 min. The obtained suspension coated a carbon paper (GDL; SIGRACET 28BC, SGL Technologies) by the drop-casting method. The solvents were evaporated overnight at room temperature. The catalyst mass loading was about 2.3 mg cm⁻².

2.5. Electrochemical Characterizations for the CO₂RR and Product Analysis

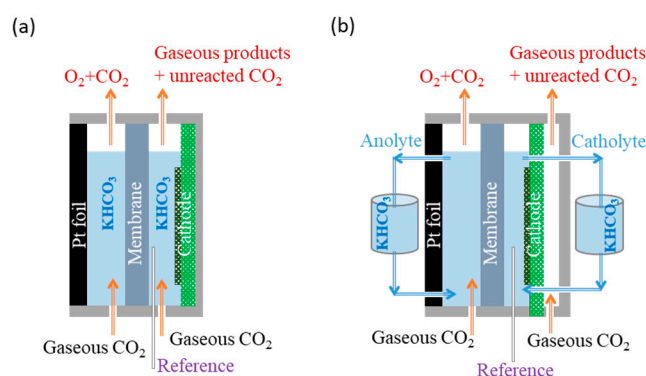
Cyclic voltammetry (CV) was performed in a three-electrode cell at room temperature with a CHI760D electrochemical workstation. The working electrode was a catalyst-coated carbon paper with a geometric area of 0.1 cm². A Pt wire was used as the counter electrode and an Ag/AgCl (3 M NaCl) was used as the reference. To study the electrocatalytic performance of the electrode, CV measurements were performed from +0.2 V to −1.3 V vs. the reversible hydrogen electrode (RHE) at a scan rate of 10 mV s⁻¹ in N₂- and CO₂-saturated 2.0 M KHCO₃ aqueous solution. Hereafter, the potentials refer to the RHE unless otherwise specified.

CO₂ electrolysis was carried out in controlled potential mode by applying chronoamperometric (CA) technique with a potentiostat (CHI760D). The potential was corrected by compensating the ohmic potential drop, of which 85% by the iR-compensation method of the instrument and 15% by manual shifting after calculation.

The CO₂RR performance of the bare biochar and biochar-BiO_x electrodes was compared in a two-compartment, three-electrode batch cell as shown in Scheme 1a. A proton exchange membrane (Nafion[™] Membrane N117, Ion Power) was used as the separator for the anodic and cathodic sides. The reference electrode was a mini Ag/AgCl (1 mm, leak-free LF-1) inserted into the catholyte. A Pt foil (3.3 \times 3.3 cm²) was used as the anode and inserted into the anolyte. The working electrode was carbon paper coated with the

synthesized catalyst (geometric area: 1.5 cm²). The electrolyte was 0.1 M KHCO₃ aqueous solution (pH 6.8). During the CA tests, a constant CO₂ flow of 15 mL min^{−1} was purged through both sides in order to supply CO₂ at the cathode and to maintain a constant pH at the anode.

Then, CA tests were conducted at the biochar–BiO_x electrode in a customized three-compartment, three-electrode semi-flow cell, as shown in Scheme 1b. In the flow cell configuration, the cathodic side was composed of two compartments, one with the catholyte flow and the other with the CO₂ gas flow. During the CA tests, both the catholyte and anolyte were 25 mL of 2.0 M KHCO₃ aqueous solution (pH 8.3) and were circulated at 2 mL min^{−1}. The CO₂ gas was purged through the anolyte at a constant flow rate of 5 mL min^{−1} in order to maintain a constant pH. A CO₂ flow of 25 mL min^{−1} was maintained at the gas compartment of the cathodic side to supply the CO₂ reactant and bring out the products.



Scheme 1. Scheme of the electrochemical cells (a) batch cell and (b) semi-flow cell.

During the CO₂ electrolysis, the gaseous products were analyzed on-line with a micro gas chromatograph (μGC, Fusion[®], INFICON). The μGC was equipped with two channels: one, a Rt-Molsieve 5A column (10 m) and the other, a Rt-Q-Bond column (8 m). A microthermal conductivity detector (micro-TCD) was used for both channels. At the end of each test, the catholyte was collected and analyzed by a high-performance liquid chromatograph (Thermo Scientific Ultimate3000 HPLC) to determine the liquid products. The HPLC was equipped with a UV–Vis detector (210 nm) and a ReproGel (300 × 8 mm²) column. A solution of 9.0 mM H₂SO₄ at a flow rate of 1.0 mL min^{−1} was used as the mobile phase.

Faradaic efficiency (FE) was the key parameter for determining the selectivity toward a specific product. The FE for a product can be calculated by dividing the charge consumed to produce a determined amount of this product by the total charge passed through the cathode as in Equation (1),

$$FE = \frac{nNF}{Q} \quad (1)$$

where Q is the total charge consumed at the cathode (coulombs, C); n is the mole of electrons required to obtain 1 mole of this product ($n = 2$ for CO, HCOOH and H₂ production); N is the mole of an identified product (mol); and F is the Faraday constant (96,485 C mol^{−1}).

The partial current density for HCOOH formation j_{HCOOH} can be obtained using Equation (2),

$$j_{\text{HCOOH}} = j_{\text{total}} * FE_{\text{HCOOH}} \quad (2)$$

where j_{total} is the total geometric current density at the cathode, and FE_{HCOOH} is the faradaic efficiency for HCOOH production.

The production rate of HCOOH (mg h^{−1} cm^{−2}) can be calculated by applying Equation (3),

$$\text{Production rate}_{\text{HCOOH}} = \frac{j_{\text{HCOOH}} * t * M}{2F} \quad (3)$$

where F is the Faraday constant ($96,485 \text{ C mol}^{-1}$); t is the seconds in one hour (3600); j_{HCOOH} is the partial current density for HCOOH formation (mA cm^{-2}); and M is the molar mass of HCOOH (g mol^{-1}).

3. Results and Discussion

3.1. Physical and Chemical Characterizations of the Synthesized Catalysts

FESEM–EDS analysis was performed on the biochar– BiO_x composite sample. Figure 1a shows the porous nature of the biochar with a pore size distribution of 100 to 1000 nm. Biochar wall thickness and length ranged from few nanometers to a micrometer. This microporous structure of biochar provided plenty of surface on which to decorate BiO_x nanoparticles (Figure 1b). The uniform layer of the granular-sized BiO_x nanoparticles on the biochar surface was observed. It is likely that all available biochar surface was completely covered by the BiO_x nanoparticles layer (Figure 1c). It confirmed that the goal of achieving uniform decoration over a biochar surface was satisfactory. The diameter of granular BiO_x nanoparticles was in nano-scale varying from 10 to 50 nm (Figure 1d), which gave them the added advantage of providing a highly catalytic surface.

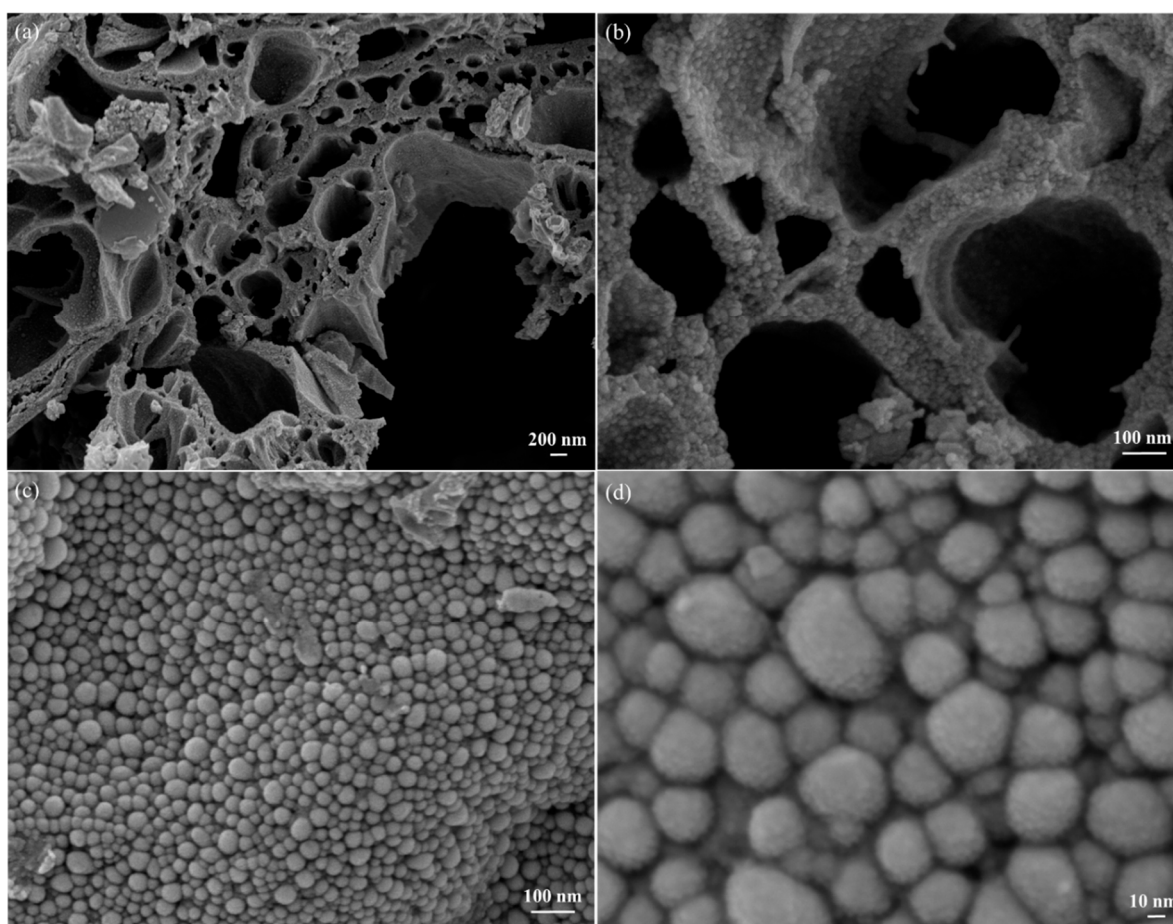


Figure 1. FESEM images of the biochar– BiO_x sample: (a) at low magnification, (b) pore channels at high magnification, (c) pore walls at high magnification and (d) surface at high magnification.

A quantitative elemental analysis of biochar– BiO_x composite was performed by EDS (Figure 2). Elemental mapping and scanning was performed on the targeted surface of the composite (Figure 2a). The mapping of the surface confirmed the presence of three elements, bismuth (Bi) in green, carbon (C) in red, and oxygen (O) in blue (Figure 2b–e). The mapping also confirmed the uniform distribution of the BiO_x layer over the carbon surface. To quantify the elemental compositions, EDS surface scanning was performed

(Figure 2f). Bismuth showed the maximum amount of 65.6 wt.% in the composite, followed by oxygen (8.9 wt.%) and carbon (25.4 wt.%).

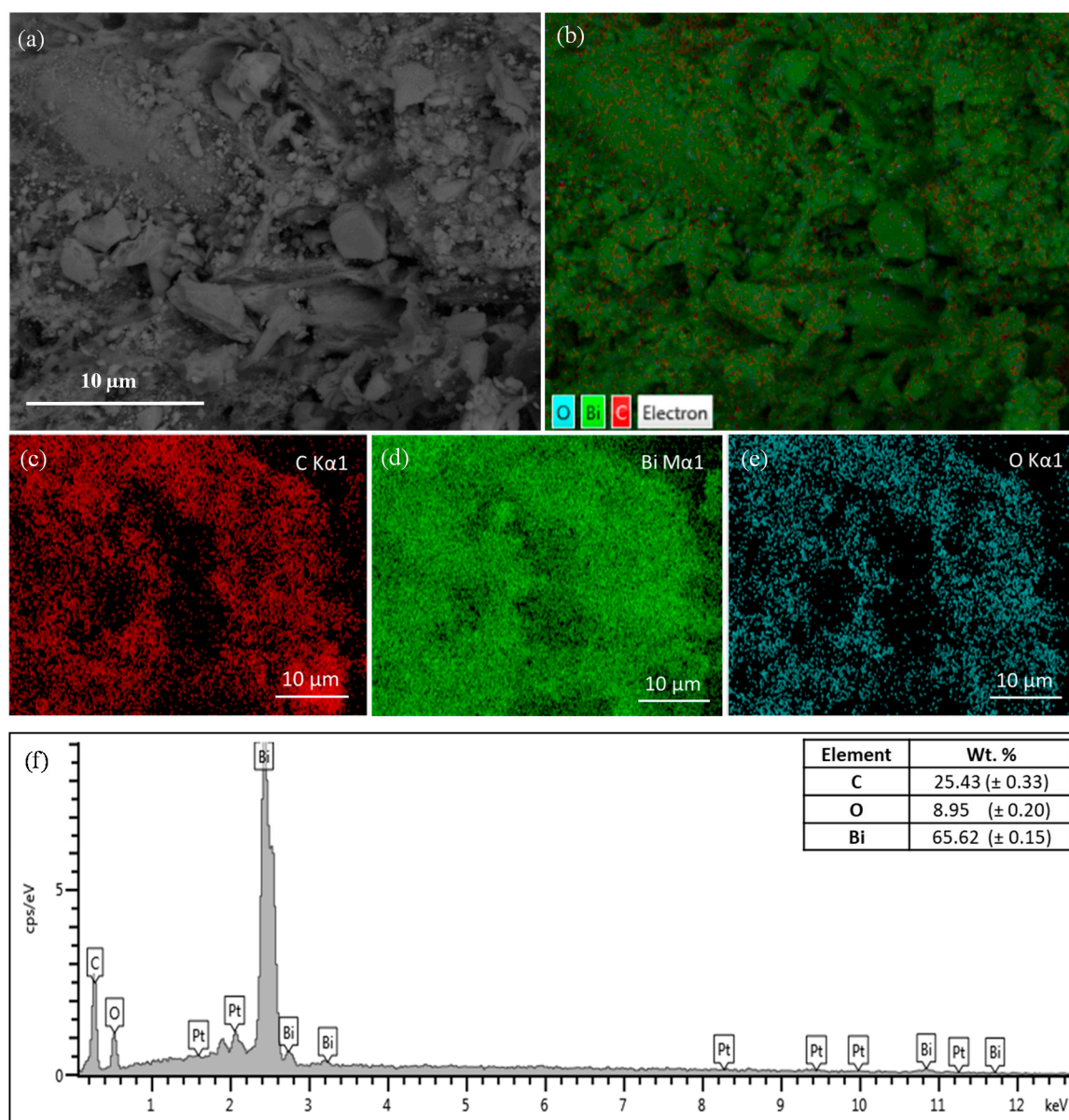


Figure 2. Biochar-BiO_x composite elemental analysis by EDS (a) targeted surface, (b–e) mapping and (f) scanning.

XRD analysis was performed on the biochar-BiO_x sample. As shown in Figure 3, metallic Bi and Bi₂O₃ were the main crystalline phases. The peaks at $2\theta = 22.6^\circ, 27.3^\circ, 38.1^\circ, 39.7^\circ, 44.8^\circ, 46.1^\circ, 48.9^\circ, 56.2^\circ, 59.6^\circ, 62.5^\circ$ and 64.7° corresponded to the reflections related to the (003), (012), (104), (110), (115), (113)/(006), (202), (024), (107), (116) and (122) planes of crystalline Bi (JCPDS 01-085-1331), while the peaks at $2\theta = 28.0^\circ, 31.7^\circ, 32.7^\circ, 46.2^\circ, 46.9^\circ, 54.2^\circ, 55.6^\circ$ and 57.8° were associated with the (201), (002), (220), (222), (400), (203), (421)/(213) and (402) planes of the crystalline Bi₂O₃ (JCPDS 01-078-1793). The broad peak centered at about $2\theta = 28.5^\circ$ was related to the graphitic carbon in the biochar.

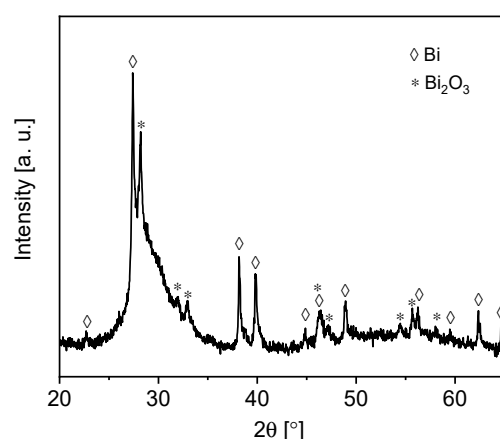


Figure 3. XRD pattern of biochar-BiO_x.

Figure 4 displays the thermal degradation in air of the biochar-BiO_x up to 900 °C. The TGA (Figure 4a) shows a slight weight loss of 0.3% below 100 °C, which was related to the desorption of some physisorbed gases. Then, a weight increase of ~2.5% associated with the oxidation of the bismuth was detected until 315 °C. This was followed by thermal degradation of the biochar-BiO_x sample, which occurred completely below 400 °C. The maximum degradation of the biochar-BiO_x sample was observed at 349 °C, and the IR spectrum acquired at this temperature is shown in Figure 4b. The gases developed were mainly CO₂ (2359 cm⁻¹) and water (3750 and 1543 cm⁻¹), being obtained by the degradation of the biochar and the release of some OH surface groups from BiO_x, respectively. CO (2180 cm⁻¹) was also observed but in a small amount (Figure 4b inset). The residue of the biochar-BiO_x degradation at 900 °C was 89.6%, which was related to the Bi₂O₃ derived from BiO_x in air at high temperature [41,42] and the residue of biochar. Figure 4c shows the TGA of the bare biochar and 12.6% of residue was found. Table 1 listed the calculated weight percentage of BiO_x in the composite sample and compared it with the nominal one. The calculated value was higher than the nominal percentage, probably due to the loss of a part of the biochar during the decoration of BiO_x.

Table 1. Weight loss percentages for biochar and biochar-BiO_x samples.

Materials	Weight Loss (%) ^a	BiO _x (wt.%) ^b	Biochar (wt.%) ^c	BiO _x (wt.%) ^d
biochar	87.4	-	100	-
biochar-BiO _x	10.4	85.3	14.7	60.1–67%

^a Weight loss (wt._{loss}) determined from TGA. ^b BiO_x weight percentage and ^c biochar weight percentage in the biochar-BiO_x sample determined from the equation: (a) wt._{loss} BiO_x × wt.% BiO_x + 2.5% × (wt.% BiO_x + wt.% biochar) + wt._{loss} biochar × wt.% biochar = wt._{loss} (biochar-BiO_x) × (wt.% BiO_x + wt.% biochar) and (b) wt.% BiO_x + wt.% biochar = 100%. The weight loss of BiO_x was considered 0% and the 2.5% × (wt.% BiO_x + wt.% biochar) was due to the weight increase attributed to BiO_x oxidation. ^d BiO_x weight percentage calculated from the precursors. If x = 0 in BiO_x and the Bi specie is metallic, the percentage of BiO_x is 60.1%; if x = 1.5 in BiO_x and the Bi specie is Bi₂O₃, the percentage of BiO_x is 67%.

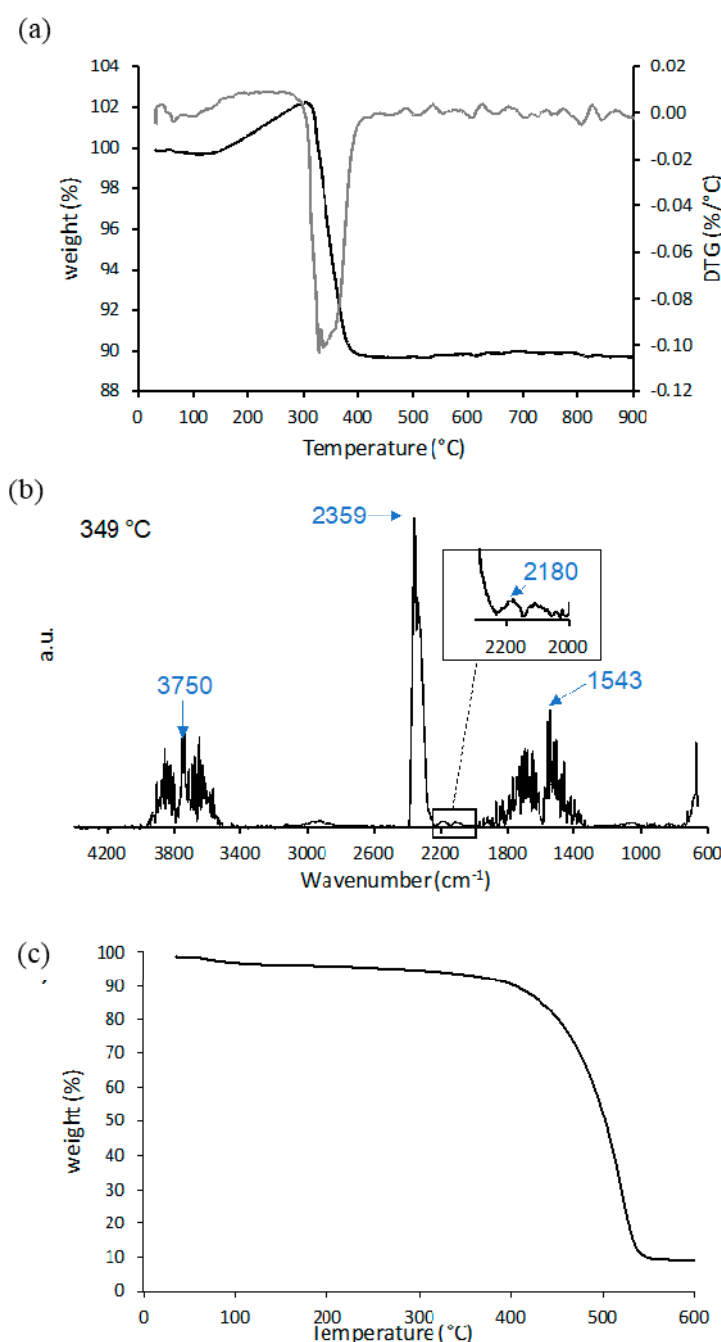


Figure 4. TGA-IR data of biochar-BiO_x. (a) TGA and first derivative, (b) IR spectra of developed gases at 349 °C and (c) TGA of bare biochar.

3.2. Electrochemical Properties of the Biochar-BiO_x Electrode toward the CO₂RR

In order to study the electrocatalytic performance of the biochar-BiO_x electrode toward the CO₂RR, CV was carried out in a N₂-purged or CO₂-saturated 2.0 M KHCO₃ aqueous solution. N₂ or CO₂ gas was bubbled through the electrolyte for 30 min before the measurements and the same gas flux was maintained during the measurements. Only the HER occurred in the N₂ atmosphere at negative potentials, while both the HER and CO₂RR happened in the CO₂ atmosphere. Hence, it is widely believed that the comparison of the geometric current densities in N₂ and CO₂ electrolytes could be suggestive for electrode selectivity between the HER and CO₂RR [11,43,44]. An electrode that generates a higher geometric current density in the CO₂ electrolyte than in the N₂ one is usually more selective for the CO₂RR with respect to the HER at the corresponding potential. As shown

in Figure 5, the current density in the CO₂ electrolyte was significantly higher than that in the N₂ one at potentials more negative than −0.7 V. This outcome indicated that the biochar–BiO_x electrode could have high activity and good selectivity for the CO₂RR at negative potentials (<−0.7 V).

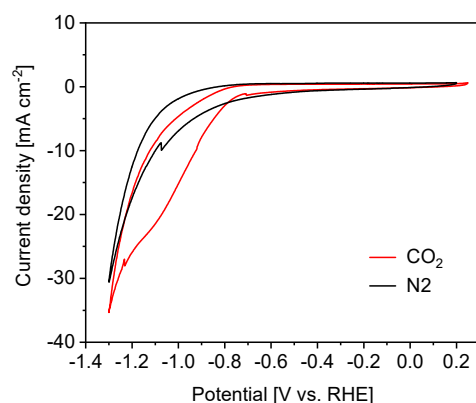


Figure 5. Cyclic voltammograms of the biochar–Bi₂O₃ electrode in CO₂- and N₂-saturated 2.0 M KHCO₃ electrolyte.

To verify the performance of the biochar–Bi₂O₃ electrode, CA tests were firstly carried out in 0.1 M KHCO₃ electrolyte at different potentials in a customized three-electrode, two-compartment batch cell. An electrode composed of bare biochar was also tested for comparison. As shown in Figure 6a, the biochar showed no selectivity toward the CO₂RR, since the faradaic efficiency for the H₂ formation was almost 100%. In contrast, the biochar–Bi₂O₃ electrode showed significant selectivity for the CO and HCOOH formation, especially at higher overpotentials (Figure 6b). However, the current density for HCOOH formation was rather low, probably due to the high resistance of the diluted buffer electrolyte.

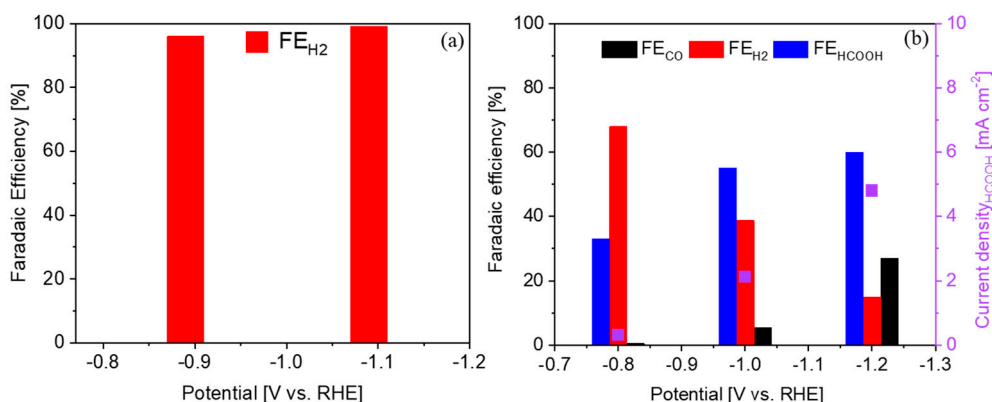


Figure 6. CO₂RR on the electrodes in 0.1 M KHCO₃ electrolyte in a batch cell: (a) faradaic efficiency of the main product on the bare biochar electrode and (b) faradaic efficiency of the main products and partial current density for HCOOH formation on the biochar–Bi₂O₃ electrode.

Hence, CA tests were further performed on the biochar–Bi₂O₃ electrode in a customized flow cell at potentials ranging from −0.6 V to −1.3 V in more concentrated electrolyte (2.0 M KHCO₃). The *i*-*t* curves at different potentials are shown in Figure 7a. The current density increased as the overpotential increased, which is in agreement with the CV result in the same electrolyte. The major products and their FE values are shown in Figure 7b. H₂ was the main product at −0.6 V and its production dramatically reduced at −0.7 V. As the electrode was polarized at more negative potentials, H₂ became a minor product and HCOOH became the predominant product, while the FE_{CO} remained less than 6% in all the investigated potential range. The selectivity for the CO₂RR significantly

outcompeted that for the HER at potentials more negatively than -0.7 V, which is consistent with the CV result. It is worth noting that the FE for CO and HCOOH production occupied more than 90% from -1.0 V to -1.3 V. As the main product, HCOOH reached good FE values ($>80\%$) in a wide potential range from -0.9 V to -1.3 V, with a maximum of 87% at -1.1 V.

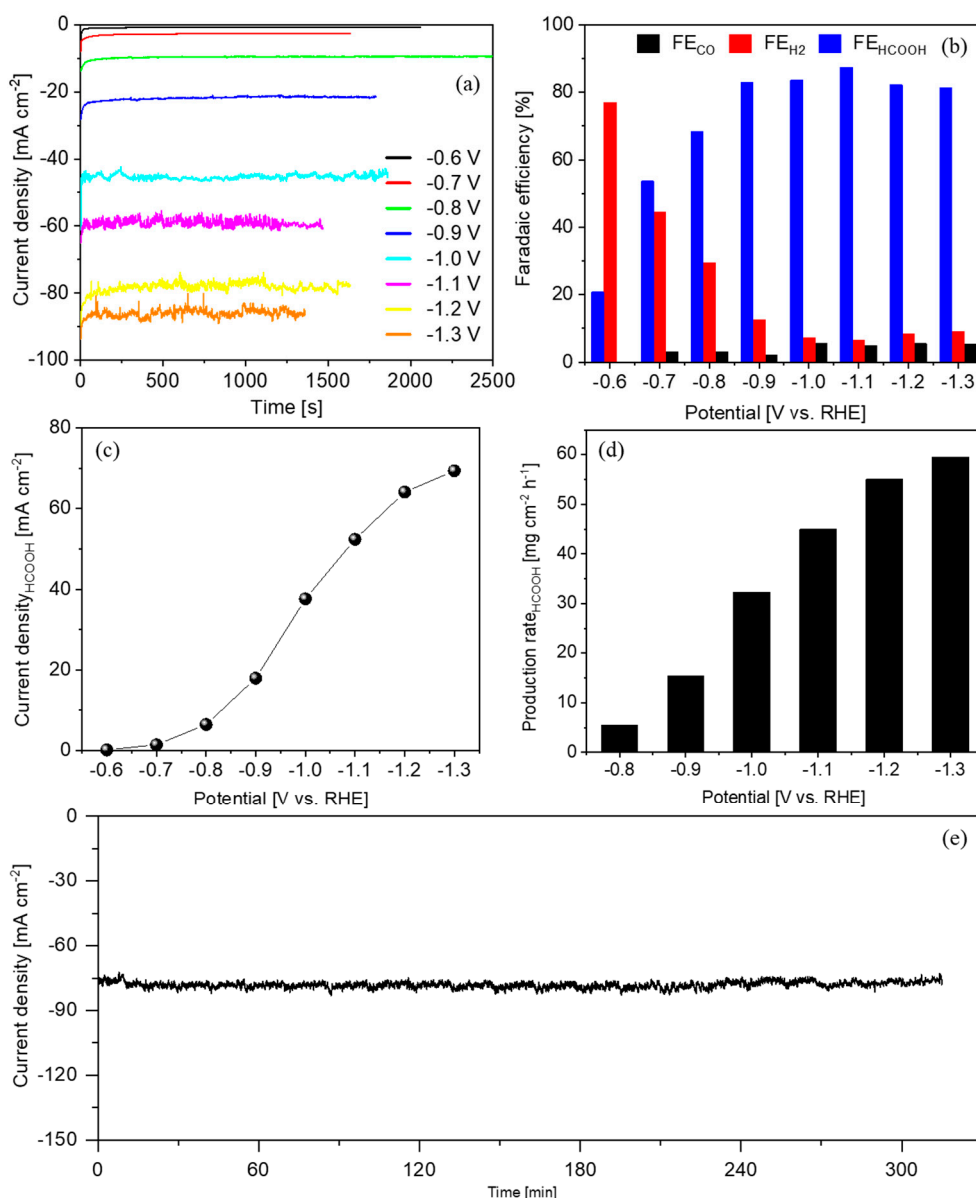


Figure 7. CO₂RR on the biochar-Bi₂O₃ electrode in 2.0 M KHCO₃ electrolyte: (a) chronoamperometric tests at various potentials, (b) faradaic efficiency of the main products, (c) partial current density for HCOOH formation, (d) the production rate of HCOOH and (e) long-term test at -1.2 V.

The partial current density for HCOOH formation as a function of potential is displayed in Figure 7c. The HCOOH production rate became significant at -0.8 V, which was $5.5 \text{ mg cm}^{-2} \text{ h}^{-1}$. It increased progressively with enhancing the driving force and ended with $59.6 \text{ mg cm}^{-2} \text{ h}^{-1}$ at -1.3 V (Figure 7d). The biochar-BiOx electrode also showed good stability during a 5-h test at a current density of 77.5 mA cm^{-2} (Figure 7e). Concisely, the biochar-Bi₂O₃ electrode showed good performance for the CO₂RR to HCOOH for both selectivity and production rate, which was in line with the best results reported on the HCOOH-selective materials as listed in Table 2.

Table 2. Comparison of different HCOOH-selective electrocatalysts in liquid-phase CO₂ electrolysis.

Electrocatalyst	Electrolyte	Potential (V vs. RHE)	Current Density (-mA/cm ²)	FE (%)	Reference
Biochar–BiO _x	2.0 M KHCO ₃	−0.90	18.0	83	This work
Biochar–BiO _x	2.0 M KHCO ₃	−1.10	52.4	87	This work
Bi Dendrite	0.5 M KHCO ₃	−0.74	2.7	89	[31]
Bi Dendrite	0.5 M KHCO ₃	−1.14	41.0	38	[31]
Bi ₂ O ₃ Nanoparticles	0.5 M NaHCO ₃	−1.00	-	82	[32]
Bi ₂ O ₃ Nanoparticles	0.5 M NaHCO ₃	−1.20	20.0	91	[32]
Bi ₂ O ₃ Nanosheets	0.1 M KHCO ₃	−0.86	1.9	65	[33]
Bi ₂ O ₃ Nanosheets	0.1 M KHCO ₃	−1.06	7.8	97	[33]
Bi ₂ O ₃	0.5 M KHCO ₃	−0.90	8.0	91	[34]
Bi ₂ O ₃	0.5 M KHCO ₃	−1.10	-	72	[34]
nano-SnO ₂ /C	0.1 M NaHCO ₃	−1.16	6.2	86	[19]
SnO ₂	0.1 M KHCO ₃	−1.06	11.0	82	[20]
Cu ₂ S	0.1 M KHCO ₃	−0.90	19.0	87	[24]
CuS _x	0.1 M KHCO ₃	−0.85	11.9	75	[25]
Cu _{1.81} S	0.5 M KHCO ₃	−0.67	3.8	82	[26]
S-modified Cu	0.1 M KHCO ₃	−0.80	20.0	80	[27]
SnS ₂ /RGO	0.5 M NaHCO ₃	−1.40 (vs. Ag/AgCl)	13.9	84	[21]
Bi ₂ S ₃ –Bi ₂ O ₃ @rGO	0.1 M KHCO ₃	−0.90	3.8	90	[30]
Sn _{56.3} Pb _{43.7}	0.5 M KHCO ₃	−2.00 (vs. Ag/AgCl)	45.7	80	[22]
AgSn/SnO _x	0.5 M NaHCO ₃	−0.80	16.0	80	[45]
Cu–Au	0.5M KHCO ₃	−0.60	10.2	81	[46]

4. Conclusions

A new material composed of BiO_x and biochar was synthesized through an easy and scalable route. The Bi species were metallic Bi and Bi₂O₃, which comprised about 87.1 wt.% of the composite. The BiO_x particles were on the nanoscale and uniformly decorated on the surface of the biochar. By careful electrochemical characterizations and product analyses, the biochar–BiO_x material is demonstrated to be a promising candidate for converting CO₂ to HCOOH, showing a maximum faradaic efficiency of 87% and the highest current density of 70 mA cm^{−2} for HCOOH production. In addition to its low price and eco-friendly nature, this proposed BiO_x-biochar catalyst could be produced on a mass scale via the simple route and be used in the electrosynthesis of HCOOH from CO₂ on an industrial scale.

Author Contributions: J.Z., P.J. and C.F.P. contributed to the conceptualization. C.F.P. contributed to the resources. P.J. and M.B. contributed to the synthesis of the material. M.A.O.L. contributed to the TGA–IR analysis and the fabrication of electrodes. J.Z. contributed to the XRD measurement and the CO₂ electrolysis experiments. M.A.F. contributed to the HPLC analysis. D.S. contributed the cyclic voltammetry experiment. All authors contributed to the manuscript composition. All authors have read and agreed to the published version of the manuscript.

Funding: This research did not receive any specific grant from funding agencies in the public, commercial or not-for-profit sectors.

Institutional Review Board Statement: Not applicable.

Informed Consent Statement: Not applicable.

Data Availability Statement: Data is contained within the article.

Acknowledgments: Mauro Raimondo is acknowledged for FESEM–EDX analyses.

Conflicts of Interest: The authors declare no conflict of interest.

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