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Method of Preparation for High-Purity Nanocrystalline Anhydrous Cesium Perrhenate

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Abstract: This paper is devoted to the preparation of high-purity anhydrous nanocrystalline cesium perrhenate, which is applied in catalyst preparation. It was found that anhydrous cesium perrhenate with a crystal size <45 nm can be obtained using cesium ion sorption and elution using aqueous solutions of perrhenic acid with subsequent crystallisation, purification, and drying. The following composition of the as-obtained product was reported: 34.7% Cs; 48.6% Re and <2 ppm Bi; <3 ppm Zn; <2 ppm As; <10 ppm Ni; < 3 ppm Mg; <5 ppm Cu; <5 ppm Mo; <5 ppm Pb; <10 ppm K; <2 ppm Na; <5 ppm Ca; <3 ppm Fe.

Keywords: cesium; caesium; rhenium; perrhenate; cation resin

1. Introduction

Both cesium and rhenium are classified as rare metals with naturally occurring deposits scattered around the world [1,2]. Molybdenites from porphyry copper deposits are the most important rhenium sources, while cesium can be mainly found in pollucite ($\text{CsAlSi}_2\text{O}_6$) [1–3]. Rhenium can be characterised by a high melting temperature (3453 K) and high density (21.0 g/cm^3). It can be used in organic and inorganic chemistry, catalysis, and pharmacy [1,3]. Recently, it has been used as an alloy and superalloy additive, while ^{186}Re and ^{188}Re isotopes have been found in radiotherapy [2]. It is catalytically active in hydrogenation and dehydrogenation reactions. It can be used in either heterogeneous or homogenous catalysis of numerous other organic reactions [4,5]. Platinum–rhenium catalysts are used in the petrochemical industry in high-octane petrol production [1]. On the contrary, cesium is soft and one of the most reactive metals, and due to its strong photoelectric effect, is mainly used in photocell production. It may also be found in infrared radiation sensors, organic synthesis, and hydrogenation reaction catalysts. The ^{137}Cs isotope is used in cancer treatment.

Synthesis, properties, and applications of cesium perrhenate are poorly described. In the past, CsReO_4 was mainly synthesised to define its physicochemical properties [6]. According to these sources, cesium perrhenate can be obtained by neutralisation of perrhenic acid aqueous solution in the presence of cesium carbonate or hydroxide [7–11], as well as by reaction of sodium perrhenate aqueous solution with cesium chloride [12,13] and by reaction of cesium hydroxide with ammonium perrhenate [14,15]. It was also reported that cesium perrhenate can be used in the synthesis of heterogeneous rhenium–cesium catalyst [16,17]. The ion-exchange technique is applied, for example, to nickel(II) and cobalt(II) perrhenate production [18]. There are also reports concerning cesium recovery from nuclear wastewater using different sorbents [19–24].

This paper presents the synthesis of high-purity anhydrous nanocrystalline cesium perrhenate, which can be used as a substrate in the preparation of organic and inorganic rhenium compounds,

including rhenium catalysts (e.g., methyltrioxorhenium (MTO), rhenium carbonyl ($\text{Re}_2(\text{CO})_{10}$) and its derivatives), as well as heterogeneous catalysts containing rhenium and cesium additives (e.g., Al_2O_3 - or SiO_2 -based). It should be noted that, currently, at the large-laboratory scale, MTO is manufactured from silver perrhenate, while $\text{Re}_2(\text{CO})_{10}$ is manufactured from ammonium perrhenate [25,26]. Cesium perrhenate has similar properties to silver and ammonium perrhenates and, therefore, it may be applied in MTO and $\text{Re}_2(\text{CO})_{10}$ synthesis. The important benefits of CsReO_4 application may be: reaction yield increase, reaction time shortening, and selection of more environmentally friendly solvents.

There are also reports on heterogenous rhenium catalysts doped with cesium. Currently, these catalysts are prepared using aqueous ammonium perrhenate solution as the rhenium source and aqueous cesium nitrate solution as a source of cesium [27]. Replacing these solutions by a single aqueous cesium perrhenate solution may result in: uniform migration of the solution in catalyst support, easier catalyst preparation, and no need to eliminate unnecessary components, like nitrate and ammonium ions.

Currently, only the nuclear industry uses cesium perrhenate on a large scale [28].

2. Experimental Section

2.1. Materials

High-purity perrhenic acid, prepared in the Institute of Non-Ferrous Metals (IMN) in Gliwice using the ion-exchange method, was used as a rhenium source [29–31]. Its chemical composition was as follows: 100.0 g/dm³ Re and <20 ppm ammonium ions and <2 ppm As; <2 ppm Bi, <3 ppm Zn, <3 ppm Mg, <3 ppm Cu, <5 ppm Mo, <5 ppm Ni, <5 ppm Pb, <10 ppm K, <3 ppm Ca, <3 ppm Fe, <3 ppm Al, <5 ppm Co [10–12]. Cesium nitrate(V) from Alfa Aesar (Chemat, Gdańsk, Poland) was used as a cesium source. Strongly acidic hydrogenated cation-exchange resins (C160, PFC100, CT169 from Purolite (Gdynia, Poland) and S100, SP112 from Bayer Chemicals (ChemTech, Warsaw, Poland) were examined. Double-distilled water with a specific conductivity of <2 $\mu\text{S}/\text{cm}$, 30% hydrogen peroxide solution (Merck, Warsaw, Poland), and acetone (Sigma Aldrich, Poznań, Poland) were used.

2.2. Cesium Perrhenate Preparation—Ion Exchange

A new method of anhydrous nanocrystalline cesium perrhenate synthesis, an alternative to the classical one, was developed. It is composed of two stages: sorption of Cs^+ ions and their elution by aqueous solution of perrhenic acid:



Tests were performed in static and dynamic conditions. Preliminary experiments at static conditions allowed the choice of the best-fitted cationite for Cs^+ sorption. Then, in further static tests, the effect of selected parameters (temperature, contact time of a solution with ion-exchange resin, and cesium concentration in an initial solution) on Cs^+ sorption efficiency of resins was determined. Ion-exchange selection was carried out using the following procedure: 10.0 g of ion-exchange resin was mixed for 30 min at ambient temperature with 100 cm³ of cesium nitrate(V) aqueous solution containing 5.0 g/dm³ Cs and followed by vacuum filtration. Solutions were analysed to estimate cesium content. Cesium sorption efficiency was calculated using the following formula:

$$W_{\text{Cs}} = \frac{m_{0\text{Cs}} - C_{\text{Cs}}V}{m_{0\text{Cs}}} \times 100\% \quad (1)$$

where: $m_{0\text{Cs}}$ is the cesium mass in a solution before sorption (g); C_{Cs} is the cesium concentration in a solution after sorption (g/dm³); and V is the volume of the solution after sorption (dm³).

The effect of temperature, contact time, and cesium concentration of an initial solution on cesium sorption efficiency was conducted in the same conditions as presented for ion-exchange selection. The examined temperature was in the range 293–333 K, with a contact time of 30–120 min, and a cesium concentration of 0.5–20.0 g/dm³. The effect of contact time and cesium concentration on sorption efficiency was examined at ambient temperature, while the effect of cesium concentration was examined for a contact time of 60 min.

In the next stage, the influence of contact time, temperature, and rhenium–cesium ratio on Cs elution was determined. Ion-exchange resins after the sorption stage contained: C160: 4.6% Cs; PFC100: 4.5% Cs; CT169: 4.7% Cs. The procedure was as follows: 5.0 g of ion-exchange resin containing sorbed cesium was magnetically mixed with 10 cm³ perrhenic acid of concentration 100.0 g/dm³ Re. Solutions were mixed for 30 min, and vacuum filtered. Then, cesium content was analysed. The following parameter ranges were examined: contact time 30–180 min, temperature 293–333 K, and rhenium-to-cesium ratio 5:1–30:1. Contact time of 60 min was used to examine the effect of temperature and Re:Cs ratio. Elution efficiency was calculated using the formula:

$$W'_{Cs} = \frac{C'_{Cs}V'}{m'_{0Cs} - C'_{Cs}V'} \times 100\% \quad (2)$$

where: m'_{0Cs} is the cesium mass in a solution before elution (g); C'_{Cs} is the cesium concentration in a solution after elution (g/dm³); and V' is the volume of solution after elution (dm³).

Dynamic tests were carried out according to the following procedure: 100 cm³ CT169 or C160 ion-exchange resins were placed in a column 2.5 cm in diameter. Then, a cesium nitrate(V) solution containing 5.0 g/dm³ Cs was added until the cesium concentration in the effluent reached the initial level. The effluent solution was separated into three portions of 100 cm³ and analysed for cesium content. After, the sorption ion-exchange resin was washed with distilled water until pH 7.0 was reached. Then, 600 cm³ perrhenic acid at a concentration of 100 g Re/dm³ was passed up the column. Ion-exchange resin was washed until neutral pH was reached and used in the next process.

Solutions from elution and the first portion of washing solutions (containing 1.0 g/dm³ Cs) were concentrated. Three operation cycles were performed in this way. The succeeding data were evaluated: cesium ion sorption efficiency: Equation (1), cesium ion elution efficiency: Equation (2), ion-exchange saturation by cesium ions in each stage: Equation (3); OZ₁: bed volume for which cesium concentration is <0.01 g/dm³; OZ₂: bed volume when the column is totally filled; OZ₃: bed volume when cesium concentration is higher than 5.0 g/dm³, for elution; OZ₄: bed volume when cesium concentration is higher than 1.0 g/dm³, for elution and washing after elution.

$$S_{Cs} = \frac{m_{0Cs} - C_{Cs}V}{m_J} \times 100\% \quad (3)$$

where: m_{0Cs} is the cesium mass in a solution before sorption (g); m_J is the mass of ion-exchange resin (cm³); C_{Cs} is the cesium concentration in solution after sorption (g/dm³); and V is the volume of solution after sorption (dm³).

Solutions from elution and washing after elution were combined in such a way as to obtain different cesium concentrations, namely 3, 5, and 10 g/dm³ in a concentration stage. The volume of solutions directed to crystallisation was constant and equal to 1.5 dm³. Concentrated solutions were intensively mixed at temperatures up to 353 K. Then, solutions were cooled at 0.5 K/min until ambient temperature. Crystallised cesium perrhenate was separated from solutions by filtration. Additionally, cesium perrhenate prepared using 10 g/dm³ Cs solution was divided into five parts and directed to the purification stage. The purification process was a single- or two-step washing of filtered crystals. The washing media were 50 cm³ portions of 5% or 10% aqueous hydrogen peroxide solutions and/or 20 cm³ portions of anhydrous acetone. Purified crystals were dried at 393 K until constant mass. Cesium and rhenium content, as well as selected contaminations (i.e., As, Cu, Zn, Ca, Fe, Mo, Ni, Pb, Na, K, Mg, Bi) were analysed and the crystallite size was evaluated.

2.3. Characterisation

All necessary analyses were performed in the Department of Analytical Chemistry at the Institute of Non-Ferrous Metals in Gliwice. Base element (Re, Cs) content was determined in obtained cesium perrhenates. A weight method using tetraphenylarsenic chloride (CTFA) as the precipitating agent was used to evaluate rhenium content and flame atomic emission spectroscopy (FAES, spectrophotometer AAS novAA400) for cesium content. In order to determine As, Cu, Zn, Ca, Fe, Mo, Ni, Pb, Na, K, Mg, Bi content, different techniques were used: GFAAS (graphite furnace atomic absorption spectroscopy with graphite cells, Z-2000, HITACHI), ICP-OES (inductively coupled plasma—optical emission spectrometer, ULTIMA 2, HORIBA Jobin-Ivon), and ICP-MS (inductively coupled plasma mass spectrometry, ICP MS NexION). Solutions were analysed to evaluate rhenium and cesium content by FAAS (flame atomic absorption spectroscopy equipped with flame module and deuter background correction, SOLAAR S4, THERMO). X-ray diffraction analysis was performed using Co K α radiation and an Fe filter (XRD 7, Seifert-FPM). Based on the XRD pattern, the crystallite size was calculated using the Scherrer method without accounting for stresses. Crystallite size can be calculated using the formula:

$$L = \frac{K\lambda}{B(2\theta)\cos\theta} \quad (4)$$

where: L is the crystallite size, K is the shape factor ca. 0.9 if crystallites are spherical, λ is the X-ray wavelength, θ is Bragg's angle, and $B(2\theta)$ is the line-broadening at half-maximum.

3. Results and Discussion

3.1. Ion-Exchange Resin Selection—Sorption Stage

In the preliminary experiments concerning ion-exchange resin selection, it was found that all materials had above 65.0% efficiency of cesium ion sorption. The highest values were recorded for PFC100 and C160—above 90.0% cesium sorption. Results are presented in Table 1. CT169, C160, and PFC100 were selected for further investigations.

Table 1. Results of ion-exchange resin selection for cesium sorption.

Ion-Exchange Resin	Volume of a Solution after Sorption	Cs Concentration in a Solution after Sorption	Cs ⁺ Sorption Efficiency
	cm ³	g/cm ³	%
S100	110	1.3	71.4
SP112	115	1.4	67.8
PFC100	125	0.4	90.0
CT169	120	0.3	92.8
C160	125	0.3	92.5

3.2. Influence of Temperature on Cesium Ion Sorption Efficiency

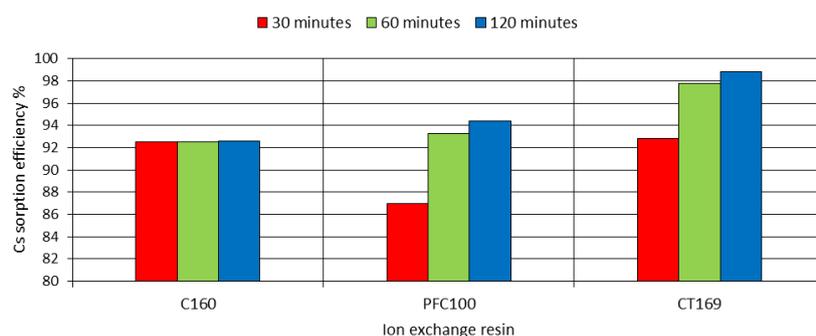
No influence of this parameter on sorption was observed for C160 ion-exchange resin. For the other materials, a slight sorption increase with increasing temperature was observed. Due to that observation, and the lower stability of ion exchange resins at elevated temperatures, all tests were performed at ambient temperature. The results are presented in Table 2.

Table 2. Results of the effect of temperature on cesium ion sorption efficiency.

Ion Exchange Resin	Temperature	Volume of a Solution after Sorption	Cs Concentration in a Solution after Sorption	Cs ⁺ Sorption Efficiency
	K			
C160	293	125	0.3	92.5
	313	120	0.3	92.8
	333	125	0.3	92.5
CT169	293	120	0.3	92.8
	313	120	0.1	97.6
	333	126	0.1	97.5
PFC100	293	125	0.4	90.0
	313	120	0.2	95.2
	333	125	0.2	95.0

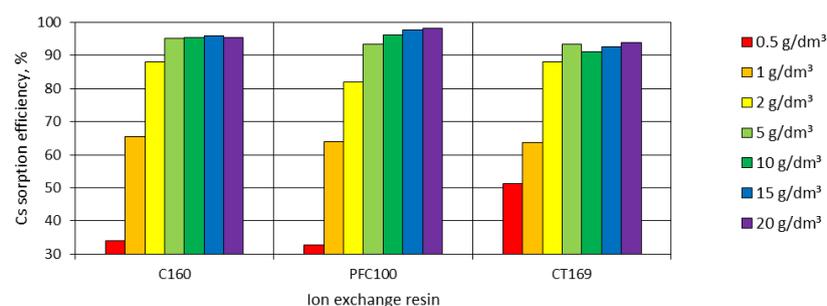
3.3. Influence of Contact Time on Cesium Ion Sorption Efficiency

No influence of contact time on cesium ion sorption efficiency was found for C160 ion-exchange resin while, for PFC100 and CT169, a significant increase was observed for longer contact times. Therefore, 60 min was selected as a required contact time for further static measurements. Results are presented in Figure 1.

**Figure 1.** Influence of contact time on cesium ion sorption efficiency.

3.4. Influence of Cesium Concentration on Cesium Ion Sorption Efficiency

Cesium concentration in nitrate(V) solutions before sorption was in the range 0.5–20.0 g/dm³. It was found that, for CT169, efficiency of cesium ions was 99% when a solution of concentration 5.0 g/dm³ was used. Additionally, C160 and PFC100 had slightly lower cesium sorption efficiencies. For cesium concentration higher than 5.0 g/dm³, about 90% sorption efficiency was observed, regardless of the type of ion-exchange resin. The results are presented in Figure 2.

**Figure 2.** Influence of cesium concentration on cesium ion sorption efficiency.

3.5. Ion-Exchange Selection—Elution Stage

Elution tests allowed the distinguishing and selection of the best appropriate ion-exchange resins. A high value of cesium ion elution was observed for C160 and CT169. It should be noted that these resins had high stability and appropriate sorption parameters. These materials were chosen for further experiments to perform cesium ion elution tests. The results are presented in Table 3.

Table 3. Results of ion-exchange resin selection for cesium sorption.

Ion-Exchange Resin	Volume of a Solution	Cs Concentration in a Solution after Elution	Cs ⁺ Elution Efficiency
	cm ³	g/cm ³	%
C160	25	8.1	88.0
CT169	24	9.4	96.0
PFC100	25	4.5	50.0

3.6. Influence of Contact Time vs. Elution Efficiency

It was established that cesium ion elution efficiency slightly increased with increase of contact time. The time of 60 min was assumed to be optimal for further research. Results are presented in Figure 3.

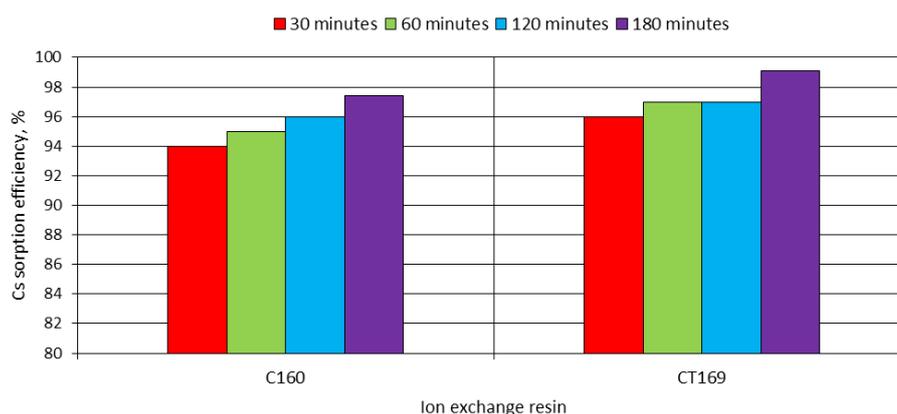


Figure 3. Influence of contact time on cesium ion elution efficiency.

3.7. Influence of Temperature on Cesium Ion Elution Efficiency

No temperature influence on elution efficiency for CT169 and C160 was observed. Results are presented in Table 4.

Table 4. Influence of temperature on cesium ions elution efficiency.

Ion-Exchange Resin	Temperature	Volume of a Solution	Cs Concentration in a Solution after Elution	Cs ⁺ Sorption Efficiency
	K	cm ³	g/cm ³	%
CT169	293	24	9.5	97.0
	313	24	9.4	96.0
	333	24	9.5	97.0
C160	293	24	9.3	97.0
	313	23	9.8	98.0
	333	23	9.8	98.0

3.8. Influence of Rhenium-to-Cesium Ratio on Cesium Ion Elution Efficiency

The rhenium-to cesium ratio increase in a solution enhances Cs^+ ion elution efficiency. It was found that at least a 10 times excess of rhenium with respect to cesium was required to exceed 90% cesium elution efficiency for CT169 resin. In the case of C160, the highest elution efficiency (92%) was reported for a rhenium-to-cesium ratio 30:1. Results are presented in Figure 4. CT169 was chosen as the best resin and it was used in dynamic tests.

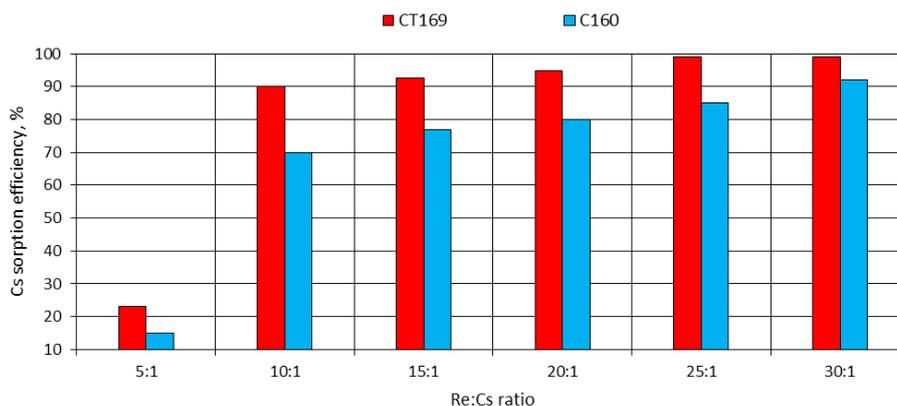


Figure 4. Influence of rhenium-to-cesium ratio on cesium ion elution efficiency.

3.9. Analysis of Three Cycles of CT169 in Dynamic Conditions

Sorption and elution tests of cesium ions allowed the determination of CT169 as the most appropriate resin for this purpose. Based on these results, optimum process parameters for dynamic tests were presented. Three operation cycles were performed using CT169. Results of cesium ion sorption are presented in Figure 5 and Table 5, while elution results are shown in Figure 6 and Table 6. High and stable sorption parameters, 75%–70% sorption efficiency and 4.5%–4.6% ion-exchange saturation, by cesium ions were obtained. Elution was very effective with an efficiency significantly above 95.0% in all cycles. It should be pointed out that a slight increase in the efficiency was observed in subsequent cycles, which indicates proper operation of the ion-exchange resin.

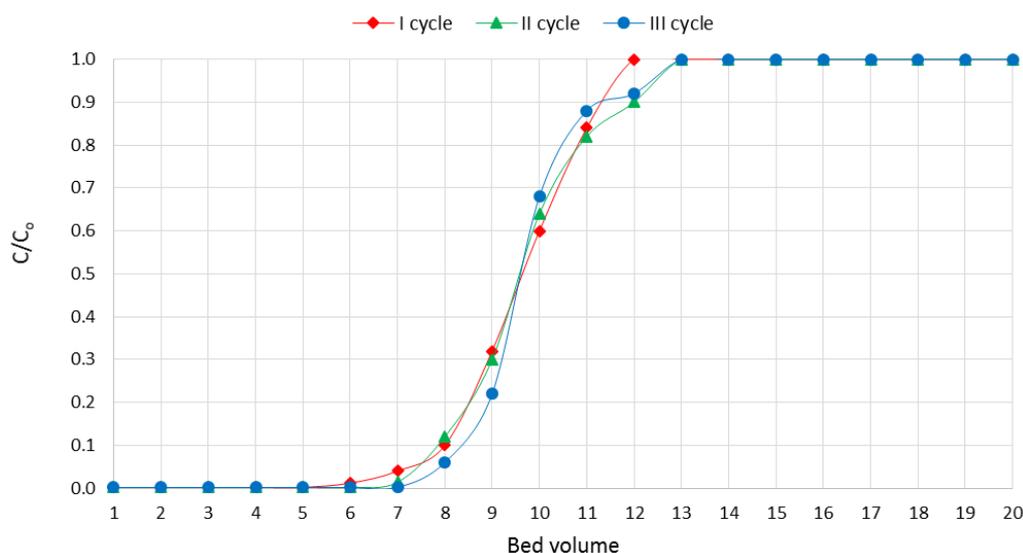


Figure 5. Characteristics of CT169 during sorption.

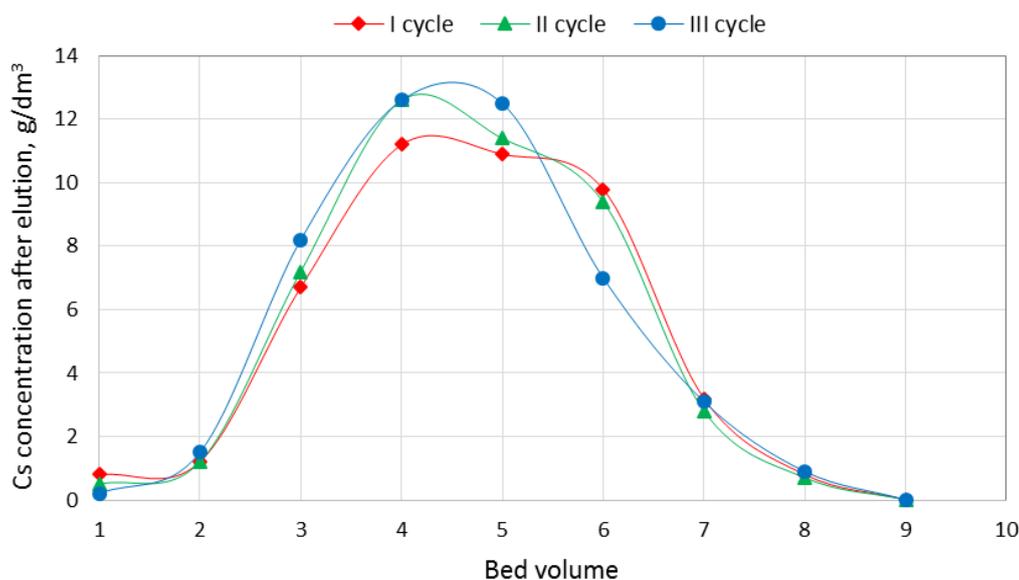


Figure 6. Characteristics of CT169 during elution.

Table 5. Parameters for cesium ion sorption—CT169.

Cycle No.	OZ ₁	OZ ₂	W _{Cs} , %	S _{Cs} , %
I	5	12	75.6	4.5
II	6	13	70.7	4.6
III	6	13	70.9	4.6

Table 6. Parameters for cesium ions elution—CT169.

Cycle No.	OZ ₃	OZ ₄	W' _{Cs} , %
I	4	6	98.3
II	4	6	99.6
III	4	6	99.7

Results of crystallisation tests are presented in Table 7. It was found that over 80% efficiency of cesium perhenate crystallisation may be achieved when cesium concentration in a solution is higher than 5.0 g/dm³ Cs. Impurity contents before and after purification are shown in Tables 8 and 9, respectively. The proposed two-step crystal purification method—with 50 cm³ 10% aq. H₂O₂ solution, followed by 20 cm³ acetone washing—allowed for obtaining nanocrystalline cesium perhenate containing stoichiometric amounts of cesium and rhenium (i.e., 34.7% Cs and 48.6% Re and <2 ppm Bi; <3 ppm Zn; <2 ppm As; <10 ppm Ni; <3 ppm Mg; <5 ppm Cu; <2 ppm Na; <5 ppm Ca; <3 ppm Fe; <5 ppm Mo; <5 ppm Pb; <10 ppm K). XRD analysis of the crystals (Figure 7) proved that cesium perhenate was prepared.

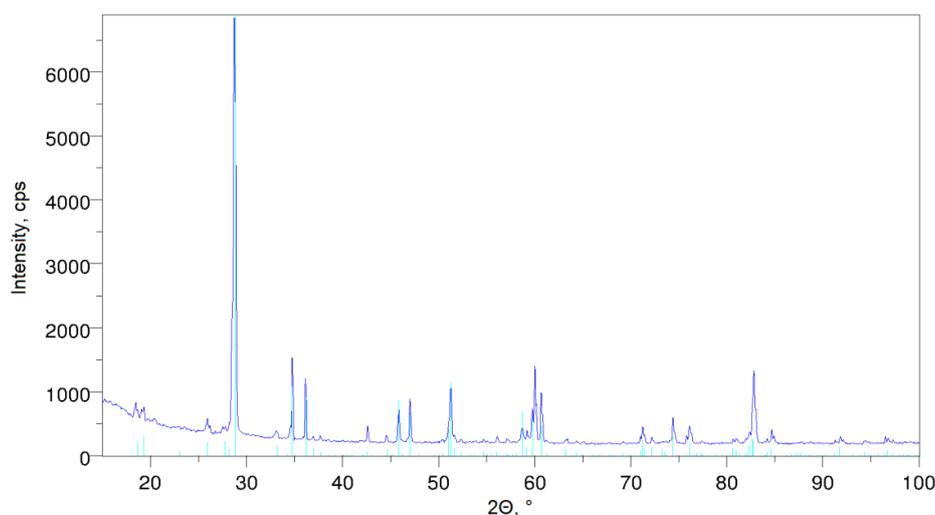


Figure 7. X-ray powder diffraction (XRD) patterns of cesium perrhenate crystals. (dark blue—sample pattern; light blue—database data).

Table 7. Results of CsReO₄ crystallisation tests.

Test No.	$c_{0\text{KCs}}$ (g/dm ³)	$c_{0\text{KRe}}$ (g/dm ³)	Crystallite Size (nm)	Crystals Mass (g)	X_{Cs} (%)	X_{Re} (%)	W_{KCsReO_4} (%)
1	0.8	56.0	- *	1.10	- *	- *	31.8
2	3.2	75.0	44	9.10	34.88	49.25	65.8
3	5.2	85.0	40	18.00	34.80	48.89	80.1
4	10.5	98.0	35	40.00	34.75	48.61	88.2

* analysis not done due to crystal mass being too small; $c_{0\text{KCs}}$: concentration of cesium in a solution directed to crystallisation; $c_{0\text{KRe}}$: concentration of rhenium in a solution directed to crystallisation; X_{Cs} : cesium content in crystals; X_{Re} : rhenium content in crystals; W_{KCsReO_4} : crystallisation efficiency (i.e., ratio of crystals' mass to theoretical mass calculated according to Cs content in a solution).

Table 8. Impurity contents in crystallised CsReO₄.

Test No.	Bi (ppm)	Na (ppm)	As (ppm)	Ni (ppm)	Mg (ppm)	Cu (ppm)	Ca (ppm)	Fe (ppm)	Zn (ppm)	Mo (ppm)	K (ppm)	Pb (ppm)
1		50		48	15			12	10			
2	<2	48	<2	38	16	<5	<5	15	12	<5	<10	<5
3		47		50	12			20	15			

Table 9. Impurity contents after purification.

	Ni (ppm)	Mg (ppm)	Na (ppm)	Fe (ppm)	Zn (ppm)
No purification	50	12	47	20	15
5% H ₂ O ₂	32	10	20	5	15
10% H ₂ O ₂	<10	10	15	<3	12
acetone	46	5	15	6	5
10% H ₂ O ₂ + acetone	<10	<3	<2	<3	<3

3.10. Scheme of the Developed Technology

The results of the experiments allowed the development of a technological scheme to produce high-purity nanocrystalline cesium perrhenate, which is patent-pending [32].

The developed method of anhydrous cesium perrhenate preparation is composed of eight technological blocks (Figure 8). Firstly, aqueous cesium nitrate solutions are directed to cesium

sorption stage, where the solution is transferred through a CT169 ion-exchange resin bed placed inside an ion-exchange column. Sorption is performed down the columns at room temperature. The solution–bed contact time is up to 60 min. After the sorption stage the ion-exchange resin is washed using demineralised water, while post-sorption and post-washing solution streams are connected and directed to the initial solution preparation. The washed column containing the bed with sorbed cesium is eluted downward with aqueous perrhenic acid solution at room temperature, so the eluent–resin contact time is above 120 min. The eluate is collected as a single portion and directed to concentration and crystallisation, while the eluted bed is washed with demineralised water. The post-washing solutions obtained after elution are divided into two portions—the first one is connected with the solution directed to the concentration stage, while the second one, together with mother liquors and post-purification solutions, is directed to waste management. The solution obtained after the second part of elution, containing more than 5.0 g/dm³ Cs, is directed to concentration and crystallisation. The concentration stage is carried out at a temperature not exceeding 80 °C, preferably under vigorous stirring. Then, the solution is quenched and cesium perrhenate is crystallised. The crystals are filtered from a solution and washed with 10% aq. hydrogen peroxide solution and anhydrous acetone. Finally, the prepared cesium perrhenate crystals (≤ 45 nm) are dried at 120 °C until constant mass.

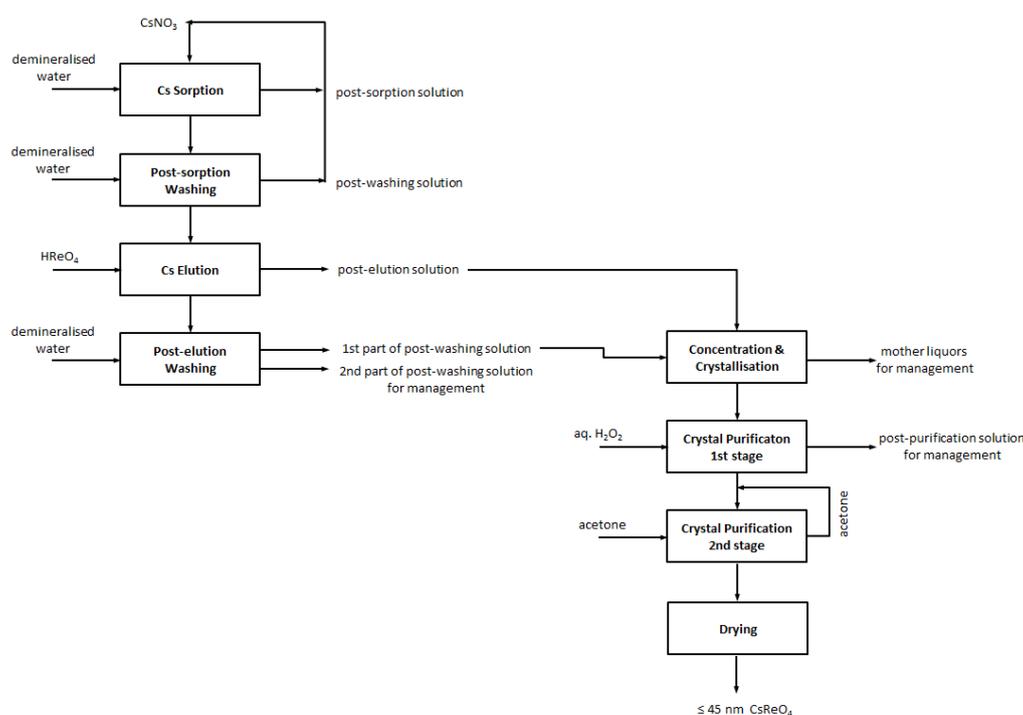


Figure 8. Technological scheme of anhydrous nanocrystalline cesium perrhenate preparation.

The proposed method, and its process conditions, allow the preparation of anhydrous nanocrystalline (≤ 45 nm) cesium perrhenate of stoichiometric composition in a reproducible way (Table 10).

Table 10. Separation and purification of cesium perrhenate.

Volume of Solution Sent for Concentration	Cs Concentration in a Solution for Concentration	Re Concentration in a Solution for Concentration	Crystallite Size	Crystallisation Efficiency
cm ³	g/dm ³	g/dm ³	nm	%
300	0.8	56.0	-	-
300	3.2	75.0	44	65.5
300	5.2	85.0	40	80.2
300	10.5	98.0	35	85.6

4. Conclusions

It was found that the ion-exchange method with strongly acidic cation-exchange resin (CT169) was able to produce extra-pure anhydrous cesium perrhenate with crystallite size <45 nm and with the following composition: 34.7% Cs; 48.6% Re; <2 ppm Bi; <3 ppm Zn; <2 ppm As; <10 ppm Ni; <3 ppm Mg; <5 ppm Cu; <2 ppm Na; <5 ppm Ca; <3 ppm Fe; <5 ppm Mo; <5 ppm Pb; <10 ppm K. Using this method, 2.0 kg cesium perrhenate was prepared.

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Conflicts of Interest: The authors declare no conflict of interest.

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