

*Supplementary materials*

# **Fabrication of Cementitious Microfiltration Membrane and Its Catalytic Ozonation for the Removal of Small Molecule Organic Pollutants**

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Text S1 Testing method of BSA rejection efficiency by CM.

The separation performance of CCMs was demonstrated by the rejected BSA based on the Bradford method [1, 2]. The BSA rejection ( $R_{BSA}$ ) experiments were performed in a dead-end filtration cell in a laboratory-scale made setup. The weight of feed solution filtering the membrane in unit time under a series of pressures (20–100 kPa) is investigated, according to following **Equation. (1)**:

$$R_{BSA} = \left(1 - \frac{C_P}{C_0}\right) \times 100\% \quad (1)$$

Where  $R_{BSA}$  is the rejection rate of BSA (%),  $C_P$  is the concentration of BSA in permeate water, and  $C_0$  is the concentration of BSA in feed water. The concentration of 100 mg L<sup>-1</sup> BSA solution was used as feed solution.

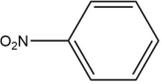
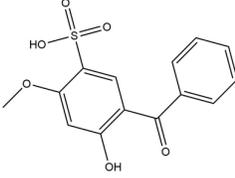
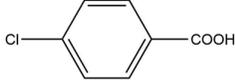
Text S2 Measured method of PWF of the membrane.

By controlling the filter pressure, the mass of pure water passed through the membrane within a certain period of time can be used to calculate the pure water flux of the membrane according to **Equation. S1**:

$$J_w = \frac{10 \times m_w}{A \times \rho \times t} \quad (1)$$

where  $J_w$  is the pure water flux of the membrane ( $\text{L m}^{-2} \text{ h}^{-1}$ ),  $m_w$  is the mass of the permeate pure water (g),  $A$  is the effective filtration area ( $\text{cm}^2$ ), which value is 7.068  $\text{cm}^2$ ,  $\rho$  is the density of water at 25 °C, which value is 0.99705  $\text{g cm}^{-3}$ , and  $t$  is the penetration time (h). The amount of deionized water that can be filtered by the membrane under a series trans-membrane pressure (40–160 kPa) was investigated. All experiments were repeated three times.

**Table S1.** Basic information of selected organic compounds.

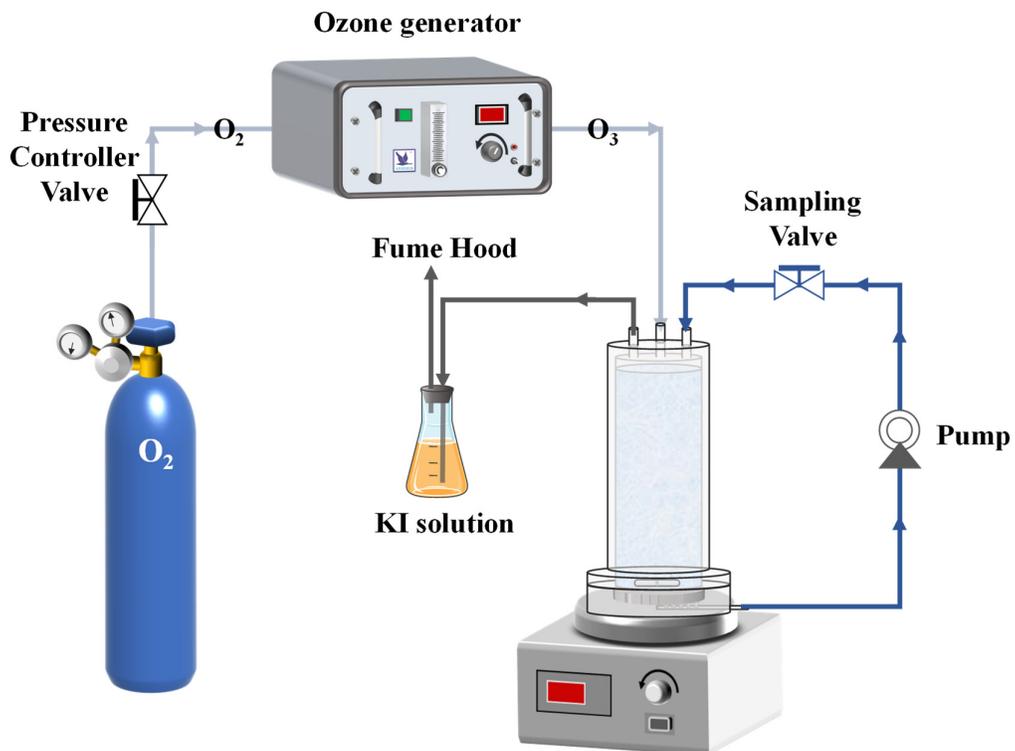
Name	Structural formula	Chemical formula	Molecular weight	pK <sub>a</sub>	logK <sub>ow</sub>
Nitrobenzene		C <sub>6</sub> H <sub>5</sub> NO <sub>2</sub>	123.111	3.98	1.85
<i>p</i> -chloroaniline		C <sub>6</sub> H <sub>6</sub> ClN	127.57	3.98 2	1.83
Benzophenone-4		C <sub>14</sub> H <sub>12</sub> O <sub>6</sub> S	308.31	5.99 3	1.89
<i>p</i> -chlorophenol		C <sub>6</sub> H <sub>5</sub> ClO	128.55	9.41	2.39
<i>p</i> -chloronitrobenzene		C <sub>6</sub> H <sub>4</sub> ClNO <sub>2</sub>	157.553	-	2.39
<i>p</i> -chlorobenzoic acid		C <sub>7</sub> H <sub>5</sub> ClO <sub>2</sub>	156.565	3.98	2.65

**Table S2.** Reaction kinetics of organic pollutants removal by CM–ozone catalyst

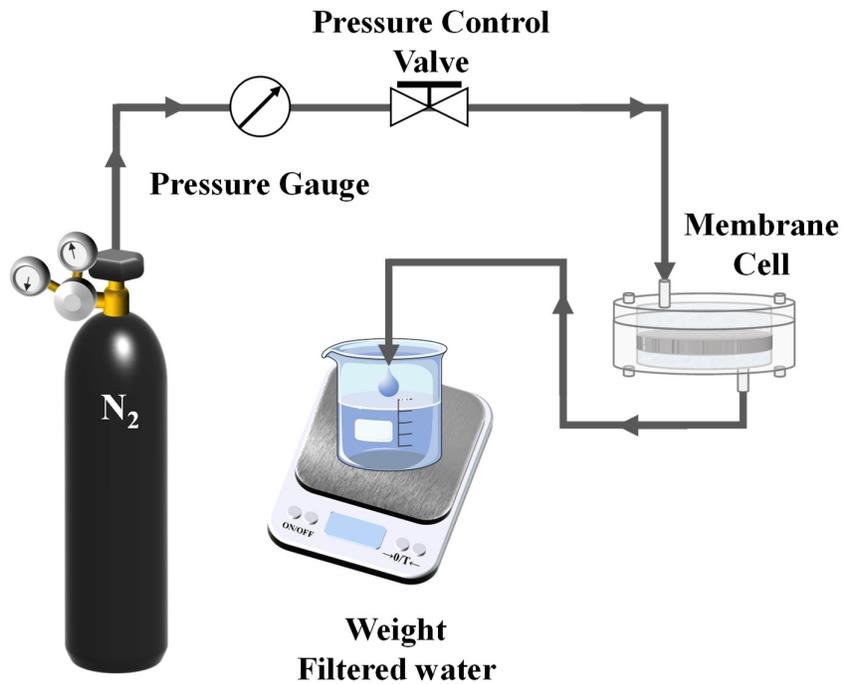
process reuse experiments.

Reuse time	Nitrobenzene		BP-4		<i>p</i> -CP		<i>p</i> -CNB		<i>p</i> -CBA	
	$k_{\text{obs}}$ $\text{min}^{-1}$	$R^2$								
1 st	0.191	0.993	0.325	0.993	0.747	0.993	0.143	0.993	0.214	0.998
2 nd	0.189	0.996	0.318	0.991	0.688	0.996	0.138	0.992	0.209	0.996
3 rd	0.180	0.997	0.303	0.995	0.683	0.995	0.137	0.990	0.216	0.989
4 th	0.177	0.995	0.283	0.998	0.634	0.997	0.134	0.986	0.207	0.993
5 th	0.185	0.995	0.273	0.997	0.571	1.000	0.133	0.986	0.203	0.992
6 th	0.184	0.995	0.262	0.999	0.559	1.000	0.130	0.988	0.203	0.998

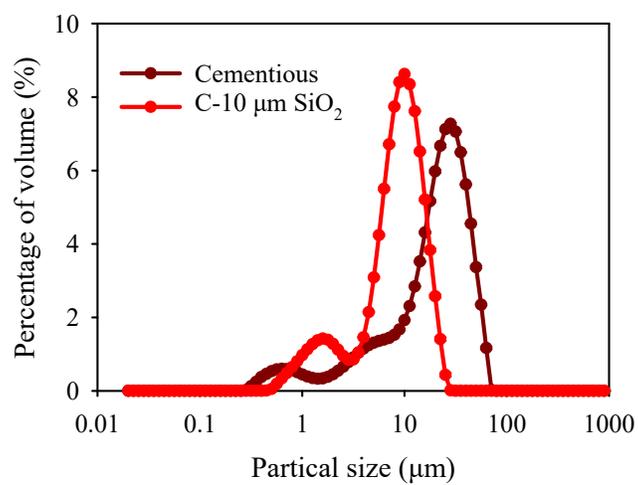
Conditions: pH = 6.9±0.1, [O<sub>3</sub>] = 0.5 mg L<sup>-1</sup>, [nitrobenzene]<sub>0</sub> = [BP-4]<sub>0</sub> = [*p*-CP]<sub>0</sub> = [*p*-CNB]<sub>0</sub> = [*p*-CBA]<sub>0</sub> = 0.064 mM.



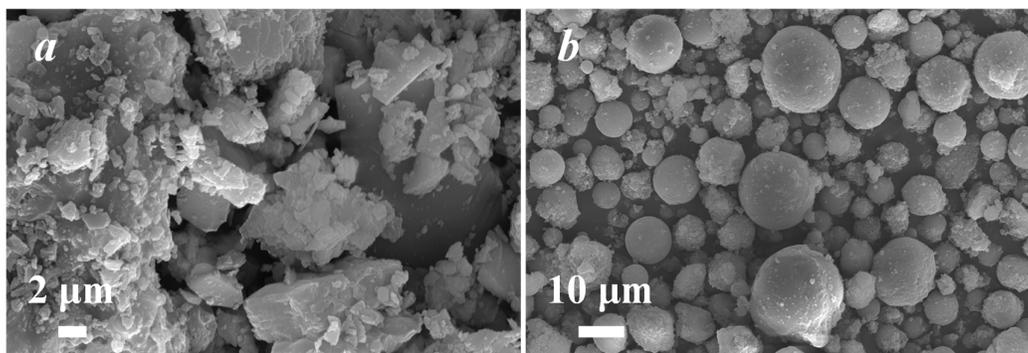
**Figure S1.** Schematic diagram of the CM-ozone coupling process to removal organic pollutant.



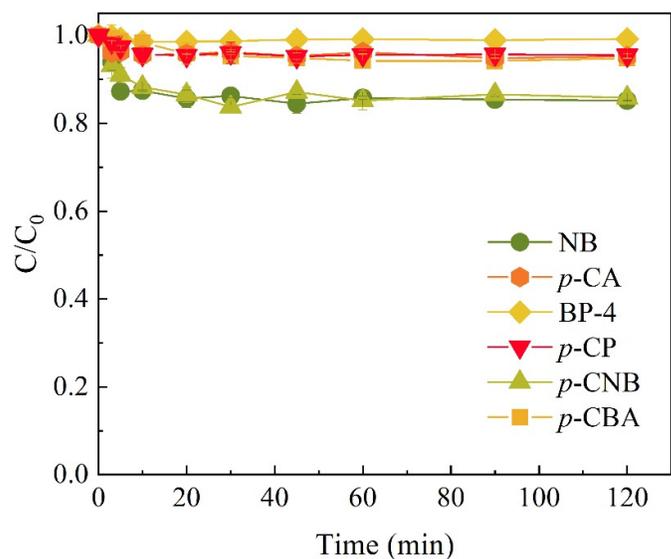
**Figure S2.** Schematic for PWF of the membrane measured.



**Figure S3.** Particle size distribution of raw materials of the fabricated cementitious membranes fabricated.

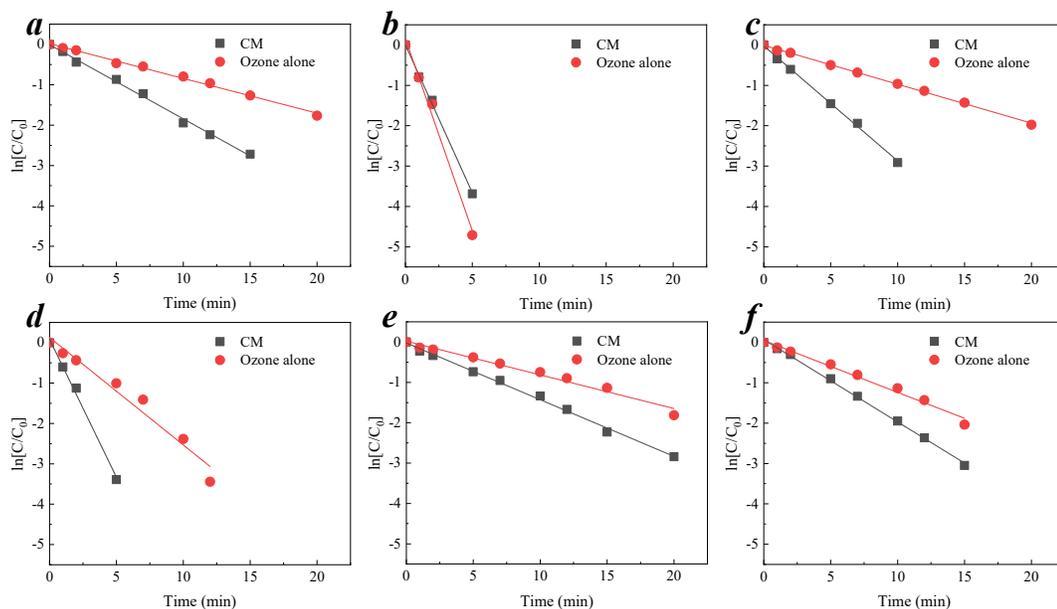


**Figure S4.** Apparent morphology of raw materials for membrane fabrication. (a) Cementitious powder; (b) SiO<sub>2</sub> particle C-10 μm.



**Figure S5.** Adsorption of the six pollutants by the CM. Conditions:  $\text{pH} = 6.9 \pm 0.1$ ,

$[\text{nitrobenzene}]_0 = [p\text{-CA}]_0 = [\text{BP-4}]_0 = [p\text{-CP}]_0 = [p\text{-CNB}]_0 = [p\text{-CBA}]_0 = 10 \text{ mg L}^{-1}$ .



**Figure S6.** Reaction kinetics of organic compound ozonation catalyzed by CMs. (a) nitrobenzene; (b) *p*-CA; (c) BP-4; (d) *p*-CP; (e) *p*-CNB; (f) *p*-CBA. Conditions:  $\text{pH} = 6.9 \pm 0.1$ ,  $[\text{O}_3] = 0.5 \text{ mg L}^{-1}$ ,  $[\text{nitrobenzene}]_0 = [p\text{-CA}]_0 = [\text{BP-4}]_0 = [p\text{-CP}]_0 = [p\text{-CNB}]_0 = [p\text{-CBA}]_0 = 0.064 \text{ mM}$ .

**References:**

1. Kim, K. S.; Sajjad, M.; Lee, J.; Park, J.; Jun, T., Variation of extracellular polymeric substances (EPS) and specific resistance to filtration in sludge granulation process to the change of influent organic loading rate. *Desalination and Water Treatment* **2014**, *52*, (22-24), 4376-4387.
2. Sedmak, J. J.; Grossberg, S. E., A rapid, sensitive, and versatile assay for protein using Coomassie brilliant blue G250. *Analytical biochemistry* **1977**, *79*, (1-2), 544-52.