

## Article

# Separation of CO<sub>2</sub> in a Solid Waste Management Incineration Facility Using Activated Carbon Derived from Pine Sawdust

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**Abstract:** The selective separation of CO<sub>2</sub> from gas mixtures representative of flue gas generated in waste incineration systems is studied on two activated carbons obtained from pine sawdust and compared to a commercial activated carbon. Dynamic adsorption experiments were conducted in a fixed-bed adsorption column using a binary mixture (N<sub>2</sub>/CO<sub>2</sub>) with a composition representative of incineration streams at temperatures from 30 to 70 °C. The adsorption behavior of humid mixtures (N<sub>2</sub>/CO<sub>2</sub>/H<sub>2</sub>O) was also evaluated in order to assess the influence of water vapor in CO<sub>2</sub> adsorption at different relative humidity in the feed gas: 22% and 60%. Moreover, CO<sub>2</sub> adsorption was studied in less favorable conditions, i.e., departing from a bed initially saturated with H<sub>2</sub>O. In addition, the effect of CO<sub>2</sub> on H<sub>2</sub>O adsorption was examined. Experimental results showed that the CO<sub>2</sub> adsorption capacity can be reduced significantly by the adsorption of H<sub>2</sub>O (up to 60% at high relative humidity conditions). On the other hand, the breakthrough tests over the adsorbent initially saturated with water vapor indicated that H<sub>2</sub>O is little affected by CO<sub>2</sub> adsorption. The experimental results pointed out the biomass based carbons as best candidates for CO<sub>2</sub> separation under incineration flue gas conditions.

**Keywords:** activated carbon; CO<sub>2</sub> capture; incineration; breakthrough curve; water vapor

## 1. Introduction

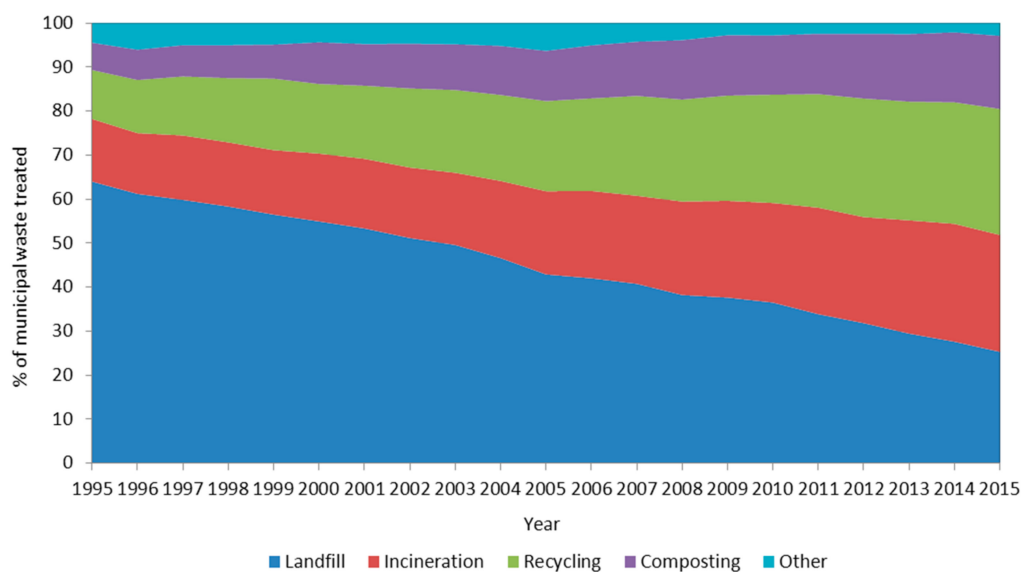
The disposal of municipal solid waste is one of the main urban issues. Development of integrated solid waste management facilities is necessary considering how to prevent, recycle, and manage solid waste in ways that most effectively protect human health and the environment. The European Legislation about waste disposal (Directive 2008/98/EC) has established a hierarchy of available technologies. It sets out five steps for dealing with waste, ranked according to the environmental impact:

- Waste minimization: Using less material in design and manufacture, keeping products for longer, re-use, using less hazardous materials.
- Reuse: Checking, cleaning, repairing, refurbishing, whole items or spare parts.
- Recycling: Turning waste into a new substance or product.
- Other recovery: Includes anaerobic digestion, incineration with energy recovery, and gasification and pyrolysis that produce energy (fuels, heat and power) and materials from waste and some backfilling operations.
- Disposal: Landfilling and incineration without energy recovery.

It is also important to bear in mind that not all waste is suitable for recycling (e.g., too contaminated materials or mixed materials), that materials degrade after repeated recycling and the existing demand for recycled products. If high quality recycling is not possible, the waste should be turned into energy, rather than being landfilled. In addition, some residues have to be incinerated such as potentially infectious clinical waste, treat meat and bone meal residues, drugs or polluted plastic packages. Therefore, the market for thermal treatment and energetic recovery of residual and other types of solid waste is growing continuously. Increasing waste amounts, shrinking landfill spaces in agglomerations and higher ecological standards stimulate this growth throughout the world.

Incineration is a disposal method in which solid organic wastes are subjected to combustion to convert them into residue and gaseous products. It reduces the waste volume and helps generate heat and/or electricity for commercial use. This generates a renewable energy source and reduces carbon emissions by offsetting the need for energy from fossil sources. The most common examples are municipal solid waste incineration plants, hazardous waste incineration plants, sewage sludge incineration plants and incineration plants for clinical waste or animal carcasses.

Each citizen of the European Union on average produces 477 kg of municipal waste per year. In 2015, that waste was managed by a variety of means: 25.3% was landfilled, 28.6% recycled, 26.6% incinerated, and 16.6% composted. It is important to note that the share of incinerated municipal waste in the EU has steadily increased over the time period, from 14.2% in 1995 to 26.6% in 2015 as it can be observed in Figure 1 (Eurostat, 2017).



**Figure 1.** Municipal waste management by type of treatment, EU-27 (%), 1995–2015 (Source: Eurostat).

According to the Confederation of European Waste-to-Energy Plants (CEWEP), in 2014, there were 482 incineration plants with energy recovery across Europe that treated about 88 million tonnes of household and similar wastes, remaining after waste prevention, reuse and recycling (hazardous waste is not included). Worldwide there are more than 2200 Waste-to-Energy plants active, with a disposal capacity of around 300 million tonnes of waste per year.

Waste incineration flue gas is mainly composed of  $N_2$ ,  $CO_2$ ,  $O_2$  and  $H_2O$  with much lesser contents of  $NO_x$ ,  $SO_2$  and  $CO$ . Waste-to-Energy plants have a range of different emission abatement equipment which aims to ensure compliance with emission regulations. The contents in the raw flue gas of  $HCl$  (and volatile metal chlorides),  $SO_2$ ,  $NO_x$ , persistent organic pollutants and heavy metals need to be removed before emission of the flue gas to the atmosphere [1–3]. However,  $CO_2$  in flue gas is not currently addressed by flue gas cleaning systems in incineration facilities although some pilot plant studies are being carried out [4]. In flue gas from municipal solid waste incinerators, the  $CO_2$

concentration was found to be about 6–12% [5]. It is then feasible to develop an integrated solution to reduce the emissions of CO<sub>2</sub> from waste incineration by recovering CO<sub>2</sub>.

In order to reduce the content of CO<sub>2</sub> in the gaseous streams, different technologies are available, such as absorption processes, membrane separation, cryogenic distillation and adsorption [6]. Adsorption-based processes that employ microporous materials constitute a very promising cost-efficient technology. Various microporous materials, such as activated carbons, zeolites and metal-organic frameworks (MOFs) have been considered to carry out the CO<sub>2</sub> separation. The capture of CO<sub>2</sub> using these types of dry sorbents involves gas-solid interactions where the porosity of the adsorbent, the temperature and partial pressure during adsorption and the easiness of the regeneration procedure are key drivers. Recent research has focused on improving the adsorbent characteristics but also on developing adsorption processes that are cost-effective for CO<sub>2</sub> capture [7–10]. The utilization of activated carbons offers advantages, when compared to other non-carbonaceous adsorbents, such as high adsorption capacity at ambient conditions, low regeneration cost, long-term stability and fast kinetics.

Many references in the literature address CO<sub>2</sub> adsorption using activated carbons in post-combustion conditions [11–15], mostly applied to coal-fired power plants flue gas. They offer advantages when compared to state-of-the-art technology, chemical absorption, owing to their stability, easy manufacture, lower cost and great potential to reduce the energy penalty associated to the regeneration step. Moreover, adsorption technology will show further advantages in an incineration facility where the volume of flue gas is significantly smaller. Incinerator flue gas has a lower partial pressure of CO<sub>2</sub> and a higher partial pressure of O<sub>2</sub> compared to coal-fired power plant flue gas [4]. To the best of authors' knowledge, adsorption with biomass based carbons has not been previously explored for the separation of CO<sub>2</sub> from incineration flue gases.

Another important aspect that is highlighted in this work has to do with the high water content in flue gases; hence, the impact of water vapor on CO<sub>2</sub> capture is a critical parameter that is particularly assessed in this work. Although the adsorption equilibrium and mechanisms of water vapor adsorption on activated carbon have been widely investigated [16–19], very few studies have used fixed-bed experiments to gain insight into the adsorption dynamics of water vapor on activated carbon [20–22]. The impact of water vapor on the capacity of biomass-based activated carbons to capture CO<sub>2</sub> has been evaluated: adsorption equilibrium data, measured for CO<sub>2</sub>, N<sub>2</sub>, O<sub>2</sub> and H<sub>2</sub>O, and breakthrough experiments performed under different relative humidity were conducted to explore the mutual influence of CO<sub>2</sub> and H<sub>2</sub>O adsorption.

## 2. Materials and Method

Two biomass-based activated carbons were prepared by single step activation with CO<sub>2</sub>. The carbon adsorbent pellets were produced from pine sawdust alone (IH3) and from blends of pine sawdust and coal tar pitch (IH2). The preparation process is thoroughly described in a previous work [23]. Further characterization of the samples was carried out in order to assess the microporous structure of the activated carbons.

These adsorbents are compared to a commercial activated carbon, Norit R2030 CO<sub>2</sub>, with a particle size of approximately 3 mm, used for the removal of carbon dioxide out of cold storage warehouses. Its characterization can be found elsewhere [24,25].

### 2.1. Chemical Characterization of the Activated Carbons

The prepared carbons were subjected to ultimate analysis in order to determine their chemical composition (see Table 1). Although both materials show a high content of carbon that of IH2 is slightly greater probably due to the addition of coal tar pitch. On the other hand, IH3 contains more oxygen. The hydrogen and nitrogen contents are similar for both activated carbons and sulfur content is negligible.

**Table 1.** Ultimate analysis of IH2 and IH3.

Sample	Elemental Analysis (wt %, daf *)			
	C	H	N	O
IH2	95.10	1.00	0.62	3.28
IH3	93.29	1.19	0.34	5.18

\* Dry ash free basis.

## 2.2. Physical Properties of the Activated Carbons

The main characteristics of the prepared adsorbents calculated from the N<sub>2</sub> adsorption isotherms at −196 °C and CO<sub>2</sub> adsorption isotherms at 0 °C are summarized in Table 2.

**Table 2.** Textural characterization by means of N<sub>2</sub> adsorption at −196 °C and CO<sub>2</sub> adsorption at 0 °C.

Sample	Adsorption Isotherm N <sub>2</sub> −196 °C						Adsorption Isotherm CO <sub>2</sub> 0 °C			
	$S_{BET}$ (m <sup>2</sup> /g)	$V_p$ (cm <sup>3</sup> /g)	$W_0$ (cm <sup>3</sup> /g)	$E_0$ (kJ/mol)	$L_0$ (nm)	$S_{mic}$ (m <sup>2</sup> /g)	$W_0$ (cm <sup>3</sup> /g)	$E_0$ (kJ/mol)	$L_0$ (nm)	$S_{mic}$ (m <sup>2</sup> /g)
IH2	563	0.23	0.21	29.85	0.59	733	0.24	30.11	0.58	836
IH3	788	0.34	0.30	25.88	0.75	811	0.30	30.06	0.58	1039
Norit R	988	0.41	0.39	18.97	1.43	549	0.26	28.02	0.65	823

CO<sub>2</sub> uptake is closely related to the textural characteristics of the adsorbents [26]. At pressures lower than 1000 kPa, the most influential parameter on CO<sub>2</sub> uptake is the micropore volume. As pressure approximates to atmospheric, the average narrow micropore width ( $L_{0,CO_2}$ ) is also relevant.

Similar features in terms of narrow micropore width are encountered in both activated carbons, although slightly higher values of narrow micropore volume,  $W_{0,CO_2}$ , and narrow micropore surface,  $S_{mic,CO_2}$ , were obtained for IH3. The characteristic energy,  $E_0$ , is practically equal for IH2 and IH3 showing the same strength in their interaction with CO<sub>2</sub>.

Comparing those activated carbons with the commercial Norit R2030, whose textural characterization has been reported previously [25], the latter presents a higher BET surface and total pore volume ( $V_p = 0.41$  cm<sup>3</sup>/g); however the narrow micropore volume ( $W_{0,CO_2} = 0.26$  cm<sup>3</sup>/g) is lower than that of IH3 and higher than IH2. It also presents lower  $E_0$  and higher  $L_{0,CO_2}$ .

## 2.3. Adsorption Isotherms

Pure component adsorption isotherms at three different temperatures (30, 50, and 70 °C) for CO<sub>2</sub>, N<sub>2</sub>, O<sub>2</sub> and H<sub>2</sub>O were collected.

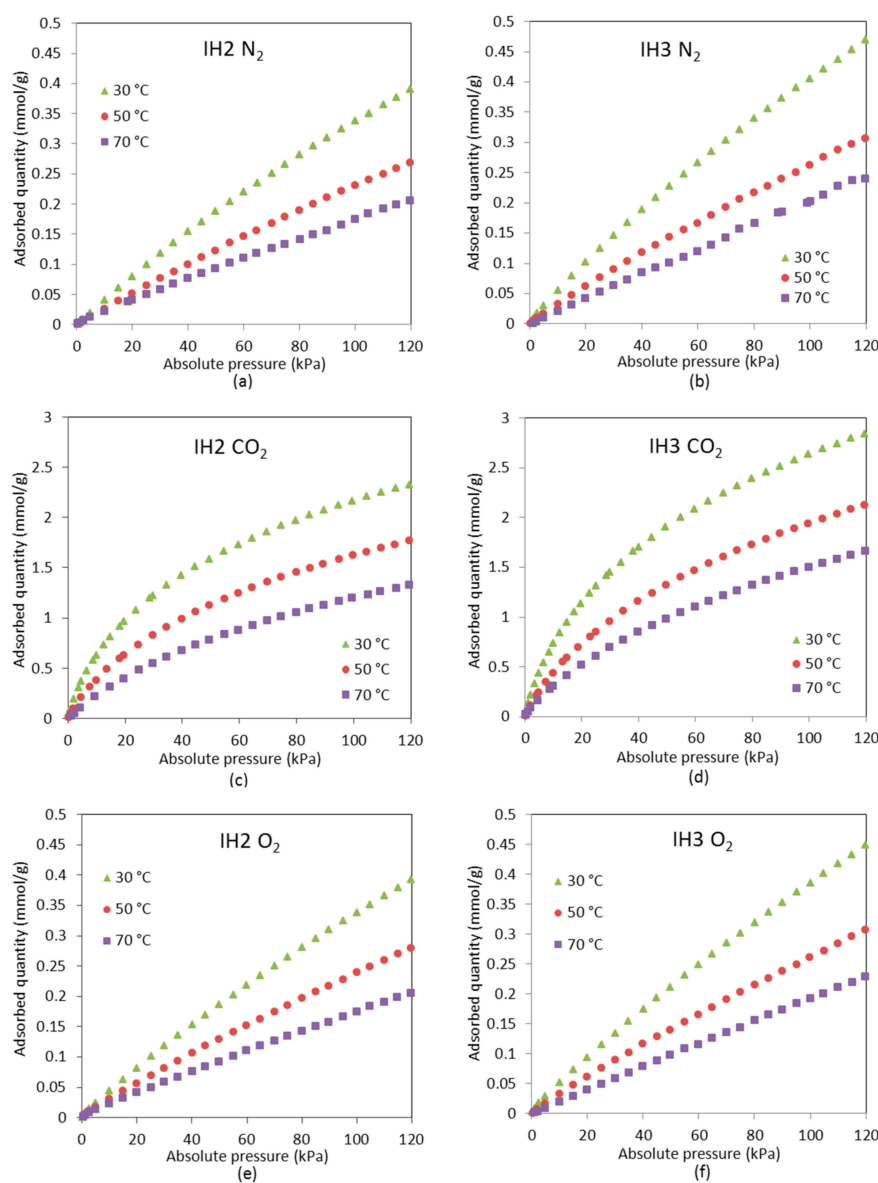
### 2.3.1. CO<sub>2</sub>, N<sub>2</sub> and O<sub>2</sub> Isotherms

The adsorption isotherms of pure CO<sub>2</sub>, N<sub>2</sub> and O<sub>2</sub> were measured up to 120 kPa using a commercial adsorption apparatus (Tristar 3000 from Micromeritics, Norcross, GA, USA). Prior to adsorption measurement, the sample was outgassed overnight under vacuum at 100 °C. The experimental data are graphically presented in Figure 2.

Both carbons present greater CO<sub>2</sub> adsorption capacity than N<sub>2</sub> and O<sub>2</sub> over the entire pressure range (subatmospheric pressures). Globally, the N<sub>2</sub> and O<sub>2</sub> isotherms exhibit linear patterns (Henry-type), characteristic of weak adsorbate-adsorbent interaction forces, whereas the CO<sub>2</sub> adsorption isotherms can be classified as Type I, i.e., they are representative of stronger adsorbate-adsorbent interactions.

Comparing the O<sub>2</sub> and N<sub>2</sub> adsorption isotherms for each sample, the similarity is very remarkable. As previously mentioned, flue gas from incineration plants contains about 7–14% O<sub>2</sub>. However, due to the similar behavior of O<sub>2</sub> and N<sub>2</sub> adsorption on samples IH2 and IH3, in this work, a gas mixture of

8% CO<sub>2</sub>/92% N<sub>2</sub> (or 8% CO<sub>2</sub>/90% N<sub>2</sub>/2% H<sub>2</sub>O in humid conditions) is to be evaluated in the dynamic study as representative of the incineration flue gas.



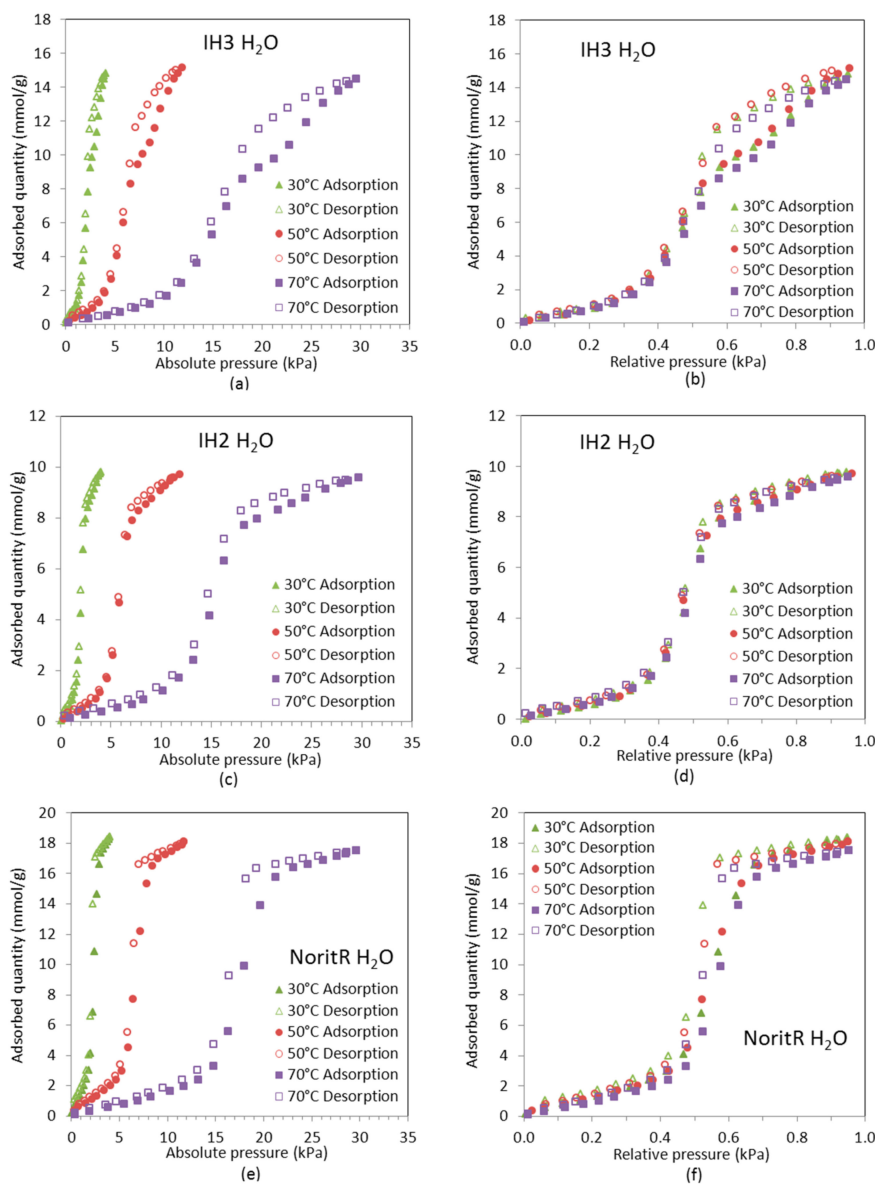
**Figure 2.** N<sub>2</sub> (a,b); CO<sub>2</sub> (c,d); and O<sub>2</sub> (e,f) adsorption isotherms at 30, 50, and 70 °C up to 120 kPa on samples IH2 and IH3.

### 2.3.2. Water Vapor Adsorption Isotherms

The adsorption isotherms of H<sub>2</sub>O at 30, 50, and 70 °C were measured up to the corresponding saturation pressure using a sorption analyzer (Hydrosorb 1000 HT, Quantachrome, Boynton Beach, FL, USA). Prior to the adsorption measurements, the sample was outgassed overnight under vacuum at 100 °C. During analysis, the temperature of the sample cell was controlled using a thermostatic circulating bath (Julabo, Seelbach, Germany). Figure 3 presents the water vapor adsorption isotherms for all the adsorbents tested.

Water vapor adsorption isotherms shown in Figure 3 present a Type V topology, according to the International Union of Pure and Applied Chemistry (IUPAC) classification, characterized by little water uptake at low relative pressures and the presence of a hysteresis loop over the majority of the

pressure range [27]. Adsorption data have been plotted in terms of absolute pressure (Figure 3a,c,e) and relative pressure (Figure 3b,d,f) for illustrative purposes. Temperature limits the pressure range for the adsorption of water vapor due to condensation. The adsorption branches of the isotherms at the different temperatures almost overlap in the entire relative pressure range. Water vapor adsorption in ultramicropores is considered practically independent of temperature [28].



**Figure 3.** Water vapor adsorption isotherms of the prepared carbons and the commercial Norit R2030 at 30, 50 and 70 °C up to the corresponding saturation pressure (full symbols correspond to the adsorption branch and empty symbols to the desorption branch): absolute pressure (a,c,e); and relative pressure (b,d,f).

For IH2, the hysteresis loop between the adsorption and desorption branches of the H<sub>2</sub>O isotherms is narrower than for IH3 and Norit R, being almost negligible at the lowest temperature evaluated, which is related to its narrower pore size and smaller pore volume.

In adsorption processes, it is common to define a separation factor as follows:



$$S_{1/2} = \frac{q_1/q_2}{p_1/p_2} \quad (1)$$

where  $q$  (mol/kg) is the uptake and  $p$  (kPa) is the partial pressure of each component. A preliminary estimation of the selectivity factor could be made based on the pure component equilibrium data for the  $\text{CO}_2/\text{N}_2$  and  $\text{H}_2\text{O}/\text{CO}_2$  separation under incineration flue gas conditions. The values calculated by means of Equation (1) are shown in Table 3.

**Table 3.** Equilibrium selectivity factors calculated for IH3, IH2 and Norit R.

Temperature	IH3		IH2		Norit R	
	$S_{\text{CO}_2/\text{N}_2}$	$S_{\text{H}_2\text{O}/\text{CO}_2}$	$S_{\text{CO}_2/\text{N}_2}$	$S_{\text{H}_2\text{O}/\text{CO}_2}$	$S_{\text{CO}_2/\text{N}_2}$	$S_{\text{H}_2\text{O}/\text{CO}_2}$
30 °C	18.31	52.97	18.78	53.19	16.90	89.60
50 °C	16.88	7.24	16.74	6.31	14.39	15.15
70 °C	14.84	4.47	13.03	4.55	13.70	6.98

According to the results in Table 3,  $\text{H}_2\text{O}$  is preferentially adsorbed on the carbons over  $\text{CO}_2$  that, at the same time, is much more adsorbed than  $\text{N}_2$ . Norit R shows the greatest selectivity to  $\text{H}_2\text{O}$  (nearly twice that of IH2 and IH3). Therefore, samples IH2 and IH3 seem more promising to separate  $\text{CO}_2$  under humid conditions. Moreover, the best  $\text{CO}_2/\text{N}_2$  separation corresponds to IH2 at 30 °C and to IH3 when the adsorption temperature is increased to 50 °C or 70 °C.

### 3. Dynamic Breakthrough Experiments

Breakthrough adsorption experiments of  $\text{CO}_2/\text{N}_2$  and  $\text{CO}_2/\text{N}_2/\text{H}_2\text{O}$  gas mixtures were conducted in a fixed-bed adsorption column. A purpose-built lab-scale rig was used; a scheme of the experimental set-up is presented in Figure 4. The adsorber (column) is a stainless steel reactor of 13 cm in height, 1.30 cm in diameter and is equipped with a porous plate located 4.7 cm from the base of the column that prevents the solids flow. A controlled evaporator mixer (CEM) coupled with a liquid delivery system (LDS) to feed liquid  $\text{H}_2\text{O}$  (LIQUI-FLOW type, from Bronkhorst High Tech, Ruurlo, The Netherlands) and two mass flow controllers (EL-FLOW type, from Bronkhorst High Tech) allow feeding a mixture of up to three components ( $\text{CO}_2$ ,  $\text{N}_2$ , and  $\text{H}_2\text{O}_{(\text{v})}$ ) with the desired flow rate and composition. This gas feeding system is purposely intended to conduct experiments with a precise control of humidity conditions. A K-type thermocouple was used to continuously monitor the column temperature with an accuracy of  $\pm 1.5$  °C. The temperature was controlled by coupling the heating element coiled around the reactor to an air-cooling device. The relative humidity of the gas mixture was monitored by means of a humidity and temperature sensor Hygroclip 2 (HT) from Rotronic (Bassersdorf, Switzerland). The bed pressure was controlled by means of a back-pressure regulator (BPR) located in the outlet pipe with a pressure range of 0–2.5 bar (WIKA, model A-10) in order to have great accuracy in the pressure range that is to be encountered in waste incineration conditions (accuracy  $\leq \pm 1\%$  of span). A vacuum system can be connected to the unit through a 3-way pneumatic valve. The vacuum line is composed of a vacuum controller coupled to a Pirani transmitter and an electromagnetic angle valve that is connected to a diaphragm pump. The mass flow rate of the effluent from the adsorbent bed was measured using a mini CORI-FLOW meter (MFM) from Bronkhorst (Ruurlo, The Netherlands). A dryer column filled with magnesium perchlorate is located after the mass flow meter, to remove water from the effluent prior to analysis by a dual channel micro-gas chromatograph (CP 4900 from Varian-Agilent Technologies, Santa Clara, CA, USA), fitted with a thermal conductivity detector (TCD) [29]. The TCD response is calibrated with  $\text{CO}_2/\text{N}_2$  mixtures of known composition. Data are recorded with a Supervisory Control and Data Acquisition (SCADA) system.

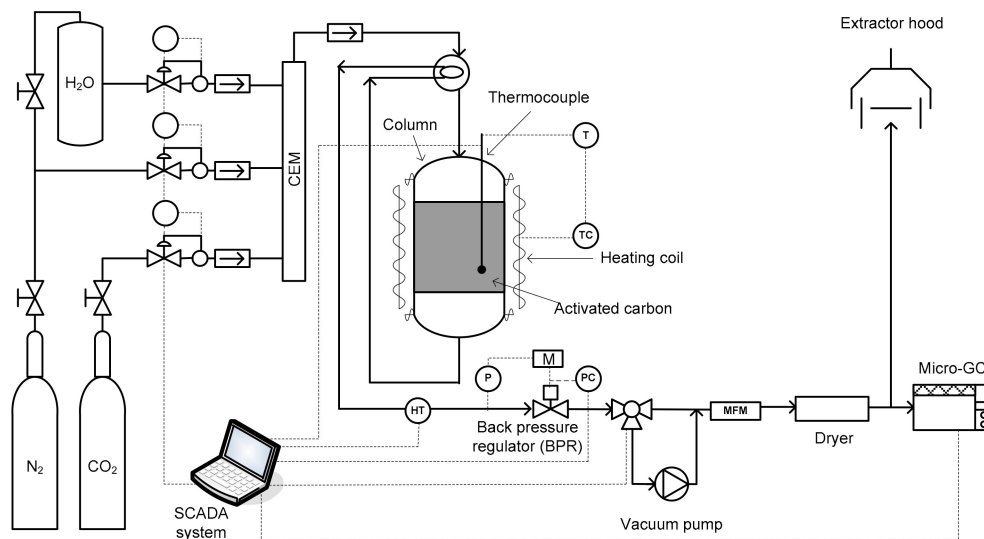


Figure 4. Experimental setup for the dynamic adsorption experiments.

### 3.1. Binary $\text{CO}_2/\text{N}_2$ Breakthrough Experiments

In order to evaluate the performance of the studied adsorbents to separate  $\text{CO}_2$  from  $\text{CO}_2/\text{N}_2$  gas streams, the fixed-bed adsorption unit was used to run breakthrough experiments with binary gas mixtures. Before each experiment the adsorbent was dried by flowing  $\text{N}_2$  ( $140 \text{ cm}^3/\text{min}$  for 60 min at  $150^\circ\text{C}$  and atmospheric pressure). The adsorbent was subjected to eight consecutive adsorption–desorption cycles to test the reproducibility of the system.

Each experimental run involved the following steps: (1) pressurization and cooling of the bed in a preconditioning step of 5 min, where  $140 \text{ cm}^3/\text{min}$  (STP) of  $\text{N}_2$  was allowed to flow through the system; (2) feed gas switch to a  $\text{CO}_2/\text{N}_2$  gas mixture to initiate the adsorption step for 30 min at the selected temperature; and (3) depressurization of the unit by lowering the pressure to sub-atmospheric and raising the temperature to  $70^\circ\text{C}$  for 25 min to fully desorb the adsorptive gases from the column. The minimum pressure of the vacuum line was set to  $0.1 \text{ kPa}$ . The outlet gas concentrations and the bed temperature profiles for the eight consecutive adsorption–desorption cycles are shown in Figure 5.

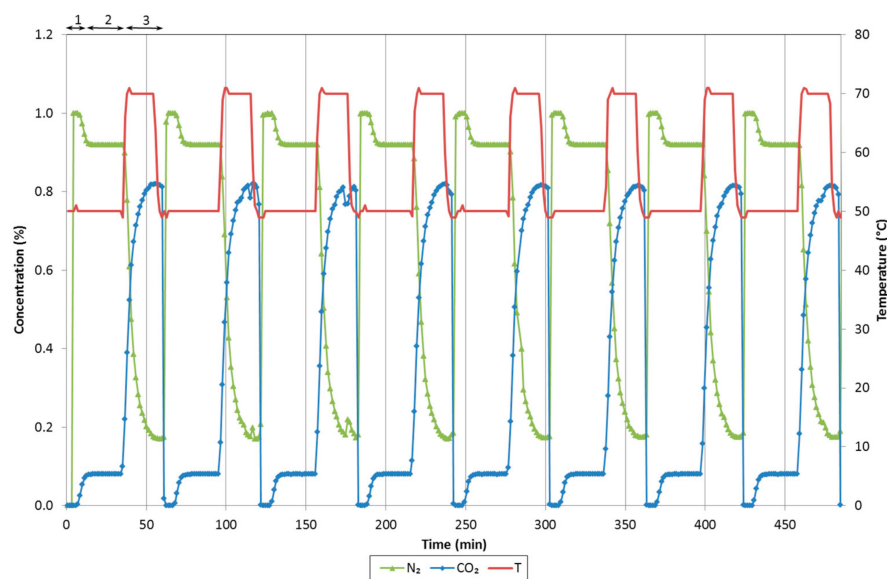


Figure 5. Example of eight adsorption–desorption cycles at  $50^\circ\text{C}$  for a mixture of 8/92 vol %  $\text{CO}_2/\text{N}_2$ .



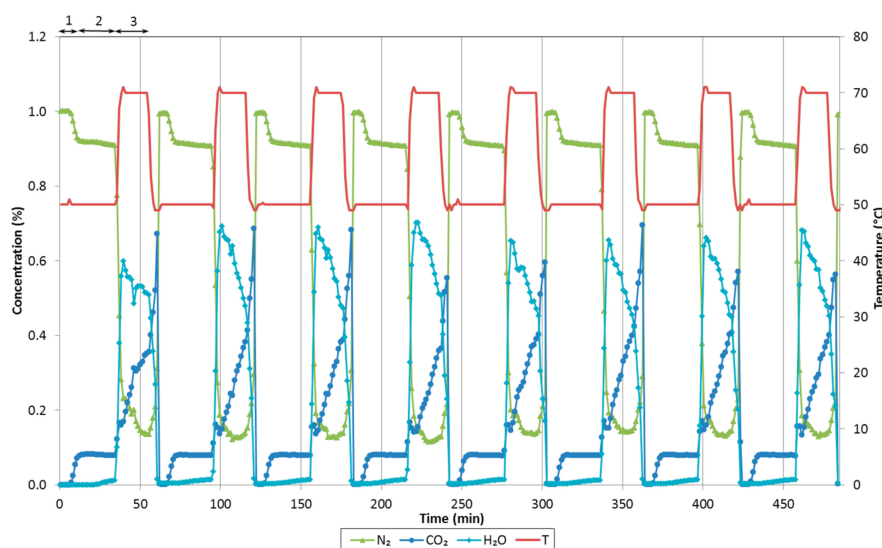
As can be observed from the concentration profile of CO<sub>2</sub> in Figure 5, CO<sub>2</sub> is not detected at the beginning of the adsorption step since it is adsorbed on the solids bed. As time goes by, the outlet CO<sub>2</sub> concentration increases until it approaches to feed concentration level, indicating that the column is saturated. During the evacuation step, the bed pressure is reduced, the temperature increased and the concentration of CO<sub>2</sub> leaving the adsorber reaches its maximum value.

The capture capacity was evaluated as an average of the adsorption performance of the adsorbents after consecutive adsorption–desorption cycles. A mass balance was applied to the bed in each adsorption–desorption cycle to calculate the specific amount of CO<sub>2</sub> adsorbed at equilibrium and the holdup of the gas phase was discounted applying correction factors. Detailed information about the calculation procedure can be found in [30]. The breakthrough times (time required for the CO<sub>2</sub> concentration front to reach the bed outlet) were taken at a relative concentration ( $C_{\text{CO}_2,\text{outlet}}/C_{\text{CO}_2,\text{feed}}$ ) of 0.05. The total amount of H<sub>2</sub>O adsorbed at the time when the saturation is reached and the breakthrough times of water vapor are estimated from the experimental breakthrough curves in the same way as CO<sub>2</sub>.

### 3.2. Ternary CO<sub>2</sub>/N<sub>2</sub>/H<sub>2</sub>O Breakthrough Experiments

Dynamic experiments were carried out in the fixed-bed adsorption unit with mixtures of N<sub>2</sub>, CO<sub>2</sub>, and H<sub>2</sub>O<sub>(v)</sub> at 50 °C in order to assess the influence of water vapor in CO<sub>2</sub> adsorption. The adsorbent was subjected to eight consecutive adsorption–desorption cycles to test the reproducibility of the system. The experimental run included the same steps as in the dry experiment. The sole difference is that during the adsorption step (second step) a ternary mixture (CO<sub>2</sub>/N<sub>2</sub>/H<sub>2</sub>O) is fed to the adsorber.

During cyclic operation in an industrial process, regeneration is expected to be interrupted before completion. Therefore, the adsorbent is not fully regenerated at the beginning of each cycle. The overlapping of experimental data in Figure 6 confirms that the cyclic steady state (the molar flow rates at any time in the cycle remain the same in all subsequent cycles) is reached at the second cycle for CO<sub>2</sub> and around the fourth cycle for water vapor.



**Figure 6.** Example of eight adsorption–desorption cycles at 50 °C for a mixture of 8/90/2 vol % CO<sub>2</sub>/N<sub>2</sub>/H<sub>2</sub>O.

In addition, to evaluate the effect of CO<sub>2</sub> on H<sub>2</sub>O adsorption at different temperatures (30 and 50 °C), the ternary N<sub>2</sub>/CO<sub>2</sub>/H<sub>2</sub>O breakthrough curve is also compared to the binary N<sub>2</sub>/H<sub>2</sub>O breakthrough curve obtained in the absence of CO<sub>2</sub>. Moreover, a ternary N<sub>2</sub>/CO<sub>2</sub>/H<sub>2</sub>O breakthrough curve was obtained departing from a bed initially saturated with H<sub>2</sub>O and N<sub>2</sub> to evaluate CO<sub>2</sub>

adsorption under less favorable conditions. In all cases, the feed flowrate was set to  $143 \text{ cm}^3/\text{min}$ , the total pressure was 140 kPa and the adsorbent was initially regenerated and in equilibrium with  $\text{N}_2$  at the temperature of the experiment.

The cumulative amount of  $\text{CO}_2$  and  $\text{H}_2\text{O}$  adsorbed calculated from the experimental data ( $q_{\text{CO}_2}$  and  $q_{\text{H}_2\text{O}}$ ) were compared to the pure component adsorption isotherm ( $q_{\text{CO}_2}^*$  and  $q_{\text{H}_2\text{O}}^*$ ), respectively.

#### 4. Results and Discussion

Dynamic breakthrough experiments to evaluate the performance of the pine sawdust based adsorbents to separate  $\text{CO}_2$  from gas mixtures representative of incineration flue gas were carried out at different temperatures in the range  $30\text{--}70^\circ\text{C}$ . The characteristics of the beds of adsorbent are presented in Table 4. In all experiments, the height of the solids bed was kept approximately constant and so the mass of adsorbent varied according to the bed density. It can be observed in Table 4 that the commercial carbon, Norit R, presents the densest bed.

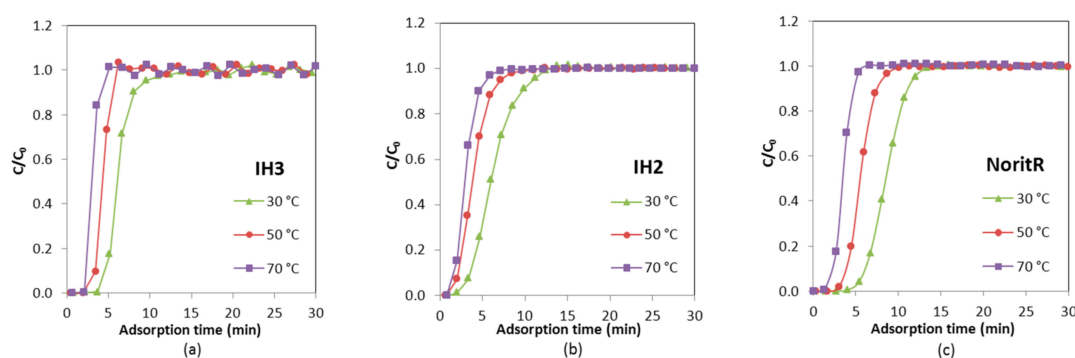
**Table 4.** Bed characteristics for the breakthrough experiments.

Bed Characteristics	IH3	IH2	Norit R
Mass of adsorbent (g)	4.482	5.186	7.888
Total porosity ( $\epsilon_T$ )	0.85	0.83	0.77
Helium density ( $\text{g}/\text{cm}^3$ ) <sup>a</sup>	1.916	1.897	2.140
Apparent density ( $\text{g}/\text{cm}^3$ ) <sup>b</sup>	0.611	0.820	0.845
Bed diameter (cm)	1.30	1.30	1.30
Bed height (cm)	12.0	12.0	12.1
Bed density ( $\text{g}/\text{cm}^3$ )	0.281	0.326	0.491

<sup>a</sup> The helium density was measured in an Accupyc 1330 apparatus at  $35^\circ\text{C}$ . <sup>b</sup> The apparent density was determined at 0.1 MPa in a Micromeritics Autopore IV 9500 mercury porosimeter.

##### 4.1. Binary $\text{CO}_2/\text{N}_2$ Breakthrough Experiments

Binary breakthrough adsorption experiments were performed at three temperatures ( $30$ ,  $50$  and  $70^\circ\text{C}$ ) feeding a gas mixture with 8% of  $\text{CO}_2$  (balance  $\text{N}_2$ ). In all cases, the feed flowrate was set to  $140 \text{ cm}^3/\text{min}$ , the total pressure was 140 kPa and the adsorbent was initially regenerated and in equilibrium with  $\text{N}_2$  at the experimental temperature. The results are shown in Figure 7, where  $C/C_0$  represents the ratio between the concentration leaving the adsorber at a given time ( $C$ ) and that in the feed ( $C_0$ ).



**Figure 7.** Breakthrough curves of  $\text{CO}_2$  at different temperatures ( $30$ ,  $50$  and  $70^\circ\text{C}$ ) for samples: (a) IH3; (b) IH2; and (c) Norit R.

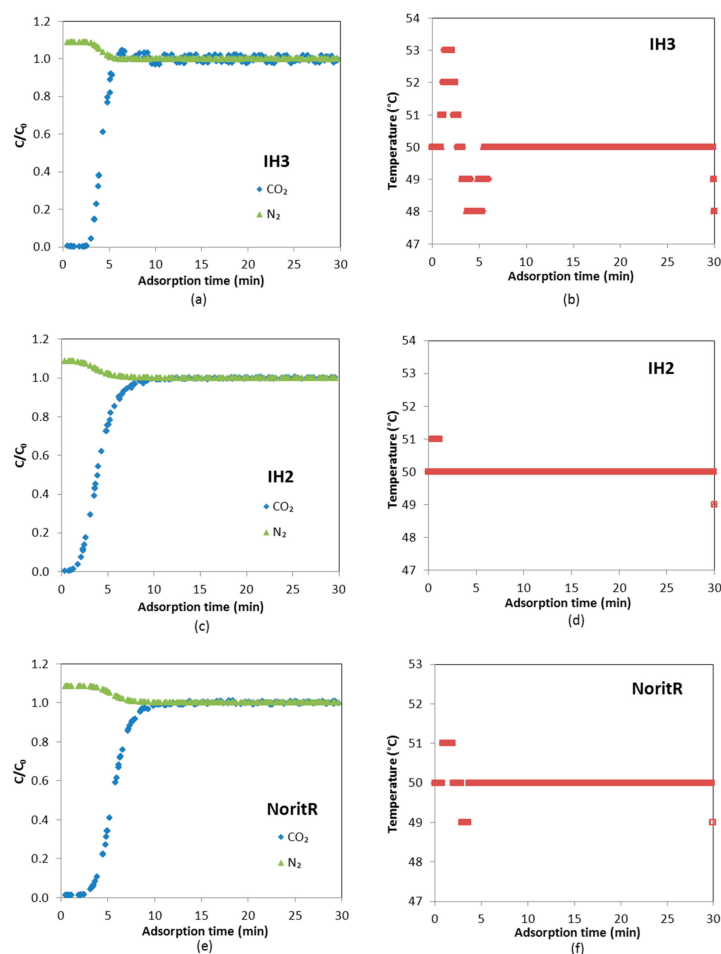
Breakthrough experiments showed steep breakthrough curves, presenting IH3 a slightly higher slope than IH2 and Norit R. This confirms that the overall mass transfer resistance for the adsorbents is small, which is highly desirable for a rapid swing adsorption process.

The obtained breakthrough curves are translated into dynamic uptake capacity by integration, with longer breakthrough times yielding higher uptake capacities. Results are shown in Table 5. As expected, CO<sub>2</sub> uptake decreases with temperature and the performance on a mass basis of IH3 is superior to that of IH2 which, in turn, is better than Norit R. Nevertheless, it can be observed that Norit R presents greater breakthrough times.

**Table 5.** CO<sub>2</sub> adsorption capacities and breakthrough times from the experiments at 30, 50 and 70 °C and 140 kPa, feeding binary CO<sub>2</sub>/N<sub>2</sub> gas mixture.

Temperature (°C)	IH3		IH2		Norit R	
	CO <sub>2</sub> uptake (mmol/g ads)	<i>t<sub>b</sub></i> ,CO <sub>2</sub> (min)	CO <sub>2</sub> uptake (mmol/g ads)	<i>t<sub>b</sub></i> ,CO <sub>2</sub> (min)	CO <sub>2</sub> uptake (mmol/g ads)	<i>t<sub>b</sub></i> ,CO <sub>2</sub> (min)
30	0.67	4.12	0.56	2.78	0.51	5.46
50	0.39	2.42	0.35	1.55	0.33	3.30
70	0.31	2.16	0.26	1.18	0.22	1.94

Figure 8 represents consecutive breakthrough curves carried out with a feed gas consisting of 8% CO<sub>2</sub> (balance N<sub>2</sub>), at 140 kPa and 50 °C. Blue symbols correspond to CO<sub>2</sub>, whereas green ones represent N<sub>2</sub>. Symbols overlap for all the consecutive curves, indicating a good repeatability and reproducibility of the adsorption–desorption cycles.



**Figure 8.** Comparison of N<sub>2</sub> and CO<sub>2</sub> concentration profiles during the binary experiments with a CO<sub>2</sub>/N<sub>2</sub> mixture (8/92 vol %) for: (a) IH3; (c) IH2; and (e) Norit R at 50 °C for the consecutive cycles; and thermal wave observed during the adsorption step for samples: (b) IH3; (d) IH2; and (f) Norit R.

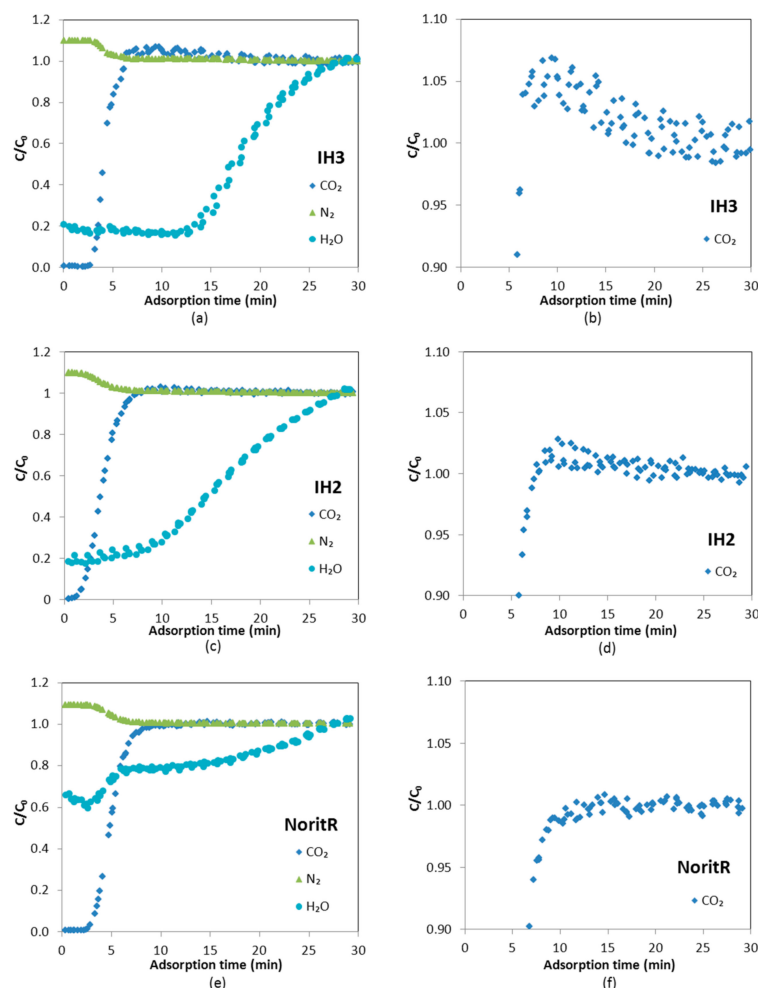
The thermal wave observed as a consequence of CO<sub>2</sub> adsorption is shown in Figure 8b,d,f. It is larger for IH3 due to the greater amount of CO<sub>2</sub> adsorbed in a slightly shorter time when compared to Norit R.

#### 4.2. Ternary CO<sub>2</sub>/N<sub>2</sub>/H<sub>2</sub>O Breakthrough Experiments

The behavior of the adsorbents under humid conditions was assessed by different experiments and compared to the previous results with dry flue gas mixtures.

##### 4.2.1. Cyclic Operation

The CO<sub>2</sub>, N<sub>2</sub> and H<sub>2</sub>O breakthrough curves were measured for the activated carbons under study and  $C/C_0$  (ratio between the outlet CO<sub>2</sub>, N<sub>2</sub> or H<sub>2</sub>O concentration at a given time and that in the feed) was plotted versus time. In Figure 9 the last five consecutive cycles (Cycles 4–8), in which concentration of both CO<sub>2</sub> and H<sub>2</sub>O can be considered stable, are represented. As far as CO<sub>2</sub> is concerned, the adsorbents reached saturation (maximum adsorption capacity of the adsorbed components) during the adsorption step and they were practically fully regenerated during the desorption step. However, water adsorption and desorption was slower. Therefore, it is observed in Figure 9 that  $C/C_0$  for H<sub>2</sub>O at  $t = 0$  is around 0.2 for IH3 and IH2 and 0.6 for Norit R.



**Figure 9.** Comparison of N<sub>2</sub>, CO<sub>2</sub> and H<sub>2</sub>O concentration profiles during the ternary experiments with a CO<sub>2</sub>/N<sub>2</sub>/H<sub>2</sub>O mixture (8/90/2 vol %) for: (a) IH3; (c) IH2; and (e) Norit R at 50 °C for the fourth to eight consecutive cycles; and detail of the CO<sub>2</sub> breakthrough curves over adsorbents: (b) IH3; (d) IH2; and (f) Norit R.

Both CO<sub>2</sub> and water vapor breakthrough curves are steeper for IH3 than for IH2, showing that the latter has greater mass transfer resistance due to diffusion based constraints associated to the less developed porosity. Norit R presents similar diffusional resistance to IH2 for CO<sub>2</sub>; however, the breakthrough times are bigger.

It is noteworthy that kinetics of H<sub>2</sub>O adsorption/desorption on Norit R are particularly slow. At 50 °C and a working partial pressure of H<sub>2</sub>O around 2.4 kPa, the adsorption capacity of Norit R, as expected from the equilibrium adsorption isotherm, is about 1.20 mmol/g. Thus, it is larger than for the other two samples: 0.79 mmol/g for IH3 and 0.56 mmol/g for IH2. Nevertheless, considerable differences can be observed between partial pressures of H<sub>2</sub>O at the beginning and the end of the adsorption stage under cyclic operation. The initial and final partial H<sub>2</sub>O pressures are 0.44 and 2.21 kPa for IH3 and 0.35 and 1.75 kPa for IH2, whereas the values obtained for Norit R are 0.56 and 0.79 kPa, respectively. This is most probably the reason why breakthrough curves of CO<sub>2</sub> for Norit R show significant differences between the dry and wet cyclic operation and, therefore, there is a greater reduction in breakthrough times.

In addition, it is found that the shape of the water breakthrough curve of Norit R under cyclic operation is very different in comparison to biomass adsorbents IH2 and IH3. The experiment with Norit R reached low water vapor partial pressure meaning that, at the end of the adsorption step, working pressure conditions correspond to the first region of the water vapor adsorption isotherm before the steep H<sub>2</sub>O uptake. Table 6 summarizes the average CO<sub>2</sub> uptakes and breakthrough times attained during the wet cyclic breakthrough experiments. Both IH3 and IH2 show enhanced performance in terms of CO<sub>2</sub> adsorption under humid conditions when compared to Norit R.

**Table 6.** Average CO<sub>2</sub> uptakes and breakthrough times calculated from the breakthrough adsorption cycles feeding ternary CO<sub>2</sub>/H<sub>2</sub>O/N<sub>2</sub> gas mixtures.

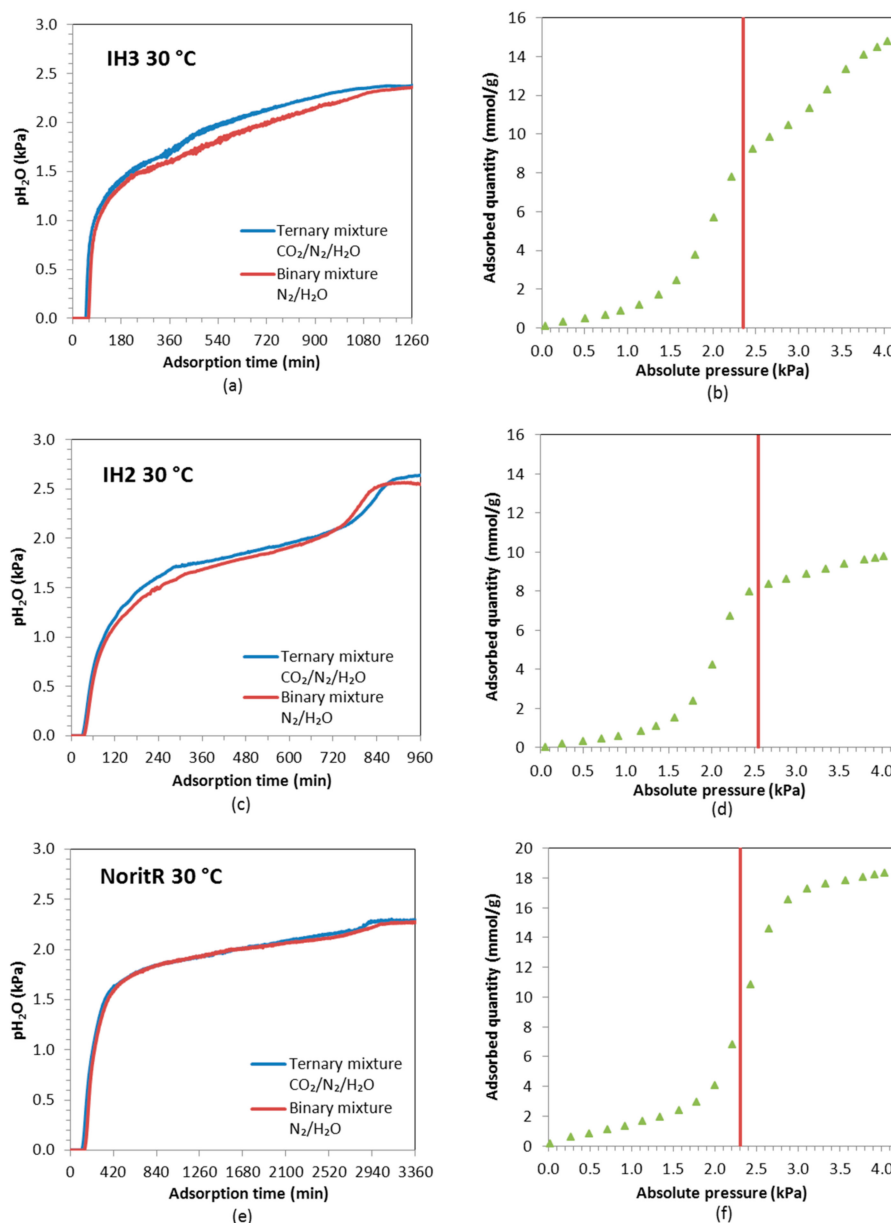
Mixture	IH3		IH2		Norit R	
	CO <sub>2</sub> Uptake (mmol/g)	<i>t<sub>b</sub></i> ,CO <sub>2</sub> (min)	CO <sub>2</sub> Uptake (mmol/g)	<i>t<sub>b</sub></i> ,CO <sub>2</sub> (min)	CO <sub>2</sub> Uptake (mmol/g)	<i>t<sub>b</sub></i> ,CO <sub>2</sub> (min)
CO <sub>2</sub> /N <sub>2</sub>	0.42	2.70	0.35	1.65	0.31	3.13
CO <sub>2</sub> /H <sub>2</sub> O/N <sub>2</sub>	0.38	2.61	0.32	1.57	0.29	2.87

A small roll-up phenomenon can be observed in the detail of the CO<sub>2</sub> curve in Figure 9: as the front of H<sub>2</sub>O travels through the bed, part of the CO<sub>2</sub> initially adsorbed is displaced by the adsorption of H<sub>2</sub>O. However, this is not observed for Norit R because the quantity adsorbed of H<sub>2</sub>O in each cycle is small given that the adsorption time is limited to 30 min.

#### 4.2.2. Non-Cyclic Operation

Ternary breakthrough curves were carried out using mixtures of N<sub>2</sub>, CO<sub>2</sub> and H<sub>2</sub>O, and these were compared to binary breakthrough curves carried out with N<sub>2</sub> and H<sub>2</sub>O mixtures to assess the influence of CO<sub>2</sub> on H<sub>2</sub>O adsorption departing from a fully regenerated adsorbent. Moreover, the influence of H<sub>2</sub>O on CO<sub>2</sub> adsorption was studied comparing the behavior of a ternary N<sub>2</sub>/H<sub>2</sub>O/CO<sub>2</sub> mixture fed to a regenerated adsorbent and to an adsorbent initially saturated with H<sub>2</sub>O. To evaluate the dynamic adsorption at different relative humidity (RH), experiments were performed at 30 °C (RH ≈ 60%) and 50 °C (RH ≈ 22%).

Dynamic adsorption experiments with gas mixtures at high relative humidity were carried out. A breakthrough curve of water vapor for a binary mixture of 98% N<sub>2</sub> and 2% H<sub>2</sub>O obtained at 30 °C is compared with the breakthrough curve for a ternary mixture of 8% CO<sub>2</sub>, 90 % N<sub>2</sub> and 2% H<sub>2</sub>O. As can be seen in Figure 10, the shape of the breakthrough curve of H<sub>2</sub>O for a binary H<sub>2</sub>O/N<sub>2</sub> mixture is similar to that of the ternary N<sub>2</sub>/CO<sub>2</sub>/H<sub>2</sub>O mixture fed to a fresh bed. These results indicate that H<sub>2</sub>O adsorption is little affected by CO<sub>2</sub> adsorption.



**Figure 10.** Water vapor breakthrough curves at 30 °C for a ternary  $N_2/CO_2/H_2O$  mixture (blue) and a binary  $H_2O/N_2$  one (red) on: (a) IH3; (c) IH2; and (e) Norit R. Water vapor adsorption isotherms at 30 °C for: (b) IH3; (d) IH2; and (f) Norit R. The red vertical line indicates the partial pressure of  $H_2O$  reached at the end of the ternary  $N_2/CO_2/H_2O$  and binary  $N_2/H_2O$  experiments.

Comparing the  $H_2O$  breakthrough curves for IH3, IH2 and Norit R, it can be observed that they have slightly different shapes. This can be directly ascribed to the shape of the equilibrium adsorption isotherms. The  $H_2O$  breakthrough curve of IH2 presents two uptake zones with an intermediate plateau, while, in the curves of IH3 and Norit R, the second uptake is not observed because, at the working partial pressure of water vapor ( $\approx 2.4$  and  $2.3$  kPa, respectively), a steep rise in the  $H_2O$  uptake is to take place according to the corresponding adsorption isotherms in Figure 10.

The experimental water vapor partial pressure at saturation is slightly lower for IH3 experiments than for IH2 ( $2.35$ – $2.4$  kPa vs.  $2.55$ – $2.6$  kPa), thus, according to the pure component adsorption isotherms of  $H_2O$ , both samples are expected to adsorb approximately the same amount of water ( $\approx 8$  mmol/g). However, water vapor adsorption on carbon IH3 is somewhat smaller than on IH2.



To the authors' knowledge this could only be due to the cumulative error associated with the integration of the experimental data over longer times in the experiments with IH3 (1260 min for IH3 vs. 920 min for IH2). Norit R presents greater water vapor adsorption capacity and slower kinetics than the pine sawdust carbon pellets, hence water vapor breakthrough is postponed and the saturation point is reached after more than 3200 min from the beginning of the water vapor feeding.

Table 7 summarizes the water adsorption uptake and breakthrough times for experiments at 30 °C ( $RH \approx 60\%$ ) and 50 °C ( $RH \approx 22\%$ ). It is worth noting that the total feed flow rate and its content in  $H_2O$  was the same for both experiments. Norit R presents the highest  $H_2O$  adsorption capacity and much longer breakthrough times both at low and at high relative humidity conditions. On the contrary, IH2 shows shorter breakthrough times and lower water vapor uptake at a relative humidity of 22%, whereas IH3 adsorbs less water vapor under high humidity conditions. The  $H_2O$  uptakes calculated for the ternary and binary mixtures at each temperature confirm the little influence of  $CO_2$  on the  $H_2O$  adsorption capacity.

**Table 7.** Adsorption capacities of water vapor and breakthrough times during non-cyclic experiments at low and high relative humidity.

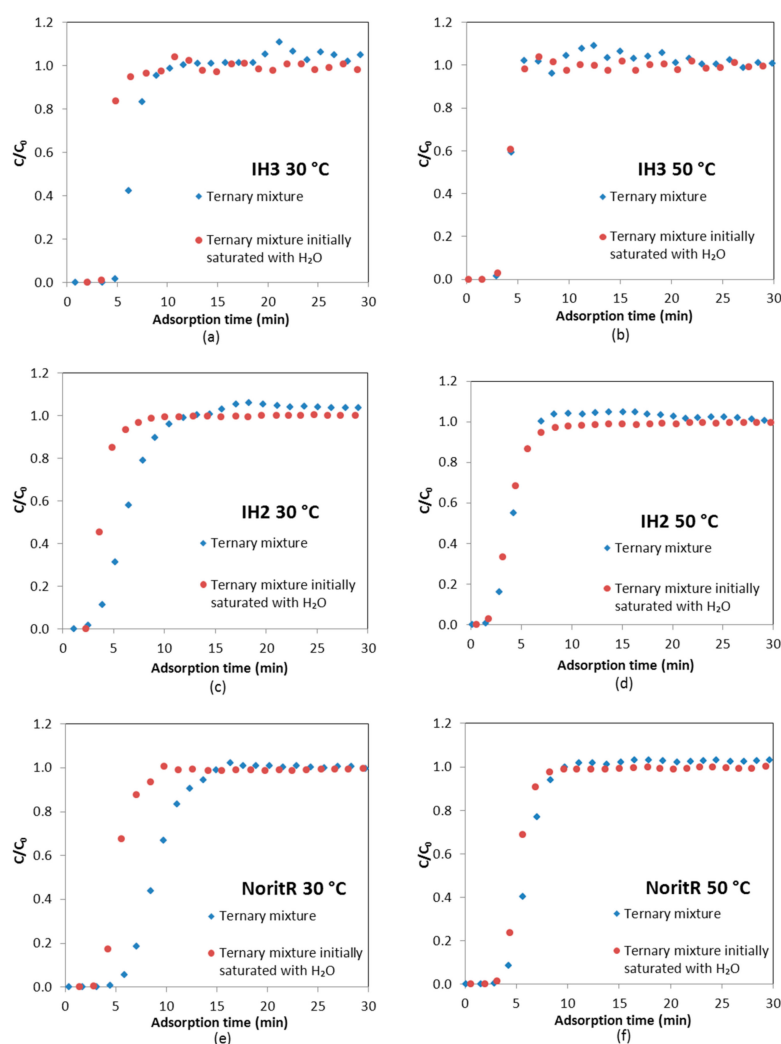
Relative Humidity	Experiment	Sample	$q_{H_2O}$ (mmol/g ads)	$t_{b,H_2O}$ (min)	$\frac{q_{H_2O}-q_{H_2O}^*}{q_{H_2O}^*} \times 100$ (%)
60%	Ternary mixture $N_2/CO_2/H_2O$	IH3	6.48	50.1	−25.8
		IH2	7.39	37.7	−10.6
		Norit R	7.70	131.6	+0.1
	Binary mixture $N_2/H_2O$	IH3	6.91	60.4	−19.8
		IH2	7.41	42.5	−9.3
		Norit R	7.75	156.4	−1.9
22%	Ternary mixture $N_2/CO_2/H_2O$	IH3	0.92	20.8	+4.4
		IH2	0.76	16.0	+16.2
		Norit R	1.37	64.2	+7.0
	Binary mixture $N_2/H_2O$	IH3	0.95	24.2	+9.6
		IH2	0.90	20.4	+39.1
		Norit R	1.52	70.3	+14.0

Under high relative humidity condition the evaluated carbons tend to lose  $H_2O$  adsorption capacity when compared to the uptake estimated from the  $H_2O$  equilibrium adsorption isotherms. On the contrary, low  $RH$  conditions seem to enhance the  $H_2O$  uptake during the breakthrough testing, although this could be due to more uncertainty on the equilibrium data for  $H_2O$  adsorption at these low partial pressures.

In addition, to further assess the influence that water holdup can represent to  $CO_2$  adsorption, a ternary mixture with 90%  $N_2$ , 8%  $CO_2$  and 2%  $H_2O$  was fed to the adsorbent bed previously saturated with  $H_2O$ . The obtained  $CO_2$  breakthrough curve is compared to that from the ternary mixture departing from a fully regenerated adsorbent bed in Figure 11.

The  $CO_2$  concentration profiles of the adsorbents under the two evaluated conditions substantially differ and point out the influence of humidity conditions of the adsorbent bed on its  $CO_2$  capture performance. However, the slope of the  $CO_2$  breakthrough curves might suggest that mass transfer would be favored when the bed is initially saturated with  $H_2O$ . A roll up effect ( $C/C_0 > 1$ ) that indicates the displacement of  $CO_2$  by the adsorption of  $H_2O$  can be observed in Figure 11 in the case of IH2 and IH3; however, Norit R requires a longer time to discern it (around 90 min) due to its higher adsorption capacity for  $H_2O$ .

Table 8 reports the  $CO_2$  uptakes and breakthrough times estimated from the ternary breakthrough experiments with initially fully wet and fresh beds. Experimental breakthrough results ( $q_{CO_2}$ ) are also compared to the data from the pure component adsorption isotherms ( $q_{CO_2}^*$ ).



**Figure 11.** CO<sub>2</sub> breakthrough curves for ternary N<sub>2</sub>/CO<sub>2</sub>/H<sub>2</sub>O mixtures over a fully regenerated adsorbent bed (blue) and ternary mixtures fed to the adsorbent initially saturated with H<sub>2</sub>O (red) at 30 °C (a,c,e) and 50 °C (b,d,f).

**Table 8.** Results of CO<sub>2</sub> uptake capacity and breakthrough times for the ternary breakthrough experiments at low and high relative humidity.

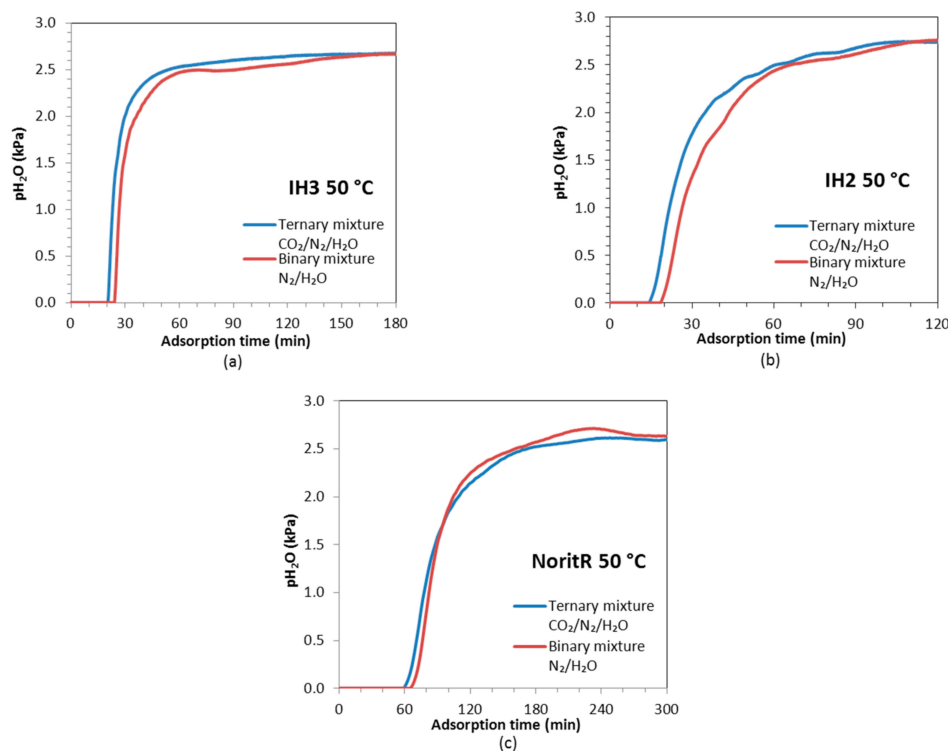
RH	Experiment	Sample	$q_{\text{CO}_2}$ (mmol/g ads)	$t_{b,\text{CO}_2}$ (min)	$\frac{q_{\text{CO}_2} - q_{\text{CO}_2}^*}{q_{\text{CO}_2}^*} \times 100$ (%)
60%	Ternary mixture N <sub>2</sub> /CO <sub>2</sub> /H <sub>2</sub> O	IH3	0.72	4.90	−8.5
		IH2	0.59	3.03	−15.3
		Norit R	0.54	5.68	−10.3
	Ternary mixture initially saturated with H <sub>2</sub> O	IH3	0.29	2.09	−62.2
		IH2	0.22	1.02	−67.3
		Norit R	0.30	3.21	−49.1
22%	Ternary mixture N <sub>2</sub> /CO <sub>2</sub> /H <sub>2</sub> O	IH3	0.40	3.10	−15.0
		IH2	0.33	1.86	−20.3
		Norit R	0.36	3.67	−2.7
	Ternary mixture initially saturated with H <sub>2</sub> O	IH3	0.41	3.08	−12.1
		IH2	0.35	1.84	−15.2
		Norit R	0.30	3.34	−17.3

The dynamic experiments carried out using ternary mixtures of  $\text{CO}_2$ ,  $\text{N}_2$  and  $\text{H}_2\text{O}$  over a fully regenerated adsorbent bed (free of  $\text{CO}_2$  and  $\text{H}_2\text{O}$  at  $t = 0$ ) show no reduction in the breakthrough time for  $\text{CO}_2$ , compared to the dry cases.

The  $\text{CO}_2$  adsorption capacities and breakthrough times are similar for both ternary experiments (bed initially saturated with water vapor and fully regenerated) at  $50^\circ\text{C}$  and to the values previously estimated for the binary mixture experiments conducted at  $50^\circ\text{C}$ . A maximum reduction of  $\approx 20\%$  when compared to the  $\text{CO}_2$  uptake estimated from the equilibrium  $\text{CO}_2$  adsorption isotherms is observed. However, at  $30^\circ\text{C}$   $\text{CO}_2$  adsorption capacity decreases to a great extent (a minimum reduction of  $\approx 50\%$  from the equilibrium  $\text{CO}_2$  adsorption isotherms) by the co-adsorption of  $\text{H}_2\text{O}$  when feeding a humid gas stream to a bed saturated with water vapor. Therefore, the breakthrough time of  $\text{CO}_2$  is also reduced. This performance accounts for the equilibrium adsorption capacity for pure water vapor at the relative humidity of the feed (e.g.,  $RH \approx 60\%$ ) that is more than 12 times greater than the adsorption capacity for pure  $\text{CO}_2$  at the partial pressure of  $\text{CO}_2$  in the feed.

Under high relative humidity conditions, which represent the worst case scenario for the adsorption of  $\text{CO}_2$ , the capacity to adsorb  $\text{CO}_2$  on Norit R seems less affected than that of IH2 and IH3 when the  $\text{CO}_2$  mixture is fed to a saturated bed. However, it can be noted that under all the evaluated conditions of relative humidity and initial conditions of the bed, IH3 generally shows higher  $\text{CO}_2$  uptake on a mass basis than the other two adsorbents.

The breakthrough curves of  $\text{H}_2\text{O}$  were also tested at low relative humidity ( $RH \approx 22\%$ ) conditions showing a single uptake zone with a steep breakthrough (see Figure 12). When the relative humidity of the feed is low, the working pressure of water vapor is located within the first region of the adsorption isotherm before the steep uptake takes place. The shape of the  $\text{H}_2\text{O}$  curves is similar for both the binary  $\text{N}_2/\text{H}_2\text{O}$  and ternary  $\text{N}_2/\text{CO}_2/\text{H}_2\text{O}$  experiments; however, the  $\text{H}_2\text{O}$  breakthrough is slightly anticipated for the ternary mixture due to the competitive co-adsorption of  $\text{CO}_2$ .



**Figure 12.** Evolution of the partial pressure of water vapor at  $50^\circ\text{C}$  over time up to saturation for a ternary  $\text{N}_2/\text{CO}_2/\text{H}_2\text{O}$  mixture (blue) and a binary  $\text{H}_2\text{O}/\text{N}_2$  one (red) for samples: (a) IH3; (b) IH2; and (c) Norit R.

The steeper slopes of the H<sub>2</sub>O breakthrough curves in the IH3 bed indicate less diffusional resistance for adsorption than in the other two adsorbent beds. On activated carbon IH2 the saturation point is reached after approximately 110 min, IH3 needs around 150 min and Norit R up to 240 min. In addition, the three samples adsorbed a little bit more of water vapor in the binary mixture experiment than in the ternary one, due to the absence of CO<sub>2</sub> adsorption competition.

At 50 °C, the relative humidity of the feed is lower ( $RH \approx 22\%$ ) so the equilibrium adsorption capacity for pure H<sub>2</sub>O is only 1.5 times greater than that for pure CO<sub>2</sub> at the partial pressure of CO<sub>2</sub> in the feed. Therefore, CO<sub>2</sub> adsorption is barely affected by H<sub>2</sub>O adsorption in low humidity conditions. In fact, as could be observed in Figure 11, the breakthrough curves of CO<sub>2</sub> over a fresh bed and over a bed saturated with H<sub>2</sub>O almost overlap for all the samples in the experiments at 50 °C. CO<sub>2</sub> breaks first through the beds of IH2 and IH3 than through Norit R; however, IH3 reaches equilibrium first meaning that adsorption kinetics is faster on this sample.

The results presented herein are in good agreement with those from a previous work [21] carried out with a biomass based carbon adsorbent evaluated under post-combustion capture conditions (14% CO<sub>2</sub> at ambient temperature). A reduction in the CO<sub>2</sub> adsorption capacity up to 64% was found for a relative humidity in the gas feed of 95%.

## 5. Conclusions

Three activated carbons were examined: two were prepared departing from pine sawdust and the other is a commercial activated carbon supplied by Norit. Their behavior was evaluated for CO<sub>2</sub> capture in a solid waste incineration facility. To the best of the authors' knowledge, no previous studies on CO<sub>2</sub> capture from waste incineration plants by means of adsorption have been reported.

Adsorption equilibrium was firstly studied for the components present in the flue gas stream, including water vapor, at different temperatures and up to 120 kPa. Dynamic adsorption–desorption studies in a fixed-bed system under dry and wet conditions revealed the importance of a multicomponent evaluation to assess the performance of adsorbents for a particular application. The maximum adsorption capacity for CO<sub>2</sub> is reached in dry conditions, whereas the minimum corresponds to the case when CO<sub>2</sub> is fed to a bed saturated with H<sub>2</sub>O at 30 °C. Therefore, the CO<sub>2</sub> adsorption capacity is proven to be reduced down to almost 54% by the co-adsorption of H<sub>2</sub>O at high relative humidity conditions when compared to the dry experiments. On the other hand, it was shown that H<sub>2</sub>O is little affected by CO<sub>2</sub> adsorption. In addition, it has been demonstrated that low humidity conditions do not affect the CO<sub>2</sub> capture performance of the adsorbents.

It is, thereupon, necessary to consider the effect of water vapor for a CO<sub>2</sub> capture process design. In swing adsorption cycles, the step times are in the order of seconds up to a few minutes. Having in mind the slow kinetics of water vapor adsorption, with a tailored cycle configuration the moisture loading can be kept low. Under cyclic operation in humid conditions, IH3 showed the highest CO<sub>2</sub> uptake per gram of adsorbent. Even though Norit R seems promising due to its larger breakthrough time, it presents the highest H<sub>2</sub>O loading at short adsorption times which is detrimental for the CO<sub>2</sub> separation.

All adsorbents evaluated showed good performance in terms of CO<sub>2</sub> adsorption capacity under conditions representative of incineration flue gas; however, IH3 outperformed in adsorption kinetics and CO<sub>2</sub> capture capacity on a mass basis.

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**Author Contributions:** Inés Durán conducted experimental work and data analysis. Inés Durán, Fernando Rubiera and Covadonga Pevida discussed the results and contributed to the writing of the manuscript.

**Conflicts of Interest:** The authors declare no conflicts of interest.

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