

Supporting Information

Simultaneous Catalysis of Sulfite Oxidation and Uptake of Heavy Metals by Bifunctional Activated Carbon Fiber in Magnesia Desulfurization

Figure S1 XPS spectrum of 40% Co-ACFs

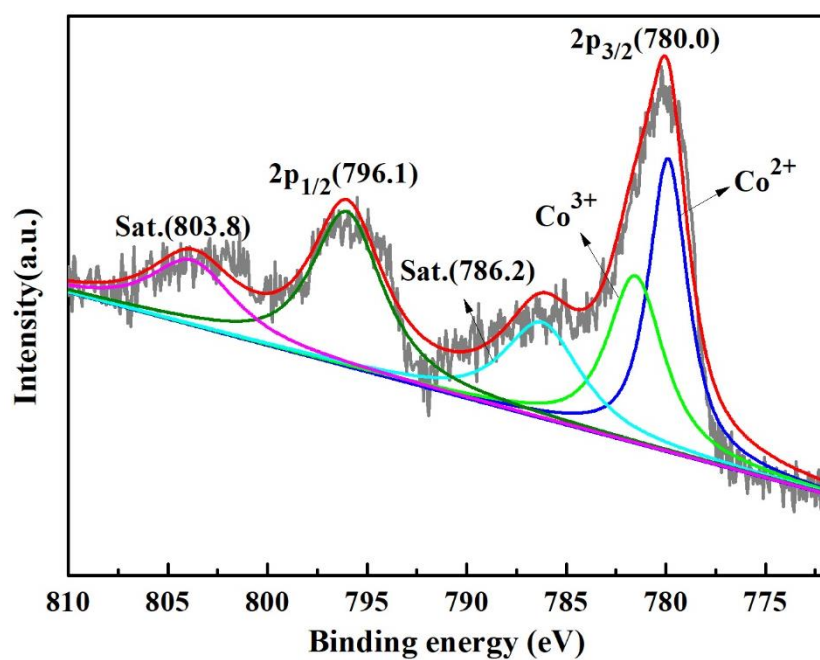


Figure S2 Catalysis performance of Co-ACFs after four cycles

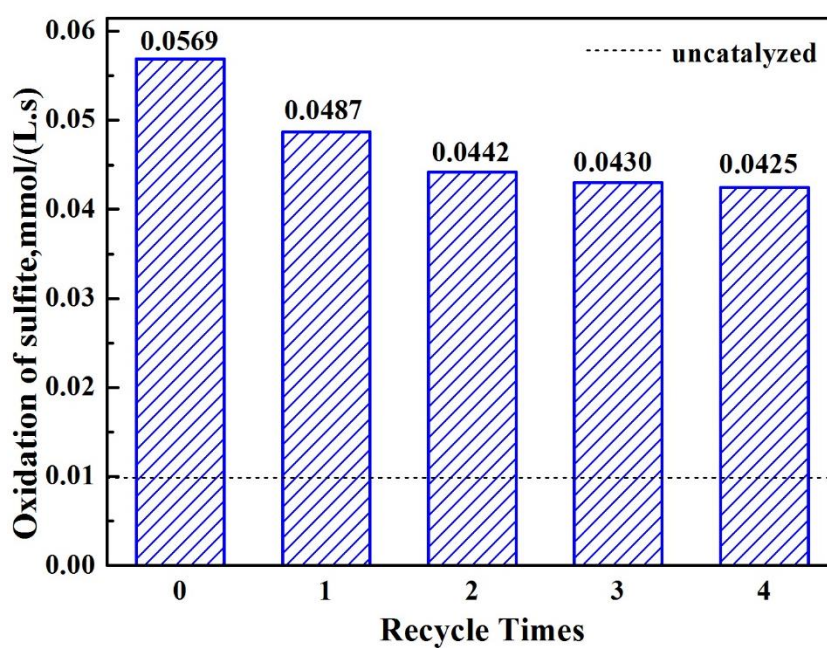


Figure S2 demonstrated that the catalytic performance of Co-ACFs declined substantially after the first and second cycles, and then remained basically the same. Although the oxidation rate decreased by about 25% from 0.0569 to 0.0425 mmol/(L·s) after four recycles, the catalytic reaction rate was still promoted by over 4.2 times compared with the uncatalyzed counterpart, which can accelerate the oxidation of magnesium sulfite. These results indicated that the developed Co-ACFs catalyst-adsorbent is stable, enabling the feasibility for application in a real desulfurization process. The calcined Co-ACFs will be squeezed into spherical particles by the binder, so that the shedding of the cobalt active component can be further prohibited.

Figure S3 Comparison of the catalysis activity of Co-based catalysts on magnesium sulfite oxidation at a concentration of 0.5 g/L

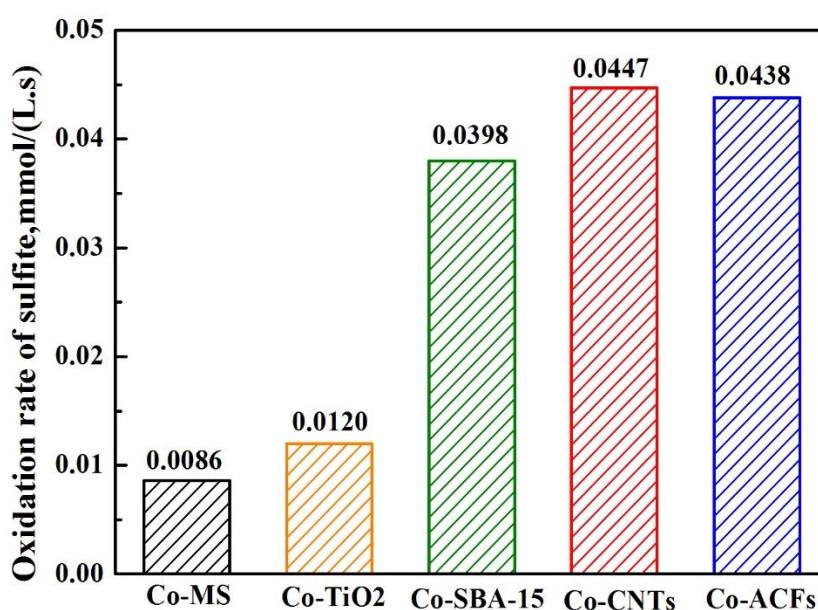


Table S1 Comparison of adsorption capacities for heavy metals onto different adsorbents.

adsorbent	ion	Fitting models	kinetic parameters		Qm(mg/g)	Ref.
			k ₂	q _e		
Co-CNTs	Hg	Langmuir	0.29	4.9	166.7	[1]
SBA-15-SH	Hg	Freundlich	0.011	189.87	195.6	[2]
Fe ₃ O ₄ NPs@AC@SO ₃ H	Cd	Langmuir	0.01	35.4	119.04	[3]
HS-mSi@MOF-5	Cd	Langmuir	0.001	91.996	98.0366	[4]
ZrO(OH) ₂ /VMT	Ni	Langmuir	0.03	80.25	90.21	[5]
HWSAC	Ni	Langmuir	0.0056	31.75	61.92	[6]
ACF	Hg	Langmuir	\	\	198.8	[7]
ACF	Cd	Langmuir	\	\	2.33	[8]
ACC	Ni	Langmuir			2.296	[9]
MO-ACF	Cd	Langmuir	0.9562	\	14.88	[8]
ACF-Ox	Ni	Langmuir	\	\	21.12	[10]
ACF modification via by sulphuric acid	Hg	Langmuir	0.019	\	300.3	[7]
	Hg	Langmuir	0.46	4	333.3	This work
Co-ACF	Cd	Langmuir	0.07	5.32	500	
	Ni	Langmuir	0.09	9.34	52.6	

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