

Supplementary Materials

Catalytic Wet Peroxide Oxidation of Anionic Pollutants over Fluorinated Fe₃O₄ Microspheres at Circumneutral pH Values

Fengxi Chen ^{1*}, Huaixiang Lv ¹, Wu Chen ², Rong Chen ^{1,3}

¹ State Key Laboratory of New Textile Materials & Advanced Processing Technologies, Wuhan Textile University, Wuhan 430200, PR China

² School of Chemistry and Environmental Engineering, Wuhan Institute of Technology, Wuhan 430205, P. R. China

³ Henan Institute of Advanced Technology, Zhengzhou University, Zhengzhou 450002, P.R. China

* Correspondence: fxchen@wtu.edu.cn (F.C.)

Table S1. Effect of initial H₂O₂ concentration on OG decolorization with H₂O₂ over F-Fe₃O₄-1 microspheres

[H ₂ O ₂] ₀ (mM)	15-min adsorption (%)	2-h DR (%)	k _{s1} /min ⁻¹ (R ²)
20	13.2	65.9	0.0082 (0.993)
40	12.3	96.8	0.0284 (0.991)
60	13.3	97.3	0.0304 (0.992)
80	10.3	90.3	0.0182 (0.994)

Reaction conditions: OG (0.1 mM, pH 6.5) and F-Fe₃O₄-1 (0.5 g/L) at 40 °C for 2 h.

Table S2. Effect of initial pH on OG decolorization with H₂O₂ over F-Fe₃O₄-1 microspheres

[pH] ₀	15-min adsorption (%)	2-h DR (%)	k _{s1} /min ⁻¹ (R ²)
3.0	22.2	97.5	0.0381 (0.991)
5.0	15.2	97.4	0.0320 (0.990)
6.5	12.3	96.8	0.0284 (0.991)
7.0	10.8	94.8	0.0252 (0.995)
9.0	8.1	88.5	0.0175 (0.992)

Reaction conditions: OG (0.1 mM) and F-Fe₃O₄-1 (0.5 g/L) at 40 °C for 2 h.

Table S3. Effect of reaction temperature on OG decolorization with H₂O₂ over F-Fe₃O₄-1 microspheres

Temperature (°C)	15-min adsorption (%)	2-h DR (%)	k _{s1} /min ⁻¹ (R ²)
25	5.5	32.6	0.0028 (0.991)
40	12.3	96.8	0.0284 (0.990)
55	13.1	99.5	0.0445 (0.991)

Reaction conditions: OG (0.1 mM, pH 6.5) and F-Fe₃O₄-1 (0.5 g/L) for 2 h.

Table S4. EDX-determined elemental composition of F-Fe₃O₄-1 microspheres

Element/wt%	O	F	Fe
Spot 1	15.19	8.24	76.57
Spot 2	16.41	7.57	76.02
Spot 3	16.71	5.48	77.82
Average/wt%	16.1±0.8	7.1±1.4	76.8±0.9

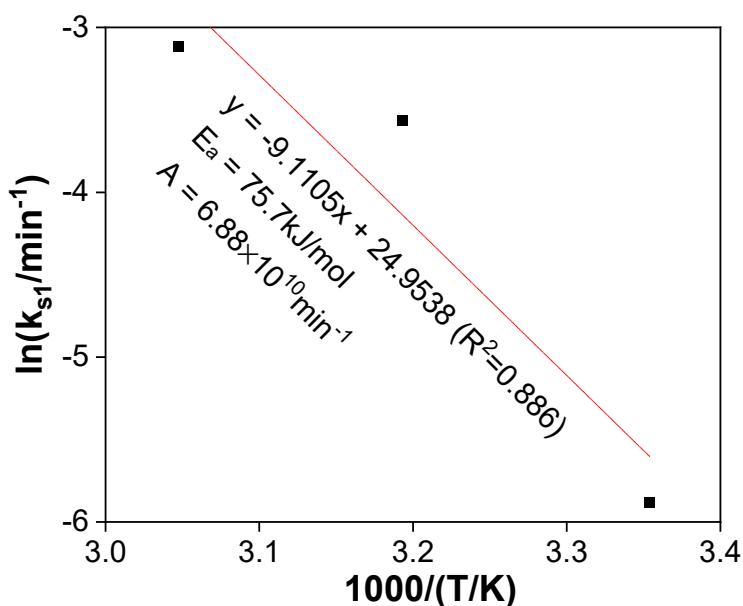


Figure S1. The Arrhenius plot of the decolorization of OG (0.1 mM, pH 6.5) with H₂O₂ on F-Fe₃O₄-1 microspheres.