

Article

Transition Metal (Fe₂O₃, Co₃O₄ and NiO)-Promoted CuO-Based α -MnO₂ Nanowire Catalysts for Low-Temperature CO Oxidation

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Abstract: As a toxic pollutant, carbon monoxide (CO) usually causes harmful effects on human health. Therefore, the thermally catalytic oxidation of CO has received extensive attention in recent years. The CuO-based catalysts have been widely investigated due to their availability. In this study, a series of transition metal oxides (Fe₂O₃, Co₃O₄ and NiO) promoted CuO-based catalysts supported on the α -MnO₂ nanowire catalysts were prepared by the deposition precipitation method for catalytic CO oxidation reactions. The effects of the loaded transition metal type, the loading amount, and the calcination temperature on the catalytic performances were systematically investigated. Further catalyst characterization showed that the CuO/ α -MnO₂ catalyst modified with 3 wt% Co₃O₄ and calcined at 400 °C performed the highest CO catalytic activity ($T_{90} = 75$ °C) among the investigated catalysts. It was supposed that the loading of the Co₃O₄ dopant not only increased the content of oxygen vacancies in the catalyst but also increased the specific surface area and pore volume of the CuO/ α -MnO₂ nanowire catalyst, which would further enhance the catalytic activity. The CuO/ α -MnO₂ catalyst modified with 3 wt% NiO and calcined at 400 °C exhibited the highest surface adsorbed oxygen content and the best normalized reaction rate, but the specific surface area limited its activity. Therefore, the appropriate loading of the Co₃O₄ modifier could greatly enhance the activity of CuO/ α -MnO₂. This research could provide a reference method for constructing efficient low-temperature CO oxidation catalysts.

Keywords: transition metal oxides; CuO-based catalysts; Co₃O₄ modification; α -MnO₂ nanowire; catalytic oxidation of CO



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1. Introduction

As a toxic pollutant in the atmosphere, carbon monoxide (CO) has widely received public attention in recent years [1]. CO is not only mainly produced during the incomplete combustion of fossil fuels and motor vehicle emissions [2], but it is also a precursor of the ozone pollution [3], which can be harmful even at low concentrations. In recent years, scientists around the world have employed different methods for the removal of CO, such as photocatalysis [4], thermocatalysis oxidation [5], etc. Thermocatalysis oxidation is the most commonly used method for the oxidation of CO. The noble metal-based catalysts (such as Au [6,7], Pt [8–10], Pd [11], Ru [12,13], Rh [14] and Ag [15]), transition metal oxide based catalysts (such as MnO₂ [16], CeO₂ [17], ZrO₂ [18], Co₃O₄ [16]), and metal

hybrid catalysts (such as MnCo_2O_4 spinel [19], Ce-Zr solid solution [20]) can be used for the thermocatalytic oxidation of CO at low temperatures without generating the secondary contaminants [16]. Transition metal-based catalysts have good potential for application due to their low cost, high stability, and good activity [21,22]. Among the transition metal oxides, manganese dioxide (MnO_2) has many advantages, such as a low price, being environmentally friendly, and abundance in nature [2,23]. As is well known, MnO_2 usually exists in different crystalline phases with different structures, such as the α -, β -, and γ -types of the one-dimensional pore structure, the δ -type of the two-dimensional pore structure, and the λ -type of the three-dimensional network structure [24], which greatly depend on the different connectivity (corner-edge sharing) exhibited by the $[\text{MnO}_6]$ octahedra [25–28]. Liang et al. [28] prepared MnO_2 nanorods with four different crystalline phases for the oxidation of CO. It was found that the order of oxidation activity of different crystalline phases with the same nanorod morphology was greatly different, in decreasing order from α - \approx δ - > γ - > β - MnO_2 (The temperature at which the conversion of CO on α - MnO_2 reached 100% was approximately 130 °C). This indicated that the oxidation activity of CO significantly depended on the phase structure and channel structure of the MnO_2 . Tian et al. [29] investigated the effects of the crystalline phases of MnO_2 (α -, β -, and ε - MnO_2) on the performances of the oxidation of CO and toluene. It was found that the β - MnO_2 performed the highest activity for CO oxidation ($T_{90} = 75$ °C) among the three crystalline phases of MnO_2 , and the α - MnO_2 behaved with the lowest activity for CO oxidation ($T_{90} = 118$ °C). The content of oxygen vacancy in the catalyst was also determined by in situ EPR spectra and the results showed that the catalytic activity of the catalyst was proportional to the concentration of oxygen vacancies, which were regarded as the active sites for the adsorption and dissociation of oxygen molecules. It was believed that the catalytic activity of MnO_2 was greatly related to the oxygen vacancy activity [29]. As is well known, the α - MnO_2 is provided with the 1D (1×1) (2×2) tunnel structures, which are attributed to the tetragonal crystal system [30].

CuO has been widely used as the active site of the low-temperature catalytic oxidation of CO due to its excellent activity and abundant availability [31]. For example, Raziye Jocar et al. [31] used CuO/ α - MnO_2 as the catalyst of the preferential oxidation of CO in the hydrogen-rich gas stream and investigated the interaction between the MnO_2 and CuO ($T_{97} = 130$ °C), the superior activity of the catalyst due to the beneficial synergistic interaction between CuO and α - MnO_2 . Meanwhile, the catalytic activity was also influenced by several factors, such as specific surface area, crystallinity, oxygen vacancies, and redox properties. Qian et al. [32] prepared a series of CuO/ MnO_2 catalysts with different CuO loading amounts by the incipient wetness impregnation method for the oxidation of CO. The catalyst activity was almost the same for the CuO loadings, from 1% to 40%. Sun et al. [33] prepared a CuO/ $\text{Cu}_{1.5}\text{Mn}_{1.5}\text{O}_4$ spinel-type composite oxide for synergistic catalysis of CO oxidation. It was found that the synergistic effect between $\text{Cu}_{1.5}\text{Mn}_{1.5}\text{O}_4$ spinel and CuO can promote the oxidation of CO, and CuO- $\text{Cu}_{1.5}\text{Mn}_{1.5}\text{O}_4$ had the best oxidation activity for CO ($T_{100} = 177$ °C).

Nowadays, for the modification of transition metal oxide-based catalysts, in addition to the carrier and active site, the promoter also plays an active role in the improvement of catalytic performance [34]. Gao et al. [35] doped transition metals (Fe, Co, Ni, and Cu) with a 1:10 molar ratio on α - MnO_2 nanowires by a one-step hydrothermal method to oxidize CO. Among the four transition metals, $\text{Cu}_{0.1}\text{MnO}_x$ had the best oxidation activity for CO ($T_{100} \sim 120$ °C). Krasimir et al. [36] investigated the effects of different molar ratios of chemical compositions on the γ - Al_2O_3 -supported CuO/ MnO_2 / Cr_2O_3 catalysts for the oxidation of CO, dimethyl ether (DME), and methanol. The results showed that the Cu-Mn-Cr/ γ - Al_2O_3 catalyst, which Cu/(Mn + Cr) has a molar ratio of 2:1 and a Mn/Cr molar ratio of 0.25, can achieve the complete oxidation of CO at 200 °C.

In order to further investigate the contribution of promoters to the catalytic performance of CuO/ MnO_2 catalysts in the oxidation of CO, in this work, the α - MnO_2 nanowire was successfully prepared by the hydrothermal method and used as the support for the

CuO-based catalysts. A series of the transition metal oxides (Fe_2O_3 , Co_3O_4 , and NiO) promoted CuO-based $\alpha\text{-MnO}_2$ nanowire catalysts were prepared by the deposition precipitation method. The effects of the type, the loading amount, and the calcination temperature of three transition metal oxides on the performance of the catalytic oxidation of CO were systematically studied. The obtained catalysts were carefully characterized by X-ray diffraction (XRD), scanning electron microscopy (SEM), nitrogen physisorption, X-ray photoelectron spectroscopy (XPS), and H_2 temperature-programmed reduction ($\text{H}_2\text{-TPR}$). These catalysts were evaluated for their catalytic performances in the oxidation of CO. The results show that the Co_3O_4 (3 wt.%) promoted 10 wt.% CuO/ $\alpha\text{-MnO}_2$ catalyst calcined at 400°C performed the greatest CO reactivity ($T_{90} = 75^\circ\text{C}$).

2. Results and Discussion

2.1. XRD Analysis

The crystalline phase structures of the catalysts were obtained using X-ray diffraction (XRD) analysis. The results of XRD are shown in the Figure 1. Figure 1a shows the XRD patterns of the transition metal oxide ($\text{MO}_x = \text{Fe}_2\text{O}_3$, Co_3O_4 , NiO) doped catalysts and the pristine 10CuO/ $\alpha\text{-MnO}_2$ -400 catalyst. It can be observed in Figure 1a that all the catalysts show wide and clear diffraction peaks at $2\theta = 12.78^\circ$, 18.11° , 25.71° , 37.52° , 41.97° , 49.86° , 56.37° , 65.11° , and 69.71° , which could be conformed to the characteristic peaks of the $\alpha\text{-MnO}_2$ (PDF#44-0141). It could be observed that the intensity of the diffraction peaks of MnO_2 increased after the loading of the second transition metal oxides, especially over the 10CuO-3 Co_3O_4 / $\alpha\text{-MnO}_2$ catalyst. This phenomenon suggests that the crystallinity of the catalysts increased after loading the transition metal oxides, especially for the catalyst loading Co_3O_4 . Meanwhile, two weak diffraction peaks could be detected at $2\theta = 35.5^\circ$ and 38.8° , which corresponded to the characteristic peaks of CuO, according to the standard card of PDF#05-0661. The diffraction peak intensity of CuO became weaker after the addition of the transition metal to the catalyst, indicating that the addition of the transition metal promoted the dispersion of CuO.

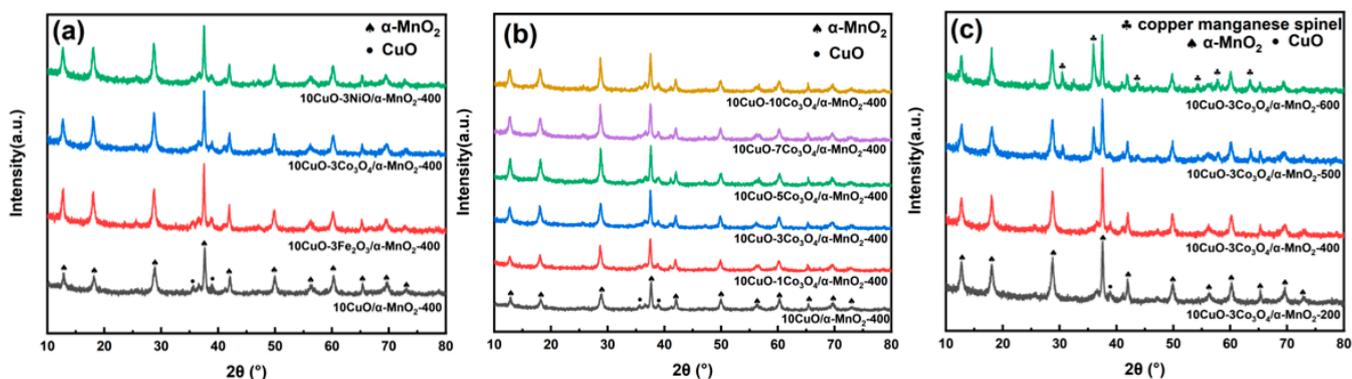


Figure 1. Powder XRD patterns of the (a) 10CuO-3 MO_x / $\alpha\text{-MnO}_2$ -400 catalysts doped with different transition metals ($\text{MO}_x = \text{Fe}_2\text{O}_3$, Co_3O_4 , NiO); (b) 10CuO- $y\text{Co}_3\text{O}_4$ / $\alpha\text{-MnO}_2$ -400 catalysts doped with different contents of Co_3O_4 ($y = 1, 3, 5, 7, 10$); (c) 10CuO-3 Co_3O_4 / $\alpha\text{-MnO}_2$ - T catalysts calcined at different temperatures ($T = 200, 400, 500, 600^\circ\text{C}$).

The XRD patterns of the 10CuO- $y\text{Co}_3\text{O}_4$ / $\alpha\text{-MnO}_2$ -400 catalysts with different contents of Co_3O_4 are shown in the Figure 1b. It can be observed that the diffraction peaks of CuO are very clear over the pristine 10CuO/ $\alpha\text{-MnO}_2$ -400 catalyst without Co_3O_4 loading. However, it can be observed that the diffraction peaks of CuO became blurry after loading Co_3O_4 from 1 wt.% to 10 wt.%. The possible reason accounting for this phenomenon was that the appropriate loading amount of Co_3O_4 could greatly enhance the dispersion of CuO on the surface of $\alpha\text{-MnO}_2$ nanowire catalyst. Nevertheless, the characteristic peaks of Co_3O_4 could be not observed in the catalysts due to the good dispersion of Co_3O_4 .

The XRD patterns of 10CuO-3Co₃O₄/α-MnO₂ catalysis calcined at different temperatures are shown in the Figure 1c. It can be recognized that the intensity of the diffraction peaks of α-MnO₂ became weak at high calcination temperatures. The reason might be due to the formation of a copper manganese spinel (JCPDS No.70-0260) [37]. Only a diffraction peak of CuO was observed at 38.8°. Meanwhile, the intensity of the diffraction peaks of CuO became sharp with the increase of the calcination temperature (only from 200 °C to 400 °C). This indicated that the dispersion of CuO on the catalyst surface gradually deteriorated and the crystal size of CuO nanoparticles grew due to the finite surface area of the MnO₂ nanowire from 200 °C to 400 °C. As the calcination temperature rose from 400 °C to 600 °C, the diffraction peak strength of copper manganese spinel increased. The diffraction peak strength of CuO decreased, indicating that the dispersion of CuO increased.

2.2. SEM Observation

The morphologies of the 10CuO-3MO_x/α-MnO₂-400 catalysts loaded with different transition metal oxides (MO_x = Fe₂O₃, Co₃O₄, NiO) and the pristine 10CuO/α-MnO₂-400 catalyst were characterized by SEM (Figure 2). It can be observed that the morphology of the α-MnO₂ nanowire support greatly changed after loading the CuO active sites and the transition metal oxides. Specifically, the length to diameter ratio of the α-MnO₂ nanowire significantly decreased compared to the pristine α-MnO₂ nanowire. This was caused by the high loading contents of the CuO and transition metal oxides. In addition, the α-MnO₂ nanowire support might experience thermal sintering and self-assemble at high calcination temperatures. The spatial distribution of the elements over the one-dimensional 10CuO-3MO_x/α-MnO₂-400 (MO_x = Fe₂O₃, Co₃O₄, NiO) nanowire catalyst was studied by energy dispersive X-ray spectroscopy mapping (EDS-mapping). As shown in Figure 2, the Cu and the doped Fe/Co/Ni elements were homogeneously distributed over these investigated catalysts. This indicated that the great dispersion of the CuO and the doped transition metal oxides on the surface of the α-MnO₂ nanowire supports could be facily achieved by the precipitation deposition method.

2.3. BET Analysis

In order to further investigate the structural properties of the catalysts, the specific surface areas, pore volumes, and pore size distributions of the catalysts were measured by nitrogen physisorption measurements. As shown in Figure 3a, all catalysts show IV isotherms with H3-shaped hysteresis loops. These results prove the presence of mesopores with a narrow slit-shape in the catalyst [38]. It is also interesting to find that the 10CuO-3MO_x/α-MnO₂ catalysts still possess the similar mesoporous structure to the 10CuO/α-MnO₂ nanowire catalysts after loading different transition metal oxides. This demonstrates that the 10CuO/α-MnO₂ nanowire catalyst was provided with good thermal stability. The pore size distribution curves of the corresponding catalysts are shown in Figure 3b. The pore diameter of the catalysts is in the range of 2–15 nm after loading with transition metals. Moreover, the specific parameters of the structural properties of these catalysts are shown in Table 1. It can be observed that the specific surface areas of both the 10CuO-3Fe₂O₃/α-MnO₂ and 10CuO-3NiO/α-MnO₂ catalysts decreased after the loading of transition metals. On the contrary, the specific surface area of the 10CuO-3Co₃O₄/α-MnO₂ catalysts increased substantially. This indicates that the Co₃O₄ on the α-MnO₂ nanowire support surface was highly dispersed. Meanwhile, the average pore sizes of all the 10CuO-3MO_x/α-MnO₂ (MO_x = Fe₂O₃, Co₃O₄, NiO) catalysts are very similar to the pristine 10CuO/α-MnO₂ nanowire. Specifically, the 10CuO-3Fe₂O₃/α-MnO₂ catalyst has a similar pore size distribution as the 10CuO/α-MnO₂ catalyst. This indicates that their mesoporous structures are not sharply impaired by the loaded metal oxides. In addition, the pore capacities of all 10CuO-3MO_x/α-MnO₂ (MO_x = Fe₂O₃, Co₃O₄, NiO) catalysts are enhanced compared to the pristine 10CuO/α-MnO₂. As for the 10CuO-3Co₃O₄/α-MnO₂ catalyst, its surface area was twice as large as the pristine 10CuO/α-MnO₂ catalyst. The higher specific surface area is beneficial for the catalyst to expose more active sites, and the larger pore

volume helps the reactant accelerate the reactant mass diffusion and has a better adsorption ability to the reactant [39].

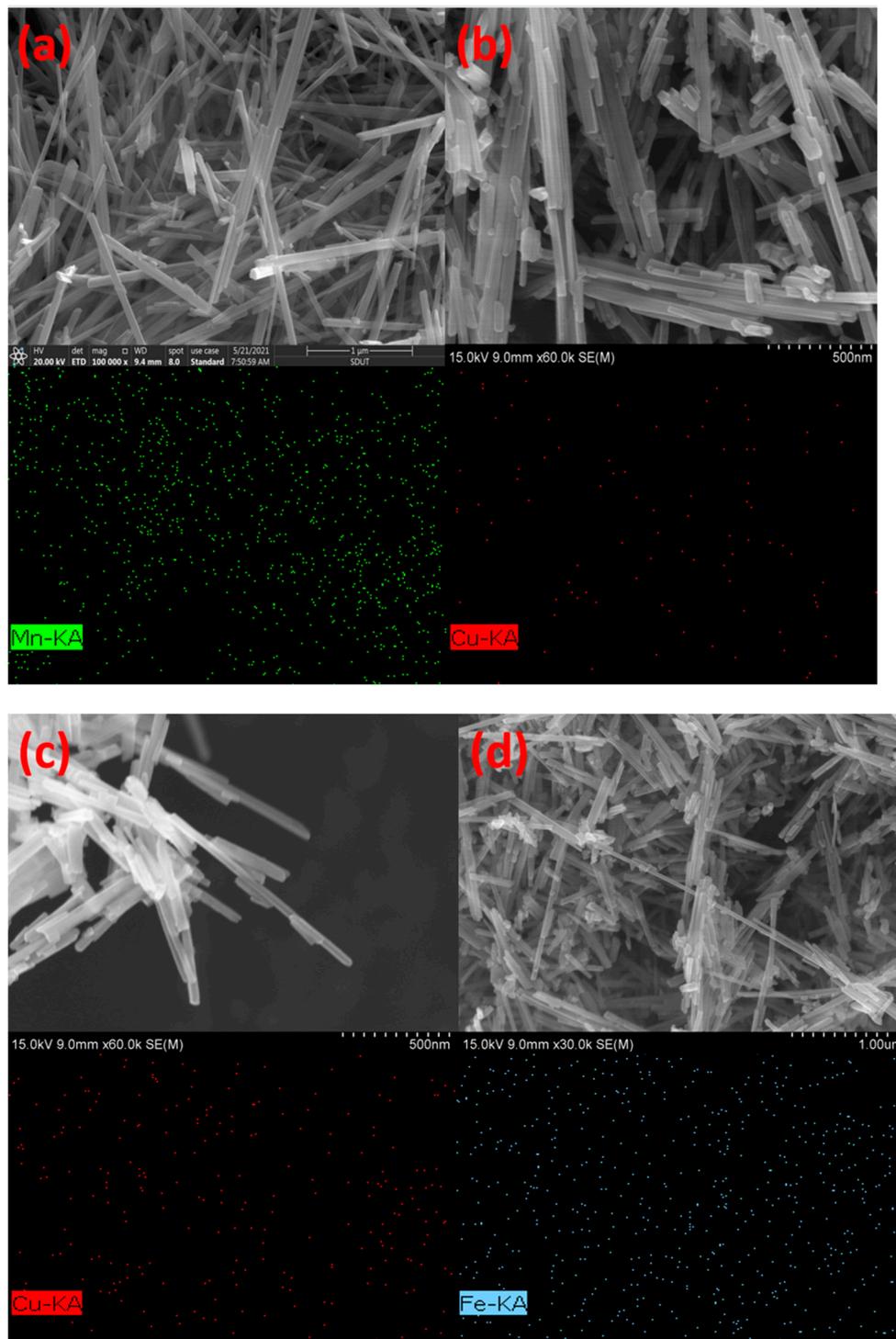


Figure 2. Cont.

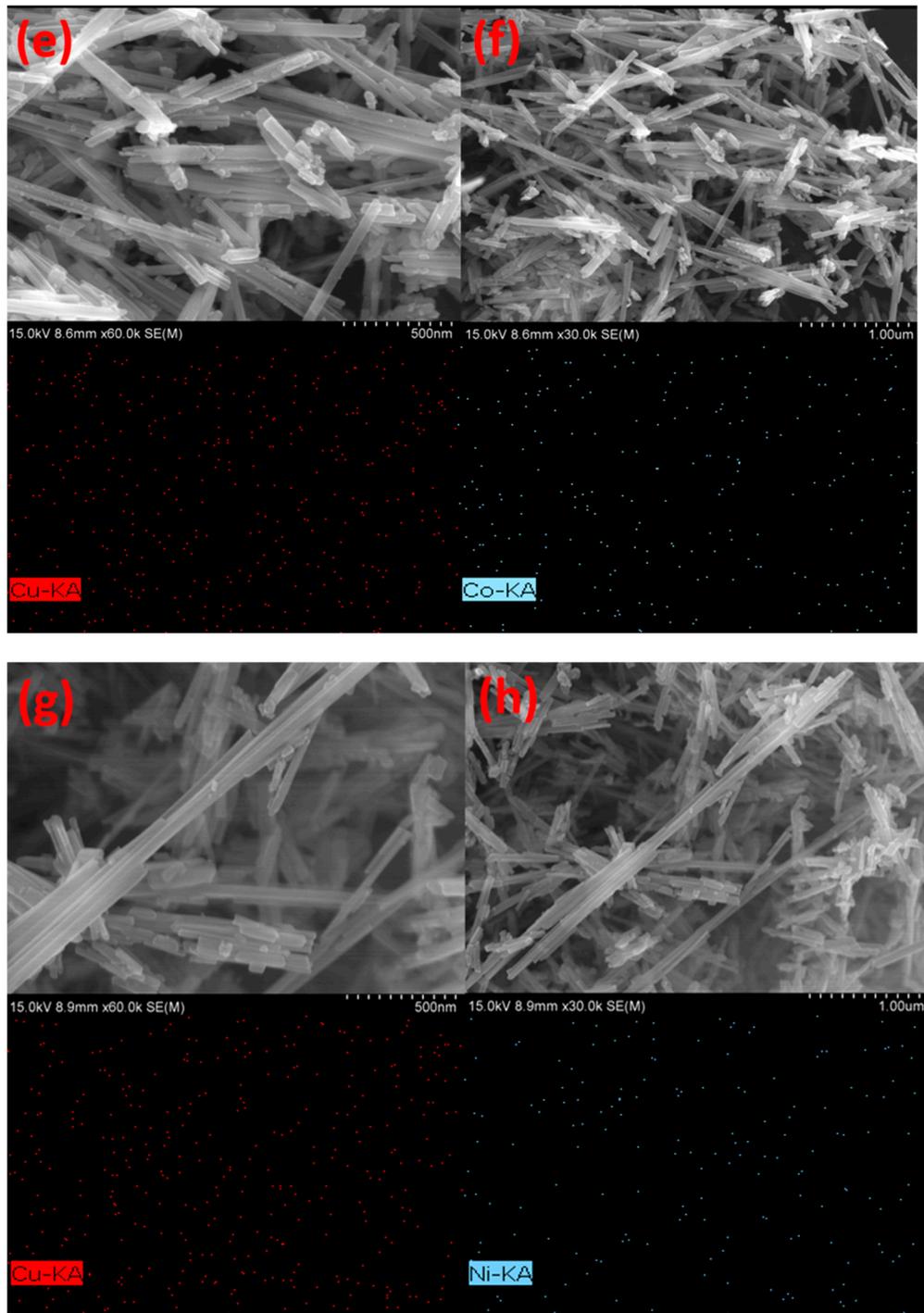


Figure 2. SEM-EDS images of the (a) α - MnO_2 nanowire, (b) $10\text{CuO}/\alpha\text{-MnO}_2\text{-400}$, (c,d) $10\text{CuO}-3\text{Fe}_2\text{O}_3/\alpha\text{-MnO}_2\text{-400}$, (e,f) $10\text{CuO}-3\text{Co}_3\text{O}_4/\alpha\text{-MnO}_2\text{-400}$, and (g,h) $10\text{CuO}-3\text{NiO}/\alpha\text{-MnO}_2\text{-400}$.

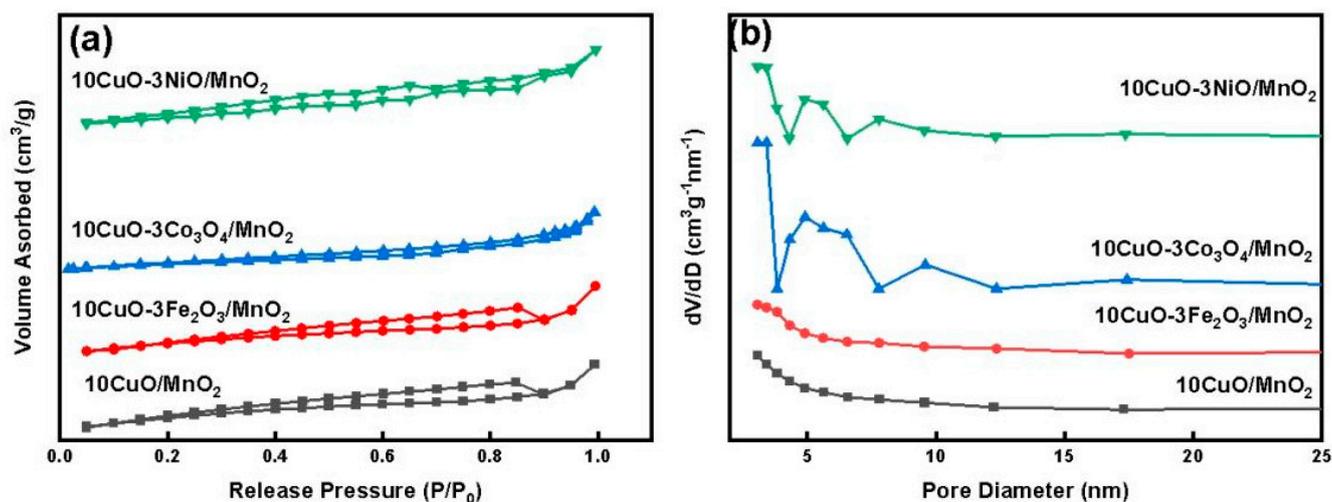


Figure 3. (a) Nitrogen adsorption-desorption isotherms and (b) BJH pore size distribution of the 10CuO/ α -MnO₂-400 and 10CuO-3MOx/ α -MnO₂-400 (MOx = Fe₂O₃, Co₃O₄, NiO) catalysts.

Table 1. Structural properties of the 10CuO/ α -MnO₂-400 and 10CuO-3MOx/ α -MnO₂-400 (MOx = Fe₂O₃, Co₃O₄, NiO) catalysts.

Catalyst	Specific Surface Area (m ² /g)	Pore Volume (cm ³ /g)	Average Pore Diameter (nm)	Isotherm Type
10CuO/MnO ₂	72	0.09	3.1	IV H3
10CuO-3Fe ₂ O ₃ /MnO ₂	64	0.10	3.1	IV H3
10CuO-3Co ₃ O ₄ /MnO ₂	84	0.18	3.1	IV H3
10CuO-3NiO/MnO ₂	52	0.10	3.1	IV H3

2.4. XPS Analysis

The coordination, composition, and valence state of the elements over the catalyst surface were investigated by X-ray photoelectron spectroscopy (XPS). The XPS of Mn 2p, O 1s, and Cu 2p of the 10CuO/ α -MnO₂-400 and the 10CuO-3MOx/ α -MnO₂-400 (MOx = Fe₂O₃, Co₃O₄, NiO) are shown in Figure 4. It can be observed in Figure 4a that 10CuO/ α -MnO₂-400 had two main peaks around 654.0 eV and 642.0 eV, which could be attributed to the binding energies of the Mn 2p_{1/2} at BE = 654.0 eV and Mn 2p_{3/2} at BE = 642.0 eV, respectively. The two main peaks were divided into four peaks after the peak fitting. The fitted peaks of Mn 2p_{3/2} at 642.0 eV and 643.5 eV indicate the existence of Mn³⁺ and Mn⁴⁺ in the 10CuO-3MOx/ α -MnO₂-400 catalyst [40–44]. The ratios of the Mn³⁺/(Mn³⁺ + Mn⁴⁺) over different catalysts followed the below order: 10CuO-3Co₃O₄/ α -MnO₂ (0.606) > 10CuO-3NiO/ α -MnO₂ (0.541) > 10CuO/ α -MnO₂ (0.414) > 10CuO-3Fe₂O₃/ α -MnO₂ (0.406). The redox electron pair in Cu-Mn oxide was the -Cu²⁺-O²⁻-Mn⁴⁺-→-Cu⁺-□-Mn³⁺+ + 1/2O₂ (□ indicates the oxygen vacancy) [45]. The content of Mn³⁺ and oxygen vacancies are proportional, or indirectly proportional, to the oxidation capacity of the catalyst [25,29,46,47]. The Mn³⁺ may cause the Jahn-Teller effect, which could prolong the Mn-O bond in [MnO₆] [48,49], thereby prolonging the distance between the oxygen pairs and causing the stretching of the Mn-O bond length [48]. As a result, the Mn-O bond was easier to break, and the mobility of oxygen became higher. Therefore, the released surface oxygen atoms are more likely to participate in the reaction and thus promote the catalytic performance. The XPS of O 1s of 10CuO-3MOx/ α -MnO₂-400 was measured to elucidate the nature of the oxygen species over the 10CuO-3MOx/ α -MnO₂-400 catalysts. As shown in Figure 4b, all the samples show double peaks of different oxygen species. Specifically, the BEs at around 529.8 eV and 531.4 eV could be ascribable to the surface lattice oxygen (O_{latt}) and surface adsorbed oxygen (O_{ads}) species [50,51], respectively. The ratios of the

$O_{\text{ads}}/O_{\text{latt}}$ are also summarized in Table 2. It can be observed that the binding energy of surface lattice oxygen (O_{latt}) shifted to a higher binding energy with the addition of transition metals. The surface oxidation states of the Cu species were also investigated to show the redox properties of the as-prepared catalysts. As shown in Figure 4c, all the catalysts displayed two main peaks of Cu $2p_{1/2}$ (953.6 eV) and Cu $2p_{3/2}$ (933.7 eV) [43]. The Cu $2p_{3/2}$ orbitals with binding energy in the range of 930.0–935.0 eV, could be divided into two peaks. Specifically, the binding energy peak at 933.4 eV was attributed to Cu^+ , and the peak at 934.4 eV was attributed to Cu^{2+} [43]. Furthermore, it is worth noting that the Cu $2p_{3/2}$ peak is accompanied by the vibrating satellite peaks in the range of 940.0–944.0 eV [52]. Combined with the results of the Mn 2p spectrum, the catalysts formed redox pairs of $\text{Cu}^+/\text{Cu}^{2+}$ and $\text{Mn}^{3+}/\text{Mn}^{4+}$, which would promote the charge transference to generate more oxygen defects [37,53].

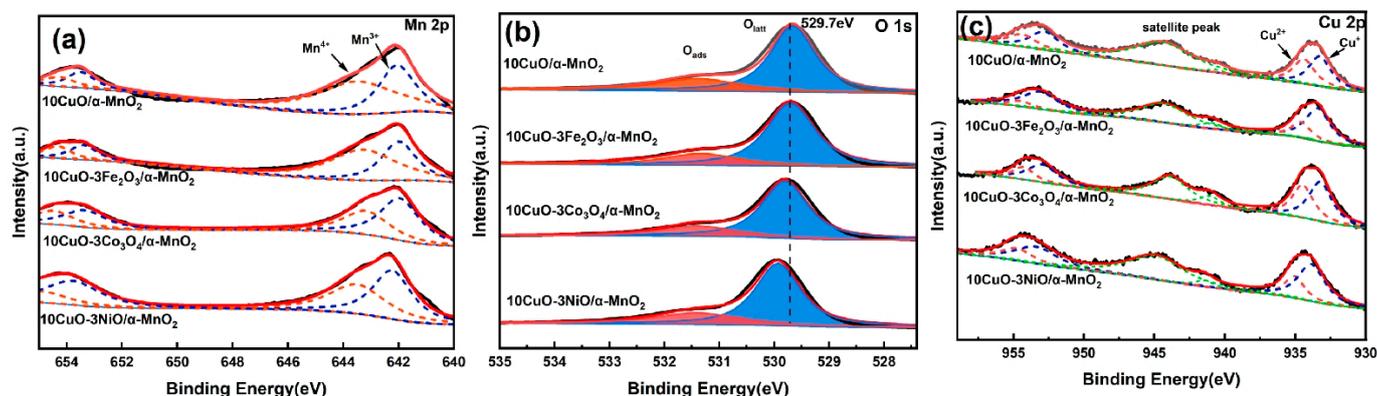


Figure 4. (a) Mn 2p spectra, (b) O 1s spectra, and (c) Cu 2p spectra of the 10CuO/ α -MnO₂, 10CuO-3Fe₂O₃/ α -MnO₂, 10CuO-3Co₃O₄/ α -MnO₂, and 10CuO-3NiO/ α -MnO₂ (T = 400 °C) catalysts.

Table 2. The ratio of $\text{Mn}^{3+}/(\text{Mn}^{3+} + \text{Mn}^{4+})$, $O_{\text{ads}}/O_{\text{latt}}$, and $\text{Cu}^+/(\text{Cu}^{2+} + \text{Cu}^+)$ of the 10CuO-3MOx/ α -MnO₂ (MOx = Fe₂O₃, Co₃O₄, NiO) and the 10CuO/ α -MnO₂ catalyst.

Sample	$\text{Mn}^{3+}/(\text{Mn}^{3+} + \text{Mn}^{4+})$	$O_{\text{ads}}/O_{\text{latt}}$	$\text{Cu}^+/(\text{Cu}^{2+} + \text{Cu}^+)$
10CuO/ α -MnO ₂	0.414	0.324	0.551
10CuO-3Fe ₂ O ₃ / α -MnO ₂	0.406	0.283	0.644
10CuO-3Co ₃ O ₄ / α -MnO ₂	0.606	0.300	0.563
10CuO-3NiO/ α -MnO ₂	0.541	0.352	0.688

To determine the valence state of the transition metals loaded on the catalyst, the XPS spectra of Fe 2p, Co 2p, and Ni 2p are determined. The XPS profile of Fe 2p over the 10CuO-3Fe₂O₃/ α -MnO₂-400 catalyst is shown in Figure 5a. The binding energies at 710.4 eV and 725.1 eV are ascribed to Fe $2p_{3/2}$ and Fe $2p_{1/2}$, respectively [54,55]. The peak of Fe $2p_{3/2}$ can be divided into two peaks (710.3 eV and 712.5 eV) [56]. In addition, a satellite peak was observed at about 718.3 eV. This indicates that the iron species existed in the form of Fe³⁺ on the surface of the 10CuO-3Fe₂O₃/ α -MnO₂-400 [54,56]. The XPS of Co 2p over the 10CuO-3Co₃O₄/ α -MnO₂-400 catalyst is shown in Figure 5b. The binding energies at 780.0 eV are ascribed to the Co $2p_{3/2}$ [57]. Meanwhile, there was no significant satellite shake-up intensity at 786 eV, indicating the dominance of Co³⁺ on the surface of Co₃O₄ [57]. The XPS of Ni 2p over the 10CuO-3NiO/ α -MnO₂-400 catalyst is shown in Figure 5c. The binding energies at 855.1 eV are ascribed to the Ni $2p_{3/2}$ [58]. The peak of Ni $2p_{3/2}$ of metallic Ni was basically at 852.6 eV, and the binding energy of NiO is about 1 eV higher than that of Ni⁰ [58]. The peak of NiO $2p_{3/2}$ is lower than that of Ni $2p_{3/2}$ in this work (854.8 eV). Therefore, the oxidation state of the indicated Ni element is mainly in the form of Ni²⁺. The higher binding energy compared to pure NiO binding energy indicates that NiO did not exist in the free form. The strong interaction was formed between the

Ni and the support. The results indicate that Fe_2O_3 , Co_3O_4 , and NiO were successfully loaded on the catalyst. The XPS electronic binding energies of the surface elements of the $10\text{CuO}-3\text{MOx}/\alpha\text{-MnO}_2-400$ ($\text{MOx} = \text{Fe}_2\text{O}_3, \text{Co}_3\text{O}_4, \text{NiO}$) are summarized in Table 3.

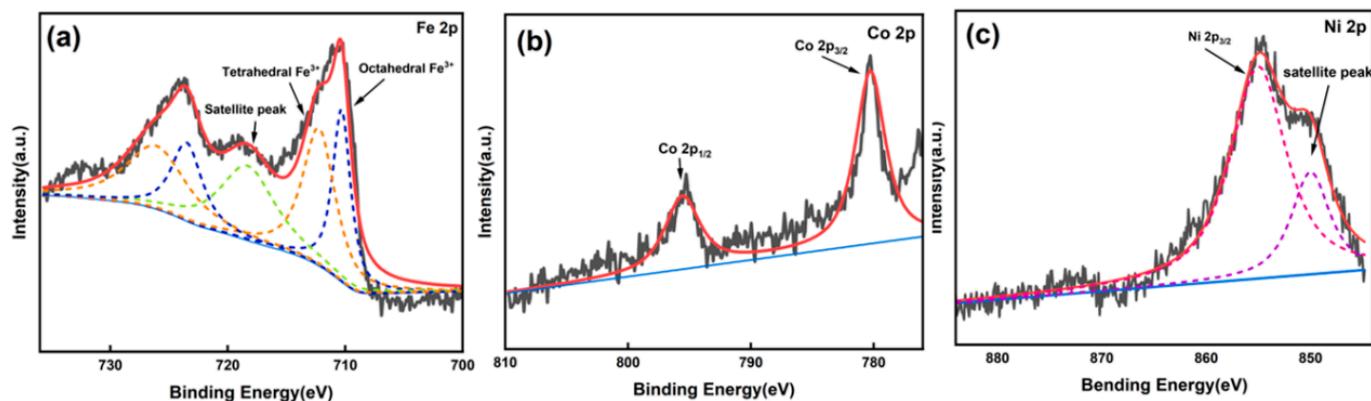


Figure 5. (a) Fe 2p, (b) Co 2p, and (c) Ni 2p spectra of the $10\text{CuO}-3\text{MOx}/\alpha\text{-MnO}_2-400$ ($\text{MOx} = \text{Fe}_2\text{O}_3, \text{Co}_3\text{O}_4, \text{NiO}$) catalysts.

Table 3. XPS electronic binding energies of surface elements of the $10\text{CuO}-3\text{MOx}/\alpha\text{-MnO}_2-400$ ($\text{MOx} = \text{Fe}_2\text{O}_3, \text{Co}_3\text{O}_4, \text{NiO}$) catalysts.

Sample	Cu 2p _{3/2}	O 1s	Mn 2p _{3/2}
10CuO/ $\alpha\text{-MnO}_2$	933.6	529.7	642.1
10CuO-3Fe ₂ O ₃ / $\alpha\text{-MnO}_2$	933.7	529.7	642.1
10CuO-3Co ₃ O ₄ / $\alpha\text{-MnO}_2$	933.7	529.8	642.3
10CuO-3NiO/ $\alpha\text{-MnO}_2$	934.2	529.9	642.4

2.5. H₂-TPR

The H₂-TPR profiles of $10\text{CuO}-3\text{MOx}/\alpha\text{-MnO}_2-400$ ($\text{MOx} = \text{Fe}_2\text{O}_3, \text{Co}_3\text{O}_4, \text{NiO}$), $10\text{CuO}-y\text{Co}_3\text{O}_4/\alpha\text{-MnO}_2-400$ ($y = 0, 1, 3, 5, 7, 10$), and $10\text{CuO}-3\text{Co}_3\text{O}_4/\alpha\text{-MnO}_2-T$ ($T = 200, 400, 500, 600$ °C) are shown in Figure 6. It can be observed that both the $10\text{CuO}-3\text{MOx}/\alpha\text{-MnO}_2-400$ catalyst and the $10\text{CuO}/\alpha\text{-MnO}_2-400$ catalyst show similar hydrogen consumption peaks in strength and shape in Figure 6a. Specifically, there were two sets of peaks in the range of 297–342 °C and 462–472 °C, which might be attributed to the hydrogen consumption derived from the gradual reduction of the $\alpha\text{-MnO}_2$ nanowire ($\text{MnO}_2 \rightarrow \text{Mn}_2\text{O}_3 \rightarrow \text{Mn}_3\text{O}_4$), according to the pioneer report [31]. Meanwhile, it is worth noting that the loading of the transition metal oxides on the $10\text{CuO}/\alpha\text{-MnO}_2$ support changed the reduction behavior of the $10\text{CuO}/\alpha\text{-MnO}_2$ catalyst. Specifically, the reduction of $10\text{CuO}/\alpha\text{-MnO}_2$ nanowires shifted to higher temperatures after loading the transition metal oxides. In general, the reducibility of the $10\text{CuO}/\alpha\text{-MnO}_2$ catalyst decreases with the addition of transition metal oxides. The H₂-TPR profiles of the $10\text{CuO}-y\text{Co}_3\text{O}_4/\alpha\text{-MnO}_2-400$ catalysts with different Co_3O_4 loading amounts are shown in Figure 6b. It is of great interest to find that the positions of the reduction peaks gradually shifted to the high-temperature region with the Co_3O_4 loading amount increasing from 1 wt% to 10 wt%. This illustrated that the reducibility of the catalysts also gradually decreases. Therefore, the reduction temperatures of the $10\text{CuO}-y\text{Co}_3\text{O}_4/\alpha\text{-MnO}_2-400$ catalyst were significantly higher than that of the $10\text{CuO}/\alpha\text{-MnO}_2$ catalyst, except for the $10\text{CuO}-7\text{Co}_3\text{O}_4/\alpha\text{-MnO}_2-400$ catalyst. The H₂-TPR profiles of the $10\text{CuO}-3\text{MOx}/\alpha\text{-MnO}_2-T$ catalyst with different calcination temperatures is shown in Figure 6c. It can be observed that the reduction temperature of the $10\text{CuO}-3\text{Co}_3\text{O}_4/\alpha\text{-MnO}_2-T$ catalyst gradually shifted to a higher reduction temperature with the increase of the calcination temperature from 200 °C to 600 °C. The positions of the two reduction peaks became closer. This suggests that the interaction between the CuO and the $\alpha\text{-MnO}_2$ nanowire support became stronger at higher temperatures.

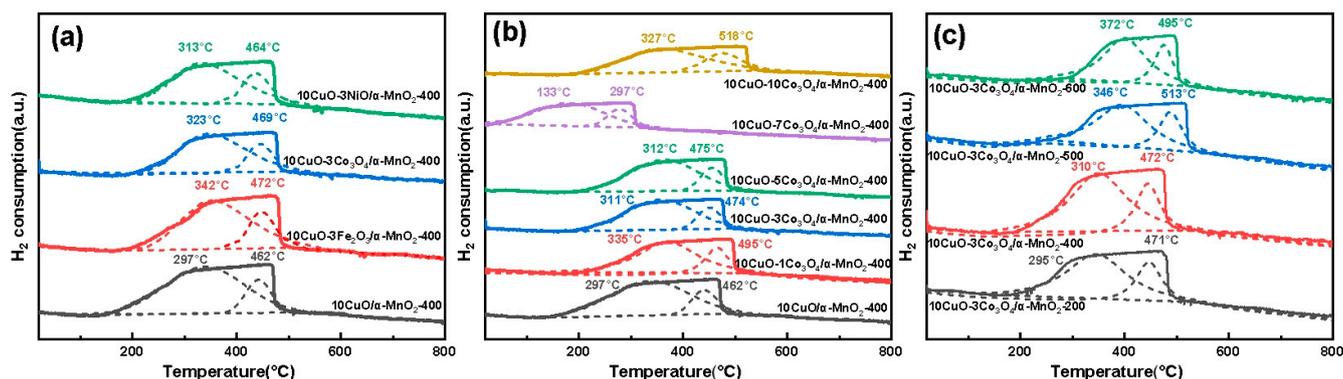


Figure 6. H₂-TPR profiles of the (a) 10CuO-3MO_x/α-MnO₂-400 (MO_x = Fe₂O₃, Co₃O₄, NiO), (b) 10CuO-yCo₃O₄/α-MnO₂-400 (y = 0, 1, 3, 5, 7, 10), and (c) 10CuO-3Co₃O₄/α-MnO₂-T (T = 200, 400, 500, 600 °C) catalysts.

2.6. Catalytic Performance of the CO Oxidation

2.6.1. The Effect of the Transition Metal Oxides (MO_x) on the Activities

The catalytic CO oxidation activities of the 10CuO-3MO_x/α-MnO₂-400 (MO_x = Fe₂O₃, Co₃O₄, NiO) catalysts were studied to investigate the effect of the dual loading of transition metals and CuO on the catalytic activity. As shown in Figure 7a, it can be observed that the CO conversion gradually increases with the increase of the reaction temperature until it reached 100%. As can be observed, the temperature of the 90% CO conversion over the 10CuO-3Co₃O₄/α-MnO₂-400 catalyst was 75 °C. The activity of the 10CuO-3NiO/α-MnO₂-400 catalyst basically shows a similar catalytic activity to that of the pristine 10CuO/α-MnO₂-400 catalyst without modification (T₉₀ = 77 °C). However, the CO oxidation activity of the 10CuO-3Fe₂O₃/α-MnO₂-400 catalyst (T₉₀ = 80 °C) was even worse than that of the pristine 10CuO/α-MnO₂-400 catalyst. Therefore, only the catalytic activity of the 10CuO-3Co₃O₄/α-MnO₂-400 was significantly improved compared with the pristine 10CuO/α-MnO₂-400 catalyst. The reason for this might be that the loading of Cu and Co could generate more oxygen vacancies and active sites to the α-MnO₂ nanowire. From the order of catalyst activity, it can be observed that the catalytic activity of the catalyst increases with the increase in of Mn³⁺ content. Meanwhile, the loading of Co₃O₄ increased the specific surface area and pore volume of the catalyst, providing more active sites for the reaction. The ratio of O_{ads}/O_{latt} was not exactly the same as the catalytic activity of the catalyst. The reasons for this phenomenon were stated in the discussion. The CO oxidation activity was significantly improved over the 10CuO-3Co₃O₄/α-MnO₂-400 catalyst. This result was well consistent with the Mn 2p XPS analysis. The results of the normalized reaction rates of the four catalysts are shown in Figure 7b. It can be observed that the normalized reaction rates gradually increase with the increase of the reaction temperature. The order of reaction rates of the catalysts per surface area was 10CuO-3NiO/MnO₂ > 10CuO/MnO₂ > 10CuO-3Co₃O₄/MnO₂ > 10CuO-3Fe₂O₃/MnO₂. The normalized reaction rates excluded the effect of specific surface area on catalytic activity and expressed the intrinsic catalytic activity of the catalysts. The order of the O_{ads}/O_{latt} ratios was consistent with the order of the normalized reaction rates.

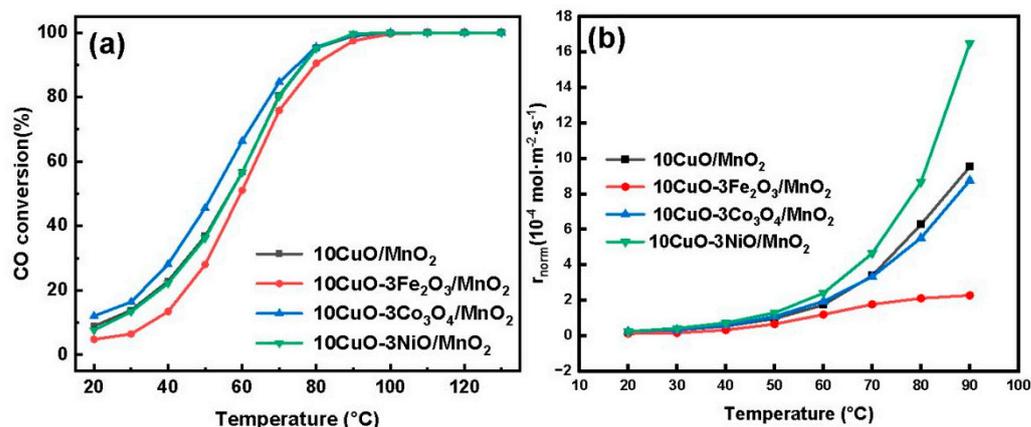


Figure 7. (a) CO conversions and (b) normalized reaction rates over the 10CuO-3MO_x/α-MnO₂-400 (MO_x = Fe₂O₃, Co₃O₄, NiO) catalysts under the reaction conditions: CO/O₂/N₂ = 1/20/79, GHSV = 12,000 mL·g⁻¹·h⁻¹, 1 atm.

2.6.2. The Effect of the Co₃O₄ Loading Amount on the Activities

The effect of Co₃O₄ loading on the catalytic activity of the CO oxidation was investigated, and the profiles of the CO conversion are shown in Figure 8a. It could be noticed that the Co₃O₄ loading amount in the range of 0 wt.%–10 wt.% had little effect on the CO oxidation activity of the 10CuO/α-MnO₂ catalyst. The order of CO catalytic activity of the catalysts is 10CuO-3Co₃O₄/α-MnO₂-400 (T₉₀ = 75 °C) > 10CuO-10Co₃O₄/α-MnO₂-400 (T₉₀ = 77 °C) ≈ 10CuO/α-MnO₂-400 > 10CuO-1Co₃O₄/α-MnO₂-400 (T₉₀ = 79 °C) ≈ 10CuO-5Co₃O₄/α-MnO₂-400 > 10CuO-7Co₃O₄/α-MnO₂-400 (T₉₀ = 84 °C). The 10CuO-3Co₃O₄/α-MnO₂-400 catalyst performed the highest activity in the low temperature region. It is shown that a certain increase in the loading of Co₃O₄ was beneficial to the catalytic activity of the catalyst.

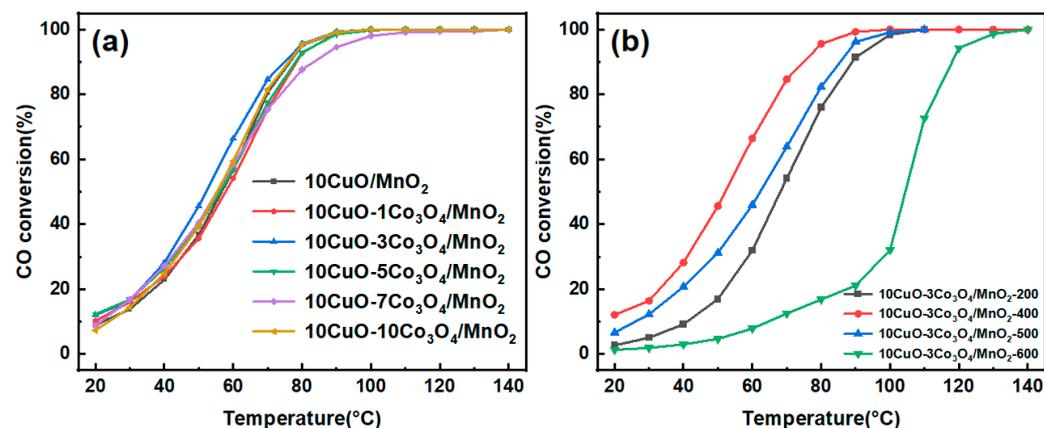


Figure 8. CO conversions over the (a) 10CuO-*y*Co₃O₄/α-MnO₂-400 (*y* = 0, 1, 3, 5, 7, 10), and (b) 10CuO-3Co₃O₄/α-MnO₂-*T* (*T* = 200, 400, 500, 600 °C) catalysts under the reaction conditions: CO/O₂/N₂ = 1/20/79, GHSV = 12,000 mL·g⁻¹·h⁻¹, 1 atm.

2.6.3. The Effect of the Calcination Temperature on the Activities

The catalytic CO oxidation of the 10CuO-3Co₃O₄/α-MnO₂-*T* (*T* = 200, 400, 500, 600 °C) catalyst was conducted to study the influence of the calcination temperature on catalytic activity, and the CO conversion profiles are shown in Figure 8b. The order of CO catalytic activity of catalysts is 10CuO-3Co₃O₄/α-MnO₂-400 (T₉₀ = 75 °C) > 10CuO-3Co₃O₄/α-MnO₂-500 (T₉₀ = 86 °C) > 10CuO-3Co₃O₄/α-MnO₂-200 (T₉₀ = 89 °C) > 10CuO-3Co₃O₄/α-MnO₂-600 (T₉₀ = 118 °C). It can be observed that catalytic activity has significantly decreased with the increase of the calcination temperature from 400 °C to 600 °C, especially over

the 10CuO-3Co₃O₄/α-MnO₂-600 catalyst. Specifically, the 10CuO-3Co₃O₄/α-MnO₂-600 catalyst had low activity of CO oxidation in the low temperature region. The possible reason is that the formation of a copper manganese spinel in the catalyst after calcination at high temperatures led to the significant decrease in the CO adsorption and oxidation capacities of the catalysts. This result is consistent with the XRD analysis. The 10CuO-3Co₃O₄/α-MnO₂-200 catalyst was also prepared for comparison. The results show that the catalytic activity of the 10CuO-3Co₃O₄/α-MnO₂-200 catalyst was significantly lower than that of the 10CuO-3Co₃O₄/α-MnO₂-400 catalyst. The possible reason is that the precursor of the Co₃O₄ could not be completely decomposed at 200 °C.

2.6.4. Stability Tests

The 12 h stability measurement was conducted over the 10CuO-3Co₃O₄/α-MnO₂-400 catalyst under the specific conditions (80 °C, CO/O₂/N₂ = 1/20/79, GHSV = 12,000 mL·g⁻¹·h⁻¹, 1 atm), and the result is shown in Figure 9a. It can be observed that the initial activity of the 10CuO-3Co₃O₄/α-MnO₂-400 catalyst can achieve 100% CO conversion in the first 2 h. Then, the CO conversion gradually decreased from 100% to 80% in the subsequent 2 h test. This indicates that the catalyst began to deactivate. However, the CO conversion could remain stable in the subsequent 8 h. This suggests that the 10CuO-3Co₃O₄/α-MnO₂-400 catalyst was provided with relatively good stability to some degree.

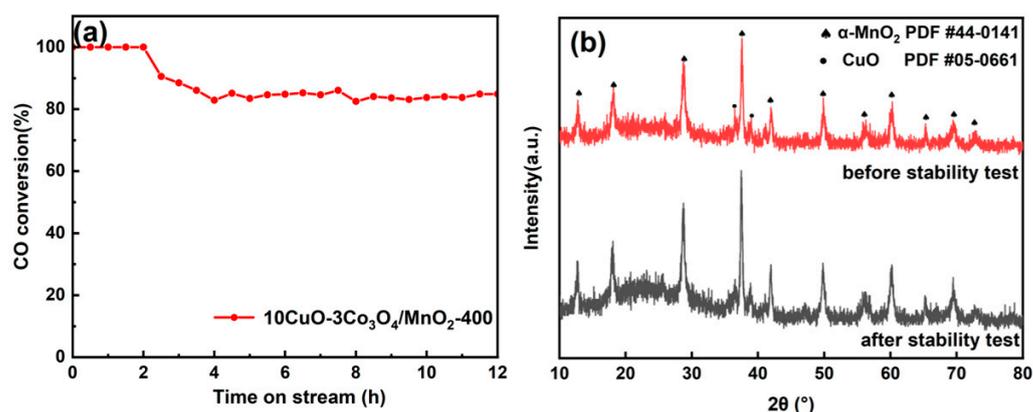


Figure 9. (a) The 12 h stability test of CO oxidation over the 10CuO-3Co₃O₄/α-MnO₂-400 catalyst under the conditions: 80 °C, CO/O₂/N₂ = 1/20/79, GHSV = 12,000 mL·g⁻¹·h⁻¹, 1 atm; (b) The XRD patterns of the fresh and the spent 10CuO-3Co₃O₄/α-MnO₂-400 catalysts before and after the stability test.

The XRD pattern of the 10CuO-3Co₃O₄/α-MnO₂-400 catalyst after the 12 h stability test is shown in Figure 9b. It can be observed that the diffraction peaks of the spent 10CuO-3Co₃O₄/α-MnO₂-400 catalyst after the 12 h stability test were a bit narrower and sharper than the fresh catalyst before the stability test. The possible reason is that the 10CuO-3Co₃O₄/α-MnO₂-400 catalyst underwent a bit of thermal agglomeration of the CuO active sites and the α-MnO₂ nanowire support during the catalytic process due to the hot spots of the catalyst bed, which could partly reduce the stability of the catalyst.

3. Discussion

The oxidation of CO on Cu-doped MnO₂ follows the Mars–van Krevelen (MvK) mechanism [35]. The reaction is divided into two parts: CO is first adsorbed on the catalyst surface and then reacts with surface-active oxygen on the catalyst surface to produce CO₂, which then generates oxygen vacancies on the catalyst surface. O₂ replenishes the depleted surface-active oxygen. After these two processes, the reaction completes a cycle [59]. There are redox electron pairs in the Cu-doped MnO₂ catalyst: -Cu²⁺-O²⁻-Mn⁴⁺ → -Cu⁺-□-Mn³⁺ + 1/2O₂ (□ indicates the oxygen vacancy) [45]. The content of Mn³⁺ on the MnO₂ catalyst is higher, presumably with more oxygen vacancies on the MnO₂ [60].

The order of the oxygen vacancy content of the catalysts is $10\text{CuO}-3\text{Co}_3\text{O}_4/\alpha\text{-MnO}_2 > 10\text{CuO}-3\text{NiO}/\alpha\text{-MnO}_2 > 10\text{CuO}/\alpha\text{-MnO}_2 > 10\text{CuO}-3\text{Fe}_2\text{O}_3/\alpha\text{-MnO}_2$. The oxygen vacancy content is consistent with the Mn^{3+} content and catalyst activity. O_2 is commonly activated near the oxygen vacancy, producing surface active oxygen species (O_{sur}) [61]. It is well known that the higher the surface oxygen vacancy, the more easily O_2 is activated to reactive oxygen species [50]. However, the XPS spectra of O 1s showed that the order of the $\text{O}_{\text{ads}}/\text{O}_{\text{latt}}$ ratios was not consistent with the oxygen vacancy content and catalyst activity. To evaluate the intrinsic activity of the catalysts, the surface area normalized reaction rates are determined. The results of the surface area normalized reaction rates show that the loading of NiO has the greatest effect on the intrinsic activity of the catalyst. The $10\text{CuO}-3\text{NiO}/\alpha\text{-MnO}_2$ catalyst did not exhibit superior catalytic activity because the specific surface area of the catalyst after NiO loading was reduced, and the effect of specific surface area on the activity of the CO oxidation reaction could not be ignored. The $10\text{CuO}-3\text{NiO}/\alpha\text{-MnO}_2$ catalyst has the most surface adsorbed oxygen and reaction rates per unit surface area, but the small specific surface area results in the catalytic activity being similar to that of the pristine $10\text{CuO}/\alpha\text{-MnO}_2-400$ catalyst. The order of the intrinsic activity of the catalyst is consistent with the order of the $\text{O}_{\text{ads}}/\text{O}_{\text{latt}}$ ratio ($10\text{CuO}-3\text{NiO}/\alpha\text{-MnO}_2 > 10\text{CuO}/\alpha\text{-MnO}_2 > 10\text{CuO}-3\text{Co}_3\text{O}_4/\alpha\text{-MnO}_2 > 10\text{CuO}-3\text{Fe}_2\text{O}_3/\alpha\text{-MnO}_2$). This suggests that the surface-adsorbed oxygen is the reactive oxygen species involved in the oxidation of CO. CO_2 was produced by CO reacting with surface-adsorbed oxygen species [62,63]. After loading different transition metals, the catalysts form different types of oxygen vacancies, which have different electron densities and affect the production of reactive oxygen species [39]. The oxygen vacancies of $10\text{CuO}-3\text{Co}_3\text{O}_4/\alpha\text{-MnO}_2$ did not activate O_2 as well, and the $10\text{CuO}-3\text{Co}_3\text{O}_4/\alpha\text{-MnO}_2$ catalyst did not form more surface adsorbed oxygen. This might be the cause of the $10\text{CuO}-3\text{Co}_3\text{O}_4/\alpha\text{-MnO}_2$ catalyst providing the highest oxygen vacancy but poor surface-adsorbed oxygen. Although the normalized reaction rate of the $10\text{CuO}-3\text{Co}_3\text{O}_4/\alpha\text{-MnO}_2$ catalyst is not the highest, its high specific surface area allows for a greater number of oxygen vacancies. The large number of oxygen vacancies of the $10\text{CuO}-3\text{Co}_3\text{O}_4/\alpha\text{-MnO}_2$ catalyst counteracted the low activity of the oxygen vacancies and performed the high catalytic activity of CO oxidation. Therefore, the $10\text{CuO}-3\text{Co}_3\text{O}_4/\alpha\text{-MnO}_2$ catalyst exhibited the highest activity owing to its maximum specific surface area. The activity of the $10\text{CuO}-3\text{Co}_3\text{O}_4/\alpha\text{-MnO}_2$ catalyst was slightly higher than that of the $10\text{CuO}-3\text{NiO}/\alpha\text{-MnO}_2$ catalyst.

4. Materials and Method

4.1. Synthesis of $\alpha\text{-MnO}_2$ Nanowire Supports

The $\alpha\text{-MnO}_2$ nanowire support was facilely prepared via the typical hydrothermal method. Typically, 2 mmol KMnO_4 (Sinopharm Chemical Reagent Co., Ltd., Shanghai, China, AR, >99.5%) and 3 mmol $\text{MnSO}_4\cdot\text{H}_2\text{O}$ (Shanghai Aladdin Biochemical Technology Co., Ltd., Shanghai, China, AR, 99%) were absolutely dissolved in 40 mL deionized water and vigorously stirred for 5 min, respectively. Then, these two solutions were mixed together by adding the KMnO_4 solution into the MnSO_4 solution. After stirring for 30 min, the obtained suspension was transferred to the 100 mL Teflon-lined stainless-steel autoclave. The autoclave was kept at 160 °C for 12 h. The obtained precipitate after the hydrothermal reaction was separated by the centrifuge and washed with ethanol (Sinopharm Chemical Reagent Co., Ltd., Shanghai, China, AR) and deionized water for three times. Then, the final powder was dried in the 100 °C oven overnight after the centrifugation. The obtained $\alpha\text{-MnO}_2$ was used as a support for subsequent experiments.

4.2. The Preparation of the Transition Metal Oxide (Fe_2O_3 , Co_3O_4 , NiO)-Doped CuO-Based Catalysts Supported on the $\alpha\text{-MnO}_2$ Nanowire

The transition metal oxides promoted CuO-based $\alpha\text{-MnO}_2$ nanowire supported catalysts were prepared by the deposition precipitation method as reported in our previous work [64]. The weight percentages of the CuO and transition metal oxides were controlled

at $x\%$ and $y\%$ ($x\% = m_{\text{CuO}} / (m_{\text{CuO}} + m_{\text{MOx}} + m_{\text{support}}) \times 100\%$; $y\% = m_{\text{MOx}} / (m_{\text{CuO}} + m_{\text{MOx}} + m_{\text{support}}) \times 100\%$) by using the $\text{Cu}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ (Shanghai Xinbao Fine Chemical Industry Factory, Shanghai, China, AR, >99.5%), $\text{Fe}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$ (Shanghai Macklin Bio-Chem Co., Ltd., Shanghai, China, AR, 98.5%), $\text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ (Shanghai Macklin Bio-Chem Co., Ltd., Shanghai, China, AR, 99%), $\text{Ni}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ (Sinopharm Chemical Reagent Co., Ltd., Shanghai, China, AR, 99%) as the precursors. For the specific procedure, the $\alpha\text{-MnO}_2$ nanowire was firstly dispersed in a $\text{Cu}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ and $\text{Fe}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O} / \text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O} / \text{Ni}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ aqueous solution. Then, the Na_2CO_3 (0.01M, Shanghai Ling Feng Chemical Reagent Co., Ltd., Shanghai, China, AR, >99.8%) solution was added by droplet into the above solution to adjust the pH to 8–9 under vigorously stirring. The resultant solution was stirred for 30 min and then kept still for 1 h. The solid powder was separated by filtration and washed by the deionized water. The obtained catalyst was dried at $120\text{ }^\circ\text{C}$ in an oven overnight and then calcinated at $400\text{ }^\circ\text{C}$ for 5 h with a heating rate of $2\text{ }^\circ\text{C}/\text{min}$. The $\alpha\text{-MnO}_2$ nanowire-supported catalysts with 10 wt.% CuO and 3 wt.% MOx (MOx = Fe_2O_3 , Co_3O_4 , NiO) were denoted as the $10\text{CuO-3MOx}/\alpha\text{-MnO}_2$ (MOx = Fe_2O_3 , Co_3O_4 , NiO). Meanwhile, the loading amount of the Co_3O_4 (wt.%) was subsequently changed in the same way. The obtained catalysts with loading amounts of 10 wt.% CuO and y wt.% Co_3O_4 were denoted as the $10\text{CuO-}y\text{Co}_3\text{O}_4/\alpha\text{-MnO}_2$ ($y = 0, 1, 3, 5, 7, 10$). As for the influence of the calcination temperature, the catalysts with loading of the 10 wt.% CuO and 3 wt.% Co_3O_4 were calcinated at different temperatures and named $10\text{CuO-3Co}_3\text{O}_4/\alpha\text{-MnO}_2\text{-T}$ ($T = 200, 400, 500, 600\text{ }^\circ\text{C}$).

4.3. Catalyst Characterizations

The X-ray diffraction (XRD) patterns of the catalysts were recorded on an X-ray power diffractometer (XRD-6100) from the Shimadzu Corporation by using the Cu K α rays (0.15046 nm), 40 KV tube voltage, and 40 mA tube current. The 2θ scanning range was from 10° to 80° , and the scanning speed was controlled at $3^\circ/\text{min}$. The scanning electron microscope (SEM) images were taken on an Apreo S Hivac instrument (Thermo Fisher Science, Waltham, MA, USA) with the accelerating voltage of 5 kV. The nitrogen physisorption was performed on an Autosorb-iQ-AG-MP instrument (Quantachrome, Boynton Beach, FL, USA) at liquid nitrogen temperature ($-196\text{ }^\circ\text{C}$). The samples were degassed at $300\text{ }^\circ\text{C}$ for 3 h to remove the surface-adsorbed water and impurities before the regular test. The specific surface areas of the catalysts were calculated by the Brunauer-Emmett-Teller (BET) method, and the pore size distribution and pore volume were calculated from the adsorption branch of the isotherm by the Barrett-Joyner-Halenda (BJH) method in the range of 0–1.0 P/P₀. The X-ray photoelectron spectroscopy (XPS) measurements were performed on a Thermo Science K-Alpha + spectrometer (Thermo Fisher Science, Waltham, MA, USA). For the XPS measurement, the penetration depth of each catalyst was about 1–2 nm. The binding energy (BE) was calibrated by using C 1s = 284.8 eV as the standard. A fixed-bed reactor was used to conduct the hydrogen temperature-programmed reduction ($\text{H}_2\text{-TPR}$) experiment. The hydrogen consumption profile was recorded and analyzed with the online LC-D200 mass spectrometer (TILON GRP TECHNOLOGY LIMITED, Shanghai, China). The mixture of H_2 (1.2 mL/min) and Ar (23.7 mL/min) was introduced into the reactor. For each test, 100 mg of catalyst was loaded. After the hydrogen signal baseline line ($m/z = 2$) was stabilized, the $\text{H}_2\text{-TPR}$ test was performed with a heating rate of $20\text{ }^\circ\text{C}/\text{min}$ from room temperature to $800\text{ }^\circ\text{C}$.

4.4. Catalyst Evaluation

The activity of the catalysts was evaluated in a fixed-bed reactor equipped with a quartz tube (i.d. = 8.00 mm). For each test, 100 mg catalyst was placed in the center of the quartz tube. The reactant gases, with a composition of 1% CO , 21% O_2 , 79% N_2 (20 mL/min), were introduced into the reactor with the gas hourly space velocity (GHSV) of $12,000\text{ mL}\cdot\text{g}^{-1}\cdot\text{h}^{-1}$. The catalytic activities of CO oxidation over different catalysts were tested in the specified temperature range. Each catalyst was tested three times at each

temperature. The outlet gases were analyzed online by using the GC-680 Perkin Elmer gas chromatograph equipped with the thermal conductivity detector (TCD). The 24 h stability tests of catalysts were carried out at 80 °C with the GHSV of 12,000 mL·g⁻¹·h⁻¹.

The catalytic performances of CO oxidation over these catalysts were stated in the form of the CO conversion. The calculated formula is listed below:

$$X_{\text{CO}} = (C_{\text{CO, Inlet}} - C_{\text{CO, Outlet}}) / C_{\text{CO, Inlet}} \times 100\% \quad (1)$$

where X_{CO} is the conversion rate of CO; and $C_{\text{CO, Inlet}}$ (ppm), and $C_{\text{CO, Outlet}}$ (ppm) are CO flowing into and out of the reactor, respectively.

In order to evaluate the intrinsic rate of CO oxidation on these catalysts, the calculated formula of the specific surface area normalization reaction rate is listed below [25]:

$$r_{\text{norm}} \left(\text{mol} \cdot \text{m}^{-2} \cdot \text{s}^{-1} \right) = \frac{C_{\text{inlet}} \cdot F}{m_{\text{cat}} \cdot S_{\text{BET}}} \cdot \ln \left(\frac{1}{1 - X_{\text{CO}}} \right) \quad (2)$$

where r_{norm} (mol·m⁻²·s⁻¹) is the normalized reaction rate, F (mol·s⁻¹) is the CO flow rate, m_{cat} (g) is the mass of catalyst, and S_{BET} (m²·g⁻¹) is the BET surface area.

5. Conclusions

In this work, the novel α -MnO₂ nanowire was prepared by the one-step hydrothermal method. A series of transition metal oxides (Fe₂O₃, Co₃O₄, NiO) promoted the CuO-based catalyst supported on the α -MnO₂ nanowire and were synthesized by the co-precipitation method for the CO oxidation reaction. The effects of the transition metal oxide type, the loading amount, and the calcination temperature on the CO oxidation reaction had been systematically investigated. It was found that the 10CuO-3Co₃O₄/ α -MnO₂-400 catalyst showed the highest reactivity with $T_{90} = 75$ °C. It was found that the 10CuO-3Co₃O₄/ α -MnO₂-400 catalyst possessed the largest specific surface area and exposed more active sites, which could further enhance the catalytic activity. The 10CuO-3NiO/ α -MnO₂-400 catalyst had the highest surface-adsorbed oxygen content and normalized reaction rate. This indicated that the surface-adsorbed oxygen was the surface-active oxygen involved in the oxidation of CO. It was found that the 10CuO-3Co₃O₄/ α -MnO₂-400 catalyst suffered from some deactivation during the 12 h stability test, which might be caused by the thermal sintering and agglomeration of the CuO active sites and α -MnO₂ nanowire support. This should be the key consideration when designing the CuO-based CO oxidation catalyst in the future.

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Data Availability Statement: The data presented in this study are available on request from the corresponding author. The data are not publicly available due to privacy.

Conflicts of Interest: The authors declare no conflict of interest.

References

1. Schnitzhofer, R.; Beauchamp, J.; Dunkl, J.; Wisthaler, A.; Weber, A.; Hansel, A. Long-term measurements of CO, NO, NO₂, benzene, toluene and PM10 at a motorway location in an Austrian valley. *Atmos. Environ.* **2008**, *42*, 1012–1024. [[CrossRef](#)]
2. Kunkalekar, R.K.; Salker, A.V. Low temperature carbon monoxide oxidation over nanosized silver doped manganese dioxide catalysts. *Catal. Commun.* **2010**, *12*, 193–196. [[CrossRef](#)]
3. Zheng, B.; de Beurs, K.M.; Owsley, B.C.; Henebry, G.M. Scaling relationship between CO pollution and population size over major US metropolitan statistical areas. *Landsc. Urban Plan.* **2019**, *187*, 191–198. [[CrossRef](#)]
4. Wu, X.; Lang, J.; Sun, Z.; Jin, F.; Hu, Y.H. Photocatalytic conversion of carbon monoxide: From pollutant removal to fuel production. *Appl. Catal. B Environ.* **2021**, *295*, 120312. [[CrossRef](#)]
5. Yoo, S.; Lee, E.; Jang, G.H.; Kim, D.H. Effect of Pd precursors on the catalytic properties of Pd/CeO₂ catalysts for CH₄ and CO oxidation. *Mol. Catal.* **2022**, *533*, 112791. [[CrossRef](#)]
6. Schubert, M.M.; Hackenberg, S.; van Veen, A.C.; Muhler, M.; Plzak, V.; Behm, R.J. CO Oxidation over Supported Gold Catalysts—“Inert” and “Active” Support Materials and Their Role for the Oxygen Supply during Reaction. *J. Catal.* **2001**, *197*, 113–122. [[CrossRef](#)]
7. Min, B.K.; Friend, C.M. Heterogeneous Gold-Based Catalysis for Green Chemistry: Low-Temperature CO Oxidation and Propene Oxidation. *ChemInform* **2007**, *107*, 2709–2724. [[CrossRef](#)]
8. Qiao, B.; Wang, A.; Yang, X.; Allard, L.F.; Jiang, Z.; Cui, Y.; Liu, J.; Li, G.; Zhang, T. Single-atom catalysis of CO oxidation using Pt₁/FeOx. *Nat. Chem.* **2011**, *3*, 634–641. [[CrossRef](#)] [[PubMed](#)]
9. Chen, G.; Zhao, Y.; Fu, G.; Duchesne, P.; Gu, L.; Zheng, Y.; Weng, X.; Chen, M.; Pao, C. Interfacial effects in iron-nickel hydroxide-platinum nanoparticles enhance catalytic oxidation. *Science* **2014**, *344*, 495–499. [[CrossRef](#)]
10. Liu, K.; Wang, A.; Zhang, T. Recent Advances in Preferential Oxidation of CO Reaction over Platinum Group Metal Catalysts. *ACS Catal.* **2012**, *2*, 1165–1178. [[CrossRef](#)]
11. Zhang, L.; Liu, H.; Huang, X.; Sun, X.; Jiang, Z.; Schlögl, R.; Su, D. Stabilization of Palladium Nanoparticles on Nanodiamond-Graphene Core-Shell Supports for CO Oxidation. *Angew. Chem.* **2015**, *127*, 16049–16052. [[CrossRef](#)]
12. Gonzalez-A, E.; Rangel, R.; Solís-García, A.; Venezia, A.M.; Zepeda, T.A. FTIR investigation under reaction conditions during CO oxidation over Ru(x)-CeO₂ catalysts. *Mol. Catal.* **2020**, *493*, 111086. [[CrossRef](#)]
13. Yadav, P.K.; Kumari, S.; Naveena, U.; Deshpande, P.A.; Sharma, S. Insights into the substitutional chemistry of La_{1-x}Sr_xCo_{1-y}MyO₃ (M = Pd, Ru, Rh, and Pt) probed by in situ DRIFTS and DFT analysis of CO oxidation. *Appl. Catal. A Gen.* **2022**, *643*, 118768. [[CrossRef](#)]
14. Bunluesin, T.; Cordatos, H.; Gorte, R.J. Study of CO Oxidation Kinetics on Rh/Ceria. *J. Catal.* **1995**, *157*, 222–226. [[CrossRef](#)]
15. Dey, S.; Dhal, G.C. Applications of silver nanocatalysts for low-temperature oxidation of carbon monoxide. *Inorg. Chem. Commun.* **2019**, *110*, 107614. [[CrossRef](#)]
16. Du, Y.; Meng, Q.; Wang, J.; Yan, J.; Fan, H.; Liu, Y.; Dai, H. Three-dimensional mesoporous manganese oxides and cobalt oxides: High-efficiency catalysts for the removal of toluene and carbon monoxide. *Micropor. Mesopor. Mat.* **2012**, *162*, 199–206. [[CrossRef](#)]
17. Xiao, M.; Zhang, X.; Yang, Y.; Cui, X.; Chen, T.; Wang, Y. M (M = Mn, Co, Cu)-CeO₂ catalysts to enhance their CO catalytic oxidation at a low temperature: Synergistic effects of the interaction between Ce³⁺-M^{x+}-Ce⁴⁺ and the oxygen vacancy defects. *Fuel* **2022**, *323*, 124379. [[CrossRef](#)]
18. Bi, F.; Zhang, X.; Xiang, S.; Wang, Y. Effect of Pd loading on ZrO₂ support resulting from pyrolysis of UiO-66: Application to CO oxidation. *J. Colloid Interf. Sci.* **2020**, *573*, 11–20. [[CrossRef](#)] [[PubMed](#)]
19. Dey, S.; Dhal, G.C. Catalytic Conversion of Carbon Monoxide into Carbon Dioxide over Spinel Catalysts: An Overview. *Mater. Sci. Energy Technol.* **2019**, *2*, 575–588. [[CrossRef](#)]
20. Kaplin, I.Y.; Lokteva, E.S.; Golubina, E.V.; Shishova, V.V.; Maslakov, K.I.; Fionov, A.V.; Isaikina, O.Y.; Lunin, V.V. Efficiency of manganese modified CTAB-templated ceria-zirconia catalysts in total CO oxidation. *Appl. Surf. Sci.* **2019**, *485*, 432–440. [[CrossRef](#)]
21. Niu, X.; Lei, Z. Copper doped manganese oxides to produce enhanced catalytic performance for CO oxidation. *J. Environ. Chem. Eng.* **2019**, *7*, 103055. [[CrossRef](#)]
22. Liu, H.; Li, X.; Dai, Q.; Zhao, H.; Chai, G.; Guo, Y.; Guo, Y.; Wang, L.; Zhan, W. Catalytic oxidation of chlorinated volatile organic compounds over Mn-Ti composite oxides catalysts: Elucidating the influence of surface acidity. *Appl. Catal. B Environ.* **2021**, *282*, 119577. [[CrossRef](#)]
23. Ye, Z.; Giraudon, J.-M.; Nuns, N.; Simon, P.; De Geyter, N.; Morent, R.; Lamonier, J.-F. Influence of the preparation method on the activity of copper-manganese oxides for toluene total oxidation. *Appl. Catal. B Environ.* **2018**, *223*, 154–166. [[CrossRef](#)]
24. Lin, Y.; Tian, H.; Qian, J.; Yu, M.; Hu, T.; Lassi, U.; Chen, Z.; Wu, Z. Biocarbon-directed vertical δ-MnO₂ nanoflakes for boosting lithium-ion diffusion kinetics. *Mater. Today Chem.* **2022**, *26*, 101023. [[CrossRef](#)]
25. Yang, W.; Su, Z.; Xu, Z.; Yang, W.; Peng, Y.; Li, J. Comparative study of α-, β-, γ- and δ-MnO₂ on toluene oxidation: Oxygen vacancies and reaction intermediates. *Appl. Catal. B Environ.* **2019**, *260*, 118150. [[CrossRef](#)]
26. Hayashi, E.; Yamaguchi, Y.; Kamata, K.; Tsunoda, N.; Kumagai, Y.; Oba, F.; Hara, M. Effect of MnO₂ Crystal Structure on Aerobic Oxidation of 5-Hydroxymethylfurfural to 2,5-Furandicarboxylic Acid. *J. Am. Chem. Soc.* **2019**, *141*, 890–900. [[CrossRef](#)] [[PubMed](#)]
27. Robinson, D.M.; Go, Y.B.; Mui, M.; Gardner, G.; Zhang, Z.; Mastrogiovanni, D.; Garfunkel, E.; Li, J.; Greenblatt, M.; Dismukes, G.C. Photochemical Water Oxidation by Crystalline Polymorphs of Manganese Oxides: Structural Requirements for Catalysis. *J. Am. Chem. Soc.* **2013**, *135*, 3494–3501. [[CrossRef](#)] [[PubMed](#)]

28. Liang, S.; Teng, F.; Bulgan, G.; Zong, R.; Zhu, Y. Effect of Phase Structure of MnO₂ Nanorod Catalyst on the Activity for CO Oxidation. *J. Phys. Chem. C* **2008**, *112*, 5307–5315. [[CrossRef](#)]
29. Tian, F.; Li, H.; Zhu, M.; Tu, W.; Lin, D.; Han, Y. Effect of MnO₂ Polymorphs' Structure on Low-Temperature Catalytic Oxidation: Crystalline Controlled Oxygen Vacancy Formation. *ACS Appl. Mater. Interfaces* **2022**, *14*, 18525–18538. [[CrossRef](#)] [[PubMed](#)]
30. Yang, R.; Fan, Y.; Ye, R.; Tang, Y.; Cao, X.; Yin, Z.; Zeng, Z. MnO₂-Based Materials for Environmental Applications. *Adv. Mater.* **2021**, *33*, 2004862. [[CrossRef](#)] [[PubMed](#)]
31. Jokar, R.; Alavi, S.M.; Rezaei, M.; Akbari, E. Catalytic performance of copper oxide supported α -MnO₂ nanowires for the CO preferential oxidation in H₂-rich stream. *Int. J. Hydrogen Energy* **2021**, *46*, 32503–32513. [[CrossRef](#)]
32. Qian, K.; Qian, Z.; Hua, Q.; Jiang, Z.; Huang, W. Structure–activity relationship of CuO/MnO₂ catalysts in CO oxidation. *Appl. Surf. Sci.* **2013**, *273*, 357–363. [[CrossRef](#)]
33. Sun, R.L.; Zhang, S.R.; An, K.; Song, P.F.; Liu, Y. Cu_{1.5}Mn_{1.5}O₄ spinel type composite oxide modified with CuO for synergistic catalysis of CO oxidation. *J. Fuel Chem. Technol.* **2021**, *49*, 799–808. [[CrossRef](#)]
34. Sophiana, I.C.; Topandi, A.; Iskandar, F.; Devianto, H.; Nishiyama, N.; Budhi, Y.W. Catalytic oxidation of benzene at low temperature over novel combination of metal oxide based catalysts: CuO, MnO₂, NiO with Ce_{0.75}Zr_{0.25}O₂ as support. *Mater. Today Chem.* **2020**, *17*, 100305. [[CrossRef](#)]
35. Gao, J.; Jia, C.; Zhang, L.; Wang, H.; Yang, Y.; Hung, S.-F.; Hsu, Y.Y.; Liu, B. Tuning chemical bonding of MnO₂ through transition-metal doping for enhanced CO oxidation. *J. Catal.* **2016**, *341*, 82–90. [[CrossRef](#)]
36. Ivanov, K.I.; Kolentsova, E.N.; Dimitrov, D.Y.; Petrova, P.T.; Tabakova, T.T. Alumina Supported Cu-Mn-Cr Catalysts for CO and VOCs oxidation. *Int. Sch. Sci. Res. Innov.* **2015**, *9*, 605–612.
37. Krämer, M.; Schmidt, T.; Stöwe, K.; Maier, W.F. Structural and catalytic aspects of sol–gel derived copper manganese oxides as low-temperature CO oxidation catalyst. *Appl. Catal. A Gen.* **2006**, *302*, 257–263. [[CrossRef](#)]
38. Huang, N.; Qu, Z.; Dong, C.; Qin, Y.; Duan, X. Superior performance of α @ β -MnO₂ for the toluene oxidation: Active interface and oxygen vacancy. *Appl. Catal. A Gen.* **2018**, *560*, 195–205. [[CrossRef](#)]
39. Wang, Y.; Wu, J.; Wang, G.; Yang, D.; Ishihara, T.; Guo, L. Oxygen vacancy engineering in Fe doped akhtenskite-type MnO₂ for low-temperature toluene oxidation. *Appl. Catal. B Environ.* **2021**, *285*, 119873. [[CrossRef](#)]
40. Rosso, J.J.; Hochella, M.F. Natural Iron and Manganese Oxide Samples by XPS. *Surf. Sci. Spectra* **1996**, *4*, 253–265. [[CrossRef](#)]
41. Li, J.; Zhu, P.; Zuo, S.; Huang, Q.; Zhou, R. Influence of Mn doping on the performance of CuO-CeO₂ catalysts for selective oxidation of CO in hydrogen-rich streams. *Appl. Catal. A Gen.* **2010**, *381*, 261–266. [[CrossRef](#)]
42. Michael, A.S. Mn₂O₃ by XPS. *Surf. Sci. Spectra* **1999**, *6*, 39–46. [[CrossRef](#)]
43. Liu, W.; Wang, S.; Cui, R.; Song, Z.; Zhang, X. Enhancement of catalytic combustion of toluene over CuMnO_x hollow spheres prepared by oxidation method. *Micropor. Mesopor. Mat.* **2021**, *326*, 111370. [[CrossRef](#)]
44. Wang, Y.; Liu, K.; Wu, J.; Hu, Z.; Huang, L.; Zhou, J.; Ishihara, T.; Guo, L. Unveiling the Effects of Alkali Metal Ions Intercalated in Layered MnO₂ for Formaldehyde Catalytic Oxidation. *ACS Catal.* **2020**, *10*, 10021–10031. [[CrossRef](#)]
45. Wang, Y.; Yang, D.; Li, S.; Zhang, L.; Zheng, G.; Guo, L. Layered copper manganese oxide for the efficient catalytic CO and VOCs oxidation. *Chem. Eng. J.* **2019**, *357*, 258–268. [[CrossRef](#)]
46. Zeng, J.; Xie, H.; Zhang, H.; Huang, M.; Liu, X.; Zhou, G.; Jiang, Y. Insight into the effects of oxygen vacancy on the toluene oxidation over α -MnO₂ catalyst. *Chemosphere* **2022**, *291*, 135890. [[CrossRef](#)] [[PubMed](#)]
47. Xu, J.; Wu, W.; Gao, E.; Zhu, J.; Yao, S.; Li, J. Revealing the role of oxygen vacancies on α -MnO₂ of different morphologies in CO oxidation using operando DRIFTS-MS. *Appl. Surf. Sci.* **2023**, *618*, 156643. [[CrossRef](#)]
48. Liu, W.; Xiang, W.; Chen, X.; Song, Z.; Gao, C.; Tsubaki, C.; Zhang, X. A novel strategy to adjust the oxygen vacancy of CuO/MnO₂ catalysts toward the catalytic oxidation of toluene. *Fuel* **2022**, *312*, 122975. [[CrossRef](#)]
49. Yang, W.; Zhu, Y.; You, F.; Yan, L.; Ma, Y.; Lu, C.; Gao, P.; Hao, Q.; Li, W. Insights into the surface-defect dependence of molecular oxygen activation over birnessite-type MnO₂. *Appl. Catal. B Environ.* **2018**, *233*, 184–193. [[CrossRef](#)]
50. Mo, S.; Zhang, Q.; Li, J.; Sun, Y.; Ren, Q.; Zou, S.; Zhang, Q.; Lu, J.; Fu, M.; Mo, D.; et al. Highly efficient mesoporous MnO₂ catalysts for the total toluene oxidation: Oxygen-Vacancy defect engineering and involved intermediates using in situ DRIFTS. *Appl. Catal. B Environ.* **2020**, *264*, 118464. [[CrossRef](#)]
51. Wang, J.; Li, J.; Jiang, C.; Zhou, P.; Zhang, P.; Yu, J. The effect of manganese vacancy in birnessite-type MnO₂ on room-temperature oxidation of formaldehyde in air. *Appl. Catal. B Environ.* **2017**, *204*, 147–155. [[CrossRef](#)]
52. Papavasiliou, J.; Avgouropoulos, G.; Ioannides, T. Combined steam reforming of methanol over Cu–Mn spinel oxide catalysts. *J. Catal.* **2007**, *251*, 7–20. [[CrossRef](#)]
53. Liu, T.; Yao, Y.; Wei, L.; Shi, Z.; Han, L.; Yuan, H.; Li, B.; Dong, L.; Wang, F.; Sun, C. Preparation and Evaluation of Copper–Manganese Oxide as a High-Efficiency Catalyst for CO Oxidation and NO Reduction by CO. *J. Phys. Chem. C* **2017**, *121*, 12757–12770. [[CrossRef](#)]
54. Yuan, H.L.; Wang, Y.Q.; Zhou, S.M.; Liu, L.S.; Chen, X.L.; Lou, S.Y.; Yuan, R.J.; Hao, Y.M.; Li, N. Low-Temperature Preparation of Superparamagnetic CoFe₂O₄ Microspheres with High Saturation Magnetization. *Nanoscale Res. Lett.* **2010**, *5*, 1817–1821. [[CrossRef](#)] [[PubMed](#)]
55. Wu, Y.; Guo, Q.; Liu, H.; Wei, S.; Wang, L. Effect of Fe doping on the surface properties of δ -MnO₂ nanomaterials and its decomposition of formaldehyde at room temperature. *J. Environ. Chem. Eng.* **2022**, *10*, 108277. [[CrossRef](#)]

56. Yadav, R.S.; Havlica, J.; Masilko, J.; Kalina, L.; Wasserbauer, J.; Hajdúchová, M.; Enev, V.; Kuřitka, I.; Kožáková, Z. Impact of Nd^{3+} in CoFe_2O_4 spinel ferrite nanoparticles on cation distribution, structural and magnetic properties. *J. Magn. Magn. Mater.* **2016**, *399*, 109–117. [[CrossRef](#)]
57. Hoang, M.; Hughes, A.E.; Turney, T.W. An XPS study of Ru-promotion for Co/CeO₂ Fischer-Tropsch catalyst. *Appl. Surf. Sci.* **1993**, *72*, 55–65. [[CrossRef](#)]
58. Oh, Y.S.; Roh, H.S.; Jun, K.W.; Baek, Y.S. A highly active catalyst, Ni/Ce–ZrO₂/θ-Al₂O₃, for on-site H₂ generation by steam methane reforming: Pretreatment effect. *Int. J. Hydrogen Energy* **2003**, *28*, 1387–1392. [[CrossRef](#)]
59. Hu, X.; Chen, J.; Li, S.; Chen, Y.; Qu, W.; Ma, Z.; Tang, X. The Promotional Effect of Copper in Catalytic Oxidation by Cu-Doped α-MnO₂ Nanorods. *J. Phys. Chem. C* **2019**, *124*, 701–708. [[CrossRef](#)]
60. Dong, C.; Wang, H.; Ren, Y.; Qu, Z. Effect of alkaline earth metal promoter on catalytic activity of MnO₂ for the complete oxidation of toluene. *J. Environ. Sci.* **2021**, *104*, 102–112. [[CrossRef](#)] [[PubMed](#)]
61. Feng, C.; Jiang, F.; Xiong, G.; Chen, C.; Wang, Z.; Pan, Y.; Fei, Z.; Lu, Y.; Li, X.; Zhang, R.; et al. Revelation of Mn^{4+} -O_{sur}- Mn^{3+} active site and combined Langmuir-Hinshelwood mechanism in propane total oxidation at low temperature over MnO₂. *Chem. Eng. J.* **2023**, *451*, 138868. [[CrossRef](#)]
62. Park, J.H.; Kang, D.C.; Park, S.J.; Shin, C.H. CO oxidation over MnO₂ catalysts prepared by a simple redox method: Influence of the Mn (II) precursors. *J. Ind. Eng. Chem.* **2015**, *25*, 250–257. [[CrossRef](#)]
63. Du, C.; Zhao, S.; Tang, X.; Yu, Q.; Gao, F.; Zhou, Y.; Liu, J.; Yi, H. Synthesis of ultra-thin 2D MnO₂ nanosheets rich in oxygen vacancy defects and the catalytic oxidation of n-heptane. *Appl. Surf. Sci.* **2022**, *606*, 154846. [[CrossRef](#)]
64. Cui, Y.; Song, H.; Shi, Y.; Ge, P.X.; Chen, M.D.; Xu, L.L. Enhancing the Low-Temperature CO Oxidation over CuO-Based α-MnO₂ Nanowire Catalysts. *Nanomaterials* **2022**, *12*, 2083. [[CrossRef](#)] [[PubMed](#)]

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