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Hydrothermally Derived Mg-Doped TiO₂ Nanostructures for Enhanced H₂ Evolution Using Photo- and Electro-Catalytic Water Splitting

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Abstract: Mg-doped TiO₂ nano-structures in different compositions (1, 2.5 and 5%) were successfully synthesized by low-temperature hydrothermal route. X-ray diffraction and electron microscopic studies were used to investigate the crystal structure, surface morphology and particle size of the as-synthesized materials. Raman studies were carried out to elucidate the phase identification and the modes of vibrations to determine the impact of dopant ion on the crystal structures. The band gap was estimated using UV-DRS studies whereas, BET surface area analysis revealed an increase in the surface area of increasing Mg²⁺ ions concentration in TiO₂ nanostructures. Among the synthesized various composition of nano-structures, 5% Mg-doped TiO₂ photocatalyst showed maximum hydrogen evolution activity (38.96 mmol g_{cat}⁻¹) in an 8-h analysis cycle. Moreover, the 2.5% Mg-doped TiO₂ nanocatalyst with tafel slopes of 123.5 and 126.7 mV/dec showed strong activity for both HER in 0.5 N H₂SO₄ and 0.1 N KOH, with an onset potential of 0.96 V (at 10 mA/cm²) and -1.38 V (at 1 mA/cm²) for HER, respectively. Experimental investigations deduced that the incorporation of Mg²⁺ ions in the TiO₂ resulted in the increase of hydrogen generation catalytic activity of titanium dioxide owing to the synergistic effect provided by the remarkable surface area and the presence of defects introduced by doping.

Keywords: nanocatalysts; Mg-doped TiO₂; water splitting; hydrogen fuel; photocatalysis; electrocatalysis



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1. Introduction

Population expansion and unregulated industrial growth have increased energy consumption, and as a consequence, hazardous chemicals and waste released into the air and rivers have caused pollution-related diseases, global warming and anomalous climate shifts [1–6]. Environmental pollution and insufficient clean, natural energy sources have become the world's most pressing issues due to their long-term disastrous effects [7–10]. Due to the unabated production of these high-level nuclear wastes, several problems have arisen, which have affected countries like, France, UK, and Germany etc. [11]. Moreover, the March 2011 earthquake and tsunami in Fukushima, Japan, damaged the nuclear reactors, raising concerns about electricity supply and environmental damage from nuclear waste and coolants. Therefore, researchers should make it their objective to aid in the development of chemical technologies by developing energy producing pathways that are less harmful to the environment and more efficient in terms of both energy use and material resources, as these are some of the most pressing problems now confronting the scientific community [12–15]. Hydrogen (H₂) as a fuel is one of the best alternatives for

non-renewable energy resources [16–18]. As opposed to the pollution and smog caused by burning fossil fuels, H₂ fuel produced by electrolysis or solar-driven photo/electrochemical routes offers a clean and environmentally friendly alternative to fossil derivatives [15,19–24]. Furthermore, a great deal of study has been conducted for the betterment of electrochemical efficiency and to engineer low-cost and effective multi-functional electrocatalytic materials for H₂ evolution reactions (HER) and oxygen evolution reactions (OER), as a substitute of noble metal led electrocatalysis [25–31]. With the discovery of electrocatalysis, transition metal-based oxides, chalcogenides and phosphides, graphene derivatives and perovskites have been extensively utilized as potentially upgraded and effective bifunctional materials for this process [32–38].

Therefore, photo/electrocatalytic water splitting for H₂ evolution has been shown to be a promising area of research, and TiO₂ based catalysts have emerged as one of the most promising catalytic systems for photocatalytic/photovoltaic applications [34,35,39–42]. In addition, materials based on TiO₂ are well-recognized as low-priced options with a wide range of possible chemical and electrical structural modifications [43,44]. However, fast charge-carrier recombination, poor photon-to-current conversion efficiency (IPCE), wide band-gap energy (>3 eV) and varying interfacial charge-transfer rates continue to restrict the practical applications of TiO₂ nanostructures [41,45,46]. The direct applications of these materials in industries are challenging due to their relatively inefficient light use. The modification of physical and chemical characteristics and properties of nanomaterials such as the morphology, crystal phase, electron lifetime, electronic structure, energy band gap and positions, surface state and defect in nanocatalysts have recently been the topic of greater focus [45,47–49]. Several methods, including coupling with other semiconductors and doping with foreign elements, have been explored to generate TiO₂ materials that lower band gaps and charge separation enhancement [13,22,50]. The addition of alkaline earth metal dopants such as Mg to TiO₂ alters its band gap, leading to better electronic properties and increased photocatalytic activity. This doping process produces lattice defects such as oxygen vacancies and generates reactive oxygen species, which reduces electron and hole recombination, thereby boosting the photocatalytic performance of TiO₂ nanoparticles [31,51]. The free electrons on a metal-doped nano-catalyst are generated by a process called surface plasmon resonance (SPR), which is triggered by metal doping of the TiO₂ host lattice [52]. By efficiently trapping electrons via solar radiation absorption, SPR improves the charge transfer efficiency of metal-doped TiO₂. In addition, metal doping results in the creation of Ti³⁺ in the TiO₂ lattice, therefore interstitial site defects and oxygen vacancies are introduced which favour the HER activity of TiO₂. Furthermore, as a result of metal, non-metal doping and heterostructured formation, an empty impurity band emerges over the valence band (VB), which facilitates kinetics of photo-induced charge carriers during photocatalysis [10,53,54].

In this paper, we have synthesized Mg-doped TiO₂ via the simple hydrothermal method, which has been characterized by various techniques such as XRD, RAMAN, TEM, FESEM, EDAX, UV-DRS and BET surface area analysis. The doping of Mg to TiO₂ improves the catalytic efficiency towards photocatalysis and electrocatalysis of water by increasing the separation of the photo generated charge carrier (electron-hole pairs) and by providing more surface-active sites with the specific surface area.

2. Results and Discussion

2.1. Powder X-ray Diffraction (PXRD) Studies

To determine the lattice structural properties of as-prepared nanostructures such as crystal structure, phase identification and purity, PXRD technique was employed. XRD patterns of 1, 2.5 and 5% Mg-doped TiO₂ nanoparticles (NPs), as seen in Figure 1, were successfully indexed in the anatase phase of titanium dioxide (TiO₂) using JCPDS card no. 71-1167 [5,55]. The synthesized material was pure, crystalline and no foreign peak related to Mg was observed, which proved that Mg²⁺ ions were completely doped in TiO₂ [56,57]. From the XRD, it can be seen that Mg-doped TiO₂ peaks have been shifted

slightly, which shows that Mg^{2+} ions have been substituted to the host of TiO_2 lattice. Additionally, the peak intensity and sharpness both increased with the increase in Mg^{2+} ion concentration to the TiO_2 host lattice, reflecting an overall rise in crystallinity [57,58]. Furthermore, the 2θ values change to lower values during 1% doping of Mg^{2+} ions, further decreasing with 2.5% and 5% doping [57,58]. Therefore, as the ionic radii of Mg^{2+} ions (0.72 Å) are significantly greater than Ti^{4+} (0.61 Å), lattice expansion must have taken place when Ti^{4+} ions were substituted by Mg^{2+} ions. The crystallinity and crystallite size of Mg-doped TiO_2 is significantly decreased, as predicted the impact of the Mg^{2+} dopant that also contributes to the growth of point defects like oxygen vacancies [51,58]. The inset of Figure 1 illustrates the Mg^{2+} , O and Ti^{4+} lattice sites in the anatase phase crystal structure of Mg-doped TiO_2 nanoparticles.

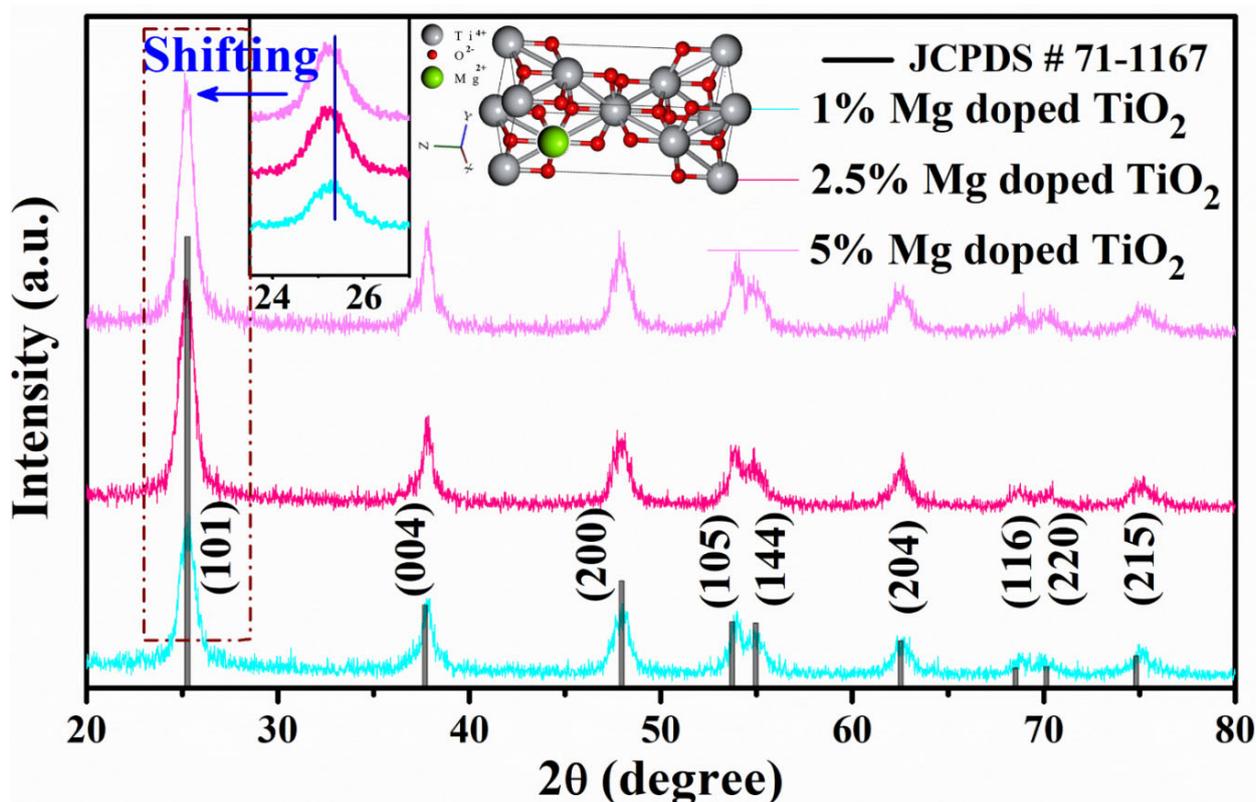


Figure 1. PXRD patterns of the as-synthesized 1, 2.5 and 5% Mg-doped TiO_2 NPs. Inset shows the crystal structure of Mg-doped TiO_2 anatase phase.

2.2. Transmission Electron Microscopy (TEM) Studies

The impact of Mg-doping on TiO_2 nanoparticles is revealed by TEM studies as shown in Figure 2a–c. TEM images reveal that Mg deposited on the TiO_2 host nanoparticles in the form of rough cubes with minor aggregation and black spots. Particle size histograms produced from the TEM images ranging from 4–24 nm showed the average particle size to be 10.45 ± 0.7 , 10.07 ± 0.6 and 9.51 ± 0.6 nm for 1%, 2.5% and 5% Mg-doped TiO_2 NPs, respectively, as shown in Figure 2d–f. Particle size also decreases when Mg^{2+} ion concentration rises, as further proven by increased surface area in BET investigations. Figure 2g–i show HRTEM micrographs of 1, 2.5, and 5% Mg-doped TiO_2 NPs, all of which exhibit well-resolved lattice fringes and have interspacing of 0.348, 0.343 and 0.341 nm, respectively, which corresponds to (101) crystallographic plane of tetragonal TiO_2 .

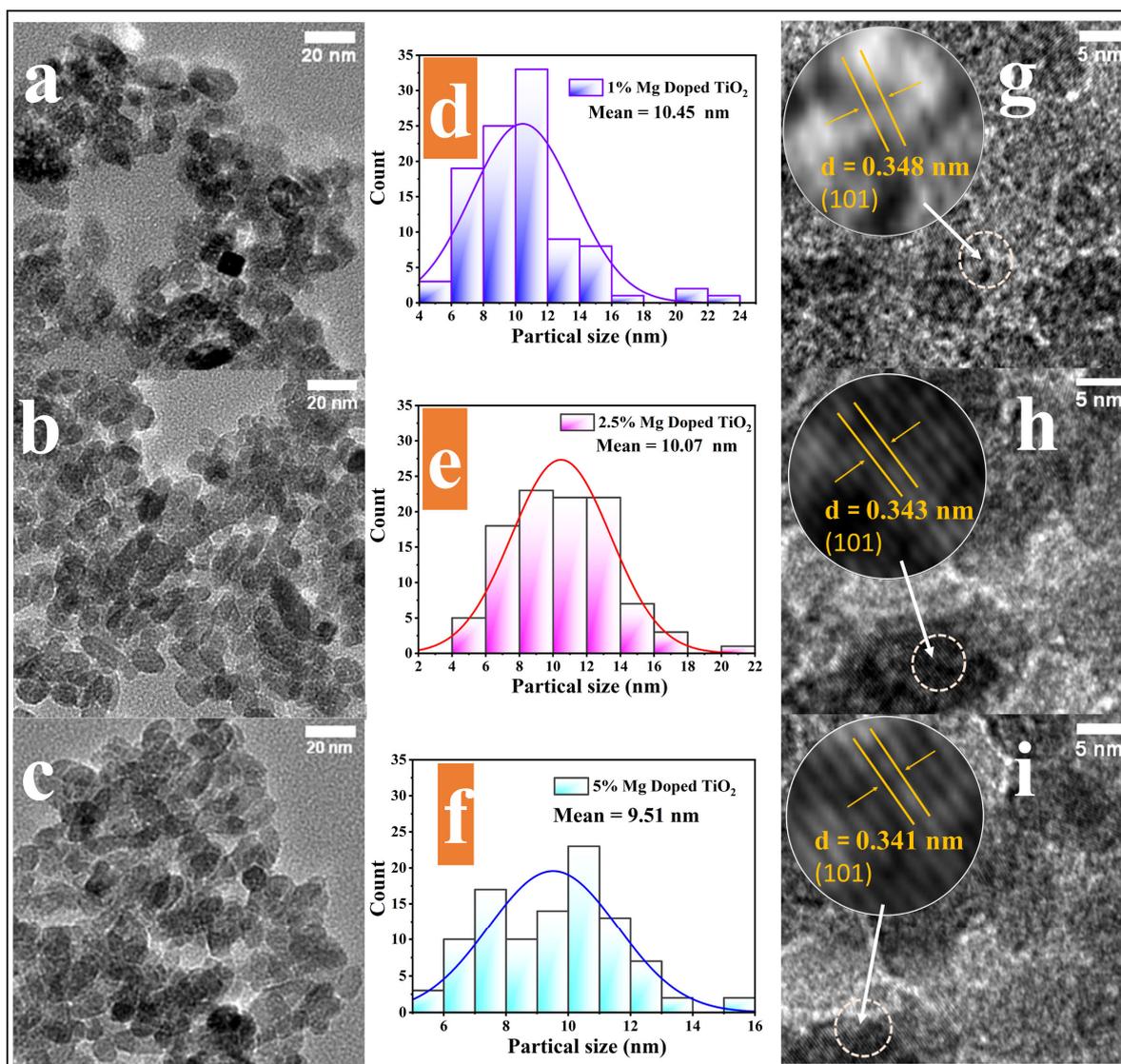


Figure 2. (a–c) TEM micrographs, (d–f) size distribution plots, and (g–i) HRTEM micrographs depicting the lattice spacing of the crystallographic planes of 1%, 2.5% and 5% Mg-doped TiO₂ NPs.

2.3. Field Emission Scanning Electron Microscopy (FESEM) Studies

The surface morphology and particle shape of the nanostructures were studied using FESEM studies. The SEM micrographs of 1, 2.5, and 5% Mg-doped TiO₂ NPs are shown in Figure 3a–c, respectively. The formation of rough, non-uniform NPs with an undefined morphology was observed in SEM micrographs. Additionally, as reported elsewhere, pure TiO₂ NPs [5] had less roughness than Mg-doped TiO₂ samples, whereas Mg-doped TiO₂ NPs exhibited less aggregation than pure TiO₂. Mg-doped TiO₂ NP's elemental distribution was analyzed by X-ray elemental mapping for more accurate identification of elements. It is evident from the X-ray elemental mapping, as shown in Figure 4a,c, that the as-synthesized nanoparticles included a high concentration of Magnesium, Titanium and Oxygen. Figure 4a,c have an inset depicting maps of the elements Mg, Ti and O with SEM micrographs. Figure 4b,d,e display the EDAX spectra of 1, 2.5, and 5% Mg-doped TiO₂, which confirms the existence of Ti and O elements in the as-synthesized sample. The spectrum also demonstrates the successful doping of the element in the host compound.

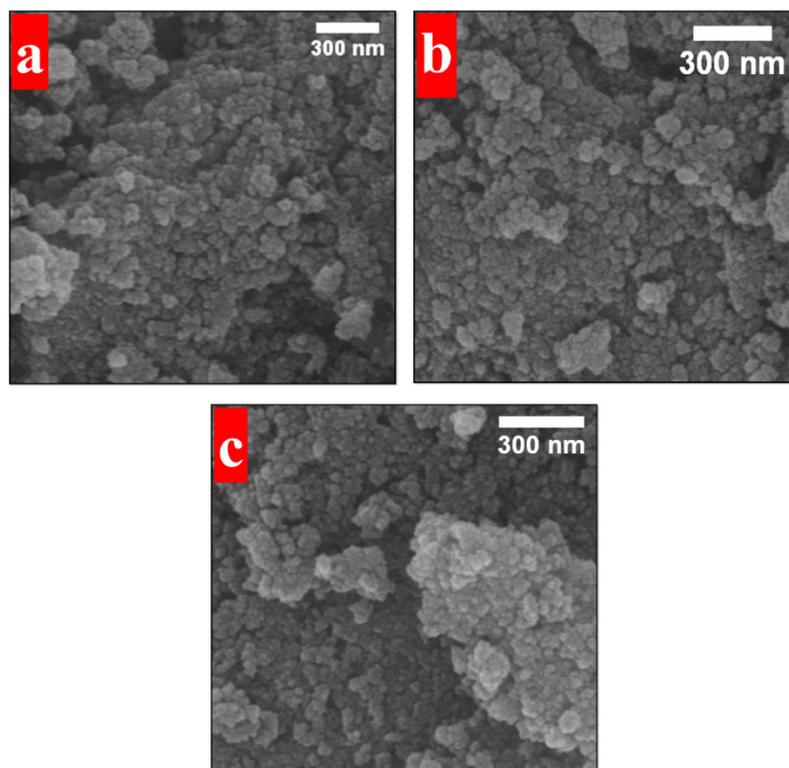


Figure 3. SEM micrographs of (a) 1%, (b) 2.5% and (c) 5% Mg-doped TiO_2 nanostructures.

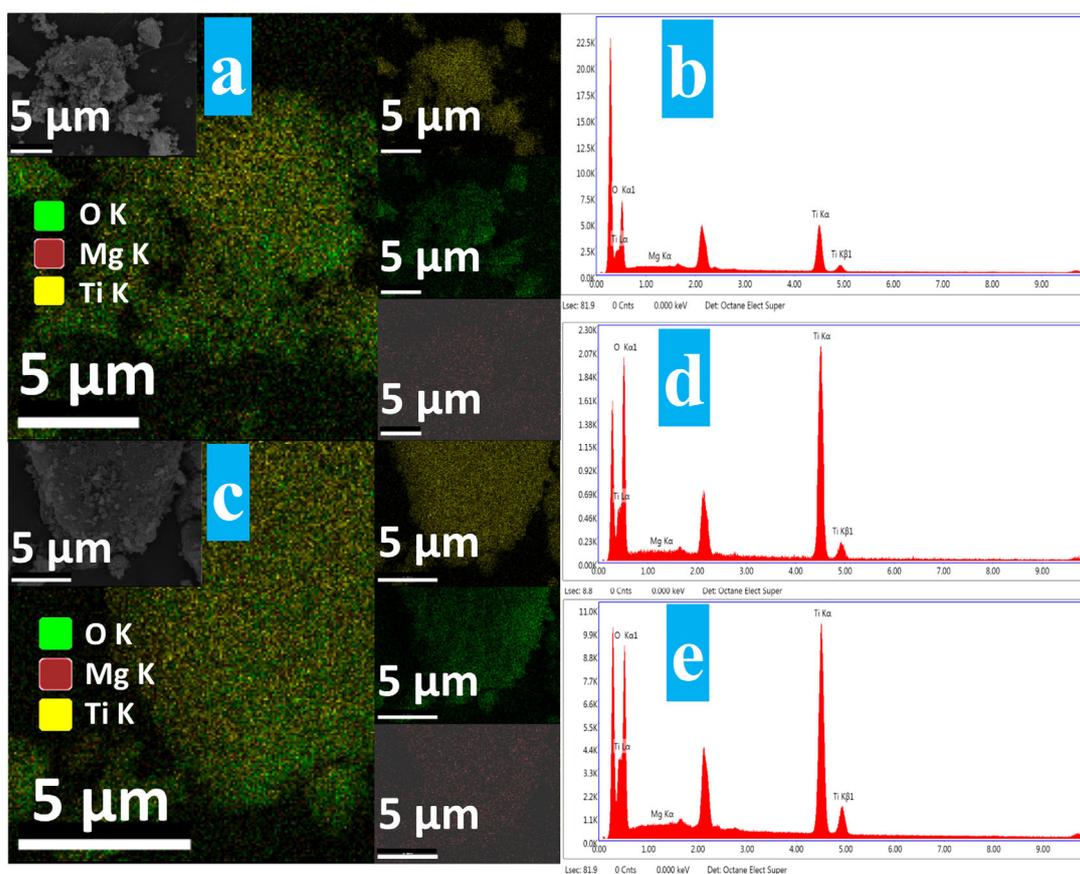


Figure 4. Elemental mapping (Inset shows maps of Ti, Mg and O with SEM micrographs) of as-synthesized (a) 1% and (c) 5% and EDAX spectra of (b) 1%, (d) 2.5% and (e) 5% Mg-doped TiO_2 nanoparticles.

The UV-visible reflectance spectra of the as-synthesized 1, 2.5 and 5% Mg-doped TiO₂ NPs were recorded in diffuse reflectance spectral mode at room temperature, as shown in Figure 5a. The spectra clearly show that the absorption coefficient drops as the wavelength increases, which is typical behavior of many semiconductor materials caused by lattice deformation and internal electric fields inside the crystal. The optical band gap was determined by employing reflectance data in the following Kubelka-Munk equation, $F(R_{\infty}) = \alpha/s = (1 - R_{\infty})^2/2R_{\infty}$ [20,23]. Using Tauc's plot $[F(R_{\infty})h\nu]^{1/n}$ Vs eV and linear extrapolation of $[F(R_{\infty})h\nu]^{1/n}$ as depicted in Figure 5b, the band gaps of 1, 2.5, and 5% Mg-doped TiO₂ nanoparticles were calculated [5,20]. The determined band gap values of 1, 2.5, 5% Mg-doped TiO₂ NPs and pure TiO₂ reported elsewhere [5] were found to be 3.18, 3.16, 3.13 and 3.19 eV, respectively. The existence of impurity or vacancy defect levels between the VB and conduction band (CB) levels may be responsible for the minute changes in the band gap [31,57]. This modifies the photo- and electro- catalytic activity of nano-structures [23,24,59].

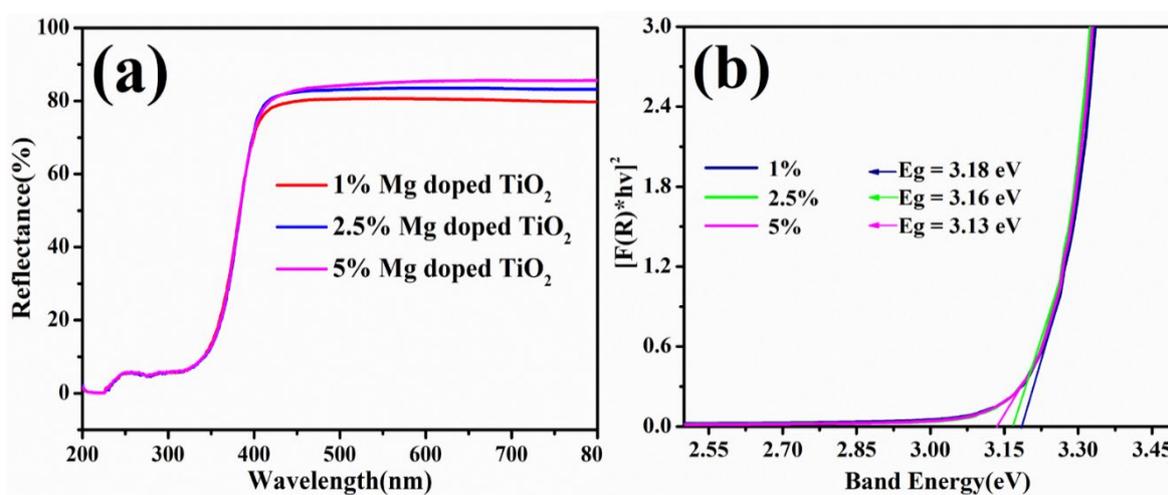


Figure 5. (a) UV-vis DRS and (b) Tauc's plot of 1, 2.5 and 5% Mg-doped TiO₂ NPs.

2.4. Raman Studies

The introduction of Mg²⁺ dopant ions into TiO₂ nanostructures and their effect on the crystal structure was examined via Raman spectroscopy. Figure 6 shows that the main peaks in the Raman spectra of TiO₂ in the anatase phase are located at 143 (E_g), 195 (E_g), 393 (B_{1g}), 513 (A_{1g} + B_{1g}), and 635 (E_g) cm⁻¹ [5,50]. Anatase TiO₂ with the E_g mode is responsible for the strong peak at 143 (E_g) [5]. Such a peak's steep profile, as discussed in the XRD results, is a direct reflection of the high degree of crystallinity present in these NPs. The E_g peak in TiO₂ is connected to the symmetric stretching vibration of O-Ti-O, the symmetric bending vibration related to the B_{1g} mode, and the antisymmetric bending vibration related to the A_{1g} mode of TiO₂ [5,57]. Mg²⁺ metal ion doping of the TiO₂ host lattice will disturb it because the ionic radii of Mg²⁺ ions (0.72 Å) are somewhat greater than those of Ti⁴⁺ (0.61 Å). As a result, oxygen vacancies are produced in the TiO₂ host lattice to balance the charge caused by the displacement of tetravalent Ti⁴⁺ ions by divalent Mg²⁺ metal ions [5,51,58]. Additionally, the Ti-O-Ti bond structure was impacted by the doping of Mg²⁺ ions with the TiO₂ host lattice, and a new bond called Mg-O-Ti was generated, which affected all Raman-active modes. Moreover, a slight 143(E_g) peak shift in the Raman spectra of Mg-doped TiO₂ nanostructures was seen as the concentration of Mg²⁺ dopant increased [5,51]. The results of the Raman investigations of 1, 2.5 and 5% Mg-doped TiO₂ nanostructures (Figure 6) show that the doping of Mg²⁺ ions was accomplished effectively, and that the Raman spectrum correlated with the PXRD pattern.

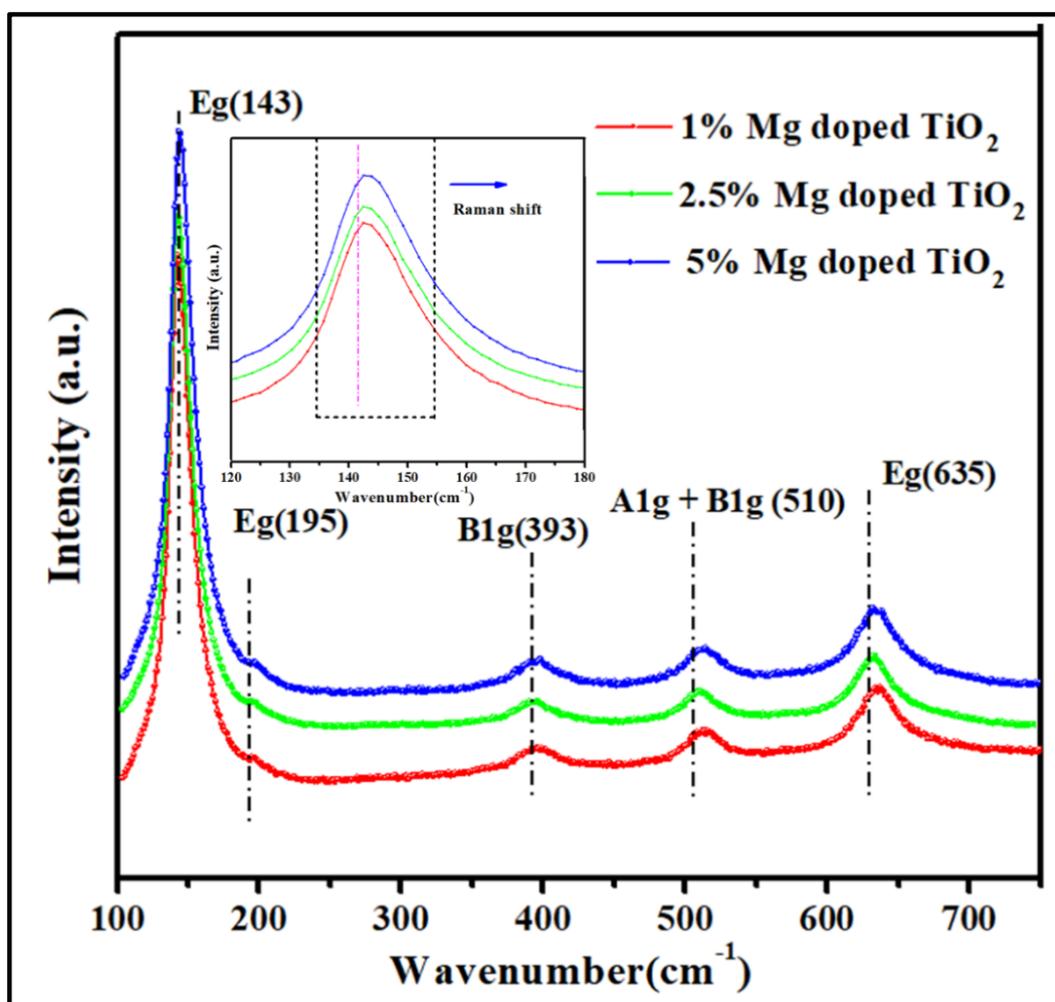


Figure 6. Raman spectra of 1, 2.5 and 5% Mg-doped TiO_2 nanoparticles.

2.5. BET Surface Area Studies

The surface area is a crucial factor in establishing the material's characteristics. Therefore, the surface-active sites present in a material with a high surface area are crucial to the success of any given operation. As a result, a material with a large surface area will have more surface-active sites accessible for any activity to occur. A Brunauer-Emmett-Teller (BET) surface area analysis was performed for 1, 2.5, and 5% Mg-doped TiO_2 nanostructures to determine their respective specific surface area, as shown in Figure 7. The specific surface area of the as-synthesized NPs was calculated using the N_2 adsorption-desorption isotherms and was found to be 152.9 , 164.2 and $198.1 \text{ m}^2 \text{ g}^{-1}$ for 1, 2.5, and 5% Mg-doped TiO_2 nanostructures, respectively. Adsorption points from the isotherm were utilized to calculate the average pore size distribution with the Barrett-Joyner-Halenda (BJH) plot. Figure 7b shows that the average pore size for 1, 2.5, and 5% Mg-doped TiO_2 NPs are 16.95 , 16.93 and 17.08 \AA , respectively, indicating that all the samples are mesoporous. Using the isotherm's adsorption and desorption points, the pore radius was determined using the Dubinin and Astakov (DA) method, as shown in Figure 7c, and was found to be 13.6 , 13.4 , and 14.1 nm for 1, 2.5, and 5% Mg-doped TiO_2 NPs, respectively. BET surface area parameters of as-synthesized materials have been summarized in Table 1.

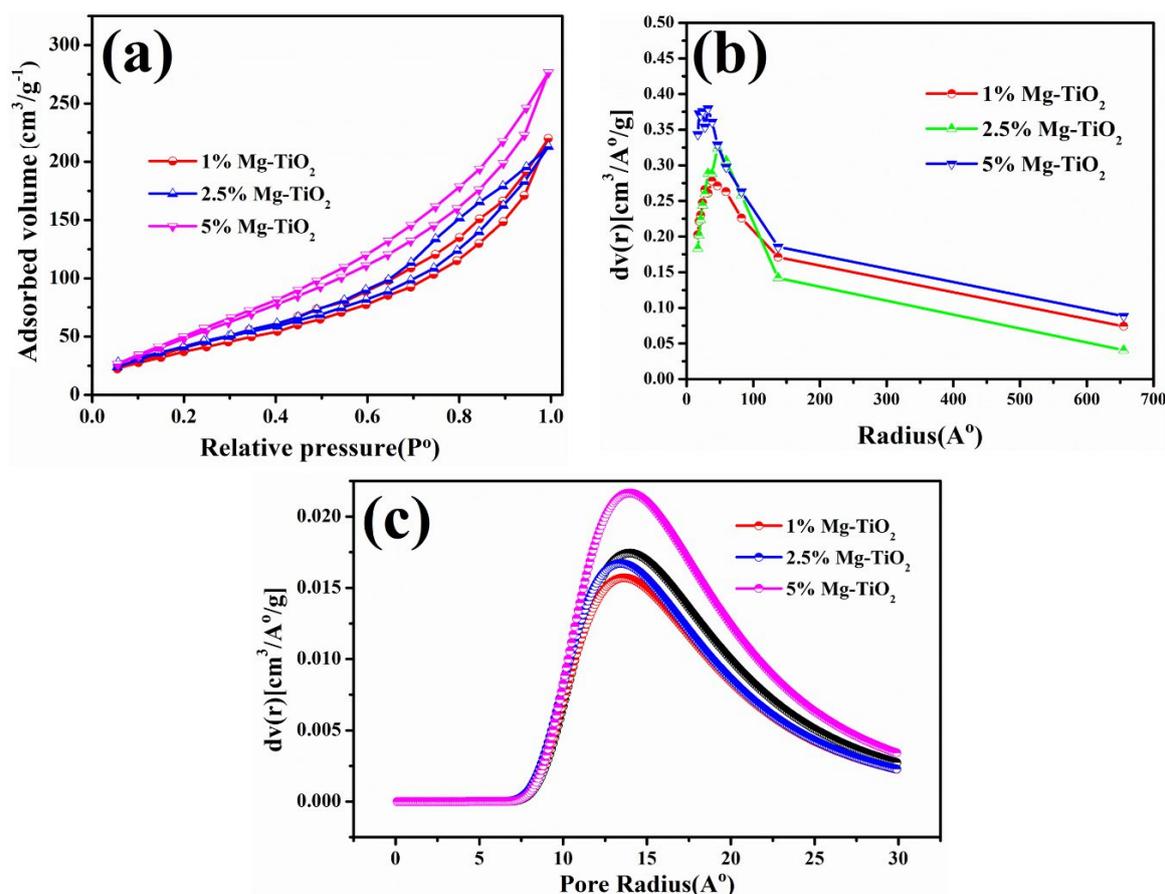


Figure 7. (a) Nitrogen adsorption–desorption isotherms, (b) BJH pore-size distribution, and (c) DA average pore-size distribution plots of 1, 2.5 and 5% Mg-doped TiO₂ NPs.

Table 1. BET surface area (S_{BET}), Average pore size and Pore volume (V_{Total}) of synthesized nanocatalysts.

Sample	$S_{\text{BET}}/(\text{m}^2 \text{g}^{-1})$	Average Pore Size/Å	$V_{\text{Total}}/(\text{cm}^3 \text{g}^{-1})$
1% Mg-doped TiO ₂	152.9	16.95	0.317
2.5% Mg-doped TiO ₂	164.2	16.93	0.391
5% Mg-doped TiO ₂	198.1	17.08	0.449

2.6. Photocatalytic Water Splitting for H₂ Evolution Studies

The H₂ evolution performance of as-synthesized nanomaterials was determined by photocatalytic water splitting studies. H₂ was produced photo-catalytically over as-prepared 1, 2.5 and 5% Mg-doped TiO₂ NPs in water using Na₂S/Na₂SO₃ as a sacrificial agent, as illustrated in Figure 8a,b. The photocatalytic efficiency of 1, 2.5 and 5% Mg-doped TiO₂ nanocatalysts was found to be approximately 17.89, 34.21 and 38.96 mmol g_{cat}⁻¹ in 8 h, which is higher than what has been reported for pure TiO₂ nanocatalyst previously [5]. Furthermore, as shown in Figure 8c, the average H₂ evolution rates for 1, 2.5, and 5% Mg-doped TiO₂ photocatalysts are 2.24, 4.28, and 4.87 mmol g_{cat}⁻¹ h⁻¹, respectively. The H₂ evolution rate is affected by Mg²⁺ dopant concentrations in host TiO₂ NPs as depicted in Figure 8a,b and the 5% Mg-doped TiO₂ photocatalyst was shown to evolve maximum H₂. The photocatalytic performance of metal-doped TiO₂ synthesized through various methods was compared with the present work and summarized the results in Table 2. Experiments on the photocatalyst's reusability and durability were conducted under ideal photocatalytic reaction conditions. Figure 8d shows a 5% Mg-doped TiO₂ nanocatalyst recycled up to three times, with an 8-h reaction period. The results show that the amount

of H₂ evolved reduced from the initial value of 38.96 mmol g_{cat}⁻¹ in the first cycle to a final value of 37.7 mmol g_{cat}⁻¹ in the third cycle. In summary, the research demonstrates that the best performance of the 5% Mg-doped TiO₂ photocatalytic H₂ evolution exhibits an average H₂ generation of 4.87 mmol g_{cat}⁻¹ h⁻¹ in the presence of sacrificial agent Na₂S/Na₂SO₃.

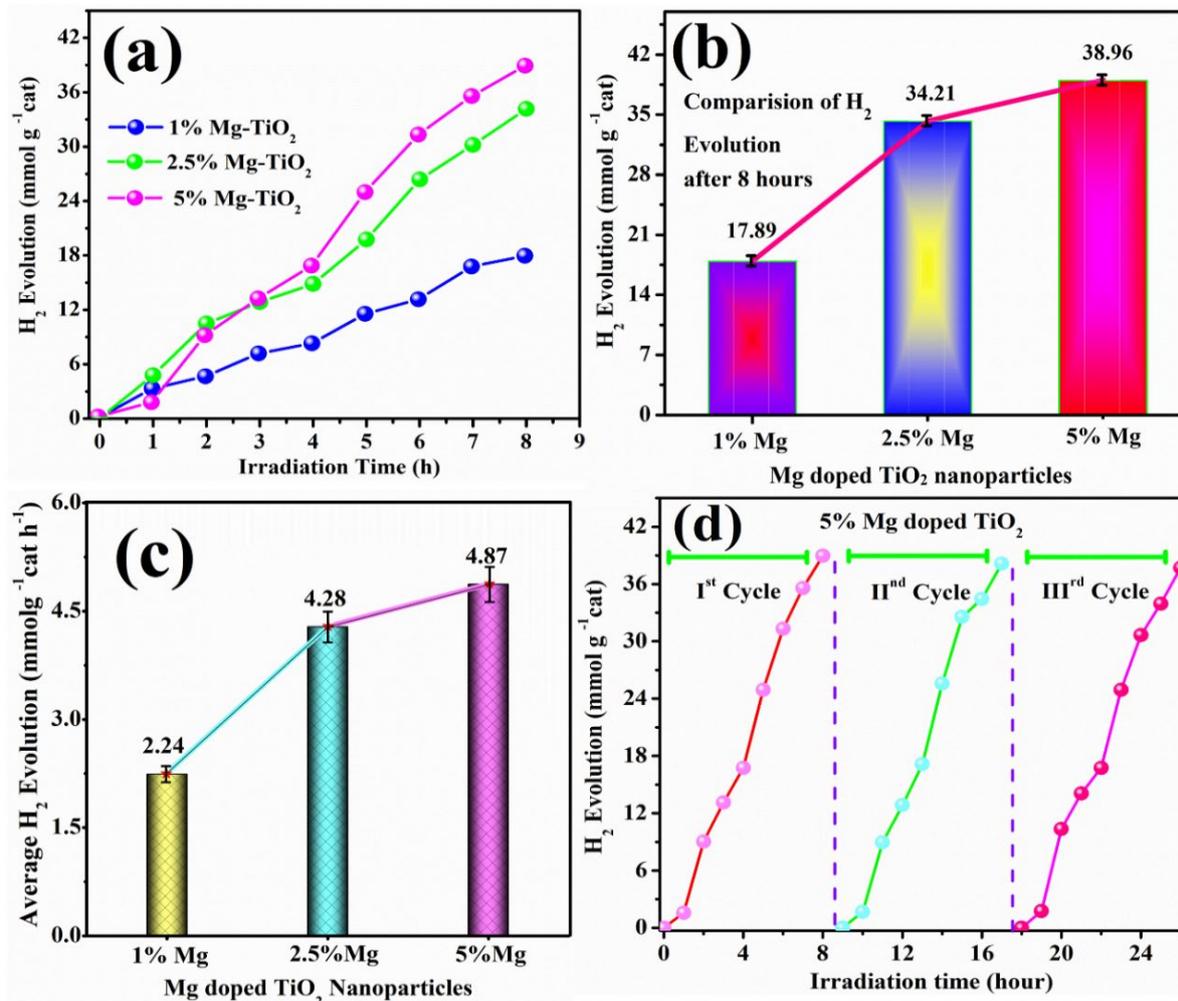


Figure 8. (a) Photocatalytic H₂ evolution curves under the UV-visible light irradiation as the function of irradiation time, (b) comparison of photocatalytic activity, (c) Average H₂ evolution per hour of 1, 2.5 and 5% Mg-doped TiO₂ NPs and (d) stability test of 5% Mg-doped TiO₂ for up to three cycles.

Table 2. Comparison of metal-doped TiO₂ photocatalytic performance, synthesized by different methods.

S. No.	Dopant	Synthesis Method	Parameters	Hydrogen Production	Ref.
1.	Platinum, Nitrogen	Photodeposition method	1150 mL Pyrex vessel, UV light	3200 μmol	[59]
2.	Nitrogen	Solid state/calcination method	UV-Vis light irradiation, Na ₂ S/Na ₂ SO ₃	18 μmol	[60]
3.	Bismuth, Nitrogen	Sol-gel method	Solar light, methanol	1800 μmol g ⁻¹	[61]
4.	Pd/0.2% K ⁺	Hydrothermal method	—	76.6 μmol h ⁻¹	[62]
5.	Strontium, silver	sol-gel method	500 W Xe arc lamp	49.4 μmol h ⁻¹	[63]
6.	Strontium	Hydrothermal method	200 W, Hg-Xe arc lamp	3.3 mmol g _{cat} ⁻¹ h ⁻¹	[5]
7.	Magnesium	Hydrothermal method	170 W, Hg-Xe arc lamp	4.87 mmol g _{cat} ⁻¹ h ⁻¹	In this work

2.7. Possible Photocatalytic Reaction Mechanism of Mg-Doped TiO₂ Nanoparticles

In Figure 9, a mechanism is proposed for the production of H₂ gas using Mg-doped TiO₂ NPs as the photocatalyst. UV-Visible light excites Mg-doped TiO₂, leading to the generation of photoexcited electrons and holes in the CB and VB, respectively [13]. Noticeably, the metallic Mg has a high reducing power as TiO₂ so that it can act as an electron donor center [39]. Specifically, the electrons in the CB of TiO₂ are transferred to metallic Mg, while the holes are transferred to the VB of TiO₂, thereby suppressing the recombination of photo generated electron-hole pairs in semiconductor materials and maintaining the strong reductive electrons and oxidative holes [34,64]. During water splitting, the strong reductive electrons in CB of metallic Mg interact with H⁺ to form H₂, while the holes retained in VB of TiO₂ are consumed by the Na₂S/Na₂SO₃ sacrificial reagent [5,20,46]. The following rule governs the reaction pathway.

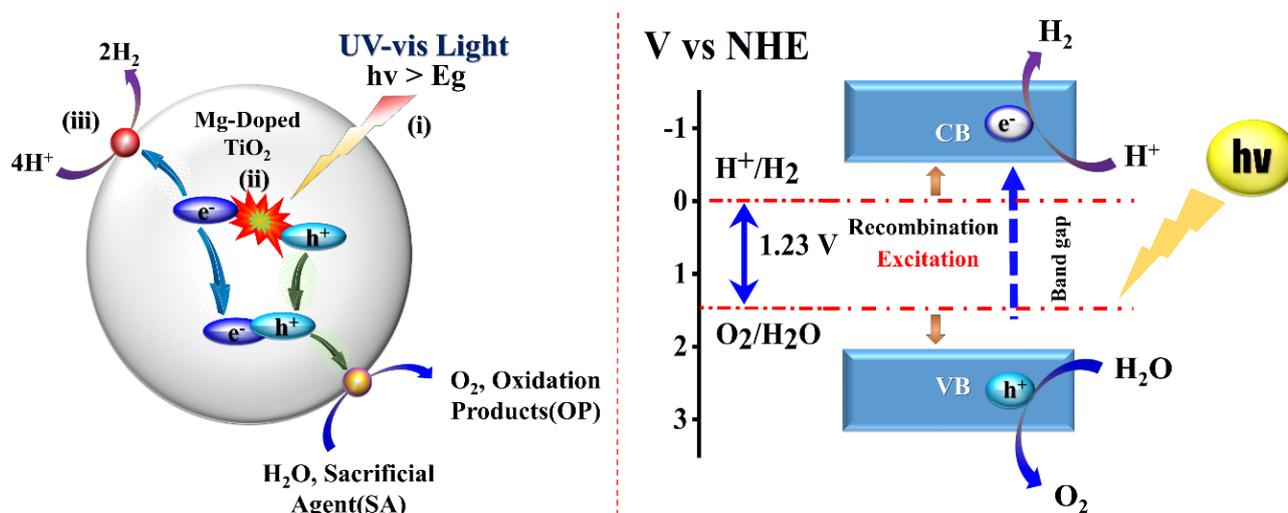
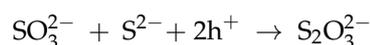
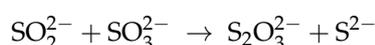
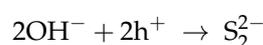
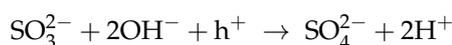


Figure 9. Schematic representation and possible electron transfer mechanism of photocatalytic H₂ evolution.

2.8. Electrocatalytic Water Splitting Studies

1, 2.5, and 5% Mg-doped TiO₂ nanostructures were electrochemically studied in order to investigate the HER responses during electrocatalytic water splitting. The 0.5 N H₂SO₄ and 0.1 N KOH electrolytic solutions were used to record both HER responses under ambient conditions. Figure 10a,b depict the anodic and cathodic sweep of linear voltammetry (LSV) and cyclic voltammetry (CV) within the required potential window with 0.5 N H₂SO₄ electrolytic solution after IR correction. Figure 10c showed cyclic CV at scan rates ranging from 10 to 100 mV/s within a potential range of +1.5 to −1.5 V after IR correction. The overpotential decreased at 10 mA cm^{−2} cathode current density as the Mg²⁺ ion concentration in the TiO₂ nanocatalyst increased up to 2.5%, demonstrating the best electrocatalytic HER activity of 2.5% Mg-doped TiO₂. At 10 mA cm^{−2}, the overpotential values of 0.98, 0.96 and 0.99 V were obtained for 1, 2.5 and 5% Mg-doped TiO₂ electrocatalysts, respectively.

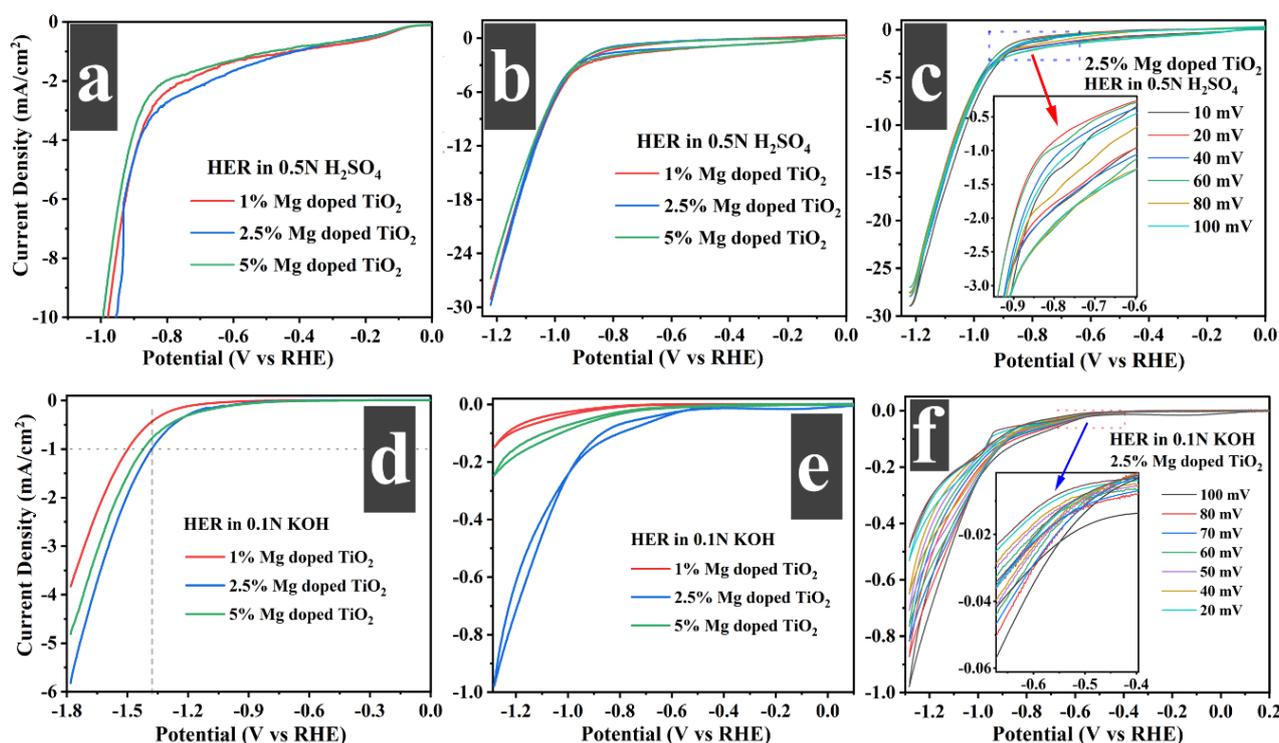


Figure 10. (a,b,d,e) show the LSV and CV plots of HER of as-synthesized 1, 2.5, 5% Mg-doped TiO₂, whereas (c,f) show the CV plots of 2.5% Mg-doped TiO₂ towards HER in 0.5 N H₂SO₄ and 0.1 N KOH electrolytic solution, respectively.

Furthermore, the electrocatalytic activity was performed with 0.1 N KOH alkaline solution to investigate the HER performance of as-synthesized 1, 2.5 and 5% Mg-doped TiO₂ electrocatalysts. The cathodic LSV and CV curves with 0.1 N KOH electrolytic solution within the necessary potential window after IR correction are shown in Figure 10d,e. From the LSV study for HER in 0.1 N KOH electrolytes solution, it was observed that 0.44, 1.00 and 0.786 mA cm⁻² current densities were generated at -1.38 V onset potential for 1, 2.5 and 5% Mg-doped TiO₂ electrocatalysts, respectively. Thus, it was found that the 2.5% Mg-doped TiO₂ nanocatalyst has a maximum current density for HER at onset potentials of -1.38 V in 0.1 N KOH electrolyte. CV scans of 2.5% Mg²⁺ ion doped TiO₂ were run at a rate of 10–100 mV/s within a potential window of +1.5 and -1.5 V after IR correction, as depicted in Figure 10f. The reaction kinetics and mechanisms for electrocatalytic activity were determined with the use of HER responses utilizing tafel analysis. The size, surface area, shape and optical properties of nanoparticles, as well as the orientation of electrocatalysts, correspond perfectly with reaction kinetics and mechanisms [5,17,30,48]. Tafel slopes of as-synthesized 1, 2.5 and 5% Mg-doped TiO₂ nanostructures for the H₂O electrolysis for HER in both 0.5 N H₂SO₄ and 0.1 N KOH electrolytes solutions are shown in Figure 11a,b and also presented in Tables 3 and 4 which show that 2.5% Mg-doped TiO₂ has the lowest tafel slope values. Therefore, we performed the chronoamperometric analysis of 2.5% Mg-doped TiO₂ electrocatalyst. According to Figure 12, the current density did not vary during the chronoamperometric (CA) experiments performed to examine the endurance and stability of a 2.5% Mg-doped TiO₂ electrocatalyst running at a voltage -1.0 V up to 4000 s in 0.5 N H₂SO₄ electrolytic solution for the HER.

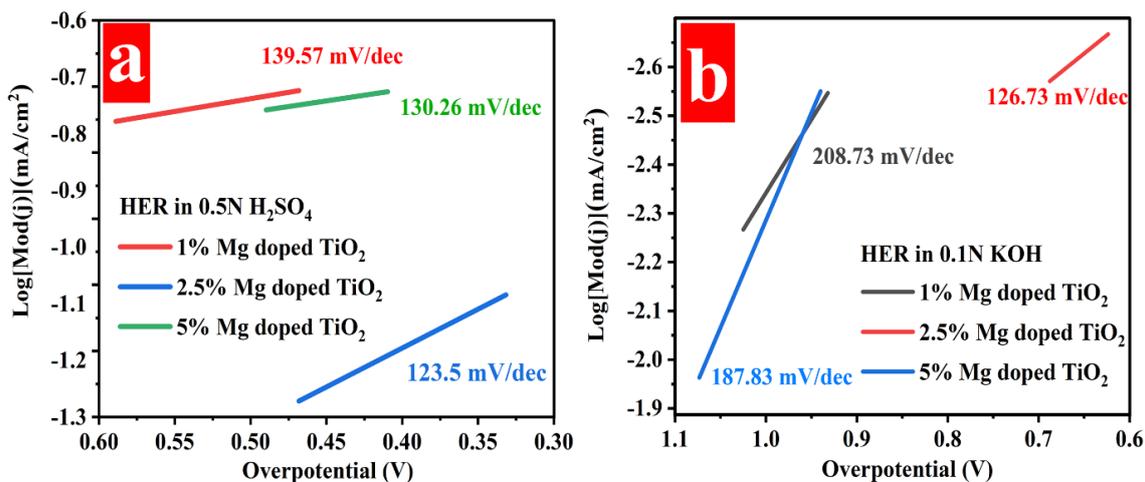


Figure 11. (a,b) shows HER Tafel plots of 1, 2.5, 5% Mg-doped TiO₂ in 0.5 N H₂SO₄ and 0.1 N KOH electrolytic solution.

Table 3. HER electrodic parameters of Mg-doped TiO₂ electrocatalysts in 0.5 N H₂SO₄ solution.

S. No.	Materials	HER		Ref.
		Overpotential (V) to Attain 10 mA/cm ²	Tafel Slope (mV/dec)	
1.	1% Mg-doped TiO ₂	0.99	139.57	This work
2.	2.5% Mg-doped TiO ₂	0.98	123.5	This work
3.	5% Mg-doped TiO ₂	0.95	130.26	This work
4.	Pristine TiO ₂	1.00	133.33	[5]

Table 4. HER electrodic parameters of Mg-doped TiO₂ electrocatalysts in 0.1 N KOH solution.

S. No.	Materials	HER	
		Cathodic Current Density (mA/cm ²) at -1.38 V	Tafel Slope (mV/dec)
1.	1% Mg-doped TiO ₂	0.436	208.73
2.	2.5% Mg-doped TiO ₂	1.00	126.30
3.	5% Mg-doped TiO ₂	0.786	187.83

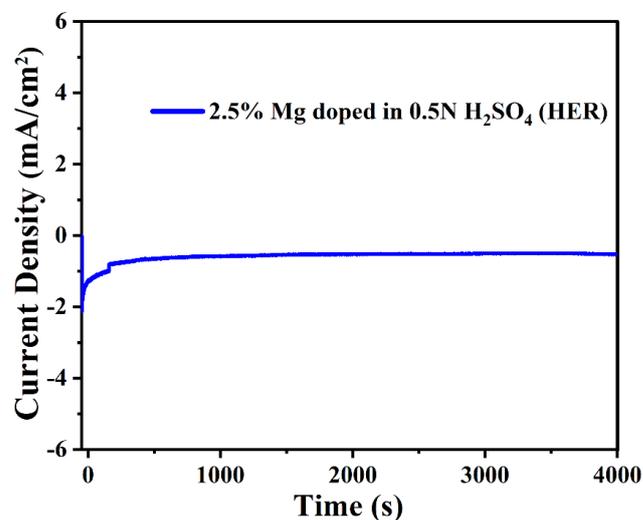


Figure 12. CA analysis at -1.0 V of 2.5% Mg-doped TiO₂ nanoparticles.

3. Experimental Section

3.1. Materials Required

The following solvents and reagents were used in the synthesis with no further purification: Titanium isopropoxide (Alfa Aesar, 97%, Heysham, Lancashire, UK), Magnesium acetate tetrahydrate (SRL, 99%, Sisco research laboratories Pvt. Ltd., Taloja, Maharashtra, India), Sodium hydroxide (Merck, 97%, Life Science Pvt. Ltd., Vikhroli, Mumbai, India), Glacial acetic acid (Fisher Scientific, Heysham, Lancashire, UK), Ethanol (Merck, 99.9%, Darmstadt, Germany), sodium sulfide (Na_2S , 60%, SRL Pvt. Ltd., Taloja, Maharashtra, India), sodium sulfite (Na_2SO_3 , Alfa Aesar, 98%, Heysham, Lancashire, UK) and extra pure deionized water.

3.2. Synthesis of Mg-Doped TiO_2 Nanoparticles

1, 2.5, and 5% Mg-doped TiO_2 NPs were produced by a hydrothermal approach using a laboratory autoclave. In order to avoid the direct oxidation of titanium isopropoxide to TiO_2 , 7.4 mL of glacial acetic acid was added to a round-bottom flask to carry out the reaction under controlled conditions and create an inert atmosphere with the purging of nitrogen gas. Then, 0.74 mL of titanium isopropoxide was added after creating an inert atmosphere with continuous stirring [5]. With the addition of 16.86 mL of distilled water, the final volume reached 25 mL, and the resultant solution is known as 0.1 M Ti^{4+} [5,65]. For doping 1% Mg ions into 0.1 M TiO_2 , 25 mL of 0.001 M Magnesium acetate solution was prepared. Subsequently, 50 mL of 0.1 M NaOH solution was added slowly while stirring the mixture on a magnetic stirrer for approximately 30 min. The autoclave containing the reaction mixture was then placed in a vacuum oven and heated for 12 h at a temperature of 150 °C. After cooling the reaction solution down to ambient temperature, it was centrifuged, rinsed many times with double distilled water and ethanol and kept at 60 °C in a vacuum oven to dry for 12 h. The dried precipitate was then ground into powder. For the 2.5 and 5% Mg^{2+} ion doping in TiO_2 nanoparticles, the method was carried out distinctly using 25 mL of 0.0025 M Magnesium acetate solution for 2.5% doping and 25 mL of 0.005 M for 5.0% doping.

3.3. Characterizations

Powder X-ray diffraction (XRD) studies in the 2θ range of 20–80° were performed using a Rigaku Japan D/max 2500 diffractometer at the scanning rate of 5°/min with Ni-filtered $\text{Cu K}\alpha_1$ radiation ($\lambda = 1.5406 \text{ \AA}$) and using a step size 0.05°/s to examine phase purity and the crystal structure of the as-prepared samples. Raman spectral investigations were analyzed utilizing an excitation wavelength of 532 nm and a He-Ne laser-equipped RENISHAW inVia Raman microscope (Leica Microsystems CMS GmbH, Wetzlar, Germany). Nanoparticles of undoped and Mg-doped TiO_2 were analyzed by TEM using 200 kV TALOS equipment. The TEM samples were prepared in ethanol and a drop casting of suspension on a copper grid (carbon-coated). FESEM was utilized to analyze the surface morphology, elemental mapping and elemental analysis using a Nova Nano SEM-450 microscope (FEI Company, Hillsboro, OR, USA) accelerated at 20 kV voltage coupled with an energy-dispersive X-ray (EDAX) system. Studies of UV-visible DRS were carried out on a PerkinElmer Lambda 365 spectrophotometer. The band gap was calculated by recording reflectance spectra of as-synthesized materials between 200 and 800 nm. Furthermore, the BET surface area studies of Mg-doped TiO_2 nanostructures were calculated using Nova 2000e, Quantachrome Instruments Limited, Boynton Beach, FL, USA, at liquid nitrogen temperature.

3.4. Photocatalytic H_2 Evolution Measurements

The activity of as-synthesized nanostructures was measured using 0.128 M/500 mg of Na_2S and 0.079 M/500 mg of Na_2SO_3 as sacrificial agents toward photocatalysis of H_2O for H_2 evolution using a UV-Visible light source [23]. An airtight photo reactor (50 mL cylindrical quartz cell) consisted of 20 mg of as-synthesized nanocatalysts, as required

sacrificial agents and 50 mL double-distilled water was used to perform H₂ evolution activity [5,30]. To achieve inert atmosphere conditions and flush out any residual gases (such as dissolved oxygen), the prepared solution was stirred while purging nitrogen gas continuously for 30 min. A quartz cell containing the solution mixture was affixed to the stirrer after being well mixed, and its flat surface was then exposed to a 170 W light source (200 W Hg-Xe arc lamp from Newport; Model 66906-200HXF-R15 Ozone-free). H₂ production was measured using a Clarus 590 GC from PerkinElmer fitted with a TCD detector and nitrogen as the carrier gas. Samples were obtained hourly using a special airtight glass syringe. After being centrifuged out of the reaction mixture, the photocatalysts were dried and evaluated for their photocatalytic response, which showed maximal H₂ generation over the course of three consecutive 8-h cycles. The photocatalysts were separated from the reaction solution via centrifugation, dried in an oven, and analyzed for their reusability, recyclability, photostability and photocatalytic response, the latter of which demonstrated maximum H₂ production over the course of three consecutive 8-h cycles.

3.5. Electrode Preparation and Electrocatalytic Measurements

Electrochemical measurements were accompanied by Metrohm Autolab PGSTAT204 with a 3-electrode electrochemical cell setup (Pt wire as counter and saturated calomel electrode as reference electrode) at room temperature in 0.5 N H₂SO₄ and 0.1 N KOH electrolyte solutions for the HER reactions. A coating of catalyst dispersion on pretreatment ITO substrates with dimensions of 1 × 1 cm² made working electrodes for electrochemical setup of as-synthesized 1, 2.5, and 5% Mg-doped TiO₂ electrocatalysts [5]. The substrate was cleaned with ethanol, acetone, and isopropanol for 15 min in each solvent before being dropped onto the substrates. In order to make catalyst suspensions, 4 mg of electrocatalysts, 300 µL of isopropanol, and 15 µL of Nafion (5 wt%, Alfa Aesar) were mixed together [5,6,35]. The resultant mixtures were subjected to a 30-min sonication process. Following the method described previously, 200 µL of dispersions were dropped onto the conducting side of an ITO substrate (the side with the oxide layer). They were then allowed to dry overnight at 60 °C in a vacuum oven. To measure the electrochemical performance of the produced nanomaterials, LSV measurements were taken at 100 mV, CV measurements were taken in the range of 10–100 mV/s, and CA measurements were taken at –1.0 V for HER.

4. Conclusions

An eco-friendly low-temperature hydrothermal method was utilized to fabricate highly crystalline tetragonal Mg-doped TiO₂ nanocatalysts. Increased photocatalytic and electrocatalytic performance can be attributed to various factors like high specific surface area, increased pore volume, the introduction of defects that facilitate Mg²⁺ ion doping. In terms of electrocatalytic and photocatalytic activity, 2.5% Mg-doped TiO₂ exhibited optimum response in terms of HER electrocatalysis in both acidic (0.5 N H₂SO₄) and alkaline (0.1 N KOH) mediums. Furthermore, 2.5% Mg-doped TiO₂ nanocatalysts showed high current density for HER in 0.5 N H₂SO₄ were 10 mA/cm² with an overpotential at 0.96 V for HER and with Tafel slope of 123.5 mV/dec. By contrast, 5% Mg-doped TiO₂ showed the highest 38.96 mmol g_{cat}⁻¹ H₂ evolution in comparison to 1% and 2.5% Mg-doped TiO₂ as photocatalysts. In this study, it was deduced that 5% Mg-doped TiO₂ and 2.5% Mg-doped TiO₂ nanocatalysts have better photo- and electro-catalytic activity, respectively. Therefore, we anticipate that these materials have the potential to upgrade the efficiency of sustainable energy resources due to their efficiency in photo/electrocatalysis for water splitting.

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