

## Article

# Synthesis and Spectroscopic Characterization of Dapagliflozin/Zn (II), Cr (III) and Se (IV) Novel Complexes That Ameliorate Hepatic Damage, Hyperglycemia and Oxidative Injury Induced by Streptozotocin-Induced Diabetic Male Rats and Their Antibacterial Activity

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**Abstract:** Diabetes mellitus (DM) causes an imbalance in the oxidative status of the human body. Three novel Dapagliflozin (Dapg) Zn (II), Cr (III) and Se (IV) complexes were prepared and characterized by elemental analysis, IR, electronic spectra, magnetic susceptibility, scanning electron microscopy (SEM) and X-ray diffraction. The molar conductance values confirmed the non-electrolytic nature of the Dapg complexes. According to spectral data, Dapg acts as a bidentate ligand. The thermal analyses of the complexes were studied using the DSC technique. The surface morphology and particle sizes of the Dapg complexes were investigated using SEM and XRD. XRD confirmed the crystalline structure for the complexity. This study investigated the effect of novel metal complexes of Dapg with the metals Zn (II), Cr (III) and Se (IV) on oxidative injury and tissue damage in the hepatic tissue of streptozotocin (STZ)-induced diabetic male rats. DM was experimentally induced in male rats. The diabetic rats received Dapg, Dapg/Zn, Dapg/Cr and Dapg/Se orally for 30 successive days. Male rats exposed to STZ showed multi-histopathological alterations in their hepatic tissue, including inflammatory and structural changes. STZ elevated oxidative stress markers in the hepatic tissue and lowered the antioxidant defense enzymes. Supplementation of Dapg with Zn, Cr or Se novel complexes significantly prevented hepatic injury and suppressed the generation of reactive oxygen species. The Dapg/Zn complex was highly effective against *Bacillus subtilis* and *Streptococcus pneumoniae*, while Dapg/Cr was highly effective against *Escherichia coli* and *Pseudomonas aeruginosa*, and Dapg/Se was highly effective against *Staphylococcus aureus*. In conclusion, Dapg novel metal complexes with Zn, Cr or Se protect against oxidative injury and the pathophysiological and bacterial complications of DM and hepatic tissue injury. The Dapg novel metal complexes improved hepatic functions, reduced blood glucose levels and enhanced the levels of antioxidant defense enzymes in diabetic male rats.

**Keywords:** drug metal complexes; dapagliflozin; chromium; oxidative stress; hepatic tissues

## 1. Introduction

Diabetes mellitus (DM) increases the rate of mortality and poses a huge burden on the economy. Globally, it is reported that approximately 4–4.6 million patients with diabetes die annually. At present, diabetes is a high-incidence chronic disease worldwide [1].

Diabetes is a disease that impairs the body's ability to control blood glucose levels. Diabetes can lead to an excessive level of sugar in the blood, which culminates in adverse pathophysiological complications, including hepatic toxicity, heart failure and kidney failure [2].

Patients with DM often have other complications, such as gastric dysfunction, insomnia and even dyslipidemia [1].

SGLT2 inhibitors are listed as a new class of oral anti-hyperglycemic medications for the pharmacological management of DM [3]. Sodium-glucose co-transporter 2 (SGLT2) inhibitors cause the kidney tissues to increase urinary glucose excretion and inhibit the hyperglycemic state, thus improving glycemic control [4]. In addition, SGLT2 targets the improved function of  $\beta$ -cells and increases insulin sensitivity [5].

Several categories of antidiabetic drugs are used for the treatment of diabetic patients for lowering blood glucose. Dapagliflozin (Dapg) is a member of the SGLT2 drug category. The chemical structures of Dapg are shown in Figure 1. Dapg is a glucose-lowering agent used in the treatment of diabetics. It is a known inhibitor of SGLT2. By inhibiting the receptor of transporter protein known as SGLT2 in the kidney, it causes a decline in glucose reabsorption in the renal tissue, which leads to excessive urinary excretion and a decline in blood glucose levels [1].

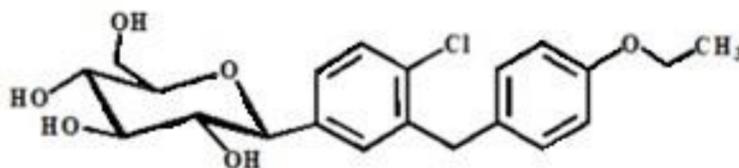


Figure 1. Chemical structure of Dapagliflozin (Dapg).

Dapg, which is the third member of the gliflozin family, has pronounced cardioprotective and neuroprotective effects in the development of insulin resistance in animals. Dapg has been shown to restore the declined mitochondrial membrane potential responsible for mitochondrial dynamics, as well as providing a marked increment in  $\text{Ca}^{2+}$  homeostasis in tested cardiomyocytes in diabetic rats [6]. Dos Santos and Filho [7] treated patients with nonalcoholic fatty liver disease with a fixed dose of 10 mg of Dapg for an average of 75 days. Liver enzymes ALT and AST, insulin hormone, insulin resistance level "HOMA-IR" and weight levels were significantly reduced after treatment, with no reports of its effect on the hepatic ecostructure.

Chromium (Cr III) is required in a small amount to maintain the vitality of the human body [1]. Cr is among the most important trace elements that the body can absorb. It is reported that Cr is the most essential stable form that can be accumulated in soft tissues, such as the liver and bone. It plays a vital role in the metabolism of both lipids and carbohydrates. It may elevate insulin sensitivity in the human body and can also inhibit cholesterol synthesis in hepatic tissues. Cr deficiency is reported to cause hypercholesterolemia and DM complications [8].

Zinc (Zn) is an essential trace element that plays a vital role in cellular metabolism. Zn is essential for enzymatic processes. Its deficiency may result in a wide spectrum of declining drug elimination, clinical manifestations and immune dysfunction [9]. The liver tissue is the main target organ that is responsible for zinc metabolism and it can be affected by hepatic diseases. Moreover, Zn deficiency may change the functions of hepatocytes and can result in several liver diseases and bacterial infections [10].

Selenium (Se) is shown to be reduced in patients with chronic liver diseases [11]. Se is an essential trace element for humans [12]; it is essential for the antioxidant defense system [13]. It was previously reported that the biosynthesis of selenoproteins is greatly affected in hepatic encephalopathy [14].

Currently, SGLT2 inhibitors are approved as good treatments for type II DM and Dapag is one of these categories of drugs and also has a good impact on hepatic enzyme levels. Without studying histological alterations in the hepatic tissues, Hazem et al. [15] reported that Dapag treatment protected the liver tissues against steatohepatitis in diabetes via inhibiting oxidative stress, inflammation and fibrosis progression. Meanwhile, the effects of Dapag with novel metal complexes on hepatic injury and metabolism and liver histological structure have not been illustrated before.

Liuran et al. [16] demonstrated that Dapag ameliorates hepatic steatosis by reducing lipogenic enzymes while inducing autophagy via the AMPK-mTOR pathway and thus improves hepatic functions and reduces hepatic lipid accumulation.

Therefore, this study aimed to investigate whether Dapag alone or its novel metal complexes (Zn, Cr and Se) could improve hepatic function parameters and insulin resistance in STZ-induced diabetic male rats and to assess the ameliorative role of complexes of metal ions Cr, Zn and Se with Dapag to enhance the hepatic functions.

## 2. Materials and Methods

### 2.1. Chemicals

Analytical-grade chemicals, solvents, zinc (II) chloride, chromium (III) chloride, selenium (IV) chloride and Dapagliflozin drug (purity 99%) were purchased from commercial suppliers (Sigma-Aldrich, St. Louis, MO, USA) and used for all experimental purposes without further purification.

### 2.2. Synthesis of Zinc (II), Chromium (III) and Selenium (IV) Complexes with Antidiabetic Drug (Dapagliflozin)

The zinc (II)–, Cr (III)– and selenium (IV)–Dapagliflozin complexes were synthesized by dissolving Dapagliflozin (1 mmol, 0.41 g) in 20 mL of methanol and then mixing with 1 mmol ZnCl<sub>2</sub>, CrCl<sub>3</sub>·6H<sub>2</sub>O and SeCl<sub>4</sub> in 20 mL CH<sub>3</sub>OH solution. The mixtures were heated at 80 °C with continuous stirring for 4 h. The mixtures were left to precipitate for 12 h, as with our previous experimental work preparation [12].

### 2.3. Characterization of Synthesized Zinc, Chromium and Selenium Dapagliflozin Complexes

#### 2.3.1. Differential Scanning Calorimetry (DSC)

The change in properties for zinc (II), Cr (III) and selenium (IV) complexes were investigated using differential scanning calorimetry (DSC) (DSC60, Shimadzu, Japan). The temperature ranged up to 350 °C and the rate of temperature was 10 °C/min at a flow rate of 20 mL/min under a N<sub>2</sub> gas atmosphere.

#### 2.3.2. Infrared Spectrophotometry (FTIR)

The FTIR analyses for zinc (II), chromium (III) and selenium (IV) complexes were carried out using an infrared Bruker spectrophotometer in the range of 400–4000 cm<sup>-1</sup>.

#### 2.3.3. Conductance Measurements

The conductance measurements with a concentration of 10<sup>-3</sup> mmol/L for the synthesized zinc (II), chromium (III) and selenium (IV) complexes in DMSO solvent were performed using the HACH conductivity meter model.

#### 2.3.4. Electronic Absorption Spectra

UV for zinc (II), chromium (III) and selenium (IV) complexes was recorded in DMSO solvent within the range of 800–200 nm using a UV2 Unicam UV/Vis Spectrophotometer fitted with a quartz cell of 1.0 cm path length.

### 2.3.5. X-ray Diffraction

The X-ray diffraction patterns for zinc (II), chromium (III) and selenium (IV) were recorded on X 'Pert PRO PANanalytical X-ray powder diffraction equipment, targeting copper with secondary monochromate.

### 2.3.6. Scanning Electron Microscopy

The surface morphologies for particles of zinc (II), chromium (III) and selenium (IV) complexes were carried out using a Quanta FEG 250 scanning microscope (SEM).

## 2.4. Experimental Animals

Sixty mature male rats (two months age), weighing 160–170 g, were kept in sensitized metal cages with free access to food ad libitum and water at a room temperature of  $25 \pm 2$  °C within a 12 h light/dark cycle. Rats were left to acclimatize for 2 weeks before the experiment. The experimental protocol was approved by the Deanship of Scientific Research at Taif University Ethical Committee, with approval number 42-0074. Rats were sacrificed under light ketamine/xylazine anesthesia, and every effort was made to reduce stress and pain. Animals involved in the experiment were fed high-fat diets for around 4 weeks before the start of the experiment to mimic type II diabetes, based on study [17]; then, experimental diabetes mellitus was induced by a single injection of STZ, as explained in detail in the next section.

Male rats were divided into 6 groups. Group I, the control group, was administered gum acacia as a vehicle. Group II, the STZ group, was given a single dose of STZ (50 mg/kg) by (I.P) injection [17]. Group III was a diabetic group given STZ and orally administrated 3 mg/kg of Dapg drug [15]. Groups IV, V and VI were the diabetic groups treated with STZ followed by oral administration of the same doses, 3 mg/kg, of Dapg/Cr, Zn and Se complexes, respectively, as shown in the experimental protocol (Figure 2).

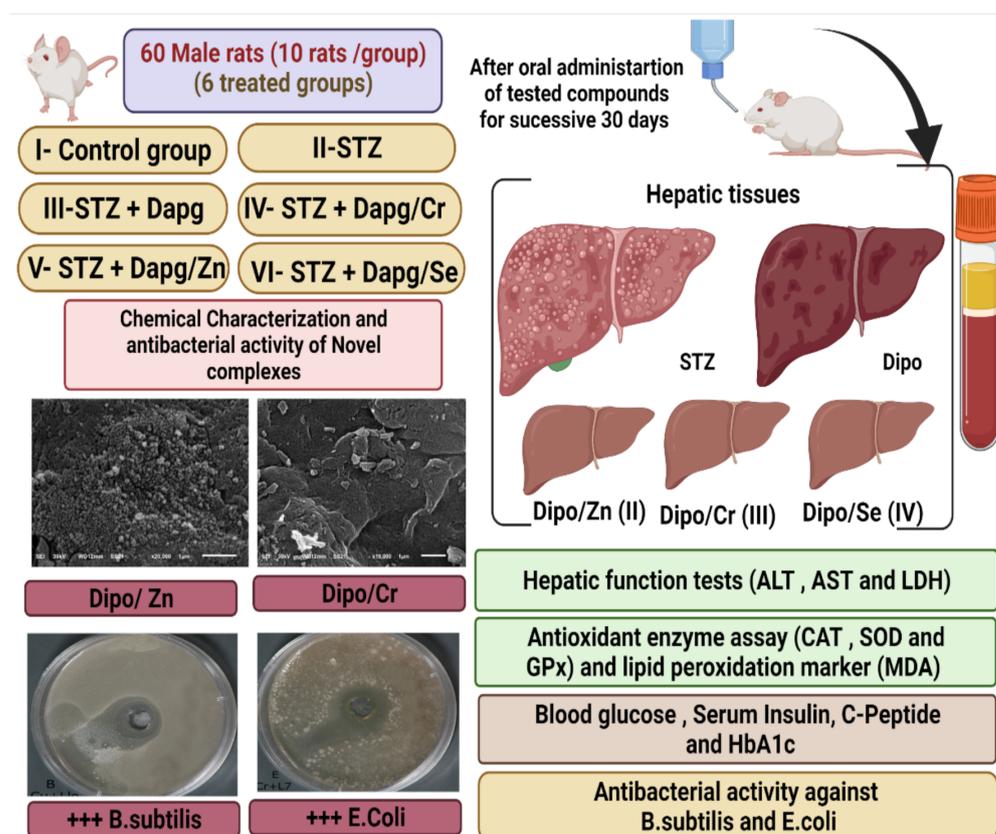


Figure 2. The experimental design and processes of the experiment.

### 2.5. Experimental Induction of DM

Freshly prepared STZ dissolved in phosphate-buffered saline (pH = 4.5) was used to induce diabetes in male rats. STZ (50 mg/kg) (I.P), which was freshly prepared in the early morning, was administered by intraperitoneal injection to male rats that had fasted for 6 h [17] following the administration of high-fat feed (composed of 66.5% commercial feed, 13.5% artificial butter and 20% sugar) for three weeks prior to the experimental induction of DM by a single injection of STZ [18]. Seventy-two hours after STZ injection, blood glucose levels were measured to evaluate the diabetic status of the animals. Subjects with blood glucose levels higher than 280 mg/dL were considered diabetic, as shown in the experimental protocol (Figure 2).

### 2.6. Blood Collection

Using capillary tubes, blood samples were collected from the eye plexus with light anesthesia for biochemical and physiological analyses. The rats were ethically decapitated. Hepatic tissue samples were kept at  $-25^{\circ}\text{C}$ .

### 2.7. Determination of the Fasting Blood Glucose Level

Fasting blood glucose levels were evaluated using commercial kits (Bio-Diagnostic Co.).

### 2.8. Measurements of Serum Insulin, C-Peptide and HbA1c

Serum insulin was evaluated using a rat ELISA kit (ALPCO Diagnostics). We used the C-peptide enzyme commercial immune assay (Sigma-Aldrich) and HbA1c kits according to the manufacturers' protocols.

### 2.9. Hepatic Function Activities and Biomarkers

ALT, AST and ALP levels were assessed by using (SENTINEL CH) kits. LDH levels were measured in accordance with the manufacturer's instructions.

### 2.10. Preparation of Hepatic Tissue Homogenates for the Determination of the Redox State

Small liver portions from hepatic tissues were used to determine oxidative injury. The tissues were immersed in 50 mmol phosphate buffer (pH 7.4); then, a protease inhibitor was added for the protection of enzymes from oxidation, and samples were centrifuged to obtain the supernatant of tissue homogenates.

### 2.11. Determination of Oxidative Stress Biomarker Activities in Hepatic Tissues

The MDA level was determined following the method of Ohkawa et al. (1979) [19]. SOD activity was determined using the technique of Marklund and Marklund (1974) [20]. CAT activity was estimated by applying the method of Aebi (1984) [21]. GRx was determined following Couri and Abdel-Rahman (1980) [22]. Glutathione peroxidase (GPx) was assayed using the technique of Hafeman et al. (1974) [23].

### 2.12. Histological Changes

Small hepatic tissue samples were fixed in 10% buffered formalin for further histological examination [24].

### 2.13. Antibacterial Activities of Dapg and Its Metal Complexes

The antimicrobial activity of the tested samples was determined by a modification of the Kirby–Bauer disc diffusion method [25]. Antibacterial activity was tested in triplicate and then the mean was calculated. In brief, 100  $\mu\text{L}$  of the best bacteria was grown in 10 mL of fresh media until they reached a count of approximately 108 cells/mL [26]. Then, 100  $\mu\text{L}$  of the microbial suspension was spread onto agar plates corresponding to the broth in which they were maintained. Isolated colonies of each organism that may play a pathogenic role were selected from the primary agar plates and tested for susceptibility by the disc diffusion method [27].

Plates were inoculated at 25 °C for 48 h. The Gram (+) bacterium *Bacillus subtilis* and the Gram (−) bacterium *Escherichia coli* were incubated at 35–37 °C for 24–48 h. Afterwards, the diameters of the inhabitation zones were measured in millimeters [28]. Standard discs of tetracycline (antibacterial agent) served as positive controls for the antimicrobial activity, and a filter disc impregnated with 10 µL solvent (dist. H<sub>2</sub>O, chloroform, DMSO) was used as a negative control.

The agar used was the Mueller–Hinton agar, which was tested continuously in terms of its pH. Further, the depth of the agar in the plate was a factor considered in the disc diffusion method. If an organism is placed on the agar, it will not grow in the area around the disc if it is susceptible to the chemical. The area of no growth around the disc is called the ‘zone of inhibition’ or ‘clear zone’. For the disc diffusion, the zone diameters were measured with slipping calipers from the National Committee for Clinical Laboratory Standards [29]. Agar-based methods, such as the Etest disc diffusion, can be good alternatives because they are simpler and faster to perform than the broth methods [30].

#### 2.14. Statistical Analysis

Data were expressed as mean ± standard error of the mean and analyzed by one-way analysis of variance using the SPSS v.22 program (SPSS Inc., Chicago, IL, USA). The significance of the mean differences was examined using the Duncan post hoc test [31].

### 3. Results

#### 3.1. Molar Conductance Data

Dapg complexes have high stability at room temperature and high solubility in DMSO with slight warming. The molar conductance data for Zn (II), Cr (III) and Se (IV) Dapg complexes are 34, 68 and 79 µs/cm, respectively. According to the conductance values of Cr (III) and Se (IV) complexes, there are one or two Cl<sup>−</sup> anions outside the chelation sphere [32]. According to the conductance values, the chromium (III) complex has 68 µs/cm, and the selenium (IV) complex has conductance equal to 79 µs/cm. This may be attributed to the presence of one Cl<sup>−</sup> anion outside the chelation sphere for the Cr (III) complex. Meanwhile, there are two ions, with one ion representing the chelation sphere. Meanwhile, for the Se (IV) complex, there are two Cl<sup>−</sup> anions outside the chelation sphere and thus the total number of ions is three, and the conductance value increases upon increasing the total number of ions [32]. By adding silver nitrate, a white precipitate is formed, which confirms the presence of chloride ions. The molar conductance for the Zn (II)–Dapg complex at 34 µs/cm and 25 °C showed that Zn (II) is non-electrolytic.

#### 3.2. Thermal Analysis

The changes in phase for Dapg and its Zn, Cr and Se complexes were investigated by differential scanning calorimetry. Dapg showed an endothermal peak at 77.98 °C, while Zn (II)–, Cr (III)– and Se (IV)–Dapg complexes showed different peaks at 110 °C, 108.79 °C and 115.58 °C, respectively, thereby confirming the formation of new complexes.

#### 3.3. Infrared

Fourier-transform infrared spectroscopy was used to characterize the stretching and molecular vibration of Dapg, which helped in the identification of the main functional groups of Dapg. By comparison with Zn (II), Cr (III) and Se (IV) Dapg complexes, the main sites of Dapg electron donation were confirmed. If the pure Dapg drug and its complex produced the same infrared spectroscopy (IR), it could be concluded that no new compounds were formed, whereas any disappearance or peak shifts confirm the formation of a new Dapg complex. The IR of pure Dapg and its Zn, Cr and Se complexes are shown in (Figure 3).

For IR of the free Dapagliflozin ligand, the assignments for stretching vibrational bands have been explained [33] as follows: the band appearing at 2940 cm<sup>−1</sup> is due to the vibrational stretching motion of aliphatic (C–H). The vibration bands appearing

at  $1100\text{--}11,050\text{ cm}^{-1}$  are assigned to  $\nu(\text{C-O})$ . The stretching vibration bands present at  $1550\text{--}1450\text{ cm}^{-1}$  are assigned to  $\nu(\text{C=C})$  aromatic rings. The bands appearing at  $1300\text{--}1200\text{ cm}^{-1}$  are assigned to  $\delta(\text{C-O-H})$ ; moreover, for the Dapagliflozin ligand, the characteristic stretching vibration for OH is observed at  $3282\text{ cm}^{-1}$  [34], while this peak is shifted to  $3294$ ,  $3295$  and  $3292\text{ cm}^{-1}$  for the Zn, Cr and Se Dapagliflozin complexes, respectively. For the zinc, chromium and selenium Dapagliflozin complexes, new bands appeared at  $419$ ,  $523$  and  $518\text{ cm}^{-1}$ , characteristic of  $\nu(\text{M-O})$  [35,36]. These changes in the IR absorption spectra confirm the chelation between the Dapagliflozin drug and metal ions.

For the IR of free Dapg, the characteristic stretching vibration for OH is observed at  $3282\text{ cm}^{-1}$  [33], whereas this peak is shifted to  $3294$ ,  $3295$  and  $3292\text{ cm}^{-1}$  for the Zn, Cr and Se Dapg complexes, respectively. The Zn, Cr and Se Dapg complexes showed new bands at  $419$ ,  $523$  and  $518\text{ cm}^{-1}$ , which corresponded to the M-O vibration mode [34,35]. These changes in the IR confirm the chelation between Dapg and the metal ions.

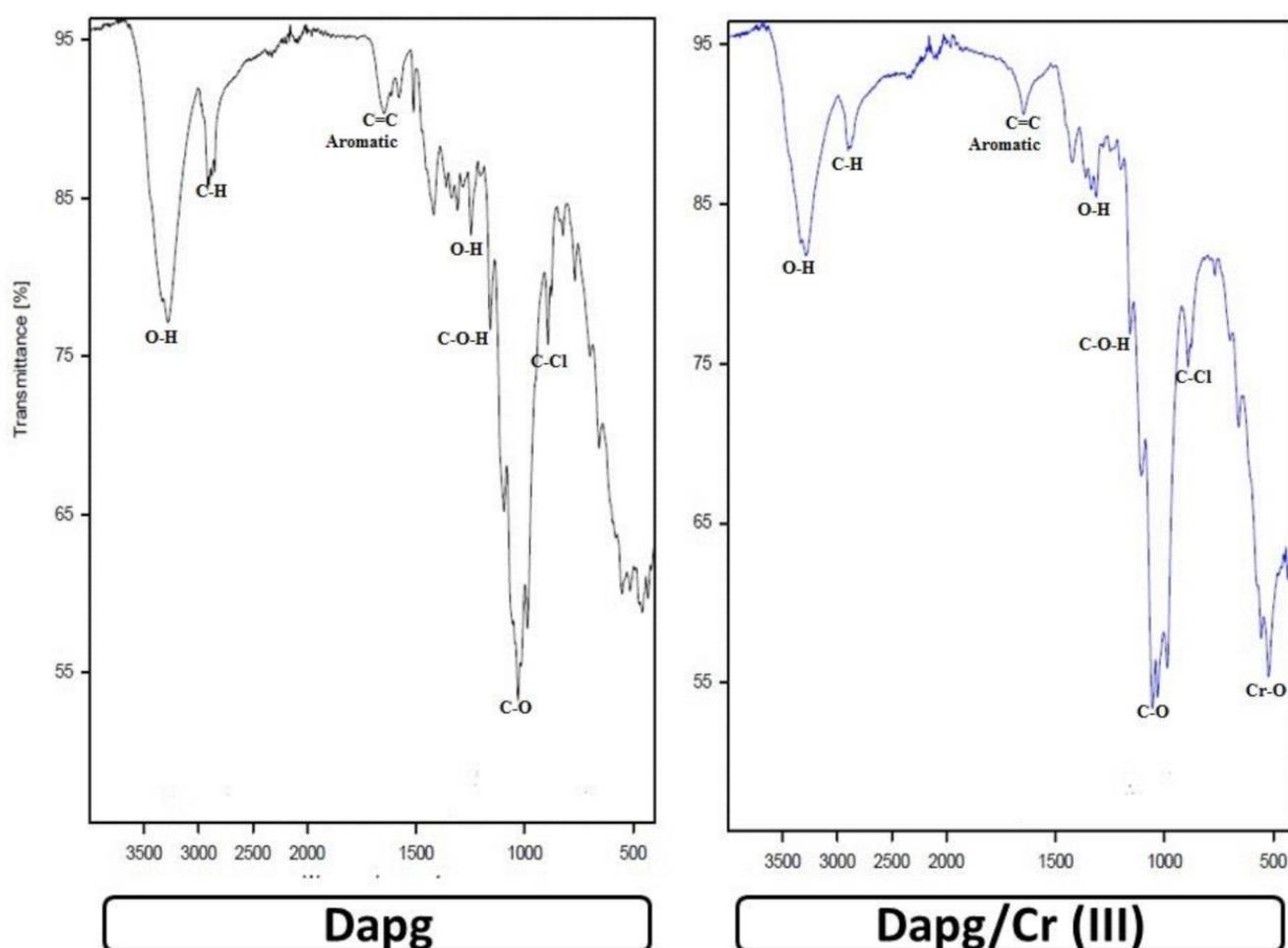
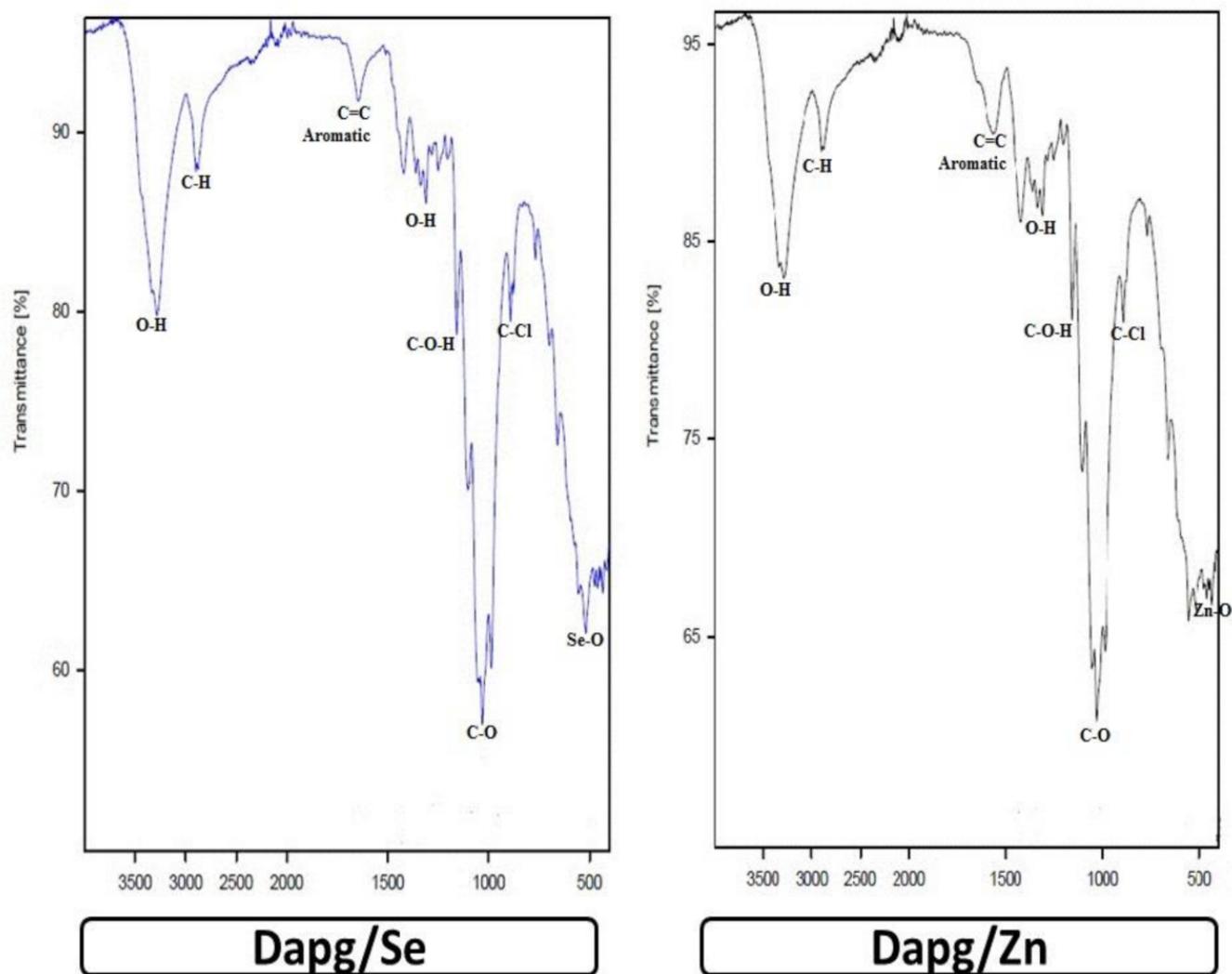


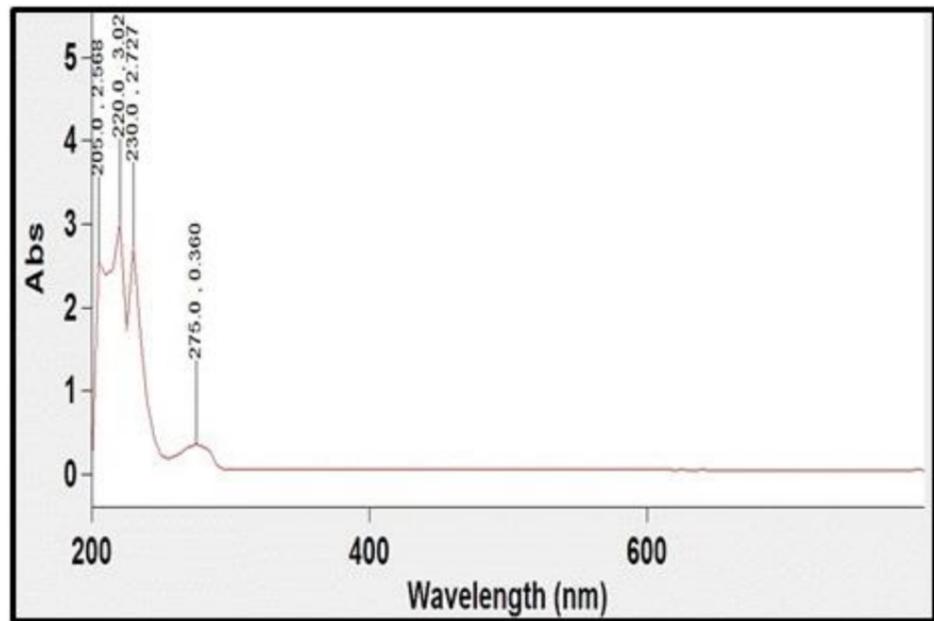
Figure 3. Cont.



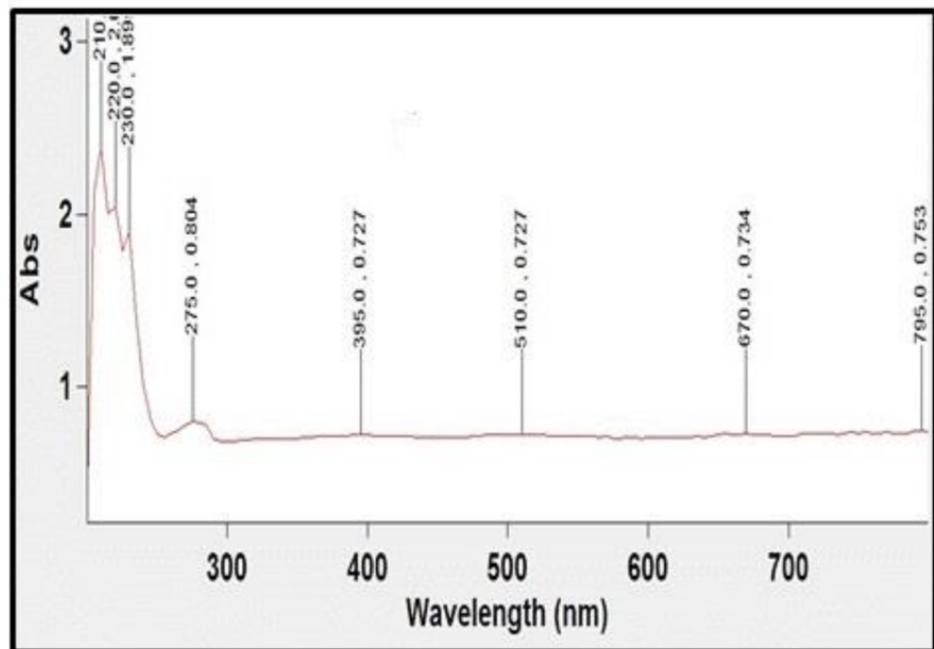
**Figure 3.** FT-IR of Dapg, Cr-Dapg, Zn-Dapg and Se-Dapg.

### 3.4. Electronic Spectra and Magnetic Measurements

For Dapg, two essential UV bands appeared at 230 and 275 nm (Figure 4), which refer to  $\pi \rightarrow \pi^*$ , for the aromatic ring and the OH and CH<sub>3</sub> groups [36] (Figure 4). For Zn (II) and Se (IV), they appeared at 275 and 335 nm, while for Cr (III), they appeared at 275 and 395 nm, which confirms the chelation of metal ions with Dapg (Figure 4). For the Cr (III) complex, the peaks appeared at 510, 670 and 795 nm, which refer to the metal–ligand charge transfer complex. The value of the magnetic moment for the Cr (III)–Dapg complex lies at 3.79 BM, which corresponds to an octahedral field [37].

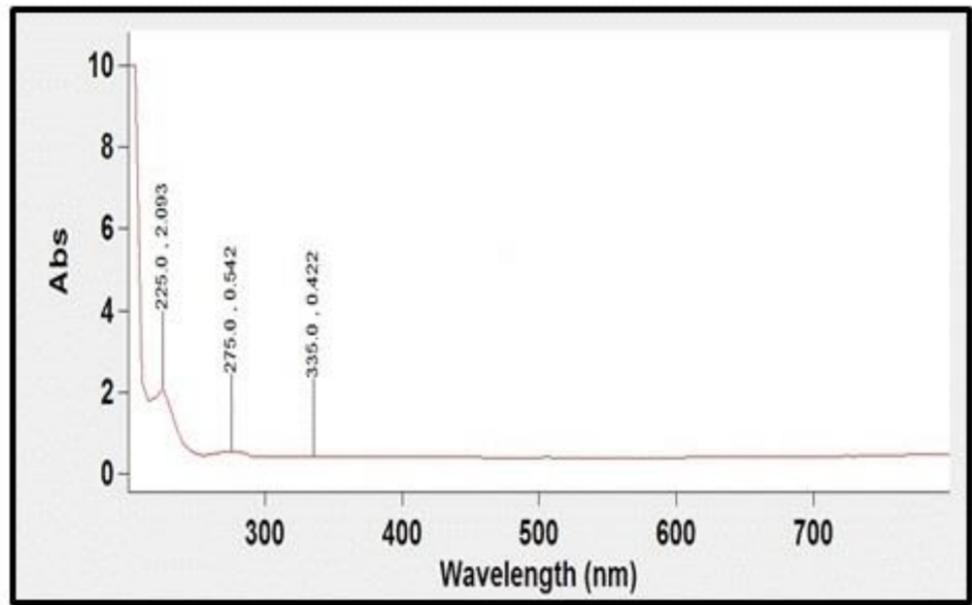


**Dapg**

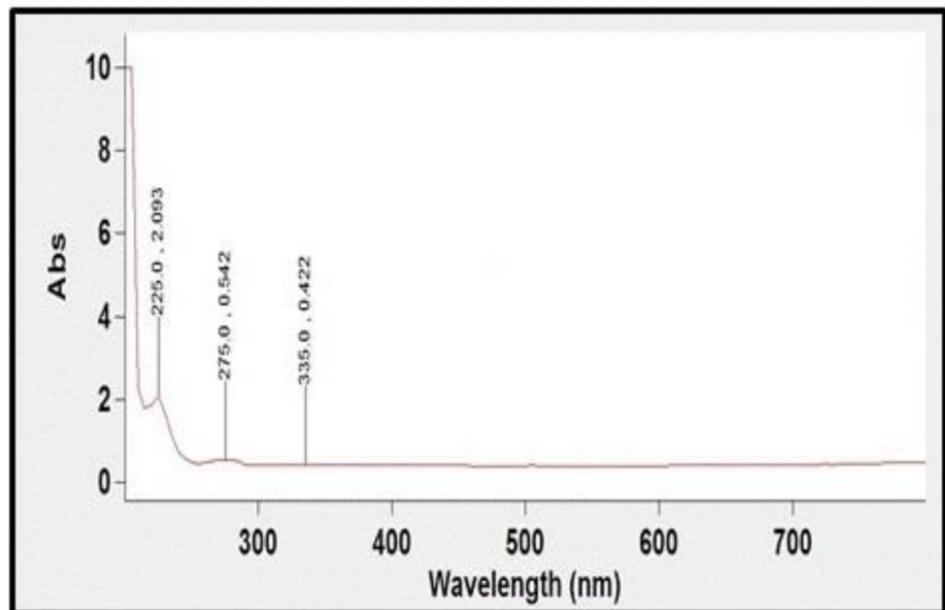


**Dapg/Cr**

Figure 4. Cont.



**Dapg/Se**



**Dapg/Zn**

Figure 4. UV of Dapg, Cr-Dapg, Zn-Dapg and Se-Dapg.

### 3.5. X-ray Diffraction (XRD)

XRD analysis was used to investigate the crystalline structure of Dapg and its metal complexes at room temperature using Cu K $\alpha$  radiation. For the synthesized Dapg metal complexes, the XRD patterns show that they are crystalline in nature [38]. The diffractograms of Zn (II), Cr (III) and Se (IV) give a sharp and strong Bragg's diffraction line. The Zn (II)-, Cr (III)- and Se (IV)-Dapg complexes showed two sharp lines at  $2\theta$  of (28, 31), (26, 28) and (27, 31), respectively. For the synthesized complexes, the crystalline size was calculated using the Scherrer formula [39,40]  $D = k\lambda / \beta \cos\theta$ , where  $k$  is a constant (0.94),  $\lambda$  is the wavelength of the X-ray used (0.154 nm), and  $\beta$  is the full width at half maxima peak of the XRD pattern. The crystalline size for Zn, Cr and Se complexes was found to be 65, 78 and 72 nm, respectively.

### 3.6. Scanning Electron Microscopy (SEM)

The images of SEM for the Dapg complexes are shown in Figure 5. The surface morphology changed according to the metal ions; some images contained a large number of irregular shapes, and some images contained regular grains. It is quite clear from the results that the average grain size estimated by SEM was significantly larger than the average grain size measured by XRD.

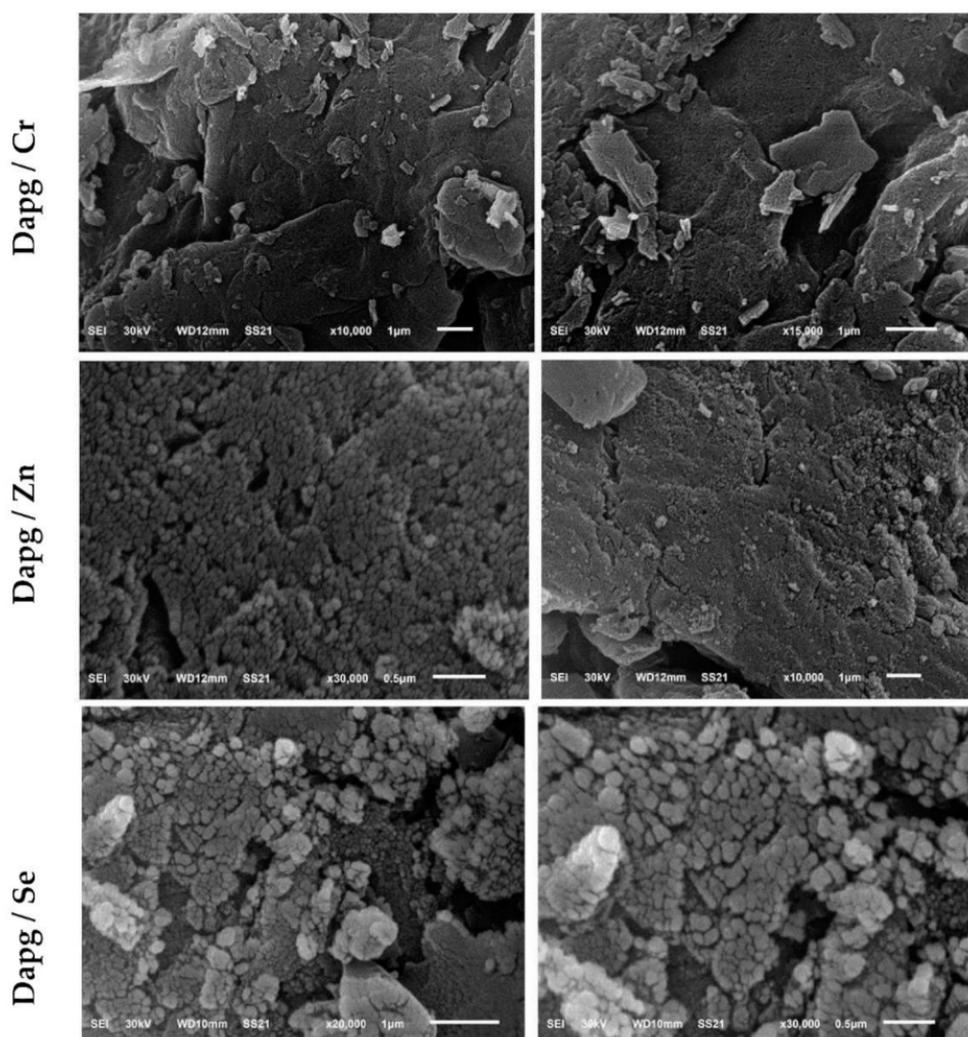


Figure 5. SEM image of Dapg complexes (Cr–Dapg, Zn–Dapg and Se–Dapg).

### 3.7. Antibacterial Activity Evaluation

Biological evaluations were performed in terms of the antimicrobial activities of the target compounds against Gram-positive (*Bacillus subtilis*, *Streptococcus pneumonia* and *Staphylococcus aureus*) and Gram-negative (*Escherichia coli* and *Pseudomonas aeruginosa*) bacteria. Results from the agar disc diffusion tests for the antimicrobial activities of the target compounds are presented in Table 1 and illustrated in Figure 6. The diameters of the zone of inhibition (in mm) of the standard drug tetracycline against Gram-positive bacteria *B. subtilis* and *S. aureus* and Gram-negative bacteria *E. coli* and *P. aeruginosa* were found to be 36, 30, 31 and 35 mm, respectively. Under identical conditions, Table 1 and Figure 6 show that all complexes were found to be efficient, with high antimicrobial activity.

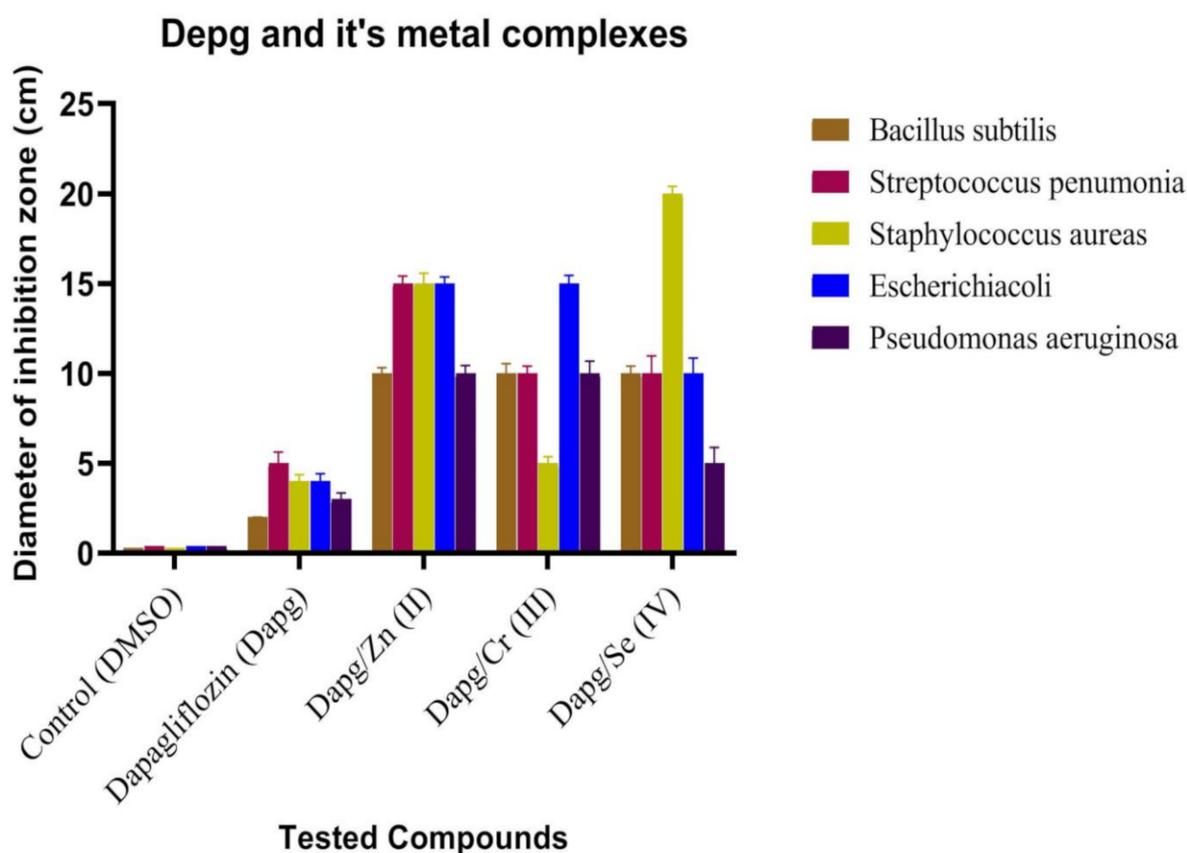


Figure 6. Antibacterial activity of Dapg and Zn (II)-, Cr (III)- and Se (IV)-Dapg metal complexes.

Table 1. Inhibition zone diameter (mm/mg sample) of Dipogliflazon and the complexes.

Sample	Inhibition Zone Diameter (mm/mg Sample)				
	<i>Bacillus subtilis</i> (G <sup>+</sup> )	<i>Streptococcus pneumonia</i> (G <sup>+</sup> )	<i>Staphylococcus aureus</i> (G <sup>+</sup> )	<i>Escherichia coli</i> (G <sup>-</sup> )	<i>Pseudomonas aeruginosa</i> (G <sup>-</sup> )
Control (DMSO)	0.0 ± 0.0 <sup>c</sup>	0.0 ± 0.0 <sup>d</sup>	0.0 ± 0.0 <sup>e</sup>	0.0 ± 0.0 <sup>d</sup>	0.0 ± 0.0 <sup>d</sup>
Dipogliflazon (Depg)	2 ± 0.01 <sup>b</sup>	5 ± 0.63 <sup>c</sup>	4 ± 0.36 <sup>d</sup>	4 ± 0.41 <sup>c</sup>	3 ± 0.35 <sup>c</sup>
Zn (II)-Dapg	10 ± 0.32 <sup>a</sup>	15 ± 0.41 <sup>a</sup>	15 ± 0.58 <sup>b</sup>	15 ± 0.36 <sup>a</sup>	10 ± 0.45 <sup>a</sup>
Cr (III)-Dapg	10 ± 0.54 <sup>a</sup>	10 ± 0.41 <sup>b</sup>	5 ± 0.36 <sup>c</sup>	15 ± 0.45 <sup>a</sup>	10 ± 0.69 <sup>a</sup>
Se (VI)-Dapg	10 ± 0.41 <sup>a</sup>	10 ± 0.98 <sup>b</sup>	20 ± 0.41 <sup>a</sup>	10 ± 0.85 <sup>b</sup>	5 ± 0.88 <sup>b</sup>

Means within the same column (mean ± SE) carrying different letters are significant at  $p \leq 0.05$  using Duncan's multiple range test, where the highest mean value has symbol (a) and those decreasing in value are assigned alphabetically.

### 3.8. Blood Glucose, Insulin and Fasting C-Peptide Levels

STZ elevated the fasting blood glucose levels, which was accompanied by a marked decline in insulin levels and serum fasting C-peptides in the STZ group as compared to the control group. As shown in Table 2, the group treated with STZ and Dapg displayed a non-marked elevation in blood glucose levels. They also demonstrated a significant decline in blood glucose levels, which was accompanied by the elevation of insulin and serum fasting C-peptide levels, as compared with the diabetic untreated group.

**Table 2.** Fasting blood glucose level, insulin hormone, HbA1C and fasting serum C-peptide results.

Groups	Fasting Blood Glucose (mg/dL)	Insulin Hormone (uIU/mL)	HbA1C (mmol/mol)	Fasting Serum C-Peptide (ng/mL)
Control group	80.81 ± 1.25 <sup>e</sup>	25.86 ± 2.15 <sup>a,b</sup>	3.12 ± 0.75 <sup>d</sup>	3.88 ± 0.19 <sup>a</sup>
STZ group	351.29 ± 8.02 <sup>a</sup>	4.70 ± 0.84 <sup>d</sup>	10.41 ± 1.26 <sup>a,b</sup>	0.22 ± 0.05 <sup>d</sup>
STZ plus Dapg group	120.96 ± 6.03 <sup>b</sup>	19.36 ± 2.85 <sup>c</sup>	5.02 ± 1.82 <sup>b</sup>	2.99 ± 0.87 <sup>c</sup>
STZ plus Cr–Dapg group	109.16 ± 4.75 <sup>c</sup>	23.03 ± 2.11 <sup>c</sup>	7.01 ± 0.87 <sup>c,d</sup>	3.76 ± 0.27 <sup>b</sup>
STZ plus Zn–Dapg group	102.26 ± 3.75 <sup>c</sup>	21.03 ± 2.91 <sup>c</sup>	5.01 ± 0.77 <sup>c,d</sup>	3.71 ± 0.77 <sup>b</sup>
STZ plus Se–Dapg group	92.27 ± 4.26 <sup>d,e</sup>	21.75 ± 1.02 <sup>b</sup>	5.54 ± 0.66 <sup>d</sup>	4.10 ± 0.76 <sup>a</sup>

Means within the same column (mean ± SE) carrying different letters are significant at  $p \leq 0.05$  using Duncan's multiple range test, where the highest mean value has symbol (a) and those decreasing in value are assigned alphabetically.

### 3.9. Oxidative Stress Biomarkers

Table 3 shows that experimental DM induced by STZ afforded a significant decline in catalase (CAT), superoxide dismutase (SOD) and glutathione peroxidase (GPx) antioxidant enzymes, while affording a highly significant elevation of the malondialdehyde (MDA) level. Diabetic groups treated with novel complexes of Cr–Dapg, Zn–Dapg and Se–Dapg showed a marked elevation of CAT, SOD and GPx levels and a significant decline in MDA levels. The group treated with Dapg in conjunction with Cr presented the best results.

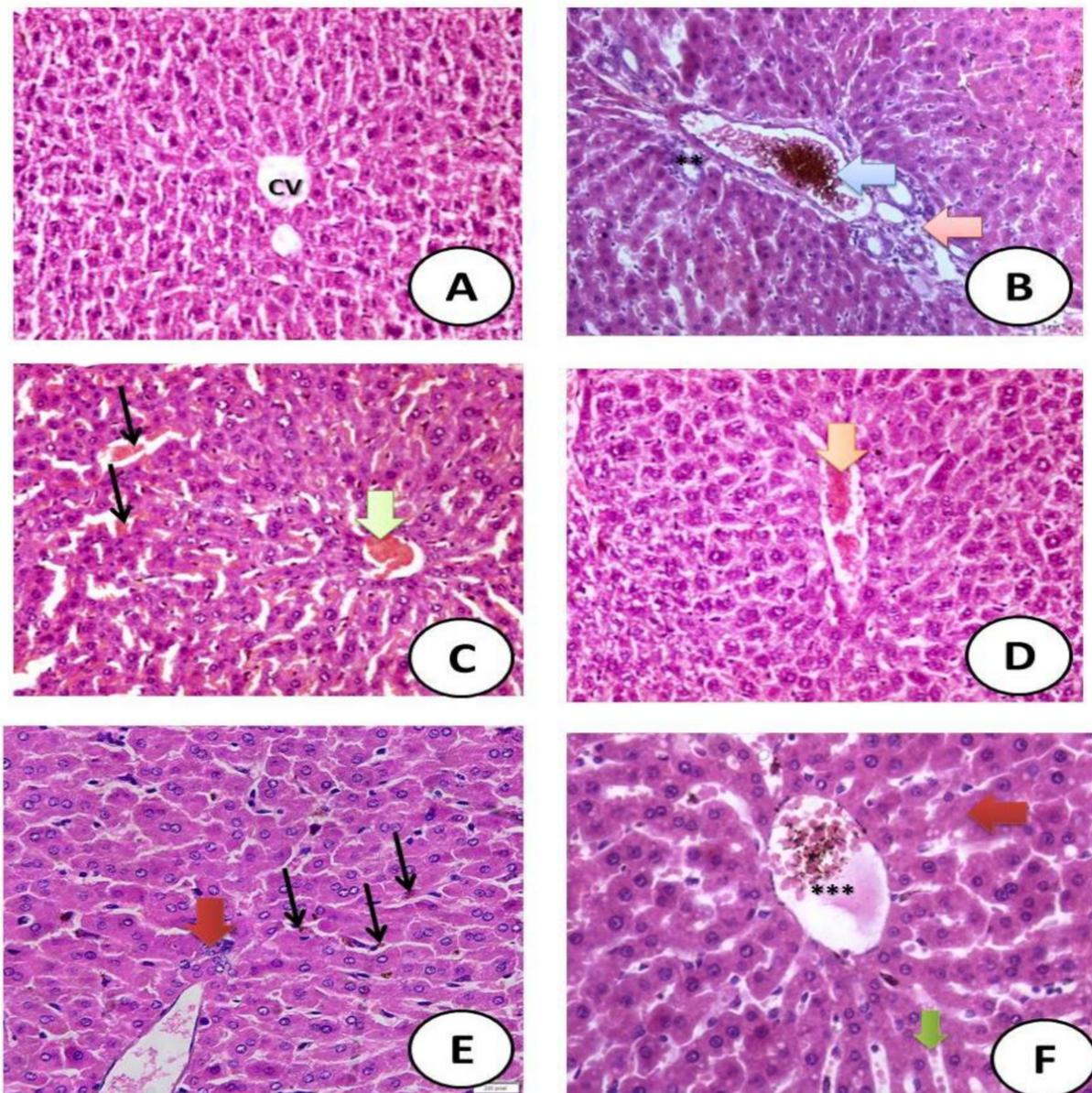
**Table 3.** Oxidative/antioxidant parameters of antioxidant enzymes in hepatic tissue of male rats.

Groups	Hepatic CAT (U/g)	Hepatic SOD (U/g)	Hepatic MDA (U/g)	Hepatic GPx (U/g)
Control group	1.88 ± 0.21 <sup>a</sup>	22.05 ± 1.15 <sup>a,b</sup>	3.05 ± 0.48 <sup>e</sup>	34.05 ± 1.85 <sup>a</sup>
STZ group	0.26 ± 0.10 <sup>d</sup>	5.22 ± 1.35 <sup>d</sup>	81.15 ± 0.96 <sup>a</sup>	7.56 ± 1.18 <sup>e</sup>
STZ plus Dapg group	1.42 ± 0.36 <sup>c</sup>	19.91 ± 1.58 <sup>c</sup>	20.42 ± 1.02 <sup>b</sup>	23.15 ± 1.15 <sup>d</sup>
STZ plus Cr–Dapg group	1.63 ± 0.48 <sup>b</sup>	20.52 ± 2.16 <sup>b</sup>	12.26 ± 1.45 <sup>c</sup>	26.41 ± 1.28 <sup>c</sup>
STZ plus Zn–Dapg group	1.74 ± 0.22 <sup>a</sup>	21.19 ± 2.25 <sup>b</sup>	8.78 ± 1.25 <sup>d</sup>	31.58 ± 1.58 <sup>b,c</sup>
STZ plus Se–Dapg group	1.88 ± 0.21 <sup>a</sup>	22.05 ± 1.15 <sup>a,b</sup>	3.05 ± 0.48 <sup>e</sup>	34.05 ± 1.85 <sup>a</sup>

Means within the same column (mean ± SE) carrying different letters are significant at  $p \leq 0.05$  using Duncan's multiple range test, where the highest mean value has symbol (a) and those decreasing in value are assigned alphabetically.

### 3.10. Histological Examination

Results showed that STZ caused significant degeneration in the hepatic structure, as well as increased eosinophilia, granular cytoplasm with hemorrhage and necrotic nuclei. Meanwhile, treatment of different groups with the novel metal complexes of Dapg led to the significant amelioration of hepatic tissue with mild toxicity, including some ballooning of hepatocytes and necrotic nuclei as shown in Figure 7.



**Figure 7.** Histological sections. (A) Control group showing normal hepatic structure with normalized central vein (CV) (H&EX400). (B) STZ-treated group showing severe hepatic toxicity in the form of hypertrophy of hepatocytes with appearance of binucleated hepatocytes and increased eosinophilia, granular cytoplasm (Blue arrow) and vesicular nuclei. The portal tract showed dilated irregular portal vein filled with blood, necrotic tissue (red arrow), focal necrosis in some hepatocytes with increased eosinophilia and nuclear disappearance and a ductal reaction (new bile duct formation) at the periphery of the portal tract with periportal fibrosis. (C) STZ + Dapg showing moderate hepatotoxicity in the form of partial degeneration of hepatic tissue with necrotic nuclei (black arrow) with bleeding (green arrow) (H&EX400). (D) STZ + Zn–Dapg-treated group showing great amelioration in the hepatic structure with mild bleeding and red blood cells (orange arrow) (H&EX400). (E) STZ + Cr–Dapg showing very mild toxicity in the form of hypertrophy of hepatocytes with some binucleated hepatocytes (black arrow) that contained mild brown particles of bilirubin indicating biliary tract obstruction (black arrow), focal necrosis of some hepatocytes (red arrow) and mild ductal reactions at the periphery of the central vein (H&EX400). (F) STZ + Se–Dapg showing very mild toxicity in the form of hypertrophy of hepatocytes with granular eosinophilic cytoplasm, with mild congested central vein (\*\*\*), ballooning degeneration in some hepatocytes (red arrow) and focal and single hepatocyte necrosis (green arrow).

#### 4. Discussion

Dapg treatment can greatly control type II DM by improving glycemic control, reducing hepatic fat accumulation, maintaining normal liver size and elevating the hepatic insulin level in the blood [41]. Thus, the current study is of great importance considering the clinically high importance of controlling blood glucose levels and ameliorating hepatic functions using Dapg with metals, including Zn, Cr and Se.

Different clinical trials for patients with DM have been conducted. These studies found a significant decline in aspartate aminotransferase (AST) and alanine aminotransferase (ALT) levels following the administration of Dapg. Previous studies concluded that such reductions were also seen in the glycemic values, which is similar to the results of the current study. However, in this study, it was found that Dapg novel metal complexes induced a greater reduction in hepatic transaminases and amelioration of the glycemic state as sitagliptin previous findings and concepts [42].

In accordance with the results of the current study, previous studies reported that Dapg induced an improvement in liver enzymes (ALT or AST) and such improvements could induce a reduction in insulin resistance [43], thereby ameliorating the homeostatic model assessment for insulin resistance. This indicates a direct effect of Dapg on the inflammatory profile of hepatic tissues. These results support the findings regarding the high improvement induced by the Dapg novel metal complexes. In addition, these metal conjugations with Dapg induced additional improvements in the drug characteristics [44].

The effects of SGLT2 on improving fat accumulation in the hepatic tissues have been reported. Previous studies reported a reduction in hepatic triglycerides and a decline in hepatic fibrosis [45]. In the current study, treatment with Dapg elicited a moderately significant alteration in hepatic indicators. Meanwhile, treatment with Dapg novel metal complexes with either Cr, Zn and Se ameliorated hepatic injury and suppressed oxidative stress [46–48]. The amelioration of biochemical hepatic functions with novel complexes is important, as [49] demonstrated that excessive lipid accumulation can lead to cellular injury and death, and this strengthens the role of Dapg metal complexes in the protection of hepatic functions and reduction of lipid accumulation.

Parallel to the current findings, treatment of *db/db* male mice with Dapg for one month significantly reduced blood glucose levels in the circulation and increased glucose excretion in the urine. These observations were associated with reduced progression of albuminuria and inflammation [50]. Additionally, liver markers of inflammation, liver enzymes, oxidative stress markers and fat content were also reduced by Dapg treatment, and the success point of our current study is that we confirmed the previously obtained results for Dapg and its additional hepatoprotective effect and higher antioxidant defense capacities against oxidative stress in the case of treatment with Dapg with either Cr, Zn and Se. A great hepatic function improvement was recorded in the Se–Dapg-treated group and this confirms the ability of Se to enhance the activity of Dapg, as reported for the role of Se in hepatic protection [51].

Lipid peroxidation is suppressed by the release of some important antioxidant enzymes, such as glutathione [52]. Glutathione plays an essential antioxidative role against excessive reactive oxygen species [53]. In the current study, a marked alteration in either glutathione reductase (GRx) or GPx levels was observed in the STZ-induced diabetic rats. Meanwhile, these parameters were improved in the diabetic rats treated with the Dapg novel metal complexes.

MDA is an essential marker for the induction of lipid peroxidation [54]. In the current study, DM caused the elevation of MDA levels. The biological and pharmacological activities of the novel complexes of Dapg with Zn, Cr and/or Se ameliorated the hepatic functions and rejuvenated the antioxidant defense system [55]. The present study demonstrated that DM induced hepatotoxicity and severe oxidative stress, which could be mitigated by the administration of Dapg metal complexes with Zn, Cr and/or Se. The Dapg metal complexes also improved the hepatic structure more than Dapg alone.

To confirm the finding that Cr with Dapg ameliorates the blood glucose level, a previous study revealed that Cr with antidiabetic drugs caused a significant decline in the blood glucose level in alloxan-induced diabetic rats as compared to non-treated diabetic rats [56].

Reactive oxygen species, such as free radicals, play essential roles in the progression of STZ-induced experimental DM, as well as in the triggering of severe oxidative injury [57]. Under excessive oxidative stress conditions, reactive oxygen species directly participate in the induction of many diseases [52]. The human body is well equipped with many antioxidant enzymes, such as SOD, CAT and glutathione (GPx and GRx) [58]. The generation of high levels of free radicals can result in severe oxidative tissue or organ injuries [57].

People who have had diabetes for a long time may have peripheral nerve damage and reduced blood flow to their extremities, which increases the chance of bacterial infection. The high sugar levels in their blood and tissues allow bacteria to grow and allow infections to develop more quickly, which will eventually affect all the organs of the body [59]. The findings of current study showed that complexes of Dapg with Zn, Cr and Se contributed to the lowering of blood glucose levels in treated groups, which is of great importance for diabetics, and these complexes also showed high antibacterial activity against different types of bacterial strains related to the digestive system and its physiological functions, such as *Escherichia coli* and *Bacillus subtilis*. Moreover, a point of novelty and importance is the antibacterial activity of Zn–Dapg against *Streptococcus pneumoniae*, which is of great importance for the protection of diabetics, especially during the COVID-19 pandemic.

The current study showed an improvement in the hepatic functions and antioxidant defense system in the group treated with the Dapg novel complexes. These effects were indicated by significant reductions in portal inflammatory cell infiltration and hemorrhage.

## 5. Conclusions

The present study provides new information on novel complexes of Dapg with Zn, Cr and/or Se and their potent hepatoprotective effects against STZ-induced experimental DM and hepatic injury. The novel Dapg metal complexes prevented histopathological alterations in DM and inhibited the excessive triggering of reactive oxygen species that induced oxidative injury in diabetic male rats. In addition, the Cr–Dapg complex upregulated the hepatic antioxidant defense system, thereby preventing severe oxidative injury. Therefore, Dapg novel complexes may be considered potent agents for attenuating the hepatic injury and pathophysiological complications of DM. In fact, chronic liver disease is an emerging risk factor for increased mortality due to COVID-19 [60]. As the COVID-19 pandemic continues, these new data may provide a potential tool to reduce mortality, and clinical studies are recommended.

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