

Article

## Role of Alloying Additions in the Solidification Kinetics and Resultant Chilling Tendency and Chill of Cast Iron

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**Abstract:** The present work describes the effect of the solidification processing and alloy chemistry on the chilling tendency index, CT, and the chill,  $w$ , of wedge-shaped castings made of cast iron. In this work, theoretical predictions were experimentally verified for the role of elements, such as C, Si, Mn, P and S, on the cast iron CT. In addition, inoculation and fading effects were considered in the experimental outcome. Accordingly, the graphite nucleation coefficients,  $N_s$ ,  $b$ , the eutectic cell growth coefficient,  $\mu$ , and the critical cooling rate,  $Q_{cr}$ , for the development of eutectic cementite (chill) were all determined as a function of the cast iron chemistry and time after inoculation. In particular, it was found that increasing the Mn and S contents, as well as the time after inoculation lowers the critical cooling rate, thus increasing the chilling tendency of the cast iron. In contrast, C, Si and P increase the critical cooling rate, and as a result, they reduce the cast iron CT and chill.

**Keywords:** chill; chilling tendency; gray cast iron; role of C; Si; Mn; P; Mn; inoculation effects

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## 1. Introduction

One of the important indices that accounts for the quality of cast iron is its chilling tendency, that is its tendency to solidify according to the Fe-C-X metastable system. The chilling tendency, CT, depends on the physical-chemical state of the liquid iron, while the chill (the fraction of eutectic cementite in the casting) formation depends additionally on the casting cooling rate. In foundry practice, the CT for the various types of cast iron is determined from comparisons of the exhibited fraction of eutectic cementite (chill) in castings solidified under similar cooling rates. Based only on these comparisons, the difference in the chilling tendency for various cast irons can be established, but the absolute CT values for given irons cannot be disclosed [1].

It is well known that the chilling tendency of cast irons determines their subsequent performance in diverse applications. In particular, cast irons possessing a high CT tend to develop zones of white or mottled iron. Considering that these regions can be extremely hard, their machinability can be severely impaired. Alternatively, if white iron is the desired structure, a relatively small chilling tendency will favor the formation of gray iron. This, in turn, leads to low hardness and poor wear properties in as-cast components. Hence, considerable efforts [1–7] have been made to correlate the inoculation practice, iron composition, pouring temperature, *etc.*, with the cast iron CT.

In the published literature, there are only a few attempts aimed at elucidating the mechanisms responsible for the chill of cast iron [3,8–11]. Besides, none of the proposed hypotheses take into account the complexity of the solidification process. In most cases, the proposed theories assume that a single factor is the determinant in establishing the final solidification structure, while the remaining factors are ignored. In addition, various numerical models have been proposed [12,13] to predict whether a given casting or a part of it will solidify according to the stable or metastable Fe-C-X system. Yet, their application is tedious due to extensive numerical calculations. Accordingly, in this work, a simple analytical model is employed to account for the mechanism or mechanisms responsible for the CT of cast iron and, hence, the exhibited chill.

### 1.1. Analysis

During cast iron solidification, two processes become relevant: (a) nucleation; and (b) growth of eutectic cells, which have a direct effect on the final microstructure.

#### (a) Nucleation of eutectic cells

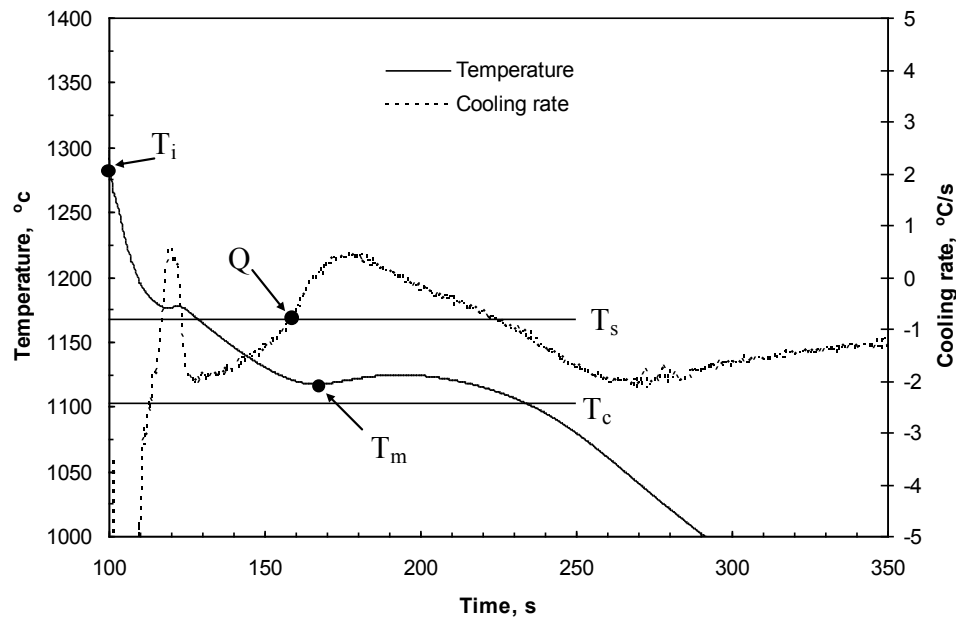
In liquid-solid transformations, nucleation is heterogeneous in nature. Accordingly, a simple model for heterogeneous nucleation of graphite in cast iron has recently been proposed [14], which describes the nucleation of eutectic cells,  $N$ , by the following expression:

$$N = N_s \exp\left(-\frac{b}{\Delta T_m}\right) \quad (1)$$

and

$$\Delta T_m = T_s - T_m \quad (2)$$

where  $N_s$  and  $b$  are the nucleation coefficients,  $\Delta T_m$  is the maximum degree of undercooling at the onset of graphite eutectic solidification,  $T_s$  is the graphite eutectic equilibrium temperature and  $T_m$  is the minimal temperature at the onset of graphite eutectic solidification (see Figure 1).



**Figure 1.** Cooling curves and cooling rate curves for cast iron.  $T_i$  = initial temperature of the metal in the mold cavity just after pouring;  $T_s$  = graphite eutectic equilibrium temperature;  $T_c$  = cementite equilibrium temperature;  $T_m$  = minimal temperature; and  $Q$  = cooling rate of cast iron at the onset of graphite eutectic solidification.

It is well known that each graphite nucleus gives rise to a single eutectic cell. Hence, it can be assumed that a measure of the graphite nuclei density is given by measurements of the eutectic cell count (cells/volume). Thus, the eutectic cell count can be related to the graphite nucleation potential through Equation (1), where the melt is characterized by the nucleation coefficients,  $N_s$ , and  $b$ .

(b) The growth rate of eutectic cells can be described by [15]:

$$u = \mu \Delta T^2 \quad (3)$$

where  $\mu$  is the growth coefficient of the eutectic cells and  $\Delta T$  is the melt undercooling.

## 1.2. Onset of Graphite Eutectic Solidification

The heat generated during the solidification of cast iron depends on both the cell count and the growth rate of the eutectic cells. Thus, by combining the heat extraction from the mold with the heat generated during the solidification of eutectic cells, the minimal temperature,  $T_m$  (Figure 1), or the maximum undercooling,  $\Delta T_m = T_s - T_m$ , at the onset of graphite eutectic solidification can be determined from [4]:

$$T_m = T_s - \Delta T_m = T_s \left[ \frac{4c_{ef}Q^3}{\pi^3 L_e N \mu^3 f} \right] \quad (4)$$

where:

$$Q = \frac{2T_s a^2}{\pi \phi c_{ef} M^2} \quad (5)$$

$$c_{ef} = c + \frac{L_\gamma}{T_{l\gamma} - T_s} \quad (6)$$

$$\phi = cB + c_{ef}B_1 \quad (7)$$

$$B = \ln \frac{T_i}{T_l}; \quad B_1 = \ln \frac{T_l}{T_s} \quad (8)$$

where  $M$  is the casting modulus,  $T_i$  is the initial liquid metal temperature just after pouring into the mold and  $T_l$ ,  $T_{l\gamma}$ ,  $a$ ,  $c$ ,  $L_e$ ,  $L_\gamma$  and  $f$  are various solidification parameters, defined in Table 1.

**Table 1.** Selected thermophysical data [14]. C, Si, P: % content of carbon, silicon and phosphorus in cast iron, respectively.

Parameter	Value and Units
Latent heat of eutectic graphite	$L_e = 2,028.8$ ; J/cm <sup>3</sup>
Latent heat of austenite	$L_\gamma = 1,904.4$ ; J/cm <sup>3</sup>
Specific heat of cast iron	$c = 5.95$ ; J/(cm <sup>3</sup> ·°C)
Material mold ability to absorb heat	$a = 0.11$ ; J/(cm <sup>2</sup> ·s <sup>1/2</sup> ·°C)
Liquidus temperature for pro-eutectic austenite	$T_l = 1,636 - 113(C + 0.25 \text{ Si} + 0.5 \text{ P})$ ; °C
Carbon content in eutectic graphite	$C_e = 4.26 - 0.30 \text{ Si} - 0.36 \text{ P}$ ; %
Maximum carbon content in austenite at $T_s$	$C_\gamma = 2.08 - 0.11 \text{ Si} - 0.35 \text{ P}$ ; %
Liquidus temperature of pre-eutectic austenite when its composition is $C_\gamma$	$T_{l\gamma} = 1,636 - 113(2.08 + 0.15 \text{ Si} + 0.14 \text{ P})$ ; °C
Weight fraction of austenite cast iron at the beginning of eutectic solidification	$g_\gamma = (C_e - C)/(C_e - C_\gamma)$
Austenite density	$\rho_\gamma = 7.51$ g/cm <sup>3</sup>
Liquid cast iron density	$\rho_l = 7.1$ g/cm <sup>3</sup>
Volume fraction of liquid in cast iron at the beginning of eutectic solidification	$f = \rho_\gamma g_l / [\rho_\gamma g_l + \rho_l (1 - g_l)]$

Combining Equations (1) and (4) yields:

$$\Delta T_m = T_s - T_m = \frac{b}{8 \text{ProductLog}[y]} \quad (9)$$

where:

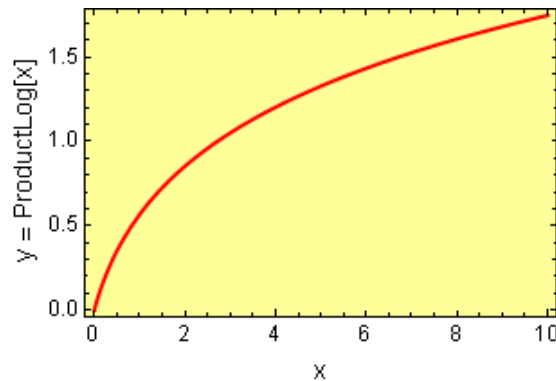
$\text{ProductLog}[y] = x$  is the Lambert function (see <http://mathworld.wolfram.com/LambertW-function.html>), also known as the omega function, graphically shown in Figure 2. This function can be easily calculated by means of the instruction  $\text{ProductLog}[y]$  in the Mathematica™ program.

Additionally:

$$y = \frac{b}{8 \cdot 2^{1/4}} \left[ \frac{\pi^3 \text{Le} N_s \mu^3 f}{c_{ef} Q^3} \right]^{1/8} \quad (10)$$

Combining Equations (1) and (9), the spatial cell count can be determined from:

$$N = \frac{N_s}{\exp[8\text{ProductLog}(y)]} \quad (11)$$



**Figure 2.** Graphic representation of the ProductLog[y] function for  $y \geq 0$ .

### 1.3. Chilling Tendency Index

Figure 3 shows the cooling curves with the  $T_m$  values, including the eutectic graphite and  $N$  along the wedge axis for two solidified cast irons with different physical-chemical states. The temperature range  $\Delta T_{sc} = T_s - T_c$  is also given in Figure 3. Notice that in the temperature range  $\Delta T_{sc} = T_s - T_c$ , eutectic graphite is the only growing structure (gray cast iron). From this figure, it is apparent that increasing  $Q$  to values equal to  $Q_{cr}$  leads to an increase of eutectic cells from  $N$  to  $N_{cr}$ . Below  $T_c$ , both graphite and eutectic cementite grow simultaneously, interfering with each other and giving rise to mottled cast iron. Accordingly,  $T_c$  can be considered as the transition temperature for the solidification of eutectic cementite from eutectic graphite or the chill formation temperature. From this figure, it is apparent that an increase of the cooling rate from  $Q$  to  $Q_{cr}$  leads to a reduction in  $T_m$  to  $T_c$  and, hence, to the formation of eutectic cementite (chill development). Thus, determination of the critical cooling rate,  $Q_{cr}$ , is key in establishing the critical conditions for the development of eutectic cementite (chill) in castings.

When  $Q = Q_{cr}$  or  $M = M_{cr}$  and  $\Delta T_m = \Delta T_{sc}$  (Figure 3), the critical cooling rate can be estimated from Equations (1) and (4) as:

$$Q_{cr} = \pi\mu \left[ \frac{L_e N_s f \Delta T_{sc}^8}{4c_{ef}} \exp\left(-\frac{b}{\Delta T_{sc}}\right) \right] \quad (12)$$

Taking into account Equations (1) and (4) (for  $\Delta T_m = \Delta T_{sc}$ ) and (5), the critical casting modulus,  $M_{cr}$ , under which it is possible to develop a chill yields:

$$M_{cr} = pCT \quad (13)$$

where:

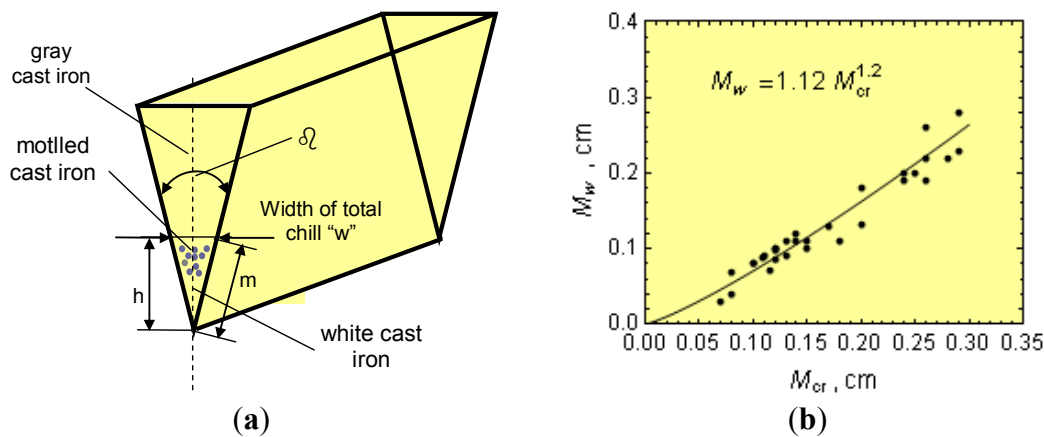
$$p = \frac{a}{\pi} \left( \frac{32T_s^3}{L_e c_{ef}^2 \phi 3} \right)^{1/6} \quad (14)$$



$$M_w = \frac{F_{ch}}{m} = \frac{\frac{1}{2} h \frac{w}{2}}{\frac{h}{\cos(\beta/2)}} = w \frac{\cos(\beta/2)}{4} \quad (16)$$

$$M_{cr} = p; CT = M_w \quad (17)$$

In the above expression,  $\beta$  is the wedge angle,  $F_{ch}$  is the half surface area chill triangle and  $h_{cr}$  and  $m$  are the critical chill height and length (Figure 4a). Taking into account Equations (14) and (17), an expression is found that relates  $M_{cr}$  to  $M_w$ .



**Figure 4.** (a) Wedge geometry and (b) relationship between the casting modulus of the wedge with the critical casting modulus.

Nevertheless, the wedge cannot be considered as a simple body. The wedge thin sections are heated up from the thicker sections, and this is not taken into account in Equation (16). Accordingly, Equation (17) is corrected to reflect this effect. Considering the experimental  $w$  and  $\beta$  values and plugging them into Equation (16), the modulus  $M_w$  can be estimated. Figure 4b shows the relationship between the modulus  $M_w$  and  $M_{cr}$ . Notice from this figure that  $M_w \neq M_{cr}$  and that the function  $M_w = f(M_{cr})$  can be described by:

$$M_w = 1.12 M_{cr}^{1.2} \quad (18)$$

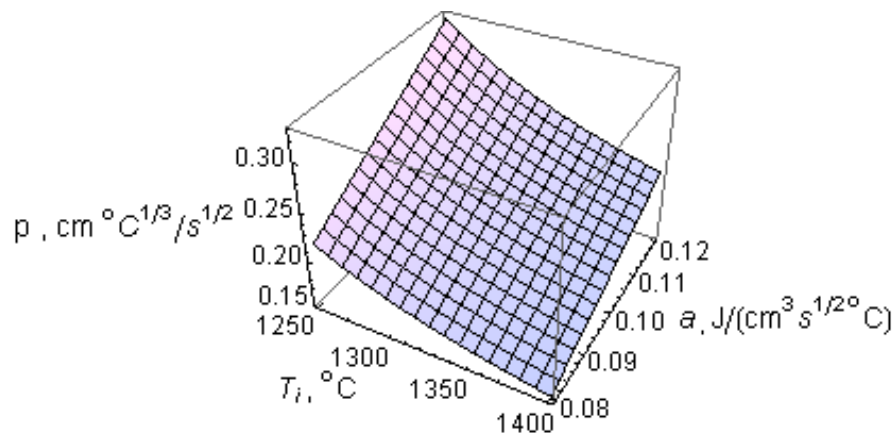
Thus, by combining Equations (16) to (18), the “ $w$ ” can be expressed as a function of the CT.

$$w = \frac{4.46 p^{1.2}}{\cos(\beta/2)} CT^{1.2} \quad (19)$$

Notice that  $w$  depends additionally on the  $p$  coefficient, which includes parameters related to the cooling rate (Equation (4)), that is:

- The ability for the mold to absorb heat,  $a$ ;
- The  $\phi$  parameter, Equation (7), which depends on the  $B$  and  $B_1$  values (Equation (8)), that is on the initial temperature,  $T_i$ , of the cast iron just after pouring the wedges.

The effect of  $T_i$  and the ability of the wedge mold to absorb heat,  $a$ , on the  $p$  parameter is shown in Figure 5. Notice that the  $p$  parameter increases as the pouring temperature  $T_p$  (and, hence,  $T_i$ ) decreases, and the ability of the wedge mold to absorb heat increases. As a result, wedge removal at lower temperatures from molds with mold materials having an increasing ability to absorb heat,  $a$  (Table 2), leads to increasing chills in order to keep CT constant.



**Figure 5.** Effect of initial liquid metal temperature just after pouring into the mold,  $T_i$ , and the ability of the mold to absorb heat,  $a$ , on the  $p$  parameter.

**Table 2.** Reported  $a$  values for mold materials [11,16,17].

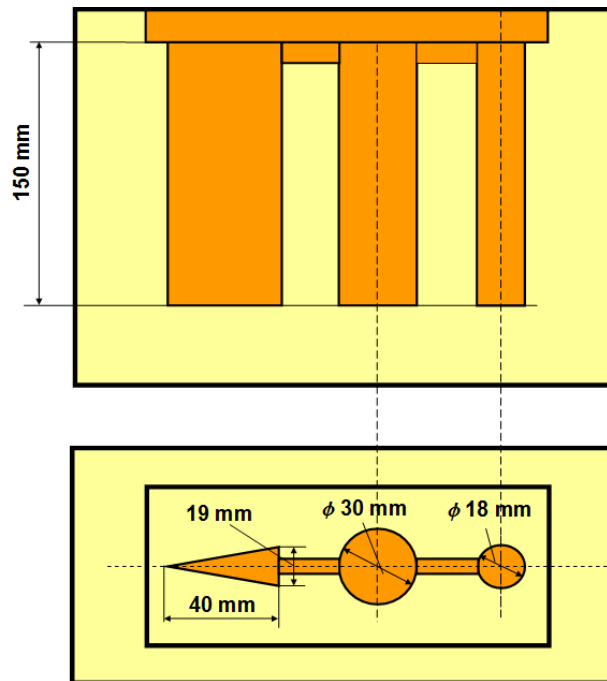
Mold Material	Material Mold Ability to Absorb Heat, $a$ , J/(cm <sup>2</sup> ·°C·s <sup>1/2</sup> )
Chromite	0.15
Zircon	0.13–0.15
Quartz sand	0.08–0.12
Olivine	0.10
Chemically-bonded urethane molding sand	0.10–0.12
Chemically-bonded furan molding sand	0.08–0.10
Chemically-bonded shell molding sand	0.10–0.11
LDASC	0.018–0.03

## 2. Experimental Section

Experimental melts were made in two stages. In the first stage, 40-mm diameter bars containing pig iron and scrap steel were prepared. The scrap composition consisted of Sorelmetal (4.27% C, 0.132% Si, 0.13% Mn, 0.026% P, 0.005% S) and steel scrap (0.2% C, 0.02% Si). In the second stage, the bars, together with various additives, such as carbon, Si, Fe-P and Fe-S, were charged into an induction furnace with a capacity of 1.5 kg. Melting was carried out at low temperatures (*i.e.*, below the equilibrium temperature for the  $2C + SiO_2 = Si + 2CO$  reaction in order to avoid C losses as CO to the atmosphere [18]). In addition, low S and Mn contents were employed in order to minimize the formation of MnS compounds [19,20]. In this case, a manganese to sulfur ratio (Mn/S) of 1.7 was assumed to be “balanced” [20]. The aim of using these Mn/S ratios was to investigate the effect of the excess Mn and S on the CT and chill.



After melting and overheating to 1653 K (1380 °C), molten iron was cast into resin-bound foundry molds with bar shapes of 18 and 28 mm in diameter, as well as wedges (see Figure 6). From each melt, a sample was taken for chemical analysis.



**Figure 6.** Mold for casting bars and wedges.

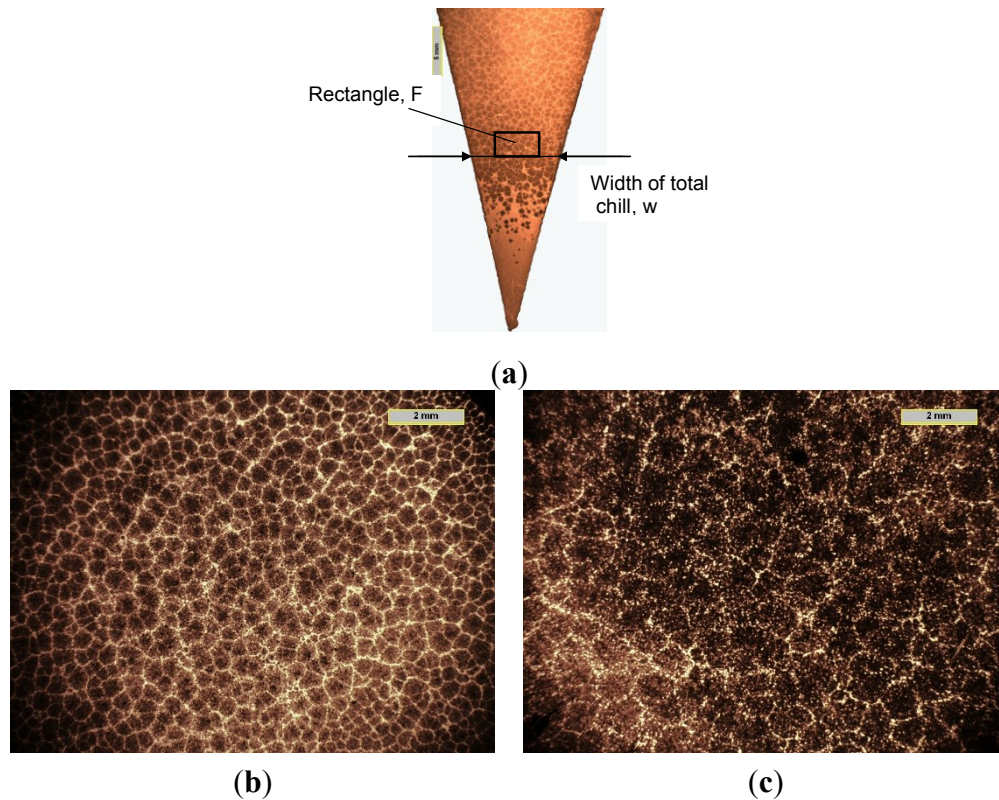
Pt-PtRh10 thermocouples were inserted in the center of the wedge and bar cavities of the sand molds. An Agilent 34970A electronic module was employed for numerical temperature recording. Figure 1 shows a typical cooling curve. The cooling curves were used for determinations of  $T_i$  just after filling the molds,  $T_m$ ,  $\Delta T_m = T_s - T_m$  and  $Q$  at the graphite eutectic equilibrium temperature,  $T_s$ . After solidification, specimens for metallographic examination were taken from the wedges and bars. Metallographic examinations were made on specimen cross-sections. After etching the wedges (using nital), the width ( $w$ ) of the total chill was measured at the gray cast iron-chilled iron intersections (see Figure 7). The wedges and bars were polished and etched using Stead reagent to reveal the graphite eutectic cell boundaries (Figure 7). Determinations of planar cell counts,  $N_F$ , were made according to the so-called Variant II of the Jeffries method and by applying the Saltykov formula as an unbiased estimator for the rectangle of observation [21]:

$$N_F = \frac{N_i + 0.5 N_r + 1}{F} \quad (20)$$

where  $N_i$  is the number of eutectic cells inside the measuring rectangle,  $N_r$  is the eutectic cell count that intersects the sides of the rectangle, but not its corners, and  $F$  is the surface area of the rectangle. The graphite eutectic cells have a granular morphology; hence, it can be assumed that the spatial grain configurations follow the so-called Poisson–Voronoi model [22]. A stereological formula can then be used for determinations of the volumetric cell count  $N$ , which yields the average number of eutectic cells per volume [22].

$$N = 0.568 (N_F)^{3/2} \quad (21)$$

The cell count,  $N_{F,cr}$ , measured in wedges from the rectangular surface  $F$  (Figure 7a) at the gray cast iron-chilled iron intersections is considered as the critical cell count,  $N_{F,cr}$ , and it is converted into a volumetric cell count,  $N_{cr}$  (Equation (19)). An average of 10 readings was used per data point.



**Figure 7.** (a) Eutectic cells in a wedge-shaped casting; (b) bar  $\phi$  18 mm and (c) bar  $\phi$  28 mm.

The experimental data and using Equation (1) in the logarithmic form yields:

$$\ln N = \ln N_s - b \Delta T_m^{-1} \quad (22)$$

where the experimental points  $(\Delta T_{m,1}, N_1)$ ,  $(\Delta T_{m,2}, N_2)$  and  $(\Delta T_{sc}, N_{cr})$  were determined for bars of 30 and 18 mm in diameter and wedges, respectively. The data are graphically plotted as a  $\ln N - \Delta T_i^{-1}$  plot (where  $i = 1, 2, cr$ ), and the data points fit a straight line (Figure 8). From this plot, experimental values of  $N_s = e^{11.9} = 147,267 \text{ cm}^{-3}$  and  $b = 97.1 \text{ }^\circ\text{C}$  were obtained.

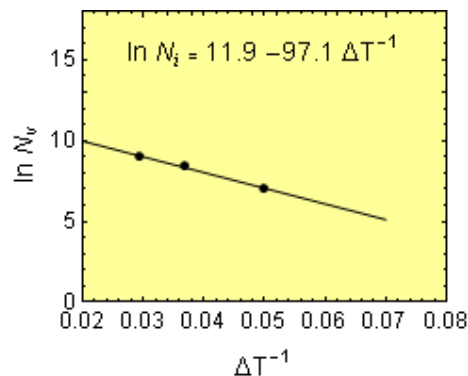
The growth coefficient for the experimental melts can be estimated as follows: for bars of diameter “d” and the length “h”, the casting modulus can be given as:

$$M = \frac{d h}{2(d + 2 h)} \quad (23)$$

Thus, for 1.8- and 3.0-cm diameter bars, the casting modulus is 0.42 and 0.65 cm, respectively. In addition, considering Equations (3)–(5),  $\mu$  can be determined from:

$$\mu = \frac{2 T_s a^2}{\pi^2 \phi M^2} \left[ \frac{4}{f_1 Le N c_{ef}^2 \Delta T_m^8} \right]^{1/3} \quad (24)$$

Furthermore, knowing  $M$ ,  $T_i$ ,  $N$  and  $\Delta T_m$ , as well as the melt thermophysical data, the effect of C, Mn, P and S on  $\mu$  can be estimated from Equation (24). In this work, the initial calculations were made separately for 18 and 30 mm bars and then averaged out.



**Figure 8.** Relationship between undercooling,  $\Delta T_i^{-1}$ , and volumetric cell count,  $\ln N_i$ , for melt P1.

The temperature range,  $\Delta T_{sc}$ , was estimated as follows: There have been numerous experimental investigations that have involved the determinations of  $T_c$  temperatures with variable results [23,24], depending on the chemistry of the experimental iron melts. In fact,  $T_s$  is the equilibrium temperature when  $Q \rightarrow 0$ , making it difficult to determine. In this work, determination of the  $T_s$  temperature was possible by using selected thermodynamic data [25], as well as the work by Heine [26].

$$T_s = 1154 + 5.25 \text{ Si} - 2.36 \text{ Mn} - 14.88 \text{ P} + 35.0 \text{ S} \quad (25)$$

Furthermore, the  $T_c$  temperature can be described by [27]:

$$T_c = 1130.0 + 4.0 \text{ C} - 13.3 \text{ Si} - 50.3 \text{ P} \quad (26)$$

Hence:

$$\Delta T_{sc} = 24 - 4.0 \text{ C} + 18.6 \text{ Si} - 2.3 \text{ Mn} + 35.5 \text{ P} + 35.0 \text{ S} \quad (27)$$

It is worth mentioning that in a binary eutectic Fe-C alloy containing 4.26% C, Equations (25) and (16) yield  $T_s$ ,  $T_c$  and  $\Delta T_{sc}$  values of 1,427 K (1,154 °C), 1,420 K (1,147 °C) and 7 K (7 °C), respectively, in agreement with the Fe-C diagram.

### 3. Results and Discussion

#### 3.1. Nucleation Coefficients $N_s$ , $b$

Different particles with various sizes that act as graphite nucleation substrates are found to be present in cast iron melts. The type, number and size of these substrates are in general influenced by numerous factors, including the chemistry of the cast iron, inoculation, charge materials, melt precondition [28], temperature, bath holding times, furnace atmosphere and the type of slags. In general, the number of nucleation sites from these substrates can be described by the following size distribution function [29]:

$$n(l) = \frac{N_s}{l_a} \exp\left(-\frac{l}{l_a}\right) \quad (28)$$

where  $N_s$  and  $l_a$  are the number of all of the potential graphite nucleation sites in the melt and the average site size, respectively;  $l$  is the size of the nucleation sites (Figure 9a). This function is schematically shown for two melt states  $n(l)$  and  $n_i(l)$  (see Figure 9b). The number of all of the graphite nucleation sites in the melt is determined by the nucleation coefficient,  $N_s$  (*i.e.*, the area under the  $n(l)$  and  $n_i(l)$  curves for the  $l_{\min}$ ,  $l_{\min,1}$  to  $\infty$  range (Figure 9b)). However, not all of the substrate sites take an active role in the nucleation process. The substrate site size,  $l$ , from which a nucleus of graphite can grow is determined by (see Figure 9a):

$$AB = l = 2r^* \sin \theta \quad (29)$$

where  $\theta$  is the wetting angle and  $r^*$  is the critical size of a graphite nucleus:

$$r^* \geq \frac{2 \sigma T_s}{\Delta H \Delta T} \quad (30)$$

and  $\sigma$  is the nucleus-melt interfacial energy,  $\Delta H$  is the latent heat of solidification and  $\Delta T$  is the degree of undercooling.

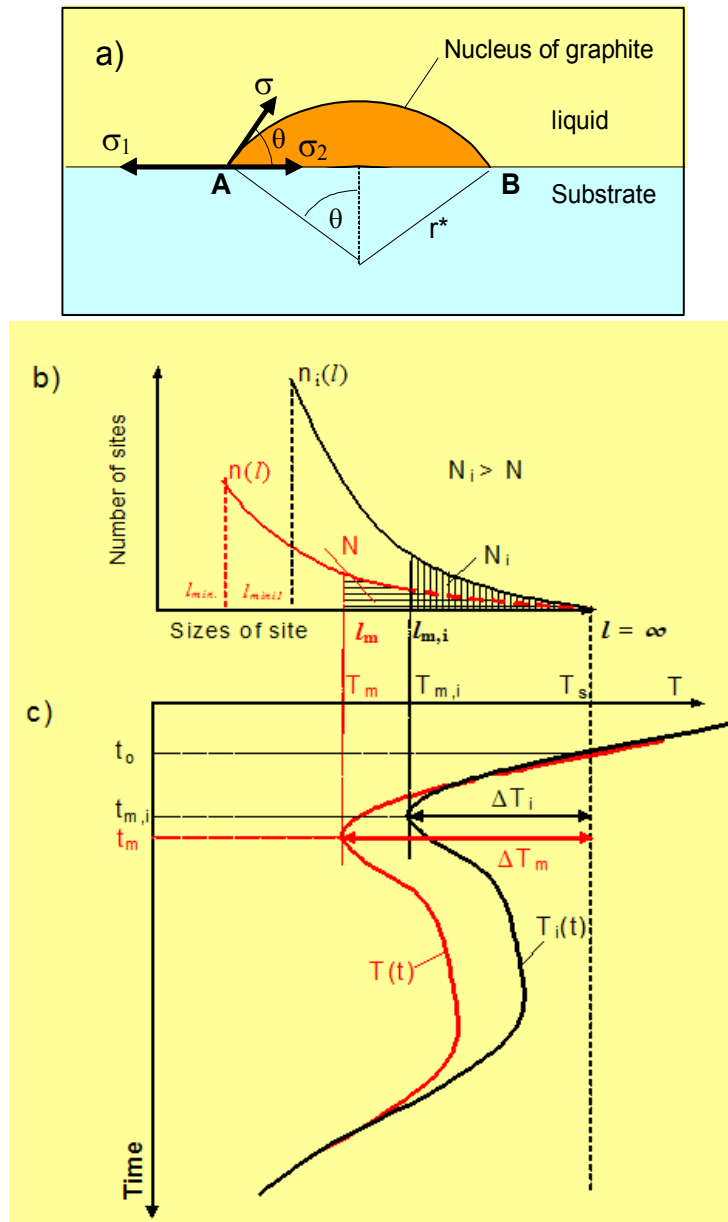
Taking into account Equations (29) and (30), the graphite substrate site size at a given undercooling,  $\Delta T$ , can be calculated from:

$$l \geq \frac{4 \sigma T_s \sin \theta}{\Delta H \Delta T} \quad (31)$$

From Equation (31), it is found that as the undercooling,  $\Delta T$ , increases, the nucleation size,  $l$ , decreases. Figure 9c shows a cooling curve,  $T(t)$ , where the arrow indicates the maximum undercooling,  $\Delta T_m$ . Notice that within the time interval  $t_b$  to  $t_m$  or  $\Delta T = 0$  to  $\Delta T = \Delta T_m$ , all of the sites with  $l_m \leq l \leq \infty$  sizes become active for nucleation purposes. At the time  $t_m$ , the degree of undercooling drops, enabling graphite particles and substrates with  $l \geq l_m$  sizes to be active. In turn, this indicates that for  $t > t_m$ , no active sites are present in the set of graphite substrates. As a result, any nucleation events stop at  $t_m$ . Thus, the number of active nucleation sites,  $N$ , is given by the area below the  $n(l)$  curve in  $l_m \leq l \leq \infty$  ( $0 \leq \Delta T \leq \Delta T_m$ ). Plugging Equation (31) into Equation (28) and integrating from  $l_m$  to  $\infty$ , the second nucleation coefficient  $b$  is obtained:

$$b = \frac{4 T_s \sigma \sin \theta}{l_a \Delta H} \quad (32)$$

The effect of the cast iron chemistry on  $N_s$ ,  $\sigma$ ,  $\theta$  and  $l_a$  is not known, and therefore, it is not possible to determine its effect on the nucleation coefficients  $N_s$  and  $b$ .



**Figure 9.** (a) Graphite nucleus on a substrate (b), size distribution for graphite nucleation sites and (c) cooling curves  $T(l)$  and  $T_i(l)$  for two physical-chemical states of cast iron melts.

### 3.2. The Growth Coefficient, $\mu$

Little is known about  $\mu$ , and reported values for  $\mu$  in pure Fe-C and Fe-C-X alloys are given in Table 3 [30]. From this table, it is apparent that  $\mu$  is highly sensitive to very small changes in the cast iron chemistry.

In Fe-C alloys,  $\mu$  can be analytically derived [15], but the reported expressions are not simple. In general,  $\mu$  is a constant depending on the composition of the cast iron [11] through: (1) the slope of the liquidus lines; (2) the volume fraction; (3) the Gibbs–Thomson coefficients for eutectic phases; (4) the contact angles between eutectic solid phases and the liquid; and (5) the liquid diffusivity. Unfortunately, there is a lack of information in the published literature on this subject. Thus, it is not possible to know the effect of the chemical composition on  $\mu$ .

**Table 3.** Reported values for the eutectic growth coefficient,  $\mu$  [30].

Alloy	Eutectic Growth Coefficient, $\text{cm}/(^{\circ}\text{C}^2\cdot\text{s})$
Fe-C	$9.18 \times 10^{-6}$
Fe-C-0.1% Si	$7.30 \times 10^{-6}$
Fe-C-0.5% Si	$5.94 \times 10^{-6}$
Fe-C-0.01% P	$8.16 \times 10^{-6}$
Fe-C-0.1% P	$6.94 \times 10^{-5}$
Fe-C-0.1% Mn	$9.76 \times 10^{-6}$
Fe-C-0.6% Mn	$1.47 \times 10^{-5}$
Fe-C-0.005% S	$7.30 \times 10^{-6}$
Fe-C-0.01% S	$5.94 \times 10^{-6}$

### 3.3. Effect of Carbon

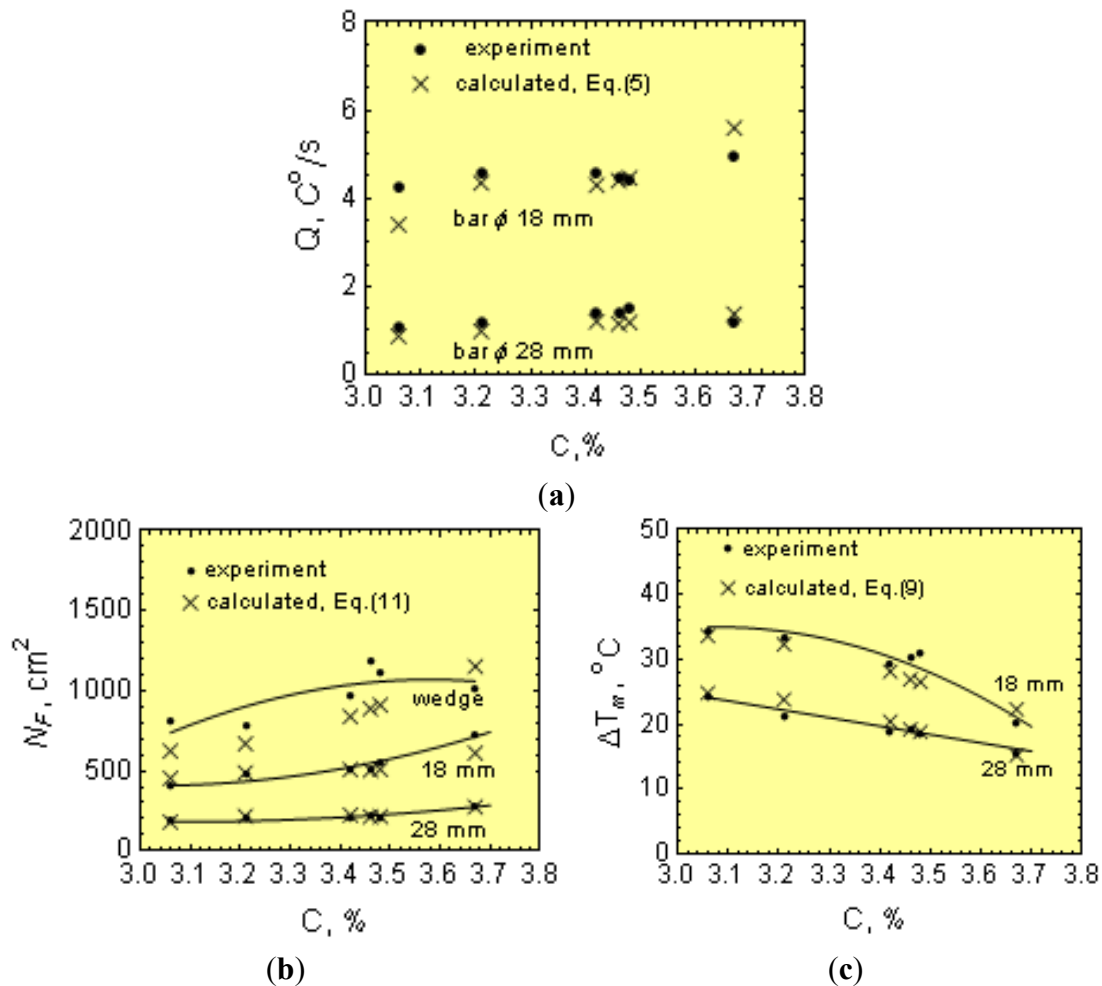
Table 4 shows the results on the exhibited  $T_i$  for various carbon contents in the experimental melts [11]. In addition, Figure 10 shows the exhibited  $Q$ ,  $\Delta T_m$  and  $N_F$  values.

Figure 10a shows the relationship between  $Q$  and the  $T_s$  estimated from Equation (6), as well as the experimental outcome from the cooling curves. From this figure, it is apparent that carbon increases the cooling rate. This is a result of the carbon effect on lowering the liquidus temperature,  $T_l$ , which, in turn, reduces the  $B_l$  (Equation (8)) and  $\phi$  (Equation (9)) values. Moreover, carbon increases the number of eutectic cells in the cast bars and wedges (Figure 10b) and reduces the undercooling (Figure 10c). Alternatively, the increase in cooling rates can be attributed to the increase in thermal conductivity associated with the increasing graphite content. These experimental correlations are described by regression equations for  $N_F$ ,  $N_{F,1}$ ,  $N_{cr}$ ,  $\Delta T_m$  and  $\Delta T_{m,1}$ , and they are given in Table 4. Notice in particular that the experimental outcome is in good agreement with the predictions of Equations (5), (9) and (11).

**Table 4.** Carbon content in cast iron and experimental  $T_i$  values.

No. Melt *	C, %	Initial Liquid Metal Temperature just after Pouring into the Mold, $T_i$ , $^{\circ}\text{C}$		
		Bar 2.8 cm	Bar 1.8 cm	Wedge
C1	3.06	1,375	1,251	1,250
C2	3.21	1,357	1,274	1,272
C3	3.42	1,335	1,260	1,272
C4	3.46	1,355	1,262	1,272
C5	3.48	1,352	1,263	1,276
C6	3.67	1,350	1,260	1,279
Cell count, $\text{cm}^{-2}$	bar 2.8 cm	$N_F = 2,730.6 - 1,653.9 C + 268.3 C^2$		
Cell count, $\text{cm}^{-2}$	bar 1.8 cm	$N_{F,1} = 7,001.2 - 4,368.1 C + 723.4 C^2$		
Critical cell count		$N_{cr} = -14,717 + 8,822.0 C - 1,232.5 C^2$		
Undercooling, $^{\circ}\text{C}$	bar 2.8 cm	$\Delta T_m = 63.9 - 13.0 C$		
Undercooling, $^{\circ}\text{C}$	bar 1.8 m	$\Delta T_{m,1} = -347.8 + 248.4 C - 40.3 C^2$		
Wedge width of the chill, cm		$w = 3.05 - 0.66 C$		
Graphite eutectic growth coefficient, $\mu$ , $\text{cm}/(\text{s}\cdot^{\circ}\text{C}^2)$		$\mu = 10^{-5} \cdot (16.152 - 10.25 C + 1.644 C^2)$		
Nucleation coefficients				
$^{\circ}\text{C}$		$b = -574.73 + 501.33 Cz - 89.60 C^2$		
$\text{cm}^{-3}$		$N_s = 10^6 \cdot (-2.11 + 1.49 Cz - 0.241 C^2)$		

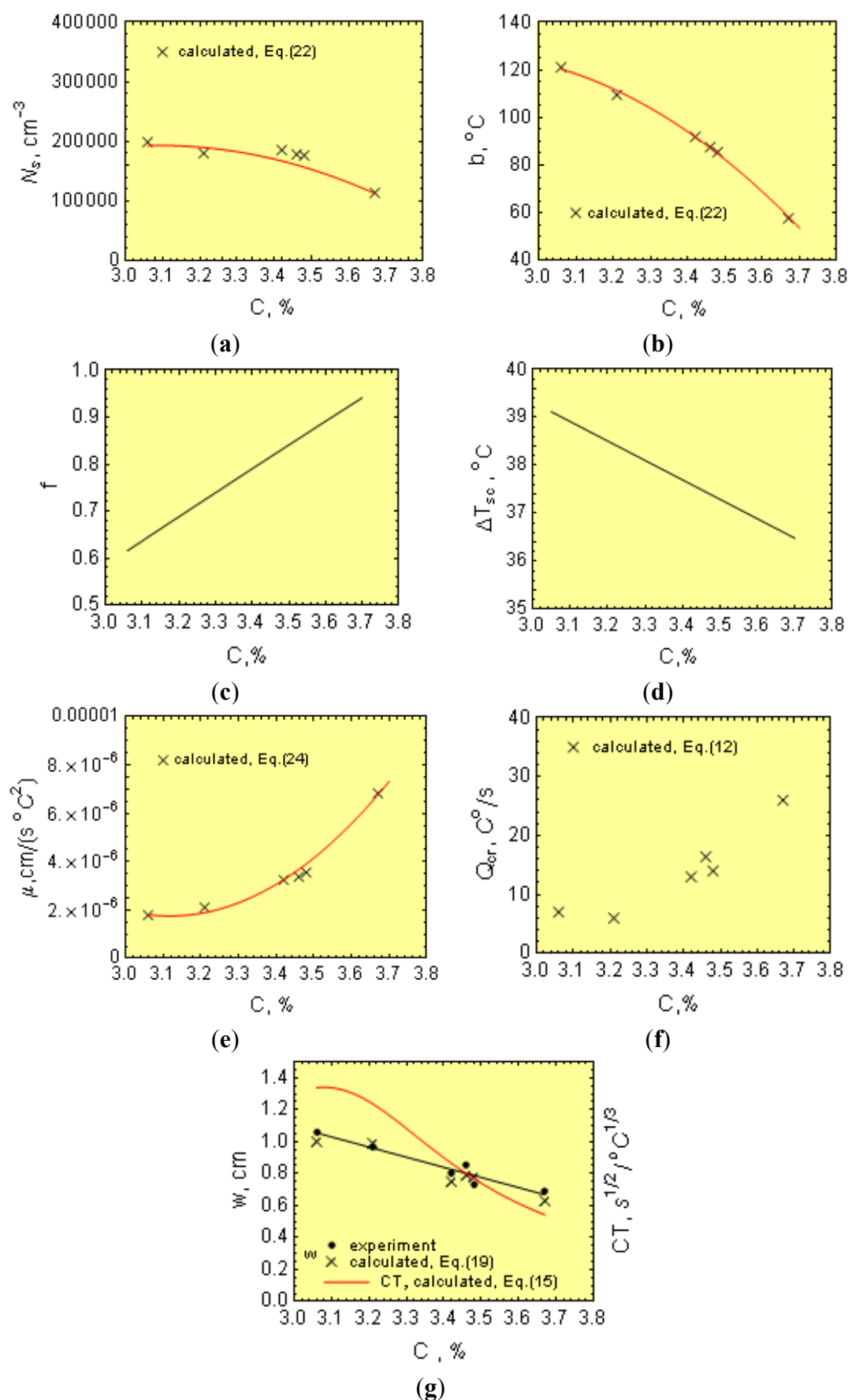
\* Si = 1.40%–1.46%; Mn = 0.070%–0.076%; P = 0.025%–0.034%; S = 0.016%–0.019%.



**Figure 10.** Effect of carbon on (a) the cooling rate,  $Q$ , of cast iron at the onset of solidification, (b) cell count and (c) undercooling.

Figure 11 shows the effect of carbon on the nucleation coefficients (Figure 11a,b) and on the graphite eutectic growth coefficient (Figure 11e). These correlations are described by regression expressions for  $\mu$ ,  $N_s$  and  $b$  in Table 4. Figure 11 also shows the volume fraction of liquid in the cast iron at the onset of eutectic solidification (Figure 11c) and the temperature range  $\Delta T_{sc}$  (Figure 11d). Notice from Figure 11 that carbon increases  $N_s$ ,  $\mu$  and  $f$  while lowering  $b$  and  $\Delta T_{sc}$ . As a result, carbon reduces the CT (Figure 11g) and increases  $Q_{cr}$  (Figure 11f), thus reducing the chill (Figure 11g). Notice that the experimentally-determined chill width and the one predicted using Equation (4) are rather similar.

From the theory on eutectic growth [15] for pure Fe-C alloys, it is found that carbon does not have any influence on  $\mu$ . However, Figure 11e shows that carbon increases  $\mu$ . This can be explained by considering the segregation of Si during solidification. Silicon is soluble in austenite, and it progressively decreases in the liquid phase during the solidification of the pro-eutectic austenite. This agrees with reported data (Table 3) for Si, which shows that as the Si content decreases, the graphite eutectic growth coefficient increases in magnitude.

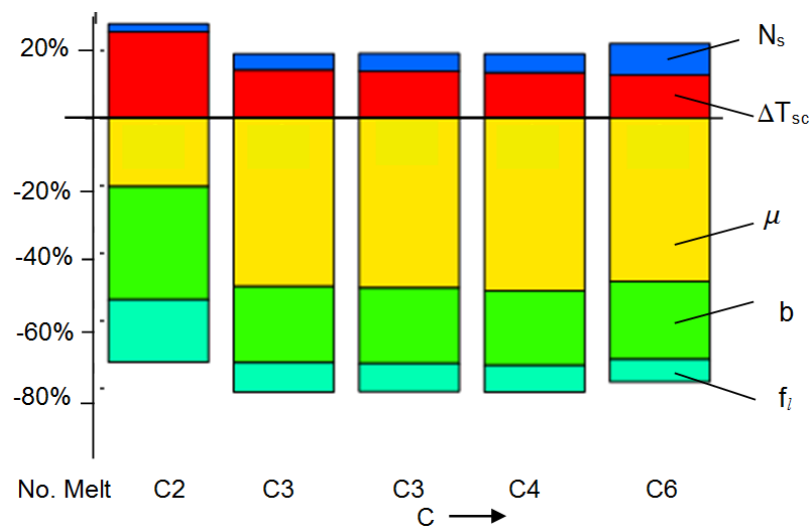


**Figure 11.** Effect of carbon on (a)  $N_s$ , (b)  $b$ , (c)  $f$ , (d)  $\Delta T_{sc}$ , (e)  $\mu$ , (f)  $Q_{cr}$  and (g) the chilling tendency (CT) index and  $w$ .



In order to establish the influence of carbon on the CT, the following calculations were made. The CT was estimated for a carbon content  $C = 3.06\%$  using Equation (15) yielding a  $CT = 1.21 \text{ s}^{1/2}/^\circ\text{C}^{1/3}$ . This value was assumed as a reference value. Next, calculations were made on CT values as they are influenced by carbon through  $f_i$ ,  $\mu$ ,  $\Delta T_{sc}$ ,  $N_s$  and  $b$ . The results of these calculations were normalized to 100%, and they are shown in Figure 12.

Notice from this figure that the intensity effect of carbon through  $\mu$  on CT is very strong. When the carbon content increases, the influence of carbon on the CT through  $\mu$  and  $N_s$  increases, while it decreases through  $b$ . In addition, the effect of carbon on CT through  $f_i$ ,  $\Delta T_{sc}$ , is almost balanced.



**Figure 12.** Carbon intensity effect on the CT index through  $f_i$ ,  $\mu$ ,  $\Delta T_{sc}$ ,  $N_s$  and  $b$  (– CT decreases; + CT increases).

### 3.4. Effect of Silicon

The results of the experimental measurements of Si in the cast iron and the initial liquid metal temperature just after pouring into the mold,  $T_i$ , are given in Table 5, while the cooling rate,  $Q$ , of cast iron at the onset of solidification, the cell count and undercooling are shown in Figure 13. These results are reported in some detail elsewhere [31].

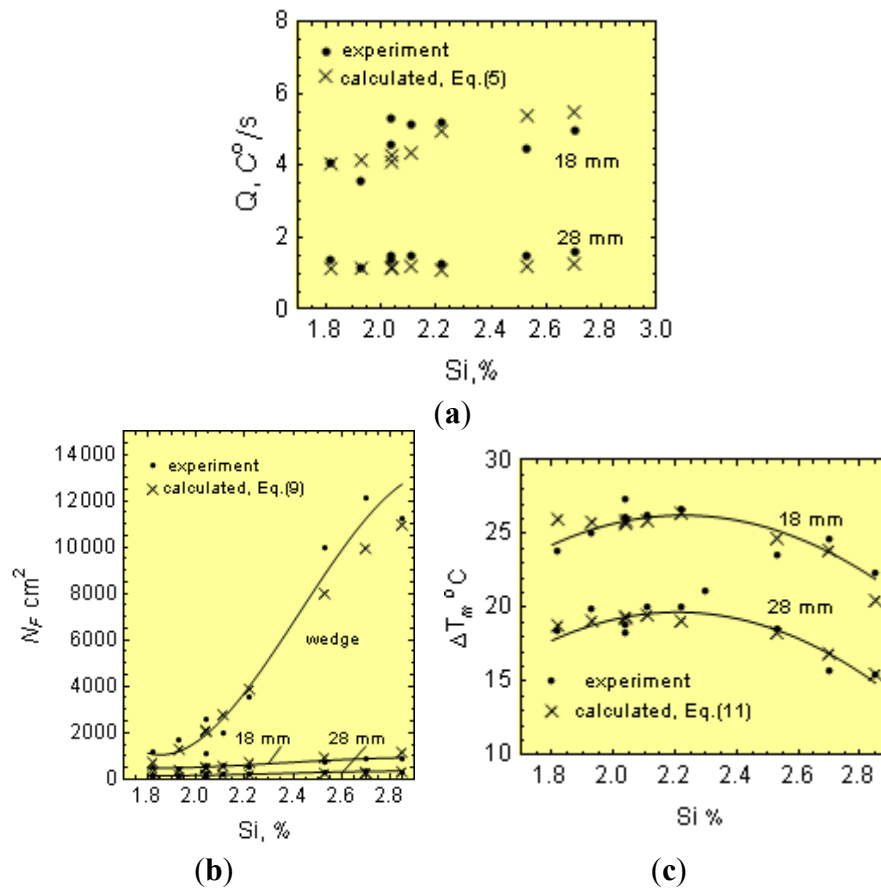
From Figure 13a, it is apparent that Si increases the cooling rate as a result of a reduction in the liquidus temperature,  $T_l$ . (just as with carbon). Si increases the number of eutectic cells in the experimental bars and wedges (Figure 13b), while reducing the undercooling (Figure 13c). The experimental results are described by the regression equations for  $N_F$ ,  $N_{F,1}$ ,  $N_{cr}$ ,  $\Delta T_m$  and  $\Delta T_{m,1}$ , and they are given in Table 5. Notice in particular that the experimental outcome is in good agreement with the predictions of Equations (5), (9) and (11) (see Figure 13).

Figure 14 shows the effect of silicon on the nucleation coefficients (Figure 14a,b),  $f$  (Figure 14c),  $\Delta T_{sc}$  (Figure 14d) and  $\mu$  (Figure 14e). Notice from these figures that in the range of up to 2.4% Si, there is an increase in the magnitude of  $N_s$  and  $b$ . At approximately 2.4% Si,  $N_s$  and  $b$  reach a maximum, and then, they decrease. In contrast,  $\mu$  decreases in the Si range of up to 2.4% Si, reaches a minimum and then increases. These correlations are described by the regression equations for  $\mu$ ,  $N_s$  and  $b$  given in Table 5.

**Table 5.** Effect of silicon on the resultant  $T_i$ .

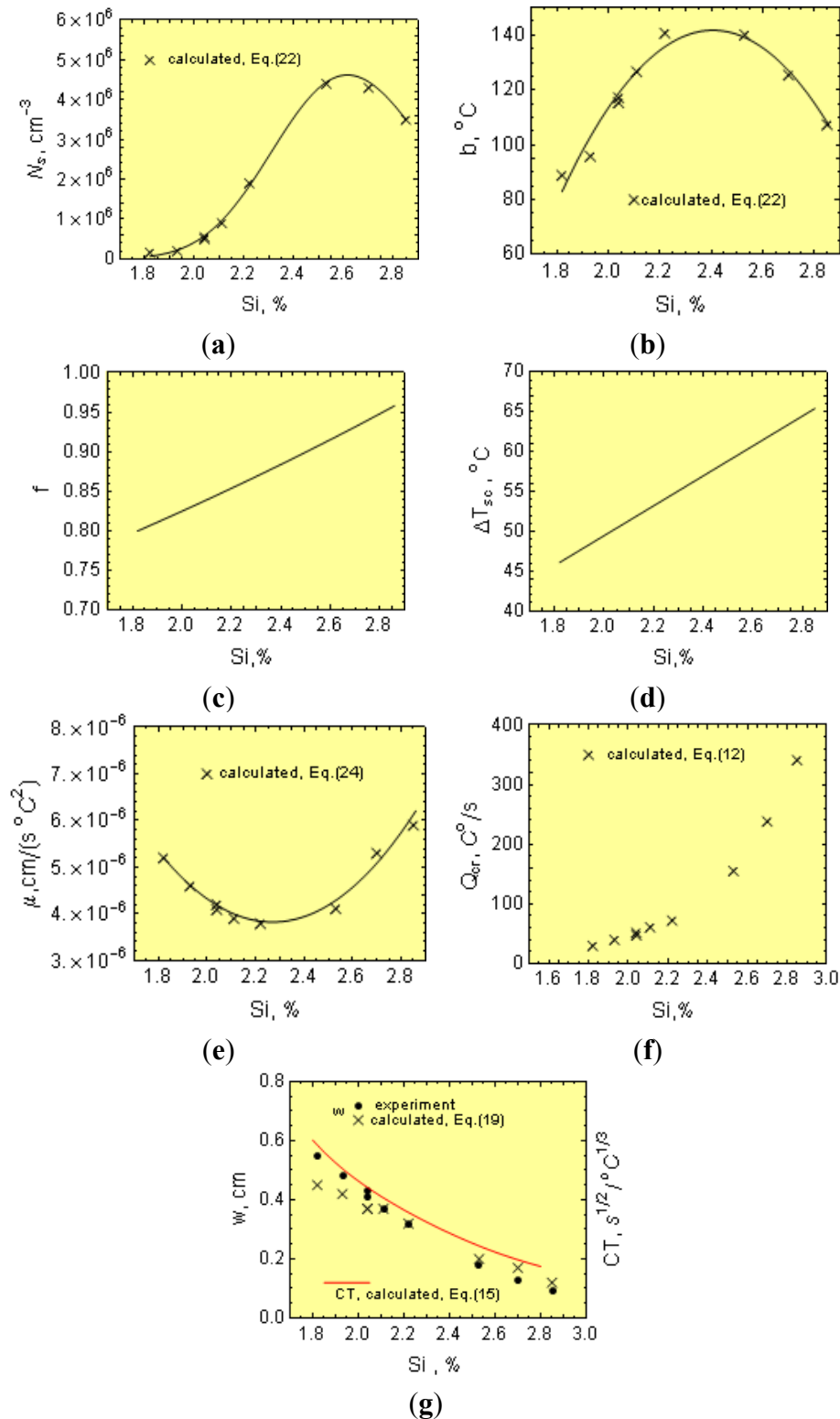
No. Melt *	Si, %	Initial Liquid Metal Temperature just after Pouring into the Mold, $T_i$ , °C		
		Bar 2.8 cm	Bar 1.8 cm	Wedge
Si1	1.82	1,348	1,268	1,249
Si2	1.93	1,348	1,268	1,265
Si3	2.04	1,348	1,268	1,261
Si4	2.11	1,348	1,268	1,241
Si5	2.22	1,374	1,251	1,247
Si6	2.53	1,351	1,253	1,252
Si7	2.70	1,352	1,235	1,239
Si8	2.85	1,354	1,279	1,276
Cell count, $\text{cm}^{-2}$	bar 2.8 cm	$N_F = 4,040.8 - 5,474.5 \text{ Si} + 2,475.5 \text{ Si}^2 - 354.6 \text{ Si}^3$		
Cell count, $\text{cm}^{-2}$	bar 1.8 cm	$N_{F,1} = 8,787.4 - 11,547 \text{ Si} + 5,165.8 \text{ Si}^2 - 731.1 \text{ Si}^3$		
Critical cell count, $\text{cm}^{-2}$		$N_{cr} = 200,666.7 - 271,316.3 \text{ Si} + 118,435.2 \text{ Si}^2 - 16,272.5 \text{ Si}^3$		
Undercooling, °C	bar 2.8 cm	$\Delta T_m = 51.2 \text{ Si} - 11.6 \text{ Si}^2 - 36.8$		
Undercooling, °C	bar 1.8 cm	$\Delta T_{m,1} = 49.4 \text{ Si} - 11.1 \text{ Si}^2 - 28.7$		
wedge width of the chill, cm		$w = 2.5 - 1.45 \text{ Si} + 0.21 \text{ Si}^2$		
Graphite eutectic growth coefficient $\mu$ , $\text{cm}/(\text{s} \cdot ^\circ\text{C}^2)$		$\mu_s = 10^{-5} (3.9 - 3.1 \text{ Si} + 0.683 \text{ Si}^2)$		
Nucleation coefficients °C		$b = 822.5 \text{ Si} - 170.9$		
$\text{cm}^{-3}$		$N_s = 7.7 \times 10^7 \text{ Si}^{72.7} \exp(-27.8 \text{ Si})$		

\* C = 3.20%–3.27%; Mn = 0.072%–0.077%; P = 0.018%–0.024%; S = 0.010%–0.016%.

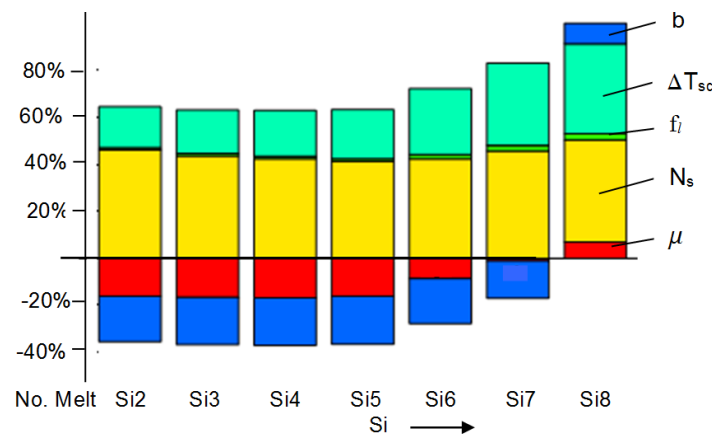


**Figure 13.** Effect of Si on (a)  $Q$ , (b)  $N_F$  and (c)  $\Delta T_m$ .

Notice that the Si tendency to lower the magnitude of  $\mu$  at low Si contents agrees with the data given in Table 3. Moreover, Si increases  $f$  and  $\Delta T_{sc}$ . As a result, Si lowers the CT (Figure 14g) and increases  $Q_{cr}$  (Figure 14f), diminishing the chill (Figure 14g). Notice that the experimental width of the chill and the one predicted using Equation (5) are rather similar (see Figure 14g).

**Figure 14.** Effect of silicon on (a)  $N_s$ , (b)  $b$ , (c)  $f$ , (d)  $\Delta T_{sc}$ , (e)  $\mu$ , (f)  $Q_{cr}$  and (g) the CT index and  $w$ .

Similar to the effect of carbon, the influence of Si on the CT index through  $f$ ,  $\mu$ ,  $\Delta T_{sc}$ ,  $N_s$  and  $b$  was determined (Figure 15).



**Figure 15.** Silicon intensity effect on the CT index, through  $f$ ,  $\mu$ ,  $\Delta T_{sc}$ ,  $N_s$  and  $b$  (– CT decreases; + CT increases).

From Figure 15, it is found that the intensity effect of the Si through  $N_s$  and  $\Delta T_{sc}$  on the CT is strong. When the Si content increases, the intensity of the Si influence on CT through  $\Delta T_{sc}$  also increases, while it diminishes through  $b$  and  $\mu$ . The effect of Si on the CT through  $f$  is rather small.

### 3.5. Effect of Manganese

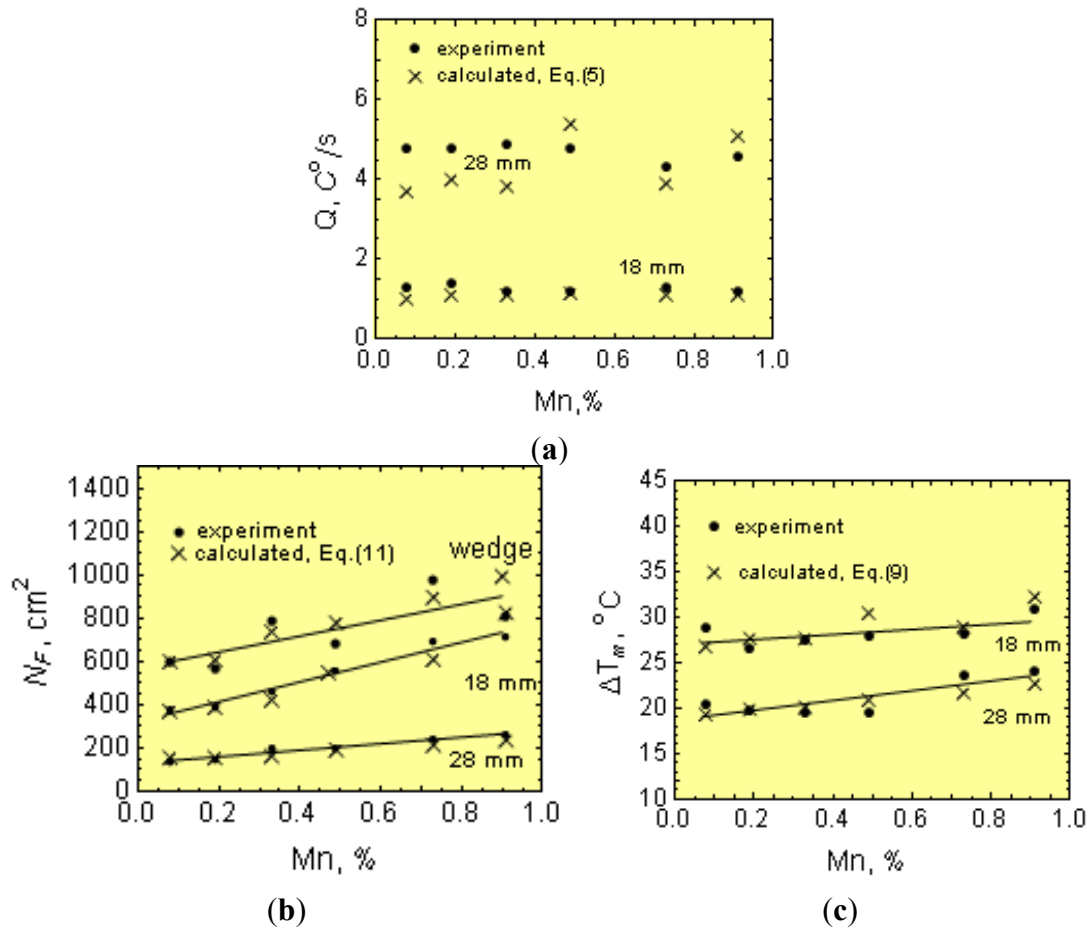
The experimental results on the effect of Mn on the various nucleation and growth factors in the cast iron are given in Table 6. In addition, Figure 16 shows the effect of Mn on  $Q$ ,  $N_F$  and  $\Delta T_m$ .

**Table 6.** Effect of Mn on cast iron  $T_i$ .

No. Melt *	Mn, %	Initial Liquid Metal Temperature just after Pouring into the Mold, $T_i$ , °C		
		Bar 2.8 cm	Bar 1.8 cm	Wedge
Mn1	0.078	1,350	1,248	1,242
Mn2	0.19	1,356	1,254	1,231
Mn3	0.33	1,360	1,260	1,256
Mn4	0.49	1,341	1,210	1,242
Mn5	0.73	1,348	1,264	1,244
Mn6	0.91	1,341	1,210	1,252
Cell count, $\text{cm}^{-2}$		bar 2.8 cm	$N_F = 115,587.8 - 119,801.3 \text{ Mn} + 1.37 \times 10^6 \text{ Mn}^2$	
Cell count, $\text{cm}^{-2}$		bar 1.8 cm	$N_{F,1} = 322.3 + 457.5 \text{ Mn}$	
Critical cell count, $\text{cm}^{-2}$			$N_{cr} = 569.5 + 367.9 \text{ Mn}$	
Undercooling, °C		bar 2.8 cm	$\Delta T_m = 18.7 + 5.4 \text{ Mn}$	
Undercooling, °C		bar 1.8 cm	$\Delta T_{m,1} = 27.02 + 2.8 \text{ Mn}$	
Wedge width of the chill, cm			$w = 0.92 + 0.22 \text{ Mn}$	
Graphite eutectic growth coefficient, $\mu$ $\text{cm}/(\text{s} \cdot ^\circ\text{C}^2)$			$\mu_s = 10^{-6} (4.2 - 2.1 \text{ Mn})$	
Nucleation coefficients °C			$b = 84.6 + 64.6 \text{ Mn}$	
$\text{cm}^{-3}$			$N_s = 115,587.8 - 119,801.3 \text{ Mn} + 1.37 \times 10^6 \text{ Mn}^2$	

\* C = 3.29%–3.37%; Si = 1.21%–1.29%; P = 0.020%–0.03%; S = 0.01%–0.02%.

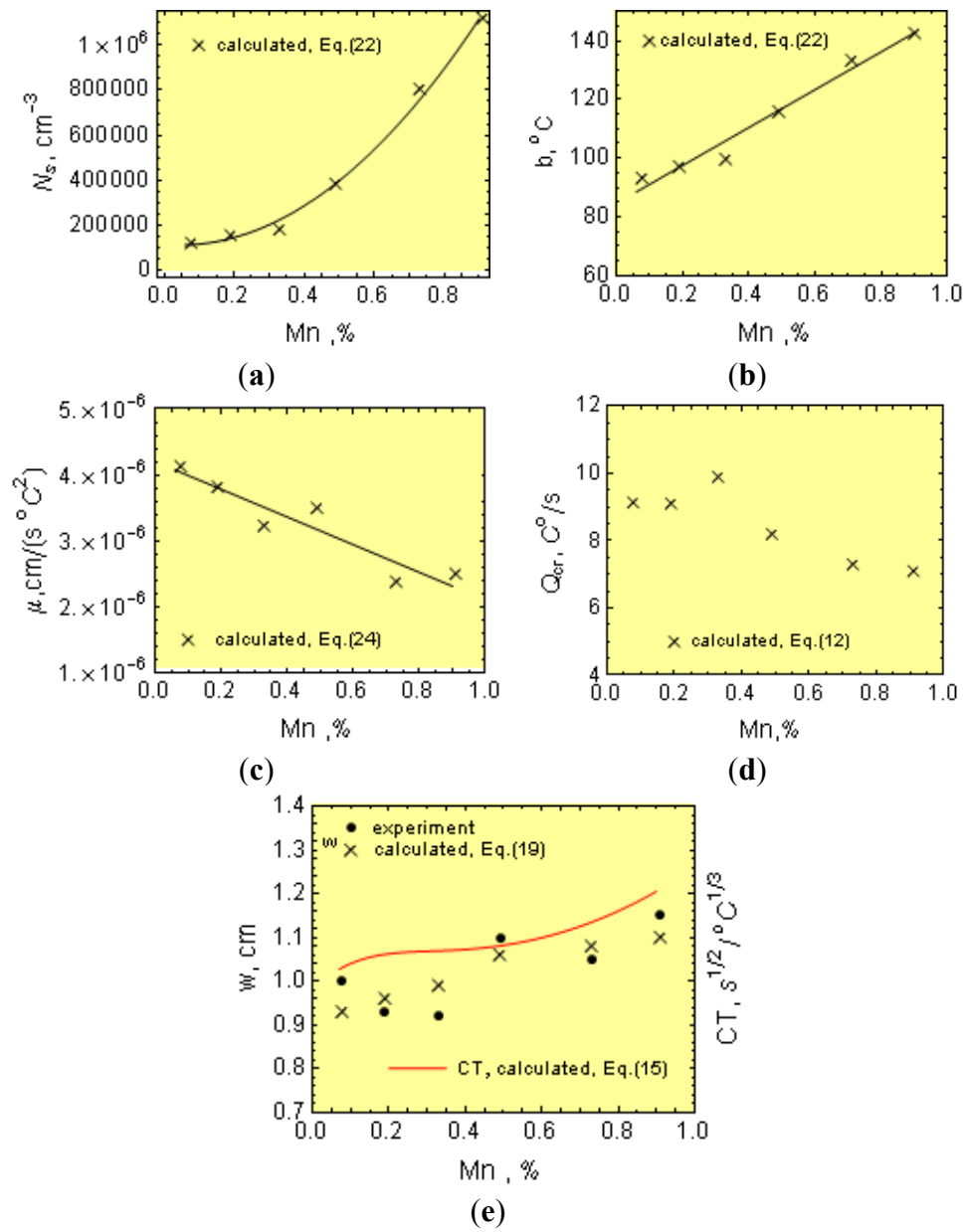
Manganese increases the number of eutectic cells in bars and wedges (Figure 16b) and the undercooling (Figure 16c). These results are described by regression equations for  $N_F$ ,  $N_{F,1}$ ,  $N_{cr}$ ,  $\Delta T_m$  and  $\Delta T_{m,1}$ , and they are given in Table 6. Notice that the experimental and calculated values from Equations (5), (9) and (11) are very close to each other.



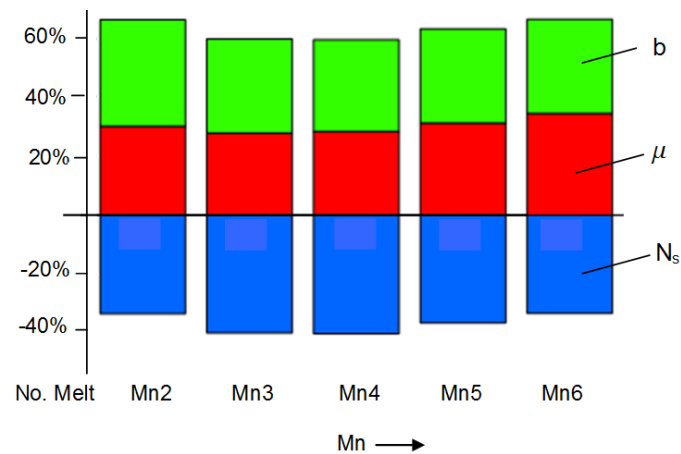
**Figure 16.** Effect of manganese on (a)  $Q$ , (b)  $N_F$  and (c)  $\Delta T_m$ .

Manganese does not have a significant effect on  $f$  and  $\Delta T_{sc}$ . Figure 17 shows the effect of Mn on the nucleation coefficients (Figure 17a,b) and on  $\mu$  (Figure 17c) (see the regression equations for  $\mu$ ,  $N_s$  and  $b$  given in Table 6). Notice from Figure 17 that Mn increases the nucleation coefficients,  $N_s$  and  $b$ , while decreasing  $\mu$ . In particular, the Mn tendency to lower  $\mu$  agrees with the data given in Table 3. Mn increases the CT index (Figure 17e) and  $Q_{cr}$  (Figure 17d), reducing the chill (Figure 17e). Notice that there is good agreement between the experimental outcome and the predictions of Equation (5) for  $w$  (see Figure 17e).

The manganese intensity effect on the CT index through  $N_s$ ,  $b$  and  $\mu$  is shown in Figure 18. Notice that the most significant effect is on  $N_s$ . In addition, the intensity of the influence of Mn on the CT tendency index, through  $\mu$  and  $b$ , is rather similar.



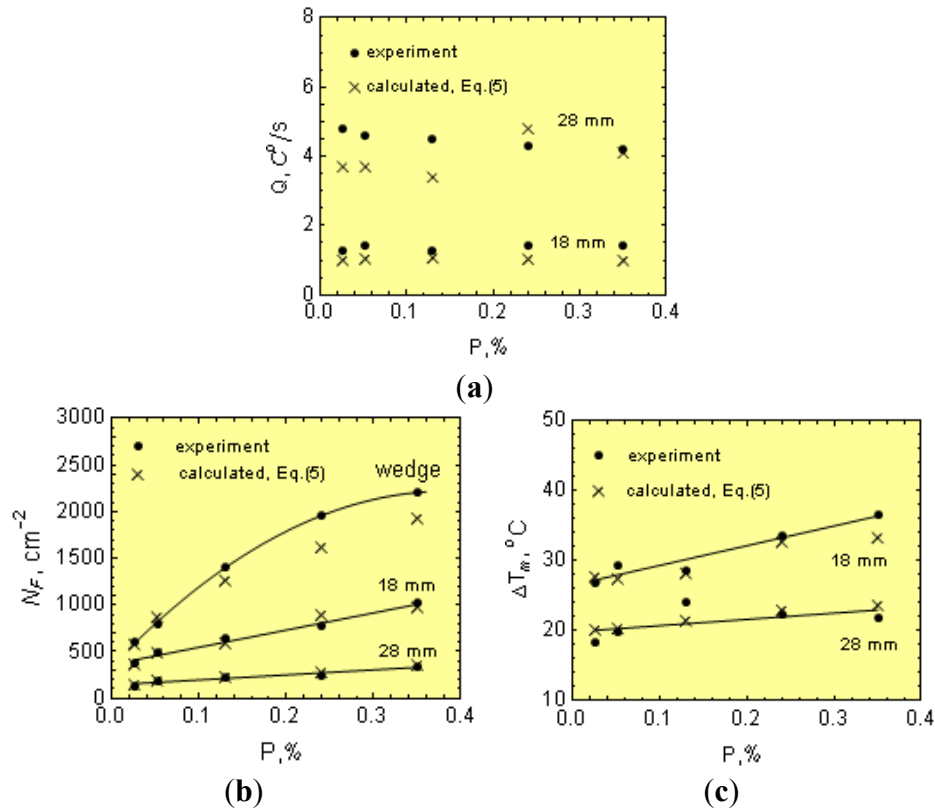
**Figure 17.** Effect of manganese on (a)  $N_s$ , (b)  $b$ , (c)  $\mu$ , (d) CT index and (e)  $w$ .



**Figure 18.** Manganese intensity effect on the CT index through  $\mu$ ,  $N_s$  and  $b$  (– CT decreases; + CT increases).

### 3.6. Effect of Phosphorous

The experimental results on the effect of P on the various nucleation and growth factors in the cast iron are given in Table 7. In addition, Figure 19 shows the effect of P on  $Q$ ,  $N_F$  and  $\Delta T_m$ .



**Figure 19.** Effect of phosphorous on (a)  $Q$ , (b)  $N_F$  and (c)  $\Delta T_m$ .

**Table 7.** Effect of P on cast iron  $T_i$ .

No. Melt *	P, %	Initial Liquid Metal Temperature just after Pouring into the Mold, $T_i$ , $^{\circ}\text{C}$		
		Bar 2.8 cm	Bar 1.8 cm	Wedge
P1	0.026	1,350	1,248	1,247
P2	0.052	1,340	1,251	1,249
P3	0.13	1,351	1,273	1,262
P4	0.24	1,351	1,213	1,244
P5	0.35	1,368	1,243	1,268
Cell count, $\text{cm}^{-2}$	bar 2.8 cm	$N_F = 143.1 + 541.6 P$		
Cell count, $\text{cm}^{-2}$	bar 1.8 cm	$N_{F,1} = 362.8 + 1,844.9 P$		
Critical cell count, $\text{cm}^{-2}$		$N_{cr} = 341.7 + 9,797.2 P - 12,769.5 P^2$		
Undercooling, $^{\circ}\text{C}$	bar 2.8 cm	$\Delta T_m = 19.8 + 8.98 P$		
Undercooling, $^{\circ}\text{C}$	bar 1.8 cm	$\Delta T_{m,1} = 26.5 + 28.3 P$		
wedge width of the chill, cm		$w = 1.02 - 1.66 P + 1.94 P^2$		
Graphite eutectic growth coefficient, $\mu$ $\text{cm}/(\text{s} \cdot ^{\circ}\text{C}^2)$		$\mu_s = 10^{-6} (6.12 - 6.02 P^{0.25})$		
Nucleation coefficients $^{\circ}\text{C}$		$b = 91.8 + 358.3 P - 770.4 P^2$		
$\text{cm}^{-3}$		$N_s = 10^6 (7.7 P - 16 P^2 - 0.04)$		

\* C = 3.29%–3.37%; Si = 1.21%–1.29%; P = 0.020%–0.03%; S = 0.01%–0.02%.

Phosphorous increases the number of eutectic cells in the experimental bars and wedges (Figure 19b) and the undercooling (Figure 19c). These results are described by regression equations for  $N_F$ ,  $N_{F,1}$ ,  $N_{cr}$ ,  $\Delta T_m$  and  $\Delta T_{m,1}$ , and they are given in Table 7. Notice that the experimental and calculated values from Equations (5), (9) and (11) are rather similar.

Figure 20 shows the effect of P on the nucleation coefficients (Figure 20a,b) and on  $\mu$  (Figure 20e) (see the regression equations for  $\mu$ ,  $N_s$  and  $b$  in Table 7). These correlations are described by regression equations for  $\mu$ ,  $N_s$  and  $b$  (Table 7). Notice from Figure 20a,b that for up to 0.25% P, phosphorous increases the nucleation coefficients,  $N_s$  and  $b$ , until they reach a maximum at approximately  $P = 0.25\%$ . Furthermore, from Figure 20e, it is found that P lowers  $\mu$  in agreement with the data given in Table 3. In addition, P increases  $f$  (Figure 20c) and  $\Delta T_{sc}$  (Figure 20d). P reduces the CT index (Figure 20g) and increases  $Q_{cr}$  (Figure 20f), reducing the chill (Figure 20g). Notice that there is good agreement between the experimental outcome and the predictions of Equation (5) for  $w$  (see Figure 20g).

Notice from this figure that increasing the P content in the cast iron, the intensity influence of P on the CT index through  $\mu$  and  $\Delta T_{sc}$  increases, while decreasing through  $N_s$  and  $b$ . The effect of the phosphorus through  $f$  on CT is negligible.

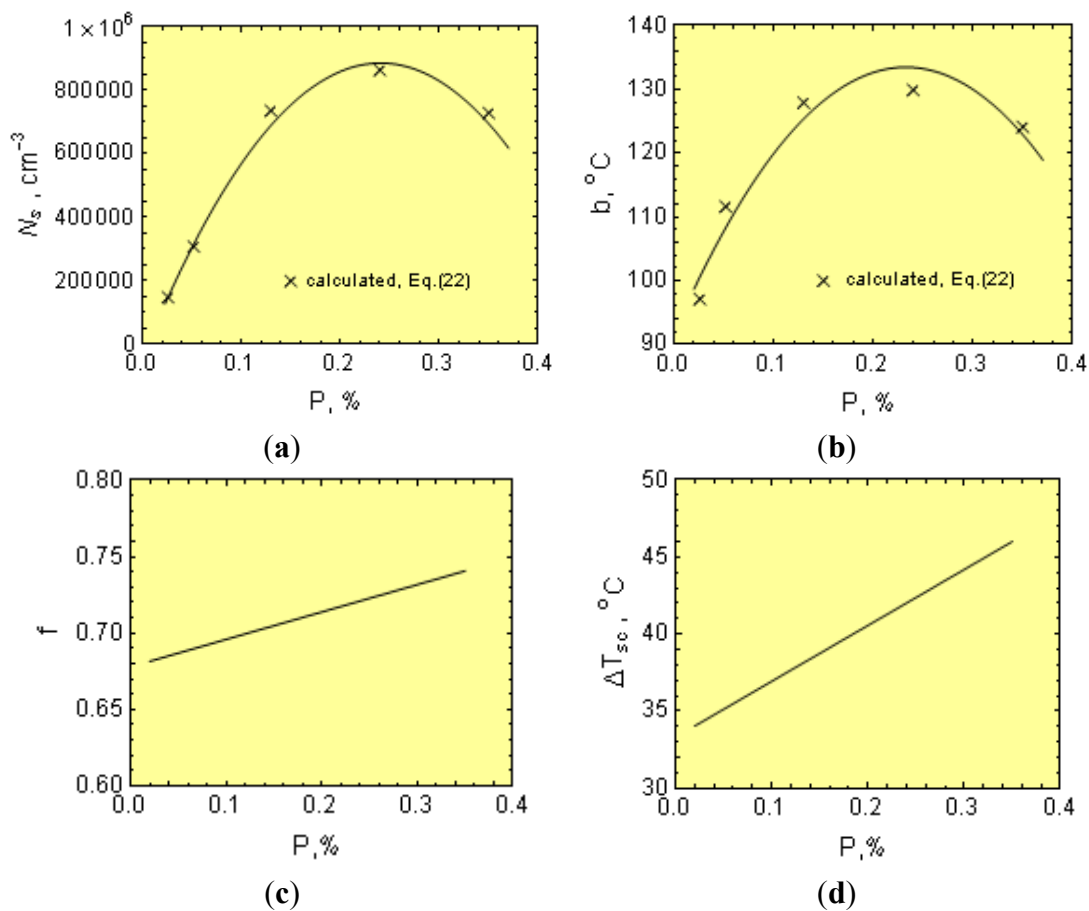
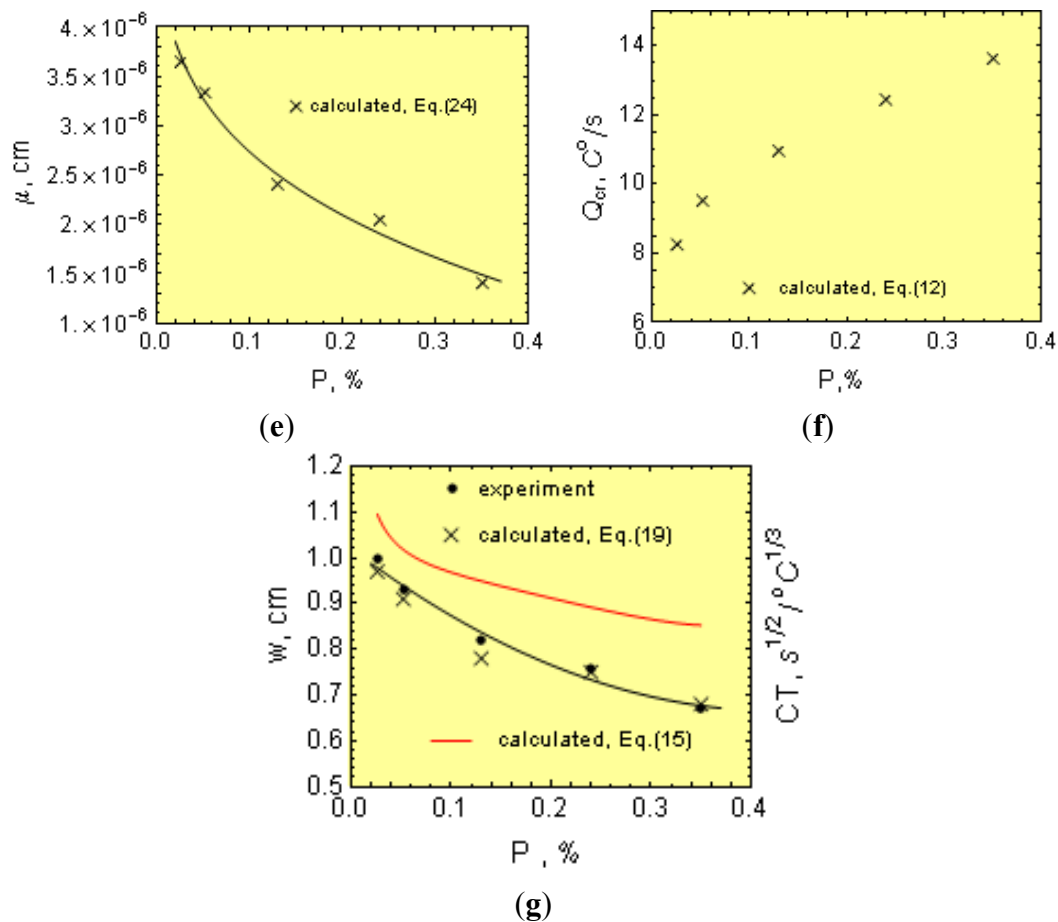


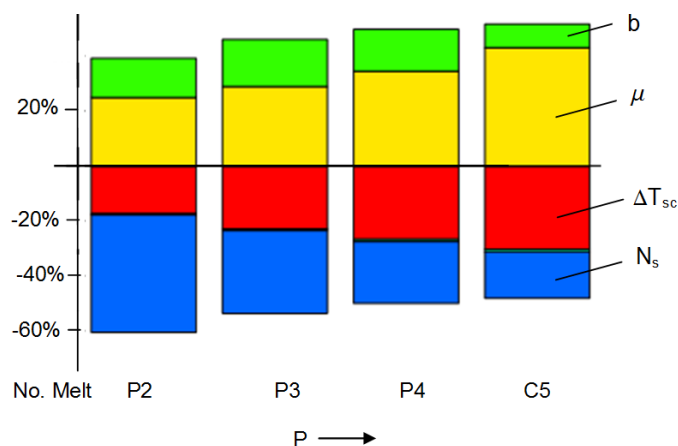
Figure 20. Cont.





**Figure 20.** Effect of phosphorous on (a)  $N_s$ , (b)  $b$ , (c)  $f$ , (d)  $\Delta T_{sc}$ , (e)  $\mu$ , (f)  $Q_{cr}$  and (g)  $w$ .

The phosphorous intensity effect on the CT index through  $f$ ,  $\mu$ ,  $\Delta T_{sc}$ ,  $N_s$  and  $b$  is shown in Figure 21.



**Figure 21.** Phosphorous intensity effect on the CT index, through  $\mu$ ,  $N_s$ ,  $b$ ,  $\Delta T_{sc}$  and  $f$  (– CT decreases; + CT increases).

### 3.7. Effect of Sulfur

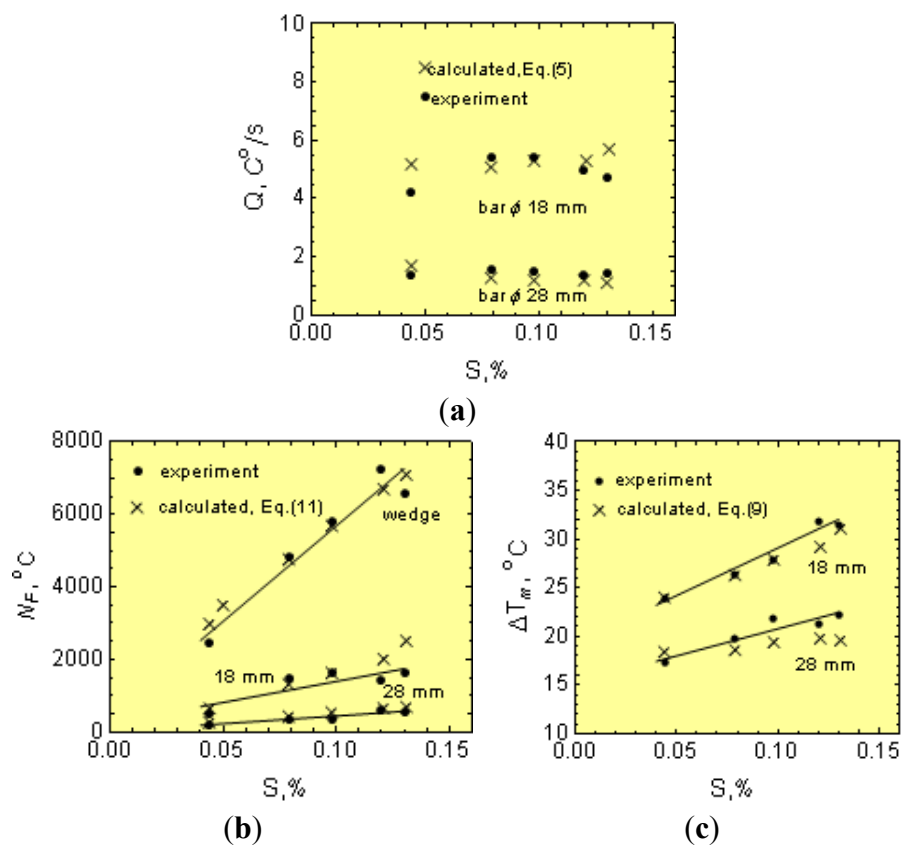
The experimental results on the effect of S on the various nucleation and growth factors in the cast iron are given in Table 8. In addition, Figure 22 shows the effect of S on  $Q$ ,  $N_F$  and  $\Delta T_m$ . These results are reported in some detail elsewhere [32,33].

Sulfur increases the number of eutectic cells in the experimental bars and wedges (Figure 22b) and the undercooling (Figure 22c). These results are described by regression equations for  $N_F$ ,  $N_{F,1}$ ,  $N_{cr}$ ,  $\Delta T_m$  and  $\Delta T_{m,1}$ , and they are given in Table 8. Notice that the experimental outcome and the calculated values from Equations (5), (9) and (11) are rather similar.

**Table 8.** Effect of S on the cast iron  $T_i$ .

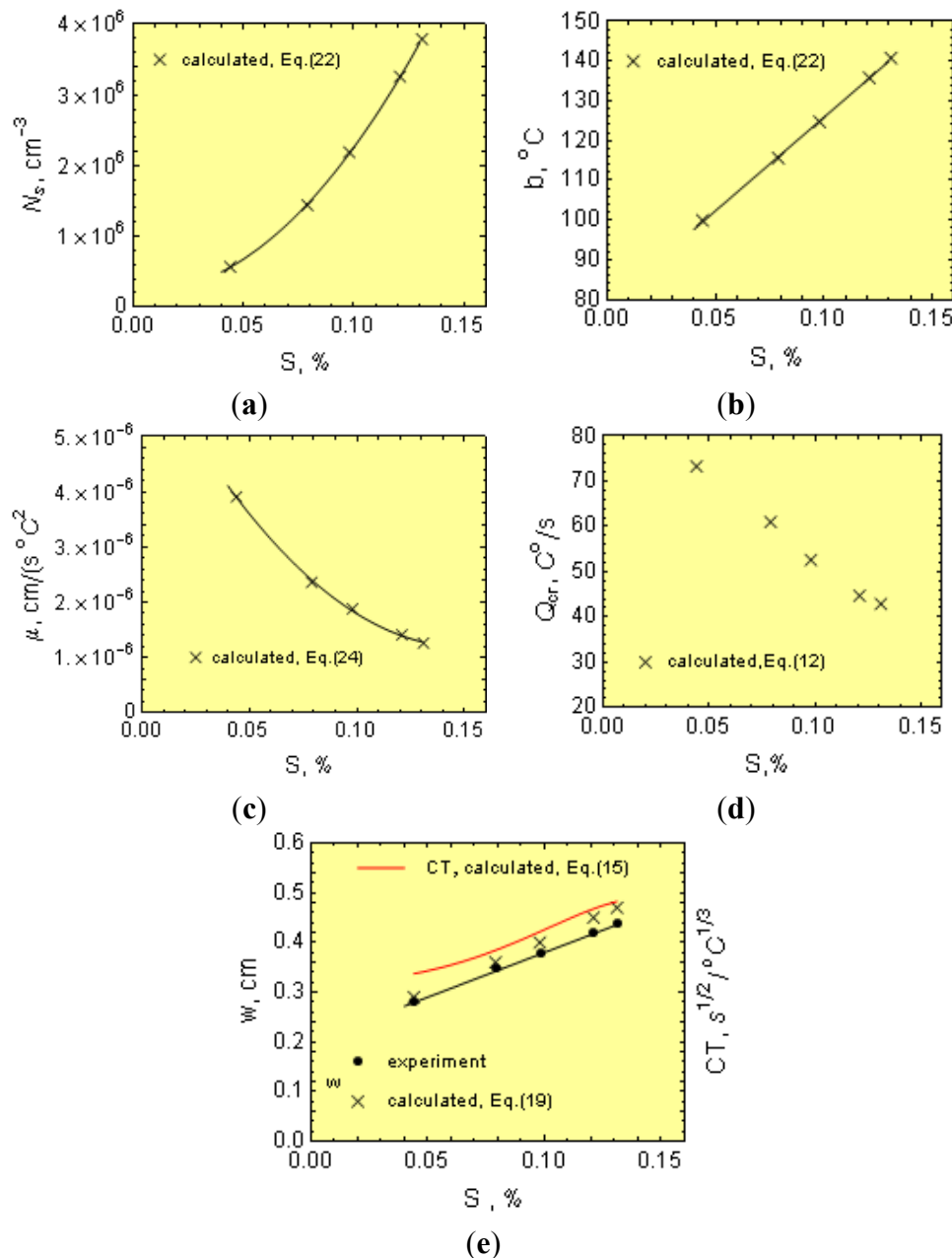
No. Melt *	S, %	Initial Liquid Metal Temperature just after Pouring into the Mold, $T_i$ , °C		
		Bar 2.8 cm	Bar 1.8 cm	Wedge
S1	0.044	1,368	1,300	1,275
S2	0.079	1,354	1,245	1,247
S3	0.098	1,349	1,248	1,245
S4	0.121	1,355	1,251	1,242
S5	0.131	1,357	1,224	1,241
Cell count, $\text{cm}^{-2}$	bar 2.8 cm	$N_F = 9.4 + 4,267.2 S$		
Cell count, $\text{cm}^{-2}$	bar 1.8 cm	$N_{F,1} = 231.1 + 11,595.1 S$		
Critical cell count		$N_{cr} = 417.2 + 52,694.6 S$		
Undercooling, °C	bar 2.8 cm	$\Delta T_m = 15.2 + 55.8 S$		
Undercooling, °C	bar 1.8 cm	$\Delta T_{m,1} = 19.2 + 96.1 S$		
wedge width of the chill, cm		$w = 0.2 + 1.8 S$		
Graphite eutectic growth coefficient, $\mu$ $\text{cm}/(\text{s} \cdot ^\circ\text{C}^2)$		$\mu_s = 10^{-6} (6.62 - 72.5 S + 242.0 S^2)$		
Nucleation coefficients °C		$b = 78.8 + 470.5 S$		
$\text{cm}^{-3}$		$N_s = 10^6 (0.214843 - 1.888 S + 223.206 S^2)$		

\* C = 3.29%–3.35%; Si = 2.21%–2.27%; Mn = 0.077%–0.082%; P = 0.020%–0.03%.



**Figure 22.** Effect of sulfur on (a)  $Q$ , (b)  $N_F$  and (c)  $\Delta T_m$ .

Sulfur does not have a significant effect on  $f$  nor on  $\Delta T_{sc}$ . Figure 23 shows the S effect on  $N_s$ ,  $b$  (Figure 23a,b) and  $\mu$  (Figure 23c). These correlations are described by regression equations for  $\mu$ ,  $N_s$  and  $b$  in Table 8. Notice from Figure 23 that S increases  $N_s$  and  $b$ , while lowering  $\mu$ . The S tendency to lower  $\mu$  is in agreement with the data given in Table 3. Sulfur increases the CT index (Figure 23e) and lowers  $Q_{cr}$  (Figure 23d), thus diminishing the chill (Figure 23e). Notice that there is good agreement between the experimental and the predictions of Equation (5) for  $w$  (see Figure 23e).



**Figure 23.** Effect of sulfur on (a)  $N_s$ , (b)  $b$ , (c)  $\mu$ , (d) CT index and (e)  $w$ .

Figure 24 shows the intensity effect of S on the CT index through  $\mu$ ,  $N_s$  and  $b$ . Notice that the most significant effect of S is through  $\mu$ . When the S content in the cast iron increases, the intensity influence of S on the CT index increases through  $\mu$ , while it is reduced through  $N_s$  and  $b$ .



Table 9. Cont.

No. Melt	1	2	3	4	5	6	7
	Base Iron			Inoculated Iron			
Nucleation coefficients							
base cast iron	$b = 76.8, ^\circ\text{C}$ $N_s = 1.6 \times 10^6, \text{cm}^{-3}$						
inoculated cast iron	$b = 96.9 + 122.6 t_r - 59.2 t_r^2, ^\circ\text{C}$ $N_s = 10^6 \cdot [6.5 - 0.8 t_r - 5.3 t_r^2], \text{cm}^{-3}$						
Chemistry							
base cast iron	C = 3.25%; Si = 1.17%; Mn = 0.13%; P = 0.08%; S = 0.047%						
inoculated cast iron	C = 3.18%; Si = 1.91%; Mn = 0.13%; P = 0.09%; S = 0.064%						
inoculant	Si = 73%–75%; Al = 0.75%; Ca = 0.75%–1.25%; Ba = 0.75%–1.25%						

Figure 25 shows that the cell count (Figure 25a) diminishes while the undercooling increases (Figure 25b) with  $t_r$ . In addition, Figure 25c,d shows the effect of  $t_r$  after inoculation on the nucleation coefficients. Notice from these figures that  $N_s$  increases, while  $b$  decreases with  $t_r$ . As a result, the CT index increases (Figure 25f), and the  $Q_{cr}$  is reduced (Figure 25e), lowering the chill (Figure 25f).

Figure 26 shows the  $t_r$  intensity effect on the CT index through  $N_s$  and  $b$ . Notice that  $t_r$  does not have an effect on  $f$ ,  $\Delta T_{sc}$  and  $\mu$  (*i.e.*, the chemistry of the cast iron is a constant). Apparently, the CT index is affected by the nucleation parameters  $N_s$  and  $b$  with  $t_r$  (increasing through  $N_s$  and decreasing through  $b$ ).

From the graphite nucleation potential curves, the following arguments can be made for the nucleation of graphite cells. Consider two melts (base and inoculated irons). For the base melt, the nucleation potential is given by the curve  $n(l)$ , which is relatively small when compared with that for inoculated iron,  $n_i(l)$  (Figure 9b). Although,  $\Delta T_m$  for the base melt is higher than for the inoculated melt,  $\Delta T_{m,i}$ , the area below the curve  $n(l)$  for  $l_m \leq l \leq \infty$  ( $0 \leq \Delta T \leq \Delta T_m$ ) is smaller than the area below the curve  $n_i(l)$  in the range  $l_{m,i} \leq l \leq \infty$  ( $0 \leq \Delta T \leq \Delta T_{m,i}$ ). In turn, this indicates that during solidification, the inoculated melt contains a higher nuclei density than the base melt.

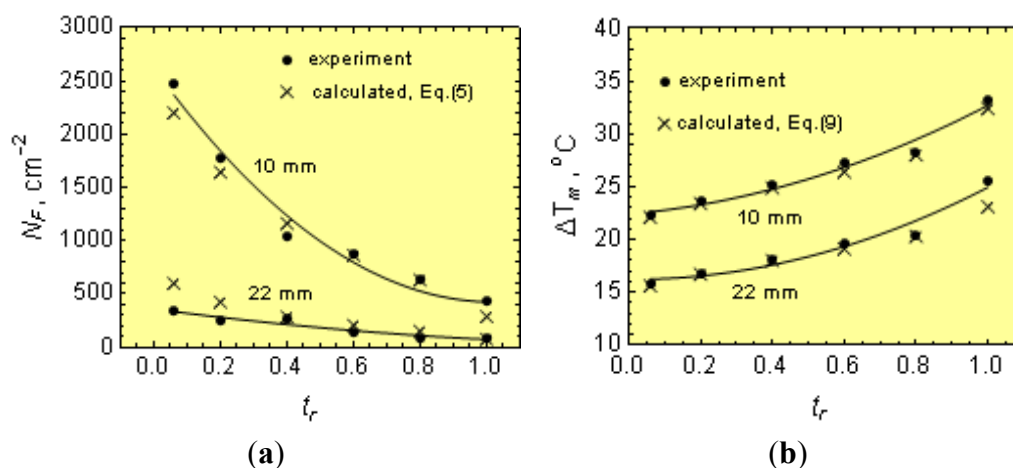
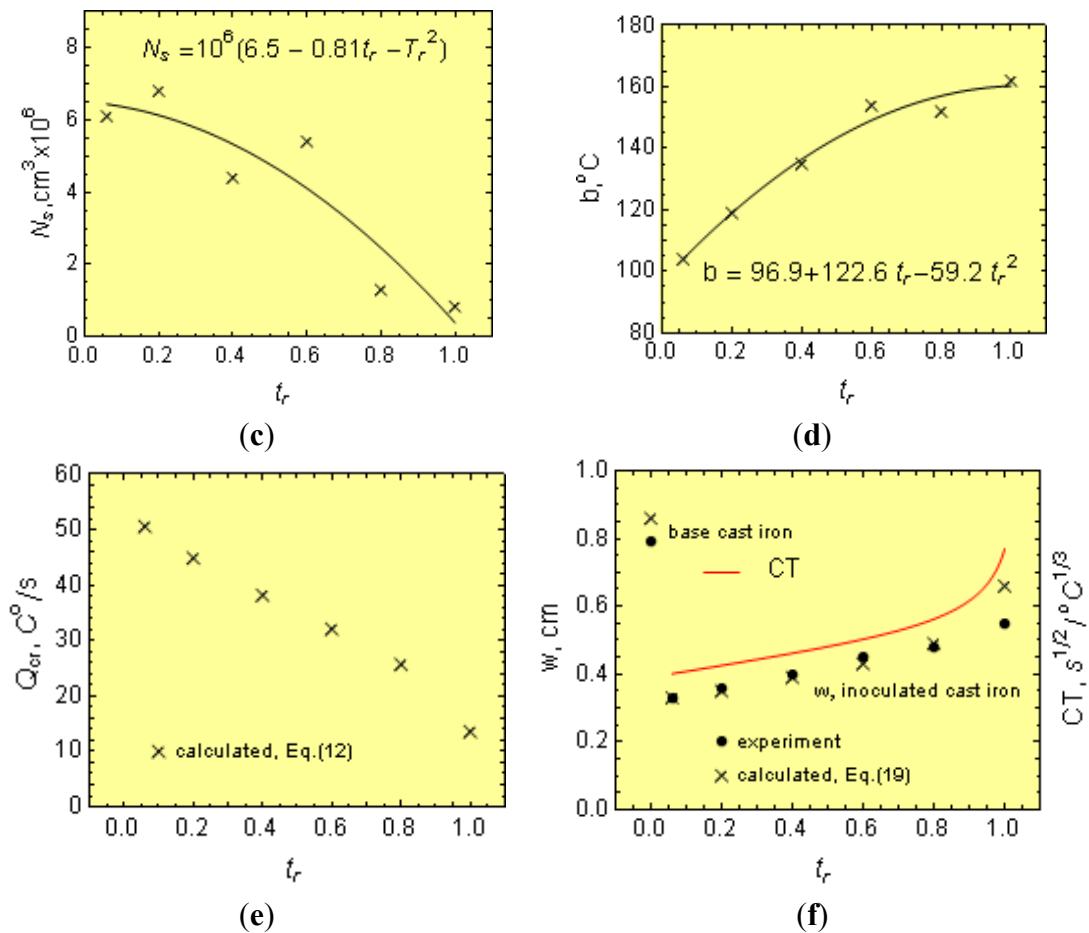
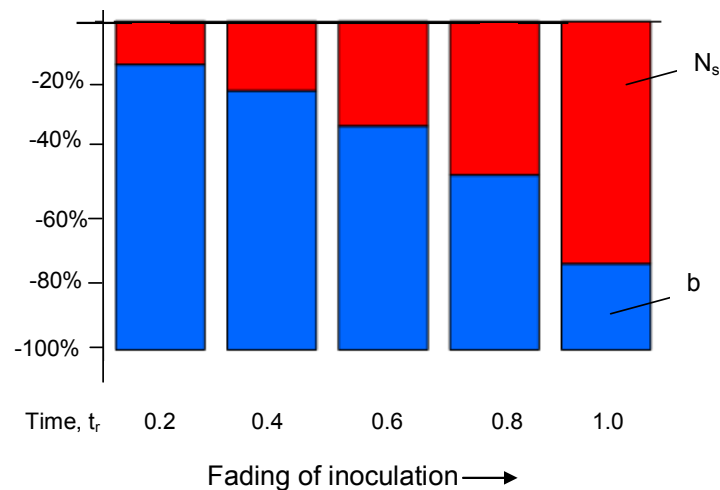


Figure 25. Cont.

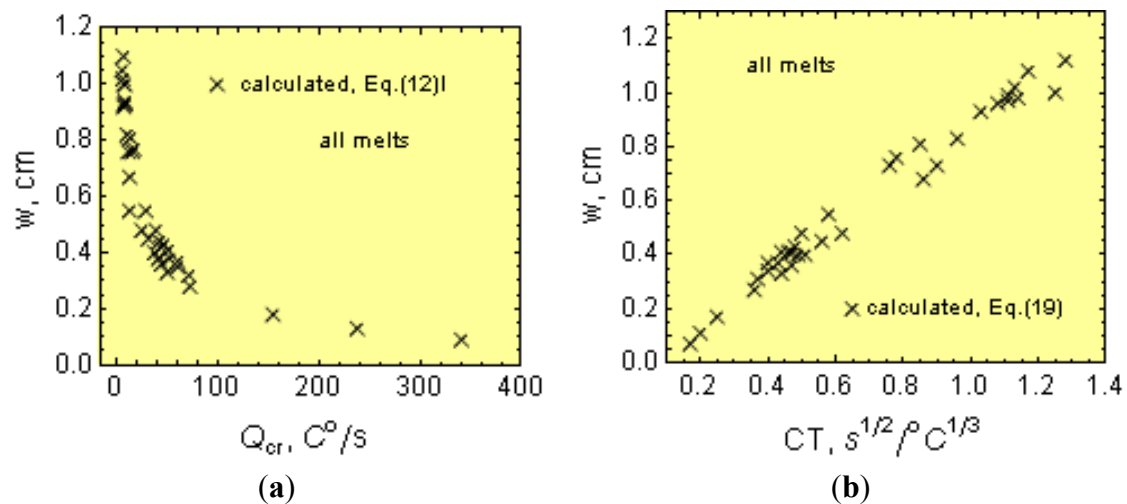


**Figure 25.** Effect of  $t_r$  after inoculation on (a)  $N_F$ , (b)  $\Delta T_m$ , (c)  $N_s$ , (d)  $b$ , (e) the CT index and (f)  $w$ .



**Figure 26.** Intensity effect of  $t_r$  on the CT index through  $N_s$  and  $b$  (– CT decreases).

Finally, in the present work, the role of various technological parameters on the CT index and the chill of cast iron can be explained by means of a common analytical model of general validity. This theory was experimentally verified using C, Si, Mn, P and S, and it includes fading effects as examples. Consequently, general relations can be found between  $Q_{cr}$ , the CT index and the chill (see Figure 27).



**Figure 27.** Relationship between (a)  $Q_{cr}$  and the chill,  $w$ , and (b) the CT and  $w$  for all melts. Experimental and calculated values were found using Equation (19).

#### 4. Conclusions

In this work, experimental and theoretical results on cast iron solidification and chill tendencies, CT, are found to be in good agreement with each other, and they are summarized as follows:

- Carbon decreases  $N_s$ ,  $b$  and  $\Delta T_{sc}$  and increases  $\mu$ ,  $f$  and  $Q_{cr}$ ; as a result,  $\Delta T_m$  decreases,  $Q_{cr}$  and  $N_F$  increase, while the CT and the  $w$  decrease.
- Silicon increases  $f$  and  $\Delta T_{sc}$  for the entire range of silicon content; it also increases  $N_s$  and  $b$  for up to 2.4% Si followed by a reduction in its effect. Furthermore, Si lowers  $\mu$  in the range of 0–2.4% Si, and this is followed by an increase in  $\mu$ . As a result,  $N_F$  increases, while  $\Delta T_m$  changes from increasing to decreasing for Si contents beyond 2.4%. Thus,  $Q_{cr}$  increases, while the CT index and  $w$  both decrease.
- Manganese increases  $N_s$  and  $b$  and decreases  $\mu$ ; consequently, it increases  $\Delta T_m$  and  $N_F$  and decreases  $Q_{cr}$ , resulting in an increase in the CT and the  $w$ .
- Phosphorous increases  $f$  and  $\Delta T_{sc}$  and decreases  $\mu$  for the entire P content range. It also increases  $N_s$  and  $b$  for up to 0.25% Si, and then, its effect diminishes. Consequently, there is an increase in  $N_F$  and  $\Delta T_m$ , including  $Q_{cr}$ , while the CT and the  $w$  both decrease.
- Sulfur increases  $N_s$  and  $b$  and decreases  $\mu$ ; this leads to an increase in  $\Delta T_m$  and  $N_F$  and a reduction in  $Q_{cr}$ , resulting in an increase in the CT, as well as the  $w$ .
- Inoculation increases  $N_s$  and  $b$  and decreases  $\mu$ ; consequently, inoculation decreases the  $\Delta T_m$  and increases  $N_F$ , leading to a decrease in the CT, as well as the  $w$ .
- The time after inoculation decreases  $N_s$  and  $b$ ; consequently,  $\Delta T_m$  increases, while  $N_F$  decreases over time, leading to a reduction in  $Q_{cr}$  and an increase in the CT, as well as the  $w$ .

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## Author Contributions

Hugo F. Lopez wrote and edited the paper and contributed to all activities. Marcin Gorny and Magdalena Kawalec performed the experiments and analyzed the results.

## Conflicts of Interest

The authors declare no conflict of interest.

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