

Article

Study of Mid-Pressure Ar Radiofrequency Plasma Used in Plasma-Enhanced Atomic Layer Deposition of α -Al₂O₃

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Abstract: This study investigated the characteristics of radiofrequency, middle-pressure argon plasma used in the atomic layer deposition (ALD) of Al₂O₃ films. Based on the electrical characteristics—the current, voltage, and phase shift between them—and the stability of the plasma plume, the optimum plasma power, allowing reliable switching on of the plasma for any step of an ALD cycle, was determined. Spectral measurements were performed to determine the gas temperature and reactive species that could be important in the ALD process. The density of metastable argon atoms was estimated using tunable laser absorption spectroscopy. It was concluded that plasma heating of substrates did not affect film growth. The crystallization-enhancing effect of plasma observed in these experiments was due to the action of OH radicals produced in the plasma.

Keywords: capacitively coupled plasma; argon plasma; plasma-enhanced atomic layer deposition; optical emission spectroscopy

1. Introduction

Extensive research has been conducted on aluminum oxide as a thin-film material with diverse applications, including protective coatings [1,2], diffusion barriers [3–5], and electronic and optical devices [6–9]. Crystalline Al₂O₃ films offer clear advantages in applications that require elevated chemical stability [2,10], specific optical or electronic attributes [7,11,12], and superior mechanical properties [13,14]. However, the deposition of crystalline Al₂O₃ thin films using traditional thermal atomic layer deposition (ALD), often used for coating surfaces with complex shapes and/or deposition of ultrathin films, is challenging [10,14]. The use of plasma during the ALD process (PEALD) has the potential to facilitate the process, as it can reduce energy consumption and enable the growth of crystalline thin films with higher density at a reduced growth temperature (T_G) compared to thermal ALD processes [15]. The main functions of plasma in ALD process are as follows:

- (i) Generation of reactive species that participate in surface reactions [16–21];
- (ii) Energy delivery to the treated surface.

Reactive plasma species, such as metastable state atoms and molecules, and vibrationally excited molecules, can promote film growth by delivering their potential energy to the film via Auger processes [22], thereby reducing the activation energy of the reactions [16]. When the ALD process occurs in the active plasma region, energy transfer to the film surface can also occur via ions and heavy particles impacting the surface with high kinetic energy. In this case, the temperature of the impact site increases locally because of the momentum exchange between these particles and film lattice atoms, enabling crystalline film growth at lower substrate temperatures [15]. Plasmas offer the flexibility to produce specific reactive species, for example, by choosing a suitable gas composition and gas pressure and tailoring gas-phase chemistry with ALD surface reactions to obtain thin films with specific properties [23].

Capacitively coupled plasma is a commonly used plasma source in PEALD reactors, which, based on geometry, can be divided into direct and indirect treatment devices [15]. In



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the case of direct treatment, the substrate is placed directly into the plasma on an electrode; therefore, the deposited films experience reactions of neutral plasma species diffused onto the film as well as the impact of ions. Indirect-treatment devices utilize remote plasma sources, and substrates are placed in the plasma particle effluent, which mostly consists of long-lifetime reactive species. The production of reactive species (e.g., O, O₃, H, OH, N, N₂(A), and NH_x) occurs in the plasma of molecular gas or inert gas with a molecular gas admixture [23–26], whereas higher concentrations of molecular gas have been found to be favorable with respect to the reactive species yield [27]. At the same time, the increase in the molecular content impedes discharge ignition, which forces an increase in the voltage needed for discharge ignition and to sustain this process [28], which, in turn, could increase the plasma switching time jitter and force the plasma to run in the high-current, γ -mode. The high-current mode has been found to worsen the film uniformity compared to that produced using a low-current α -mode [29]. A possible solution for lowering the plasma ignition voltage while still producing reactive species is to use remote inert gas plasma mixed with molecular gases downstream from the plasma source. Such an arrangement has rarely been studied, although it has been found to provide a more stable discharge and even increase the yield of reactive species [30].

In our previous study [31], we investigated the influence of capacitively coupled Ar plasma applied during different steps of the ALD process on the growth of Al₂O₃ from trimethylaluminum (TMA) and water (H₂O) vapor. The Ar plasma was mixed with the precursors downstream of the plasma source. According to X-ray reflection (XRR) studies of the films deposited on Si(100) at a substrate temperature of $T_G = 650$ K, the application of plasma during the H₂O pulse and during the purge periods following the TMA and H₂O pulses caused an increase in the densities of the films compared with the densities of the films deposited via thermal ALD, that is, without plasma excitation. However, all the films obtained on bare Si(100) substrates were amorphous, and the increase in density caused by plasma excitation did not exceed 2%. More remarkable changes were caused by plasma excitation to the properties of the Al₂O₃ thin films deposited on the α -Cr₂O₃ seed layers. According to X-ray diffraction (GIXRD) studies, the films deposited at $T_G \geq 600$ K on α -Cr₂O₃ seed layers using Ar plasma during the purge after the H₂O pulse contained the α -Al₂O₃ phase, whereas the films deposited using thermal ALD were amorphous.

As the growth of α -Al₂O₃ at these relatively low T_G values is of marked importance for many applications, a more extensive characterization of plasma processes causing this effect was performed in this study. Electrical measurements, spatiotemporal optical emission spectroscopy, and tunable diode laser absorption spectroscopy (TDLAS) were used to determine the plasma parameters that could influence the PEALD process.

2. Materials and Methods

The experimental setup (Figure 1) is described in our previous study [18]. Here, the more specific details are focused on. The plasma was ignited using a 13.6 MHz RF generator coupled with a BDS-AMN 750 automatic matching network (both produced by the BDISCOM SRL).

The generator output power, fixed at 30 W, was switched on and off using a signal from the ALD process controller. At this generator power, the output voltage was sufficiently high for instant ignition of the plasma. A coaxial electrode configuration was used for plasma generation in our experiments. The diameter of the powered central stainless-steel electrode was 6.4 mm, and the inner diameter of the grounded surrounding stainless-steel tube was 16 mm (hereafter, “plasma tube”). The length of each electrode was 500 mm. The diameter of the quartz tube isolating the reactor from ambient air was 47 mm. The overall flow of Ar (99.999%) through the reactor was 300 standard cubic centimeters per minute (sccm), whereas the flow through the plasma tube was 150 sccm. The gas pressure in the reactor was 1.9–2.1 Torr. TMA (98%) and H₂O were used as precursors in the ALD film deposition and carried into the reaction zone by the Ar flow. The durations of the TMA and H₂O pulses and the TMA purge were 2 s, whereas that of the H₂O purge was 5 s.

Al_2O_3 films were deposited on a Si(100) substrate covered with an $\alpha\text{-Cr}_2\text{O}_3$ seed layer. The deposition of the seed layer is described in a previous study [13]. The samples were inserted into the reaction chamber on a SiO_2 holder. The temperature of the samples was recorded, prior to the ALD cycles and after temperature stabilization, using a K-type thermocouple.

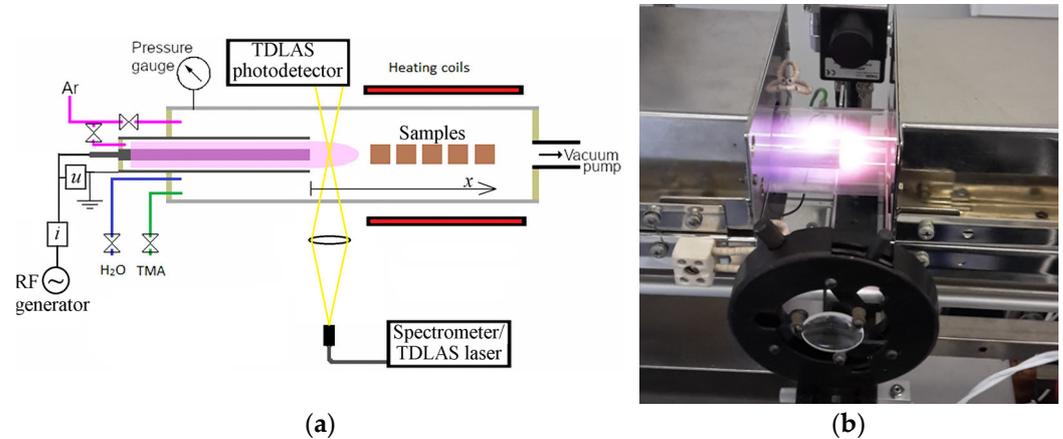


Figure 1. (a) Schematic representation of the experimental device; (b) Photo of PEALD.

The phase composition, thickness, density, and roughness of the films were measured using a SmartLab (Rigaku) diffractometer and analyzed using the GIXRD and XRR methods. The optical emission spectra of the plasma were recorded along the plasma axis at different distances (x) from the plasma tube orifice (Figure 1a) using Ocean Insight HR4 PRO (spectral range 274–414 nm, resolution approximately 0.1 nm) and Ocean Optics USB4000 (spectral range 185–850 nm, resolution approximately 1 nm) spectrometers. Because of the moderate spectral resolution of the USB4000 spectrometer, some spectral lines were not resolved and were deconvoluted using a nonoverlapped Ar spectral line at 696.5 nm as an instrumental function. An example of the deconvoluted lines is shown in Figure 2.

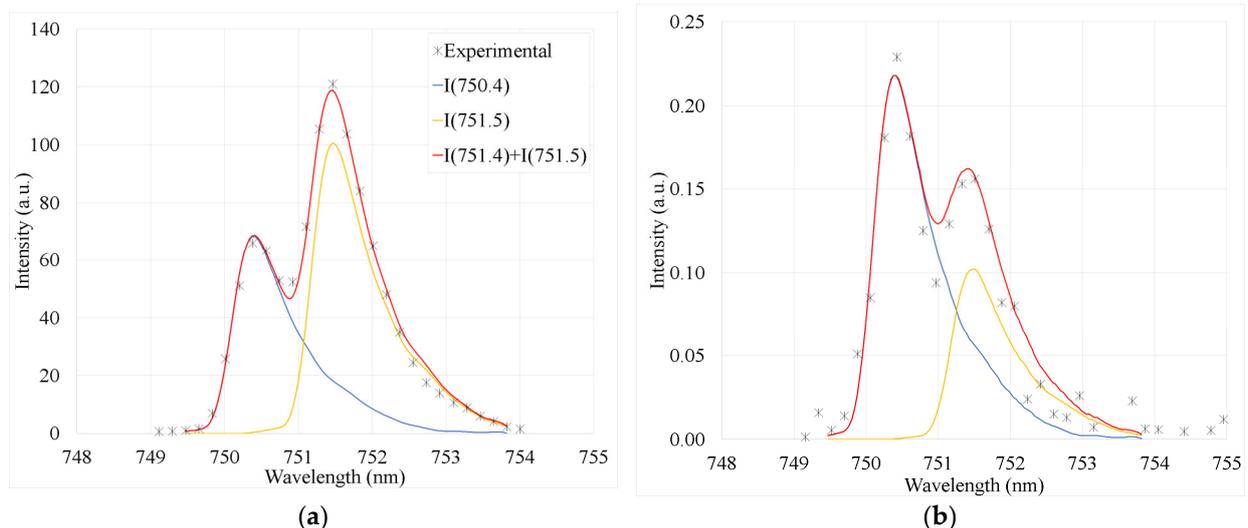


Figure 2. Example of experimental and deconvoluted Ar spectral lines at 750.4 and 751.5 nm: (a) in Ar discharge; (b) during TMA pulse (2 s after opening TMA valve).

The fitting of the experimental and calculated spectra used to estimate the rotational temperature and the concentration of Ar atoms in the metastable state $1s_5$ (Paschen notation) was performed using the MathCad 15.0 software function Minimize. Electrical characteristics were recorded using an oscilloscope Tektronix TDS-540B. The voltage (u) was measured using a 1:100 Tektronix voltage probe P5100, whereas the current (i) was

measured using a McPherson current monitor 6585. The concentration of Ar $1s_5$ state atoms was estimated using a TDLAS unit. The scan range of the Littman-type external cavity diode laser Thorlabs TLK-L780M with a power of 50 mW was 740–800 nm, and the line width was <130 kHz. The free spectral range and finesse of Fabry–Perot interferometer Thorlabs SA-200, used to check the laser operation stability and to calibrate the relative wavelength, were 1.5 GHz and 200, respectively. A photodetector Thorlabs APD110A2 coupled with an interference filter with a central wavelength of 763.5 nm and a full width at half maximum of 40 nm was used to register the laser signal.

3. Results and Discussion

3.1. Plasma Characteristics

3.1.1. Discharge Appearance and the Choice of Generator Power

The appearance of the discharge depended on the generator power. At low power values, the radiation intensity in the plasma tube was low, and the space between the electrodes was only partially filled with plasma. An increase in the generator power caused the gradual filling of the tube with plasma and, finally, the emergence of a plasma plume outside the tube. At a generator power of 30 W and after a long-term Ar purge, the bright plasma plume of the Ar discharge extended outside the tube to a distance of 20–30 mm, whereas the diameter of the plume was somewhat larger than that of the outer electrode, as shown in Figure 1b. A low-intensity plasma column was observed at longer distances (up to 100 mm). By further increasing the generator power, randomly located irregular plasma sparks, in addition to homogenous plasma, appeared between the electrodes. With the appearance of sparks, instabilities were observed in the plasma plume outside the tube. Therefore, ALD was conducted at a generator power of 30 W. Outside the tube, the discharge appearance also depended on the ALD step. The TMA and H_2O pulses caused a rapid reduction in the extent of the plasma plume outside the plasma tube and changed the plasma color. During the following purge, plasma extension was gradually restored.

3.1.2. Electrical Characteristics

Plasma power was calculated from the current and voltage waveforms as described in our previous study [32]. When the plasma was produced at a low generator power, the waveforms exhibited a sinusoidal shape, similar to that observed without discharge. At a generator power of 30 W, the current waveform deviated significantly from a sinusoidal waveform, as shown in Figure 3. The irregular shape of the current waveform can be explained by the different areas of powered and grounded electrodes.

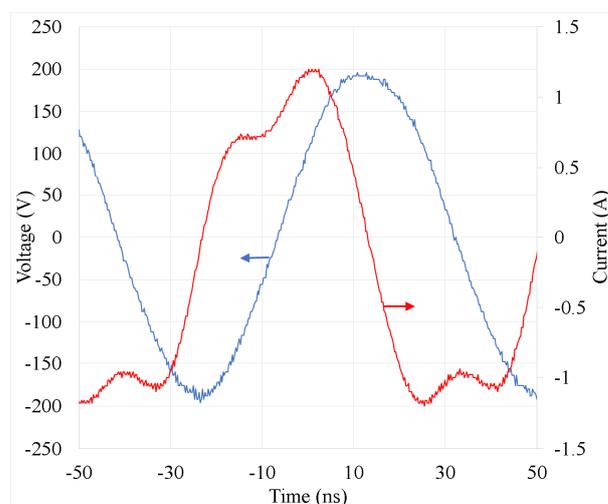


Figure 3. An example of measured current and voltage waveforms. The phase shift component due to different lengths of connecting cables and parasitic capacitance was considered in the graph.

Plasma power was estimated using the formula $P = \frac{1}{T} \int_0^T i_p(t) u_p(t) dt$. The phase shift component caused by different lengths of the connecting cables, parasitic capacitance, etc., was determined from the measured u and i waveforms without discharge, and it was considered in the plasma power calculation. At a generator power of 30 W, the phase shift between i and u was $84.2 \pm 1.0^\circ$, and the plasma power was 11.6 ± 1.8 W. The plasma power did not depend on the Ar plasma gas flow rate in the plasma tube (tested in the range of 100–300 sccm) or on the gas composition outside the plasma tube in the reactor. Hence, the plasma power was completely determined by the Ar discharge inside the plasma tube.

3.1.3. Plasma Spectrum

The optical emission spectra recorded at distances of 5 and 50 mm are shown in Figure 4a,b, respectively. In the case of the Ar plasma, the most intense spectral lines belonged to Ar. Because of the impurities in Ar (due to leakage, etc.), a weak atomic oxygen line at 777 nm, and bands of OH at 308 nm, a N₂ first positive system (FPS) at 550–700 nm and a N₂ second positive system (SPS) at 300–400 nm were observed. Deconvolution of the N₂(C-B,0-0) band also revealed the presence of an NH(A-X,0-0) band at 336 nm (Figure S1). The start of the TMA pulse caused a rapid decrease in the intensity of the aforementioned molecular bands and atomic lines (Figure 4c, time interval $\Delta t = 10$ –15 s). As a result of TMA (Al₂(CH₃)₆) decomposition, bright Al lines at 308.22, 309.27, 394.40, and 396.15 nm [33] appeared in the spectra. During the following purge ($\Delta t = 15$ –25 s), the intensity of the Ar lines and N₂ bands gradually increased and became even more intense than before the ALD pulses, whereas the OH intensity decreased due to H₂O removal in the reaction with TMA. The emission of CN radical at 358, 388, and 416 nm, CH at 386 and 431 nm, and C₂ at 516.2 nm were identified from the molecular bands [34–37]. Interestingly, although the appearance of these bands is also related to TMA decay, they were not observed during the TMA pulse but during the TMA purge. A possible reason for this observation is the increase in electron temperature and density during the TMA purge, which enabled the production and excitation of these molecules. Similar to the TMA pulse, the H₂O pulse ($\Delta t = 25$ –30 s) reduced the intensities of all the spectral lines and N₂ bands, with the exception of the OH band. The increased production of OH excited-state molecules can be attributed primarily to the reaction between Ar 1s_{2...5} atoms and H₂O molecules [38–40]. During the purge following the H₂O pulse ($\Delta t = 30$ –40 s), the intensities of the Ar lines and N₂ bands gradually increased, whereas the OH intensity reached a sharp maximum ~1 s after the end of the H₂O pulse and then started to decrease.

With an increase in the distance x , the intensities of all bands and lines decreased; however, the general trends remained similar, as described above (Figure 4b,d).

3.1.4. Gas Temperature

The plasma gas temperature (T_{gas}) was estimated using various methods. First, T_{gas} was estimated based on the N₂(C) rotational temperature (T_{rot}) calculated for several vibrational transitions, as described in our previous study [41]. In the case of electron impact excitation of N₂(C), $T_{\text{rot}} \approx T_{\text{gas}}$. However, in Ar discharges, T_{rot} of these molecules could be influenced by the energy transfer from the metastable Ar atoms, which can result in remarkably higher T_{rot} values than T_{gas} [41]. The rotational temperatures of the N₂ SPS vibrational transitions 0-2 (bandhead at 380 nm), 1-3 (375 nm), and 2-4 (371 nm) (Figure 5) were estimated for Ar discharge (after a long-term Ar purge) to reduce the interference of TMA/H₂O additives on the rotational spectra. The temperature near the plasma tube orifice estimated under such conditions is expected to be somewhat higher than that during ALD pulses because the plasma plume extension outside the plasma tube was the largest in the case of pure Ar discharge, and the plasma power was insensitive to the gas composition outside the plasma tube (Section 3.1.2).

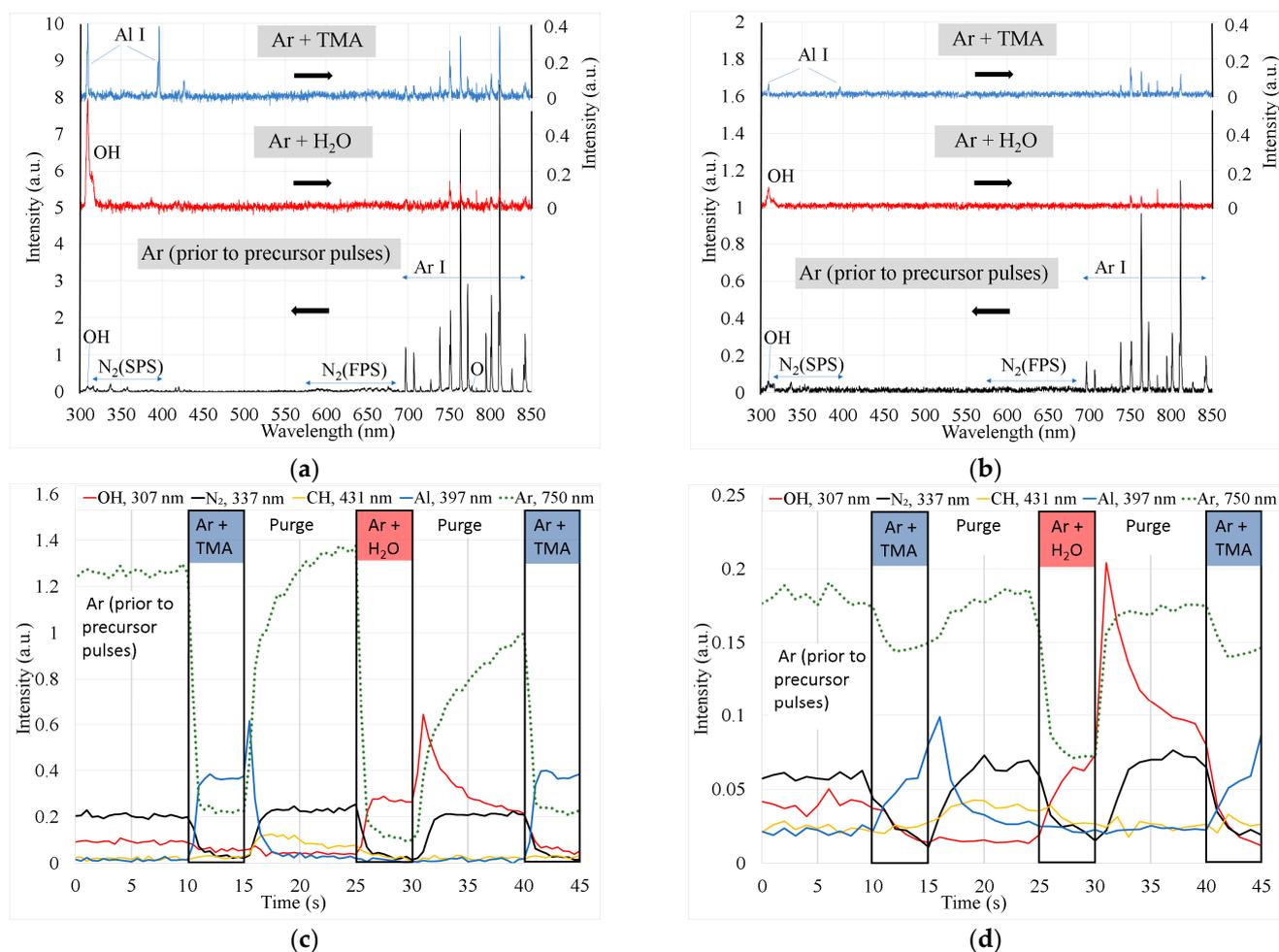


Figure 4. (a,b) Spectra of Ar plasma after long-term Ar purging and during H₂O (Ar + H₂O) or TMA (Ar + TMA) pulses recorded at $x = 5$ mm (a) and $x = 50$ mm (b). The spectra describing the effect of the H₂O and TMA pulses were recorded 2 s after the corresponding valve was opened. The intensity scale of the corresponding spectrum is indicated by the black arrow. (c,d) Intensities of OH (307 nm), N₂ (337 nm), CH (431 nm), Al (397 nm), and Ar (750 nm) lines recorded prior to ALD pulses (0–10 s) and during the TMA pulse (10–15 s); the first purge (15–25 s); H₂O pulse (25–30 s); the second purge (30–40 s); and TMA pulse (40–45 s) at $x = 5$ mm (c) and $x = 50$ mm (d). Plasma was turned on for the entire measurement period (0–45 s).

Examples of the recorded and calculated spectra of the N₂ SPS vibrational transitions 0-2 and 2-4 are shown in Figure 6a. The lowest rotational temperature, $T_{\text{rot}} \approx 450 \pm 100$ K, among the studied N₂ SPS vibrational transitions was found for N₂(C, $v = 2$) near the plasma tube orifice ($x = 0$ mm), and it diminished with increasing distance (Figure 6b). At $x > 10$ mm, the N₂(C-B, 2-4) band intensity was too low to estimate temperature reliably.

Considerably higher T_{rot} values were found for the N₂(C, $v < 2$) vibrational states, reaching ≈ 1300 K for N₂(C, $v = 0$) and ≈ 1100 K for N₂(C, $v = 1$) (Figure 6b). This dependence of T_{rot} on the number of vibrational states v can be explained by the excitation transfer from the Ar excited state atoms $\text{Ar}(1s_{2...5}) + \text{N}_2(\text{X}) \rightarrow \text{Ar} + \text{N}_2(\text{C})$, which can cause an overpopulation of higher rotational levels when comparing the population distribution corresponding to equilibrium with T_{gas} [41]. From an energetic perspective, the ability to populate higher rotational states via this reaction decreases for higher values of v , as shown in the inset of Figure 6b, and the reaction rate coefficient also decreases with the vibrational number. For example, at 300 K, the ratio of the rate coefficients for the population $v = 0, 1$, and 2 states is 1:0.25:0.1 [42]. Therefore, with trace amounts of N₂ in Ar plasmas, lower vibrational levels are more susceptible to the overpopulation of higher

rotational levels than higher vibrational levels. At higher pressures, the neutral particle collisions can still thermalize the rotational population distribution even for the vibrational state $N_2(C, v = 0)$ and enable T_{rot} to be used as a T_{gas} estimate [43]. At the low pressure used in our experiment, the thermalization is apparently too slow, and T_{rot} from $N_2(C, v < 2)$ overestimates T_{gas} .

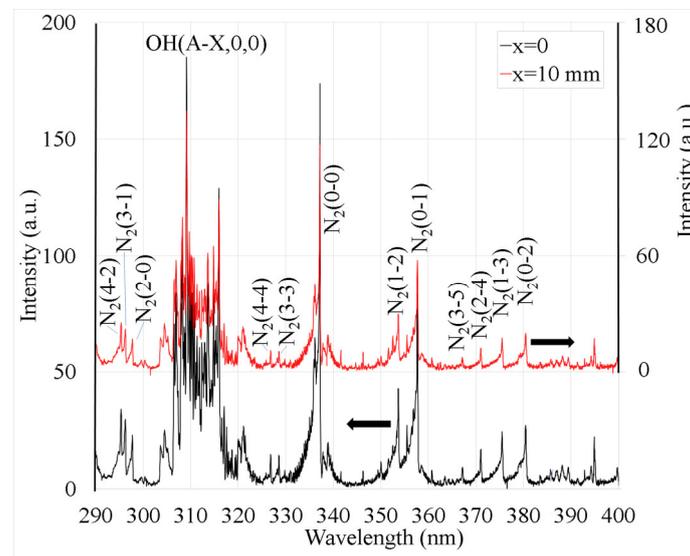


Figure 5. Spectra of OH(A-X,0-0) and several bands of N_2 SPS registered using an Ocean Insight HR4 PRO spectrometer at distances $x = 0$ and $x = 10$ mm. The reaction chamber heater was turned off. In brackets, N_2 SPS vibrational transitions are shown. The intensity scales of corresponding spectra are shown by black arrows.

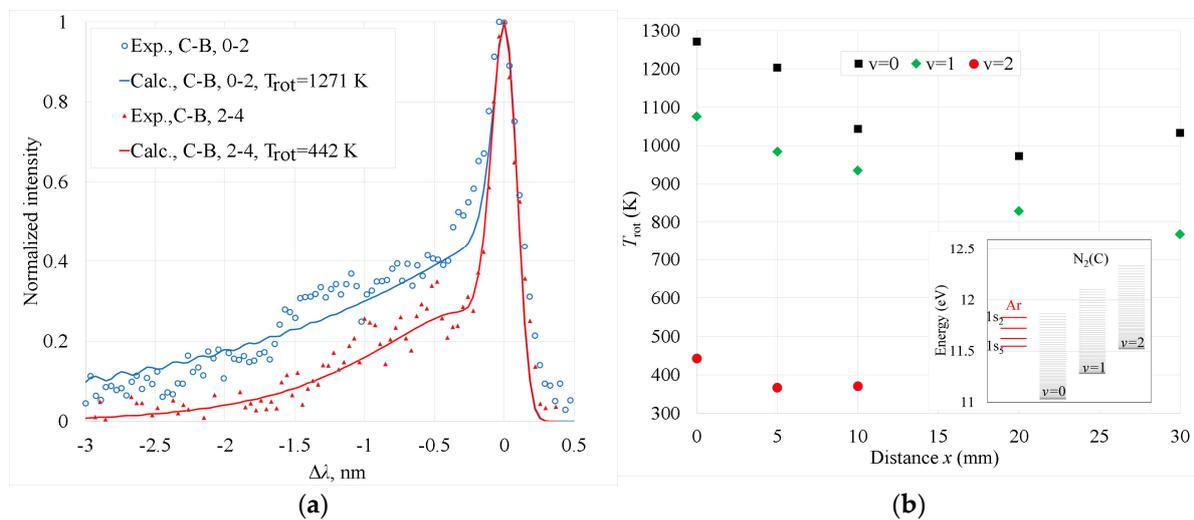


Figure 6. (a) Example of registered and calculated spectra of $N_2(C-B)$ 0-2 and 2-4 vibrational transitions at the distance $x = 0$. (b) Rotational temperature of $N_2(C, v = 0, 1, 2)$ as a function of x . Inset depicts the energy levels of Ar metastable ($1s_3$ and $1s_5$) and resonant ($1s_2$ and $1s_4$) state atoms and N_2 rotational states (up to $J = 60$) for vibrational states $v = 0, 1, 2$.

An attempt to determine T_{gas} was made using the OH(A-X,0-0) transition, which has been used to estimate T_{gas} at higher pressures [44]. In our spectra, the OH(A-X,0-0) band at 307 nm is one of the brightest peaks during the H_2O pulse and the following purge (Figure 4). Similar to $N_2(C, v < 2)$, higher rotational levels of OH(A,0) can be excited by energy transfer from excited Ar atoms, which results in T_{rot} higher than that expected from T_{gas} [39]. The radiative lifetime of OH(A, $v = 0$) is remarkably longer than that of the

N_2 SPS bands [41]. Therefore, the rotational relaxation can still be expected to achieve a population distribution equilibrium with T_{gas} . However, in our plasma, OH(A, $v = 0$) T_{rot} also overestimated T_{gas} , as the temperature determined using Lifbase 2.1 software [45], $T_{\text{rot}}(x = 0) = 1300$ K, was remarkably higher than that of $N_2(C, v = 2)$. Consequently, the relaxation was still too slow at the gas pressure used in our ALD reactor.

The gas temperature was additionally estimated by analyzing the absorption shape of the 763.5 nm line, which was determined using the TDLAS technique. The shape of the line depends on T_{gas} [46,47], and the temperature estimation procedure is described below Figure S2. The remarkable fluctuation in the laser emission of our TDLAS setup (Figure S3) and the low signal-to-noise ratio prevented the reliable determination of T_{gas} as a function of distance x . At $x = 0$, the determined temperature was 700 ± 200 K (Figure S2). This coincided with T_{rot} of $N_2(C, v = 2)$ within the uncertainty margins.

Knowledge of plasma power also allows a rough estimation of T_{gas} [48]. Under stationary conditions, $T_{\text{gas}} = T_E + \frac{w}{n_0 \cdot C_p \cdot v_H}$, where T_E is the temperature of the plasma tube wall, $w \approx 0.14$ W cm⁻³ is the power density, n_0 is the gas density, $C_p \approx 3.3 \times 10^{-23}$ J/K is the specific heat of Ar at constant pressure per atom, and $v_H = \frac{8 \cdot \lambda_{Ar}}{n_0 \cdot c_p \cdot r^2} + \frac{2 \cdot v_{\text{gas}}}{L}$ is the heat removal frequency. Here, λ_{Ar} is the thermal conductivity of Ar (0.0177 W/(K·m)), $r = 8$ mm is the inner radius of the plasma tube, v_{gas} is the linear velocity of the gas (5.9 m/s) in the plasma tube, and $L = 500$ mm is the length of the plasma tube. The first and second terms in the formula used for the calculation of v_H provide the heat removal frequency by heat conduction and convective heat transport, respectively. Under our conditions, the contribution of heat removal by convective heat transport was $\approx 2.3\%$. Therefore, T_{gas} depended only slightly on n_0 . Presuming T_E is stabilized at room temperature, the calculated T_{gas} near the plasma tube orifice is $T_{\text{gas}} \approx 360$ K. Actually, T_E is probably higher owing to the poor cooling of the plasma tube; however, it is expected to be lower than the maximum temperature of the O-rings of our vacuum system (ERIKS' 51,414 green, $T_{\text{max}} = 473$ K). This gives the upper limit of T_{gas} as 540 K, which also coincides with the T_{rot} of $N_2(C, v = 2)$ within the uncertainty margins. Therefore, in the following, we used the T_{rot} of $N_2(C, v = 2)$ as T_{gas} .

3.1.5. Line-Integrated Concentration of Ar Metastable 1s₅ State Atoms

The concentration of metastable Ar 1s₅ state atoms ([Ar(1s₅)] was determined in the Ar discharge on the basis of Ar spectral line $\lambda_0 = 763.5106$ nm [33] absorption (transition 1s₅ → 2p₆) using the TDLAS technique [46]. The estimation was performed between distances of 0–10 mm, because, at $x > 10$ mm, the absorption was below the detection limit of the TDLAS apparatus. The optical depth (OD) (Figure 7a) was determined as $OD = \ln\left(\frac{I_0}{I_t}\right) = k(\lambda) \cdot l$, where I_t and I_0 are the laser intensities with and without plasma, respectively, and k and l are the absorption coefficient and absorption length, respectively.

The concentration of Ar(1s₅) state atoms was calculated using the experimentally determined OD according to the formula $[Ar(1s_5)] = \frac{8 \cdot \pi \cdot g_i \cdot c}{\lambda_0^4 \cdot g_j \cdot A_{ij} \cdot l} \int k(\lambda) \cdot l \cdot d\lambda$ [46]. Here, $g_i = 5$ and $g_j = 5$ are the statistical weights of the upper and lower states, respectively; c is the speed of light; $A_{ij} = 2.45 \times 10^7$ s⁻¹ [33] is the Einstein coefficient; and $l = 20$ mm is the absorption length. The concentration of Ar(1s₅) atoms decreased almost exponentially with increasing x from 4.5×10^{10} to 4×10^9 cm⁻³ when x increased from 0 to 10 mm.

The introduction of any ALD precursor caused a decrease in the concentration of Ar(1s₅) to below the detection limit of our TDLAS apparatus. However, the spectral measurements indicated that, during the TMA and H₂O pulses, the Ar(1s_{2...5}) concentrations were very low. Compared with the other Ar spectral lines, the intensity of the Ar line at 750.4 nm was less sensitive to the introduction of TMA and H₂O. The rate coefficient for the population of the upper state of this transition, 2p₁, via the direct electron impact excitation from the ground state, is higher than that for other 2p-1s transitions observed in the spectra; for other transitions, stepwise excitation from Ar 1s_{2...5} is more important [49]. As an example, Figure 2 shows the intensities of spectral lines at 750.4 (transition 2p₁ → 1s₂) and

751.5 nm (transition $2p_5 \rightarrow 1s_4$) registered from Ar plasma and during a TMA pulse. The 751.4 nm line was more intense in the Ar discharge prior to the ALD pulses and also during the purge periods, but during the TMA and H_2O pulse the 750.4 nm line dominated. Notably, the ratio of the rate coefficients for direct electron impact excitation to $2p_1$ and $2p_5$ was approximately 2 [49] in the electron temperature range $T_e = 2.5\text{--}3.5$ eV. This is close to the intensity ratio of the transitions $2p_1 \rightarrow 1s_2$ and $2p_5 \rightarrow 1s_4$ (Figure 2b). Considering similar Einstein coefficients of these transitions ($A_{750.4} = 4.5 \times 10^7 \text{ s}^{-1}$; $A_{751.5} = 4.0 \times 10^7 \text{ s}^{-1}$ [33]), this finding indicates negligible stepwise excitation of Ar states during the TMA and H_2O pulse due to the very efficient quenching of the $1s_{2..5}$ states.

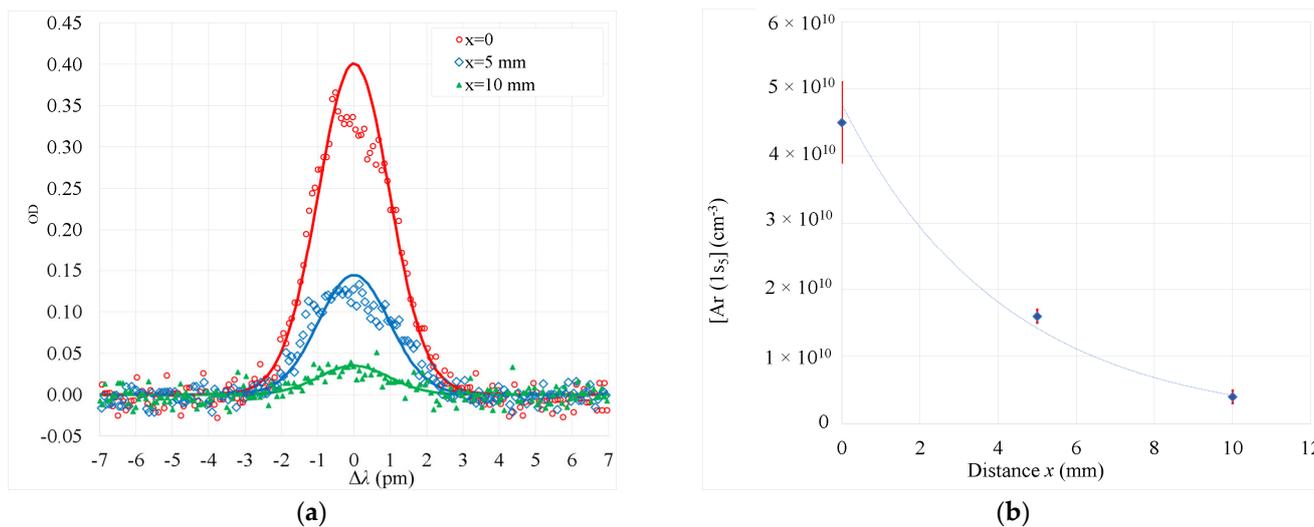


Figure 7. (a) Optical depth (OD) of Ar absorption line at 763.5 nm for different x . The absorption line profile was fitted with the Voigt function shown using solid lines. (b) Ar $1s_5$ state atom concentration $[Ar(1s_5)]$ as a function of x . The solid line represents the exponential trendline of data points. Error bars shown in the figure are found as differences of $Ar(1s_5)$ concentrations calculated on the basis of experimental OD and the Voigt fitting.

3.2. Effect of Plasma on Film Properties

The ALD process used for the deposition of films for post-growth studies consisted of 1000 ALD cycles. Each cycle contained the following steps: TMA pulse (duration 2 s), purge (2 s), H_2O pulse (2 s), and purge (5 s); it was possible to switch the plasma on for any step. Our previous study [31] revealed that plasma applied during the TMA pulse caused TMA decomposition, resulting in low-density amorphous Al_2O_3 films. It was also found that plasma had a detrimental effect on crystal growth when applied during TMA purge, whereas, in this case, the film growth per cycle (GPC) and density were similar to those of thermal ALD. As indicated by the intensity decay of the Al spectral lines (Figure 4c), the TMA residuals were removed from the gas within 1–2 s of the TMA pulse. Similarly, fast TMA removal after the TMA pulse was also concluded from Q-pod quartz crystal microbalance (QCM) measurements [31]. With a decrease in the intensity of the Al lines during the TMA purge, CH, CN, and C_2 bands appeared in the spectrum (Figure 4c) and were observed throughout the TMA purge step. The appearance of these bands, but not Al lines, indicates that plasma applied during the TMA purge caused decomposition of surface intermediate species formed during the TMA ($Al_2(CH_3)_6$) pulse. The decomposition of surface species led to the formation of excited CH, CN, and C_2 molecules via plasma–chemical reactions involving desorbed CH_x species. The Al atoms remained on the surface, which explains why the GPC and density of the film were not affected by the plasma applied during the purge following the TMA pulse.

The plasma excitation, applied during the purge following the H_2O pulse or during the H_2O pulse and the following purge, most significantly contributed to the increase in

the density and crystallization of the Al_2O_3 films grown on the $\alpha\text{-Cr}_2\text{O}_3$ seed layers. This effect enabled us to obtain $\alpha\text{-Al}_2\text{O}_3$ in films deposited at 600 K (Figure 8).

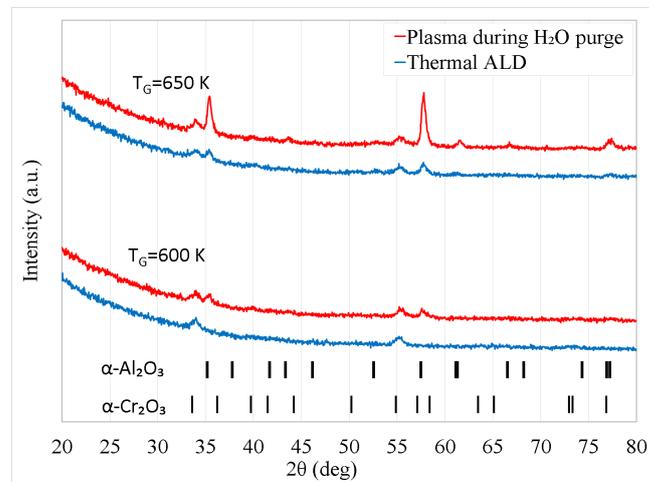


Figure 8. GIXRD diffractograms of Al_2O_3 films deposited on $\alpha\text{-Cr}_2\text{O}_3$ at substrate temperatures of 600 and 650 K via thermal ALD and PEALD (plasma turned on during the H_2O purge). During deposition, the samples were located at a distance of 20 mm from the outlet of the plasma tube. Reflections peaking at 35.3 and 57.6° indicate the formation of $\alpha\text{-Al}_2\text{O}_3$.

The film growth depended on the distance between the samples and the orifice of the plasma tube. At $x = 20$ mm, the plasma applied during the purge following the H_2O pulse significantly enhanced the growth of $\alpha\text{-Al}_2\text{O}_3$, as shown in Figure 9. In contrast, the GIXRD diffractograms of the films deposited via PEALD and thermal ALD at $x = 70$ mm were similar. Consequently, the influence of plasma on crystal growth was negligible in the latter case.

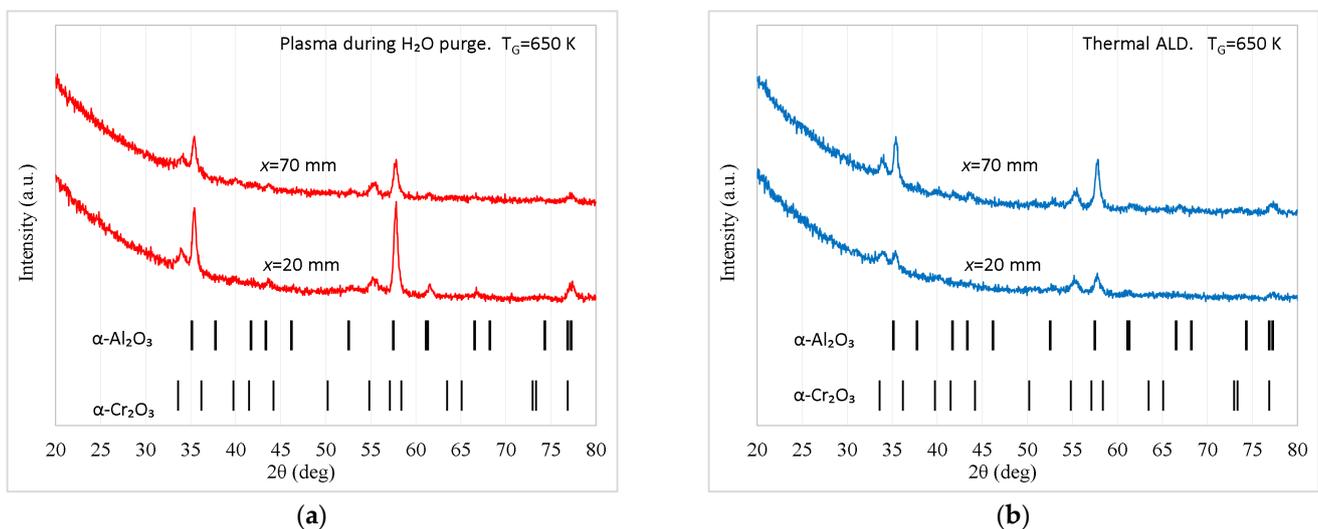
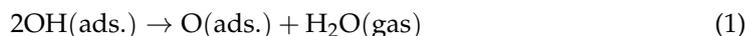


Figure 9. GIXRD diffractograms of Al_2O_3 films deposited at a substrate temperature of 650 K on $\alpha\text{-Cr}_2\text{O}_3$: (a) PEALD (plasma turned on during H_2O purge); (b) thermal ALD. During the deposition, the samples were located at distances of 20 and 70 mm from the outlet of the plasma tube.

Our previous study showed that, compared with the purge following the TMA pulse, a markedly longer purge period was needed after the H_2O pulse to obtain self-limited film growth [31]. Spectral measurements in the present study confirmed this result, as the OH band intensity decay after the H_2O pulse was substantially slower than that of the Al lines observed after the TMA pulse (Figure 4). The likely reason for this is the gradual release of

H₂O to the gas phase from the surface due to the reaction between the OH groups formed during the H₂O pulse [50]:



and the subsequent formation of excited OH in gas-phase reactions, as explained in Section 3.1.3. As the surface hydroxyl groups are absorption sites for TMA [51], the decrease in the site concentration on the surface causes a reduction in GPC. The plasma applied during the H₂O purge caused a significant (~10%) decrease in GPC compared with that obtained via thermal ALD (Table 1). However, the density of the film deposited via PEALD on the substrates located at $x = 20$ mm was approximately 7% higher than the densities of other films (Table 1). Thus, the plasma applied during the H₂O purge insignificantly influenced the mass growth per cycle (MGPC) at $x = 20$ mm; at $x = 70$ mm, the effect of plasma on MGPC was considerable.

Table 1. Growth per cycle (GPC), density (ρ), mass growth per cycle (MGPC = $\rho \cdot \text{GPC}$), and surface roughness of films grown via thermal ALD and PEALD (plasma switched on during H₂O purge) on substrates positioned at different distances from the plasma tube orifice. Films were deposited at 650 K on substrates coated with $\alpha\text{-Cr}_2\text{O}_3$.

ALD Process	x , mm	GPC, nm	ρ , g/cm ³	MGPC, ng/cm ²	Roughness, nm
Thermal ALD	20	0.104 ± 0.002	3.21 ± 0.06	33.4 ± 0.9	2.4 ± 0.1
PEALD	20	0.094 ± 0.002	3.43 ± 0.07	32 ± 1	3.3 ± 0.1
Thermal ALD	70	0.105 ± 0.002	3.27 ± 0.06	34.3 ± 0.9	1.5 ± 0.1
PEALD	70	0.092 ± 0.002	3.25 ± 0.06	29.9 ± 0.9	1.5 ± 0.1

A possible explanation for this difference is the plasma enhancement of the reaction described by Equation (1) and the re-adsorption of OH radicals formed in the plasma. At $x = 20$ mm, the concentration of OH radicals in the gas phase was probably high enough to compensate for the loss of surface OH groups; at $x = 70$ mm, this kind of compensation was evidently not obtained. As the concentration of OH in the gas phase could not be very high, even at $x = 20$ mm, the re-adsorption was obviously a site-sensitive process supporting crystallization. With increasing x , the concentration of OH in the gas phase decreased (Figure 4); therefore, the contribution of OH re-adsorption to the crystal growth also decreased and became negligible.

The effect of possible plasma heating that could also promote the crystallization and influence GPC is not probable because the GPC values of the films deposited at higher gas temperatures (inside the plasma plume at $x = 20$ mm) and at lower gas temperatures (at $x = 70$ mm) are the same; although, according to the results of our previous study, the GPC should considerably decrease with increasing T_G . For example, the increase in T_G from 620 to 650 K caused the GPC to decrease from 0.115 to 0.089 nm [31]. In addition, the plasma turned on during the purge period following the TMA pulse did not cause changes in the GPC.

A possible explanation for the fast decay of the plasma effect on crystallization, observed with increasing x , is the corresponding decrease in the concentration of reactive species and the short lifetime of those species. The formation of long-living H₂O₂ has also been observed in plasma at atmospheric pressure in the presence of water vapor [52]. However, under our conditions, the corresponding reaction is relatively slow (rate coefficient $k_1 \approx 4 \times 10^{-14}$ cm³s⁻¹ at 400 K [53]). The lifetime of OH appears to determine the concentration profile of the reactive species in the gas phase. Although the quenching of the ground state OH radical, OH(X), is slower than that of the OH(A) state [54,55], the convective transport of OH(X) from the OH-rich plasma plume region to the sample location at $x = 70$ mm is still unlikely. For example, in the case of an Ar: 9% H₂O mixture plasma at 200 Torr pressure, the OH density decay to 10% took up approximately 500 μ s, whereas the OH diffusion losses were insignificant [56]. Considering the 100 times lower

pressure and linear velocity of gas in our reactor (1.1 m/s), such decay occurs within 55 mm. Actually, this distance is expected to be even smaller as, at our low pressure, additional OH diffusion losses should be significant. Spectral measurements that revealed a rapid decrease in the OH intensity and, consequently, in the OH concentration with increasing x (Figure 4c,d) confirmed this estimation.

4. Conclusions

The present study investigated the properties of low-pressure argon RF plasma used in the plasma-enhanced atomic layer deposition of Al₂O₃ films. The plasma power, estimated from the recorded electrical characteristics, was approximately 12 W. The argon metastable atom density in pure Ar plasma decreased 10 times with increasing distance from the plasma tube orifice from 0 to 10 mm. The addition of the ALD precursors reduced the density of metastable atoms to a level below the detection limit of the apparatus used. The gas temperature was estimated using the N₂(C, $v = 2$) rotational temperature, which was near the plasma tube orifice ≈ 400 K. Plasma spectra, measured during the TMA pulse, revealed TMA disintegration, as concluded from the appearance of Al lines and CN(B-X) bands. The H₂O pulse caused a rapid increase in the intensity of the OH(A-X,0-0) band, indicating effective OH production.

In respect to Al₂O₃ film crystallization, the following conclusions are drawn:

- (i) Under our conditions, the plasma gas temperature does not affect Al₂O₃ crystallization.
- (ii) The main plasma agents that helped Al₂O₃ crystallization were OH radicals. Plasma-enhanced Al₂O₃ crystallization was observed only for samples located near the plasma tube orifice, inside the plasma plume where the OH concentration was high. Because of the rapid decay in the OH concentration with distance, the plasma effect on crystallization was negligible at 70 mm from the tube orifice.

Supplementary Materials: The following supporting information can be downloaded at: <https://www.mdpi.com/article/10.3390/pr12030612/s1>, Figure S1: NH(A-X,0-0) band spectrum. Figure S2: Temperature estimation from Ar (1s₅→2p₆) line shape analysis. Figure S3: TDLAS laser stability on the basis of Fabry–Pérot interferometer signal. (See Refs. [33,47,57]).

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