



## Article

# Screening of Amino Acids and Surfactant as Hydrate Promoter for CO<sub>2</sub> Capture from Flue Gas

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**Abstract:** In this study, the kinetics of flue gas hydrate formation in bulk water in the presence of selected amino acids and surfactants are investigated. Four amino acids (3000 ppm) are selected based on different hydropathy index. Constant-ramping and isothermal experiments at 120 bar pressure and 1 °C temperature are carried out to compare their hydrate promotion capabilities with surfactant sodium dodecyl sulfate (SDS) (500–3000 ppm) and water. Based on experimental results, we report the correlation between hydrate promotion capability of amino acids and their hydrophobicity. Hydrophobic amino acids show stronger flue gas hydrate promotion capability than water and hydrophilic amino acids. We discuss the controlling mechanisms to differentiate between promoters and inhibitors' roles among the amino acids. Between 2000–3000 ppm concentrations, hydrophobic amino acids have near similar promotion capabilities as SDS. This research highlights the potential use of amino acids as promoters or inhibitors for various applications.

**Keywords:** flue gas hydrate; amino acids; CO<sub>2</sub> capture; sodium dodecyl sulfate

## 1. Introduction

To reduce CO<sub>2</sub> emissions into the atmosphere, the scientific focus has been on the advancement of CO<sub>2</sub> capture, separation, and storage technologies from pre-combustion (CO<sub>2</sub>/H<sub>2</sub>) and post-combustion (CO<sub>2</sub>/N<sub>2</sub>) feed gases [1]. Post-combustion CO<sub>2</sub> released from thermal power plants, cement plants, steel plants, coal-based chemical plants has been considered as a significant factor behind the increase in the greenhouse effect [2]. Gas hydrate-based CO<sub>2</sub> capture, separation, and storage has been proposed for its potential as an eco-friendly and less costly technology, and it could provide an innovative solution to capture CO<sub>2</sub> [3,4]. Another advantage of using a hydrate based technique to separate CO<sub>2</sub> from flue gas is the advantage of using hydrate formation for separating CO<sub>2</sub> is due to less energy consumption than traditional means [5]. Hydrate formation is divided into three main steps, which include the dissolution of CO<sub>2</sub>-rich gas into liquid phase, hydrate nucleation, and hydrate growth phase. While the nucleation stage indicates the supersaturation of the liquid phase by CO<sub>2</sub>-rich gas, the growth phase is responsible for gas capture into the hydrate phase. Linga et al. [6] have proposed a multistage hydrate based CO<sub>2</sub> separation process from a CO<sub>2</sub>-rich gas system and showed 36–42% CO<sub>2</sub> recovery in the first step, in the form of hydrate using pure water and a CO<sub>2</sub>/N<sub>2</sub> system.

The post-combustion CO<sub>2</sub>-rich gas system includes 15–25% CO<sub>2</sub> in a CO<sub>2</sub>/N<sub>2</sub> gas mixture [6,7] and including impurities, such as SO<sub>2</sub>, H<sub>2</sub>S, and fly ash. Thermodynamic studies confirm that such a CO<sub>2</sub>/N<sub>2</sub> system would require high compression and high energy consumption during the hydrate-based gas separation process, which makes it costly to adopt at a commercial scale. For commercialization, chemicals known as promoters are starting to become considered for CO<sub>2</sub> capture, separation, and storage applications, due to their role in lowering formation pressure,

formation time, enhanced gas uptake, and improvement in kinetics. Dasthi et al. [3] have reviewed the role of different promoters in hydrate-based CO<sub>2</sub> capture.

Thermodynamic promoters such as tetrahydrofuran (THF) [8] and cyclopentane [9] have been successfully tested previously for CO<sub>2</sub> capture and separation from a gas mixture at lower formation pressure. Kinetic hydrate promoters, such as the surfactant sodium dodecyl sulfate (SDS), are known to improve formation both individually and along with other promoters, such as THF [10] and TBAB [11]. Other advantages include increased storage capacity and low dosage. Many theories have been proposed to explain the effect of SDS on the growth mechanism, and a change in interfacial properties of the solution in the presence of surfactant has been widely accepted as a governing mechanism. A review of all possible mechanisms during surfactant-based gas hydrate formation can be found from Yanhe et al. [12].

Few studies on kinetics have been done to evaluate and compare the promotion capability of surfactants for CO<sub>2</sub> capture, separation, and storage from a CO<sub>2</sub>-rich feed gas mixture [7,13,14]. Kumar et al. [14] have studied CO<sub>2</sub> capture from 15–20% CO<sub>2</sub> in a CO<sub>2</sub>/N<sub>2</sub> mixture using liquid water and found 55–57% CO<sub>2</sub> recovery in hydrate at 7.7 MPa and 0.6 °C. In the presence of impurities, such as SO<sub>2</sub>, it was found to lead to a higher temperature, lower formation pressure, and faster nucleation during CO<sub>2</sub> hydrate formation from the CO<sub>2</sub>/N<sub>2</sub> feed gas. In another study, it was found that fly ash impurities do not affect flue gas hydrate thermodynamics but enhance the formation kinetics, including CO<sub>2</sub> recovery [15]. Several reactors configurations, including fixed bed or stirred reactor, have been used to conduct laboratory-scale CO<sub>2</sub> hydrate formation and capture [4]. CO<sub>2</sub> capture from CO<sub>2</sub>-rich gas mixtures is also studied in the context of methane production from gas hydrate reservoirs [16]. Existing promoters have many disadvantages. For example, thermodynamic promoters occupy some of the cages within the hydrate cage, which leads to lower gas uptake within the structure. Many are derivative of petrochemical feedstock, thus not environmentally friendly. Kinetic promoters, e.g., surfactants, are also considered harmful to the environment and difficult to naturally degrade, as evaluated by Scott et al. [17]. SDS also form foam during the hydrate dissociation that may cause blockage and lead to additional maintenance at the industrial scale. Hence, there is a need for green, non-toxic chemicals for CO<sub>2</sub> hydrate formation with a high formation rate and high gas uptake.

Amino acids are being seen as an attractive chemical for CO<sub>2</sub> capture applications via hydrate formation because of their ability to mix with water through hydrogen bonding and because of their nontoxic nature, nonvolatility, and eco-friendliness. The absence of foam in the presence of amino acids during degassing is also an added advantage, and there would be no emissions risk from pilot plants compared to amine-based technologies. Initial research indicates that the amino acid mechanism to enhance hydrate formation is different from surfactants, and their performance in terms of flue gas hydrate promotion, CO<sub>2</sub> recovery, and hydrate stability is mostly unstudied.

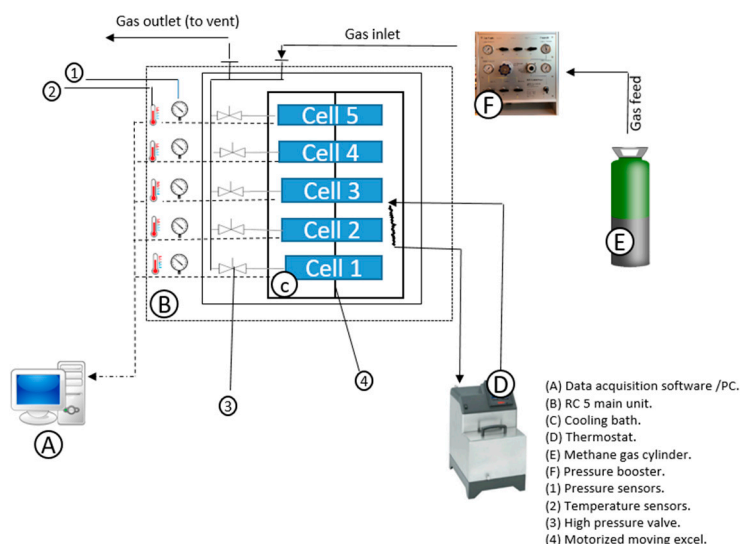
Different conclusions are reported in the literature about the role of amino acids at different capacities, such as kinetic hydrate inhibition, thermodynamic hydrate inhibitor, and kinetics hydrate promotion. However, there is no reported evidence of classifying amino acids as a thermodynamic hydrate promoter. Sa et al. [18] reported that amino acids exhibit substantial electric charges/electrostatic interactions with water as zwitterions and interact with water molecules through strong hydrogen bonding, due to their hydrophilic nature, which qualifies them as suitable inhibitors. Sa et al. [19,20] evaluated many amino acids and concluded that amino acids' thermodynamic inhibition is more visible at higher amino acid concentrations. Sa et al. [21] also categorizes a few amino acids as kinetic hydrate inhibitors and found that inhibition decreases with an increase in hydrophobicity. The inhibition effect also depends on the history of water, concentration of the additive, length, and nature of the side chain of the amino acid. Water perturbation is also expected to cause the inhibition effect [22]. Recent studies using L-valine and L-methionine [23] also show that hydrophilic certain amino acids enhance pure CO<sub>2</sub> hydrate growth kinetics. However, gas hydrate-based CO<sub>2</sub> capture using amino acids is still in the early stages of development [24]. The effect of amino acids on hydrate growth are not well understood and on the kinetics of flue gas hydrate formation is investigated, and different parameters,

such as nucleation temperature, induction time, gas uptake capacity, and CO<sub>2</sub> recovery, are calculated. Based on experimental results, amino acid performance is compared with water and SDS. This work also highlights the potential application of the rocking cell for flue gas hydrate formation and CO<sub>2</sub> capture studies. The rocking cell is traditionally used to study the hydrate inhibitors to determine onset time and onset temperature, using constant temperature and constant ramping techniques. Recently, our group has demonstrated a successful experimental analysis for methane hydrate formation kinetics in the presence of SDS with rocking cells [25].

## 2. Materials and Methods

### 2.1. Setup and Material

In this work, flue gas hydrate formation is studied using a rocking cell setup presented in Figure 1. An analytical grade of CO<sub>2</sub>/N<sub>2</sub> gas mixture (99.99% purity) with different CO<sub>2</sub> (Concentration) is obtained from Air Liquide Company. A total of 20 and 30 mole% CO<sub>2</sub> in CO<sub>2</sub>/N<sub>2</sub> mixture is used as feed gas to study the effect of higher CO<sub>2</sub> concentration. The schematic layout of a rocking cell is given in Figure 1, and details can be referred elsewhere [26]. Amino acids are supplied from Sigma Aldrich. Distilled water is used to prepare all the samples to minimize the effect of the impurities in the solution phase. All amino acids used here are of 3000 ppm concentration, while SDS concentration varied from 500 ppm to 3000 ppm. The effect of different parameters using the rocking cell-based investigation is described in detail elsewhere. Most of the studies for hydrate promoters are performed in a stirred reactor and unstirred system, such as a fixed-bed reactor, water spraying reactor, gas inducing agitated reactor, etc. The rocking cell is different from the stirred cell in the agitation method. The rocking cell is a batch system and is capable of handling pressure up to 200 MPa and temperature between −10 °C to 60 °C. Within the rocking cell setup, several rocking cells can be placed in one cooling bath, thus facilitating the parallel experiments in identical pressure & temperature conditions.



**Figure 1.** Schematic diagram of the rocking cell-5 apparatus.

According to the classification proposed by Kyte et al. [27], amino acids test into three categories: hydrophobic, polar, and charged. The structure and physical properties of four selected amino acids in the present studied are provided in Table 1.

**Table 1.** Classification of selected amino acids and their chemical formula.

#	Name	Side Chain Polarity	Side Chain	Molecular Formula	MWeight (g/mol)	Hydropathy Index Kyte et al. [27]
1	L-valine	Non-polar	-CH(CH <sub>3</sub> ) <sub>2</sub>	C <sub>5</sub> H <sub>11</sub> NO <sub>2</sub>	117.15	4.2
2	L-methionine	Non-polar	-CH <sub>2</sub> -S-(CH <sub>2</sub> ) <sub>2</sub>	C <sub>5</sub> H <sub>11</sub> NO <sub>2</sub> S	149.21	1.9
3	L-histidine	Basic polar. aromatic side chain	-CH <sub>2</sub> C <sub>3</sub> H <sub>3</sub> N <sub>2</sub>	C <sub>6</sub> H <sub>9</sub> N <sub>3</sub> O <sub>2</sub>	155.16	-3.2
4	L-arginine	Basic polar. aliphatic side chain	HN=C(NH <sub>2</sub> )-NH(-CH <sub>2</sub> ) <sub>3</sub>	C <sub>6</sub> H <sub>14</sub> N <sub>4</sub> O <sub>2</sub>	174.20	-4.5

As in the above table, L-arginine has the most negative hydropathy index (hydrophilic) among four amino acids, while L-valine has the most positive hydropathy index (hydrophobic). Hydrophobicity, acidity, and solubility of amino acid in water is strongly dependent on the side chain property. Further, hydrophobicity of the side chain and charge distribution on the amino acid molecule are considered an essential factor behind the promotion or inhibition role of amino acid during CO<sub>2</sub> hydrate formation [28]. Hydrophilic side chains in amino acids tend to form ionized or resonance structures and induce polarization of the amino acids [21]. Understanding of the structure-property relationship is also essential; for instance, L-histidine has an imidazole side chain that can be easily protonated at lightly acidic conditions (relevant for physiology, where it is often referenced), due to a pKa value of about 6 [29]. Below this pH value, one of the lone nitrogens of the imidazole ring is protonated, and the molecule becomes positively charged. Further, it has an uneven distribution of covalent bonds, making it one of the few polar amino acids, and the ring is expected to form weak hydrogen bonding. L-arginine is another such polar molecule, whereby its guanidinium group has a high pKa of 12.48, causing it to be protonated in most situations. However, because it has a relatively extended carbon backbone, which is hydrophobic, this causes it to become somewhat amphiphilic. Thus, L-histidine and L-arginine, are both expected to cause perturbation of the liquid water structure. L-valine, instead, belongs to the larger class of non-polar amino acids, and due to its bulky isopropyl side group, the molecule is fairly hydrophobic. Likewise, L-methionine shares similar properties, albeit due to its S-methyl thioether side group.

SDS shares some features of L-arginine, such as amphiphilicity, although to a much greater extent. It has a “head” comprised of a sulfur atom surrounded by four oxygens, and one of these is negatively charged, which is balanced by a nearby sodium cation. The oxygens of the molecule’s head contribute to its hydrophilic part, while the “tail” of the molecules is a long twelve-carbon backbone, which is attributed to the hydrophobic component. Due to this amphiphilicity, it is known to form micelles [30]. There have been few studies on CO<sub>2</sub> hydrate formation in the presence of SDS [31–34]. SDS-based CO<sub>2</sub> hydrate growth at the gas-water interface is capillary-driven and mass transfer-driven [32]. Table 2 shows the acidic behavior of the selected amino acids. From Table 2, it can be said that L-valine and L-methionine are neutral amino acids, while L-histidine and L-arginine are basic.

**Table 2.** Selected amino acids and their chemical properties.

Name	Water Solubility (g/L)	Acidity (pH)	Pka Alpha Carboxy	Pka Alpha Amino	Pka Side Chain	pH	Isoelectric Point
L-valine	58.5	Neutral	2.32	9.62		6.62	5.96
L-methionine	56.6	Neutral	2.28	9.21		6.76	5.74
L-histidine	45.6	Basic (weak)	1.82	9.17	6.04	7.64	7.59
L-arginine	182	Basic (strong)	2.17	9.04	12.48	10.31	11.15

## 2.2. Methods

A rocking cell with five identical pressure test cells (RC-5, PSL Systemtechnik, Osterode am Harz, Germany) is used to test the effect of SDS and amino acids on flue gas hydrate formation. Two different temperature programs, constant ramping and isothermal, are applied in this study. The constant ramping method, with initial operating pressure 120 bar and temperature 25 °C, is selected to determine

the onset of nucleation temperature ( $T_o$ ). Methodology to calculate  $T_o$  is discussed in our previous publication [25].

To calculate the induction time ( $t_o$ ), total gas uptake,  $\text{CO}_2$  recovery, isothermal tests are carried out at constant temperature schemes ( $1^\circ\text{C}$ ), and initial operating pressure (120 bar). Calculated temperature ( $T_o$ ) from constant ramping experiments shows smaller deviations compared to induction time ( $t_o$ ) measured via isothermal experiments [26]. Gas chromatography (GC) analysis is carried out at the end of the isothermal experiment to calculate the moles of  $\text{CO}_2$  and  $\text{N}_2$  remaining in the gas phase.

The constant ramping test procedure is as follows. (1) Each cell is loaded with 10 mL of solution. (2) Air inside each cell is removed via purging the cell with a 3–5 bar of flue gas. (3) The cell is pressurized with flue gas at desired pressure at  $25^\circ\text{C}$  temperature and rocked at 20 rocks per min with a rocking angle of  $35^\circ$ . (4) The cells are cooled further from  $25^\circ\text{C}$  to  $15^\circ\text{C}$  in 1 h and  $15^\circ\text{C}$  to  $1^\circ\text{C}$  in 14 h at a rate of  $1^\circ\text{C}/\text{h}$  at flow conditions. (5) The pressure and temperature of each cell and cooling bath are continuously monitored by a data acquisition (DAQ) system throughout the experiments. Table 3 below summarizes the experimental plan for the constant ramping experiment. We did not observe any hydrate formation during the constant ramping experiment for a 10%  $\text{CO}_2$ /90%  $\text{N}_2$  gas mixture. Thus, the discussion is focused around 20% and 30% concentration of  $\text{CO}_2$  in the flue gas. The typical  $\text{CO}_2$  concentration in the post-combustion gas mixture ( $\text{CO}_2/\text{N}_2$  mixture) generated from the power plant is around 15% and includes impurities like  $\text{H}_2\text{S}$  and  $\text{SO}_2$ . To study the effect of  $\text{CO}_2$  concentration on the kinetics of flue gas hydrate formation,  $\text{CO}_2/\text{N}_2$  gas mixture with different  $\text{CO}_2$  concentration is selected. In our experiments, it is noted that amino acids are effective for 20% and 30%  $\text{CO}_2$  concentration. Information received using 20% and 30%  $\text{CO}_2$  could be useful to understand the hydrate formation from pure  $\text{CO}_2$  as well as other gas mixture having higher  $\text{CO}_2$  concentrations such as fuel gas.

**Table 3.** Experimental plans for constant ramping.

Promoters	10% $\text{CO}_2$ /90% $\text{N}_2$	20% $\text{CO}_2$ /80% $\text{N}_2$	30% $\text{CO}_2$ /70% $\text{N}_2$
	Trials	Trials	Trials
Water	-	4	4
SDS	1	2	1
Amino Acid	1	2	1

The isothermal test procedure is as follows. (1) Each cell with a 10 mL sample volume is placed inside the bath, and air inside the cell is removed via purging the cell. (2) The temperature of the bath is reduced to an experimental pressure of  $1^\circ\text{C}$ . (3) Once the desired temperature is achieved, cells are pressurized with flue gas at an initial operating pressure of 120 bar. (4) Rocking is started at 20 rocks per minute, and a rotating angle of  $35^\circ$ . (5) at the end of the experiment, to calculate the  $\text{CO}_2$  and  $\text{N}_2$  mole% in the remaining gas, the ventilation valve is gently opened to collect to gas sample from each cell for GC analysis. Throughout the experiment, the pressure and temperature of each cell and cooling bath is continuously monitored by DAQ. Table 4 below summarizes the experimental plan for the isothermal experiments.

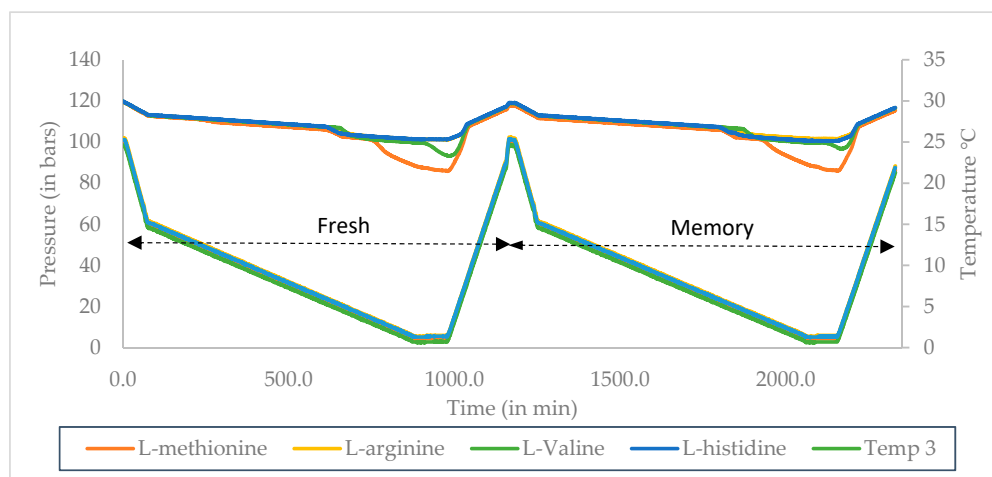
**Table 4.** Experimental plans for isothermal experiments.

		20% CO <sub>2</sub> /80% N <sub>2</sub>		30% CO <sub>2</sub> /70% N <sub>2</sub>	
	Concentration (ppm)	Fresh Runs	Fresh + Memory Run	Fresh Runs	Fresh + Memory Run
Water	0		4	4	-
SDS	500				
	1000	2	1	1	-
	2000				
	3000				
Amino Acid	L-valine				
	L-methionine	2	1	1	-
	L-histidine				
	L-arginine				

### 2.3. Experimental Data Processing

#### 2.3.1. Constant Ramping Experiments

The P-T variation during the constant ramping experiment is shown in Figure 2. The ramping test includes both fresh and memory run. In constant ramping mode, initial pressure reduction is due to thermal contraction and due to gas dissolution into the solution. At the beginning of hydrate nucleation, pressure and temperature follow linear correlation, and the onset of nucleation is observed from the point where a sudden deviation from the linear trend is found. These P-T data are plotted against time to detect the flue gas hydrate formation in the presence of different solutions.



**Figure 2.** Flue gas (30 mole% CO<sub>2</sub>) hydrate formation in the presence of amino acids: P-T variation during constant ramping scheme for selected amino acid during flue gas hydrate formation with 30% CO<sub>2</sub>.

Methodology to calculate the onset nucleation temperature,  $T_o$ , can be referred to other publications [25,26]. The rapid drop in pressure observed after  $T_o$  is the region of high growth where promoter concentration does not play a role. It is observed during the experiments that hydrate promoters increase certainty in the system during the hydrate formation process and reduce the stochastic nature of hydrate formation.

Sub-cooling temperature ( $\Delta T_{sub}$ ) is considered as the driving force behind the hydrate growth curve. It can be expressed as the difference between  $T_{eq}$  and operational temperature  $T_{op}$  [35]. The  $T_{op}$  is the temperature, primarily referred to as the temperature during the isothermal test. In Equation (1),  $T_{eq}$  is



calculated using CSMGem software for pure water and gas mixture with different molar compositions [36].

$$\Delta T_{sub} = T_{eq} - T_{op} \quad (1)$$

In a system containing hydrate promoters, an understanding of sub-cooling can be helpful in understanding the effect of concentration increase on the driving force. The maximum sub-cooling required in terms of the start of hydrate nucleation in case of hydrate promoters can be expressed as

$$\Delta T_{sub} = T_{eq} - T_o \quad (2)$$

where  $T_o$  is assumed to be  $T_{op}$  for simplicity in the calculation.

Different sub-cooling with different concentrations in the case of hydrate promoter (HP) would refer to different operational requirements. Higher sub-cooling indicates the additional condition for a more significant temperature drop below the hydrate equilibrium temperature for nucleation to start in case of flue gas hydrate with different CO<sub>2</sub> Concentration.

Further, a change in sub-cooling is calculated by computing the difference between the sub-cooling temperature of pure water and sub-cooling temperature at a different concentration of SDS

$$\Delta T_{sub,change} = \Delta T_{sub(purewater)} - \Delta T_{HP,i} \quad (3)$$

$$\Delta T_{sub,change} = (T_{eq} - T_o)_{(purewater)} - (T_{eq} - T_o)_{SDS,i} \quad (4)$$

$$T_{eq,water} = T_{eq,SDS} \quad (5)$$

$$\Delta T_{sub,change} = T_{o(SDS,i)} - T_{o(purewater)} \quad (6)$$

$$\Delta T_{sub,change} < 0 \quad (7)$$

$\Delta T_{sub,change} < 0$  for HP would indicate that the HP-based system needs a higher degree of sub-cooling in comparison to a simple water-based system while forming flue gas hydrate. According to Sloan et al. [37], the difference between  $\Delta T_{sub}$  and  $\Delta T_{op}$  provided the contrast of the operating temperature (isothermal tests) and expected onset temperature. The equation below provides a comparison between constant ramping and isothermal experiments.

$$\Delta T_{sub} - \Delta T_{op} = (T_{eq} - T_o) - (T_{eq} - T_{op}) = T_{op} - T_o \quad (8)$$

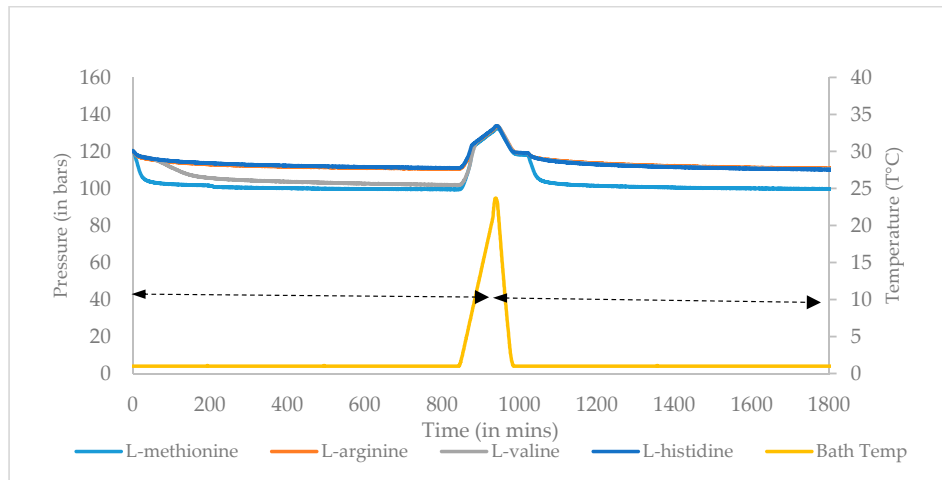
$T_{eq}$  is constant by fixing the initial operating pressure for the given gas composition and pure water case. If  $T_{op} < T_o$ , then immediate hydrate formation is expected to occur. For the case when  $T_{op} > T_o$ , a delay in hydrate formation is expected.

### 2.3.2. Isothermal Experiment

P-T variation for different amino acids during the isothermal experiment is shown in Figure 3. Under isothermal experiments, induction time, gas uptake, and CO<sub>2</sub> recovery are calculated to measure the growth phase of flue gas hydrate formation. Flue gas hydrate formation is a crystallization process involving nucleation, growth, and agglomeration stages. Mass and heat transfer are considered key factors during growth. In current studies of kinetics, we are concerned with the rate of phase transformation and factors affecting the phase transformation. While the rate of nucleation is challenging to measure the induction time, gas uptake can be experimentally measured and describes the kinetics of flue gas hydrate formation in terms of onset of crystallization and rate at which hydrate solution interface advances [13].

All kinetics experiments are carried out in hydrate formation conditions to obtain the finite rate of crystallization. During the isothermal experiments, the difference between initial and stability pressure acts as a driving force behind the hydrate formation. At the end of the isothermal experiment, the residual gas phase composition is determined by GC. P-T data recorded during the experiments are

used to measure the induction time, gas uptake, and CO<sub>2</sub>% recovery. Figure 3 shows the P-T trend during the isothermal experiment, both from fresh and memory run.



**Figure 3.** Flue gas (20 mole% CO<sub>2</sub>) hydrate formation in the presence of amino acids at 1 °C: P-T variation during isothermal experiments for selected amino acid during flue gas hydrate formation.

Total number of moles of flue gas injected into the pressure cell is calculated as

$$n_i^{fluegas} = \frac{P_1 V}{Z_1 R T} \quad (9)$$

$$n_i^{CO_2} = n_i^{fluegas} \times y_i^{CO_2} \quad n_i^{N_2} = n_i^{fluegas} \times y_i^{N_2} \quad (10)$$

$P_1$  is the initial operating pressure after flue gas is injected into the high-pressure cell. Where  $V(V = V_T - V_L)$  is the available gas volume in the reactor.  $V_T$  is the total cell volume and  $V_L$  is the HP solution volume (equal to 10 mL).  $T$  is the temperature of the isothermal test. The compressibility factor  $Z_1$  at the initial pressure and temperature for given flue gas composition is calculated using the Benedict-Webb-Rubin-Starling equation of state.  $R$  is the universal gas constant,  $8.314 \text{ J} \cdot \text{mol}^{-1} \cdot \text{K}^{-1}$ .

Assuming the process is constant volume, available gas volume remains constant after flue gas hydrate formation. Thus,  $n_f^{fluegas}$  is the number moles of flue gas hydrate formation in the gas phase after flue gas hydrate formation, and it is given by the following equation:

$$n_f^{fluegas} = \frac{P_2 V}{Z_2 R T} \quad (11)$$

where  $P_2$  is the pressure at the end of the isothermal experiment, and  $Z_2$  is the compressibility factor corresponding to  $P_2$ .  $T$  is isothermal bath temperature.  $Z_2$  is measured for a given gas vapor composition at the end of the experiment. The gas composition is measured using GC analysis. The moles of carbon dioxide and nitrogen in the equilibrium gas phase are calculated as:

$$\begin{aligned} n_f^{CO_2} &= n_f^{fluegas} \times y_f^{CO_2} \\ n_f^{N_2} &= n_f^{fluegas} \times y_f^{N_2} \end{aligned} \quad (12)$$

The change in the total number of moles of flue gas  $\Delta n_H^{fluegas}$  trapped in hydrate formation is given by

$$\Delta n_H^{fluegas} = \frac{P_1 V}{Z_1 R T} - \frac{P_2 V}{Z_2 R T} \quad (13)$$



The total moles of CO<sub>2</sub> and N<sub>2</sub> stored in the hydrate is thus calculated as given by

$$\begin{aligned} n_H^{\text{CO}_2} &= n_f^{\text{CO}_2} - n_i^{\text{CO}_2} \\ n_H^{\text{N}_2} &= n_f^{\text{N}_2} - n_i^{\text{N}_2} \end{aligned} \quad (14)$$

where  $n_H^{\text{CO}_2}$  and  $n_H^{\text{N}_2}$  are moles of CO<sub>2</sub> and N<sub>2</sub> stored in hydrates.

CO<sub>2</sub> recovery % is calculated as

$$S.Fr. = \frac{n_H^{\text{CO}_2}}{n_i^{\text{CO}_2}} \quad (15)$$

where  $n_i^{\text{CO}_2}$  and  $n_H^{\text{CO}_2}$  are the number of moles of CO<sub>2</sub> in feed and hydrate gas.

### 3. Results

Using CSMGem, pure CO<sub>2</sub> and N<sub>2</sub> systems require 14 bar and 178 bar minimum pressure, respectively, to form hydrate at 1 °C. CO<sub>2</sub> induces the formation of S-I structure, while N<sub>2</sub> induces S-II structure. The final structure of a flue gas hydrate could either be S-I or S-II, which will depend on the individual gas molecule composition in the small and large cavities. It is observed that the CO<sub>2</sub>/N<sub>2</sub> system with more than 10% CO<sub>2</sub> will be S-I in nature [38].

In this study, we have analyzed important kinetic parameters, including nucleation and growth kinetics, as well as CO<sub>2</sub> recovery in the presence of sodium dodecyl sulfate (SDS) and selected amino acids (L-valine, L-methionine, L-histidine, and L-arginine) using rocking cell. Flue gas with 20 and 30 CO<sub>2</sub> mole% is used during the studies. Experimental results are divided into different subsections, and a discussion is followed.

#### 3.1. Onset Temperature ( $T_0$ ) and Sub Cooling ( $\Delta T_{sub}$ ) for Surfactant and Amino Acid

Nucleation kinetics are studied to observe the onset of nucleation temperature ( $T_0$ ). Hydrate nucleation depends on many factors, including subcooling temperature, the composition of the gas, agitation, and geometry of pressure cells [36]. Therefore, to study nucleation in the laboratory, the number of variables need to be monitored and controlled. Due to the standardization of the rocking cell, it is possible to control most of the parameters while calculating the nucleation temperature. Measurement of nucleation temperature is preferred due to lower deviation observed; therefore, subcooling temperature measurement is reported in place of induction time to evaluate the nucleation kinetics, including promotion and inhibition of different solutions. The effect of change in CO<sub>2</sub> concentration in flue gas is studied for different SDS concentrations and different amino acids at 3000 ppm concentration. Nucleation temperature is recorded for both fresh and memory samples. The initial solution volume used is 10 mL, and the initial operating pressure is 120 bar.  $\Delta T_{sub}$  for all the solutions is calculated as  $\Delta T_{sub} = T_{eq} - T_0$ , and for simplicity,  $T_{eq}$  is calculated using CSMGem for pure water and flue gas with 20 and 30 mole% CO<sub>2</sub> at 120 bar pressure.

Results including  $T_0$  and  $\Delta T_{sub}$  are summarized in Table 5 below. Results include data from fresh and memory run for both 20 mole% and 30 mole% CO<sub>2</sub> with pure water, SDS (500–3000 ppm), and amino acids (3000 ppm). Results indicate that the subcooling requirement for a lower concentration of SDS and hydrophilic amino acids is higher than water, while the subcooling requirement of SDS (2000, 3000 ppm) at higher concentration and hydrophobic amino acids is found to be lower than water. A higher subcooling requirement indicates a delay in nucleation with respect to water. Higher subcooling requirements for individual amino acids reflect the perturbation of liquid water structure correlated with hydrophobicity [39]. L-histidine was found to be more effective in inhibition due to the presence of charge at the side chain. On the other hand, L-valine and L-methionine were able to improve the water structure, thus reducing the subcooling requirement. An increase in CO<sub>2</sub> mole% leads to an increase in nucleation temperature recorded for all solutions, including water;

however, it was observed that at higher CO<sub>2</sub> mole%, the subcooling requirement for all solutions was higher compared to 20 mole% CO<sub>2</sub>. Overall, variation in  $T_o$  trend can be correlated with their hydrophathy index.

**Table 5.** Effect of change in CO<sub>2</sub> concentration on onset temperature ( $T_o$ ) during fresh (Fr) and memory (Me) run for 10 mL solution at initial pressure 120 bar in constant ramping temperature scheme.

Onset Nucleation Temperature and Sucooling ( $\Delta T_o$ ) in °C									
		20% CO <sub>2</sub> /80% N <sub>2</sub>				30% CO <sub>2</sub> /70% N <sub>2</sub>			
		Fr	Me	Fr	Me	Fr	Me	Fr	Me
$T_{eq}$		6.34							
	Concentration (ppm)			$\Delta T_{sub,fr}$	$\Delta T_{sub,me}$			$\Delta T_{sub,fr}$	$\Delta T_{sub,me}$
water	0	4.10	4.10	2.24	2.24	5.80	5.70	2.52	2.62
SDS	500	3.85	3.65	2.49	2.69	4.90	4.80	3.42	3.52
	1000	4.05	3.80	2.29	2.54	5.90	5.50	2.42	2.82
	2000	5.80	5.20	0.54	1.14	6.00	5.70	2.32	2.62
	3000	4.25	3.80	2.09	2.54	5.00	4.90	3.32	3.42
Amino Acid	L-valine	5.40	5.30	0.94	1.04	5.80	5.70	2.52	2.62
	L-methionine	4.60	4.65	1.74	1.69	5.70	5.40	2.62	2.92
	L-histidine	3.90	4.05	2.44	2.29	5.90	5.80	2.42	2.52
	L-arginine	4.25	4.25	2.09	2.09	5.80	5.70	2.52	2.62

During the memory run, different concentrations of SDS in solution are shown to eliminate the memory effect, reflected by lower  $T_o$  values observed during the memory run at different CO<sub>2</sub> concentrations. SDS has shown similar behavior during the methane hydrate formation in the presence of SDS [25]. It is likely that a surfactant-based hydrate promoter eliminates memory effect during the nucleation, independent of the host molecule.

$T_o$  values of hydrophobic amino acids are distinctively different from hydrophilic amino acids at 20% mole CO<sub>2</sub>. An increase in CO<sub>2</sub> concentration to 30 mole% in flue gas leads to nearly a similar  $T_o$  values for all four amino acids as CO<sub>2</sub>% increases,  $T_o$  increases for all solutions. This is because the CO<sub>2</sub> concentration increase leads to a higher driving force due to  $T_{eq}$  moving from 6.3 to 8.32 °C; thus, nucleation initiates at a higher temperature. For SDS, the nucleation temperature is lower and close to water nucleation temperature between 500 ppm to 1000 ppm concentrations. This reflects that at a lower concentration, SDS delays the nucleation. Further, as SDS concentration increases from 500 ppm to 2000 ppm, nucleation temperature increases until 2000 ppm and decreases after 2000 ppm. A similar trend is observed for methane hydrate formation in Pandey et al. [25]. Change in  $T_o$  between 2000 to 3000 ppm reflects the presence of the critical micelle concentration (CMC) of SDS, which is also supported by Roosta et al. [31]. Based on the observations in this paper and a previous study by Pandey et al. [25] for methane hydrate formation in the presence of SDS, it is likely that the CMC of hydrate-promoting surfactant is independent of the guest molecule, while the dependent on surfactant type.

In the case of amino acids, for flue gas with 20 and 30% mole CO<sub>2</sub>,  $T_o$ , for all amino acids is found to be higher or closer to the pure water case, except in L-histidine case, where  $T_o$  is recorded as 3.9 °C and 5.9 °C. This indicates L-histidine functions as an inhibitor, thus delaying the nucleation compared to water. This could be because of a charge present on the side chain in L-histidine that leads to perturbation of water structure around the side chain and reduces the hydrogen bonding capability of the water molecule. L-valine and L-methionine, which are hydrophobic amino acids, showed a higher  $T_o$  compared to the hydrophilic amino acids. As CO<sub>2</sub> concentration increases in flue gas from 20% to 30%, the memory effect becomes weaker in the case of amino acids, and the deviation between

$T_o$  values in fresh and memory cases reduces. It can, therefore, infer that the concentration of  $\text{CO}_2$  in flue gas plays an important role behind hydrate nucleation temperatures.

Figure 4 shows a graphical representation of the subcooling temperature for different combinations. The surfactant and amino acids are seen to eliminate the memory effect as subcooling requirements in memory runs are higher than in the fresh run. SDS at lower concentrations (500 and 1000 ppm) behaves as an inhibitor compares to pure water case, as suggested by higher subcooling requirements. At higher concentration, (2000 and 3000 ppm) subcooling requirement is lower than water. It can be concluded that SDS show promotion effect at higher concentration during flue gas hydrate formation. An increase in  $\text{CO}_2$  concentration in flue gas leads to higher nucleation temperature as well as higher subcooling requirements, as shown in Figure 4b. It is likely that an increase in  $\text{CO}_2$  concentration in flue gas would play a role in inhibition. Figure 4c discusses the subcooling requirement of amino acids. Hydrophobic amino acids have lower subcooling requirements compare to hydrophilic and pure water; therefore, hydrophobic amino acids are considered as a promoter during flue gas hydrate formation. Hydrophobic and hydrophilic amino acids do not show a strong memory effect, as suggested by higher subcooling value in memory run in Figure 4c. An increase in  $\text{CO}_2$  concentration leads to an increase in subcooling requirement for all amino acids, as shown in Figure 4d.



**Figure 4.** Subcooling kinetics in presence of amino acids and surfactant. Subcooling temperatures at the onset of the flue gas hydrate formation with amino acids and surfactant. (a) Subcooling temperature for 20% mole  $\text{CO}_2$  for sodium dodecyl sulfate (SDS) at different concentration (fresh and memory). (b) Subcooling temperature for 20 and 30 mole%  $\text{CO}_2$  for SDS at different concentration (fresh). (c) Subcooling temperature for 20% mole  $\text{CO}_2$  for amino acids at different concentration (fresh and memory). (d) Subcooling temperature for 20 and 30 mole%  $\text{CO}_2$  for amino acids at different concentration (fresh).

### 3.2. Induction Time Measurement for Surfactant and Amino Acids

Due to low formation rates and excessive energy consumption, hydrate-based CO<sub>2</sub> capture, separation, and storage are uneconomical. Understanding of induction times and growth rates during the flue gas hydrate formation plays an essential factor in the selection of appropriate promoters. The measurement of induction time is summarized in Table 6. At 120 bar and 274.15 K for the pure water case, no rapid hydrate formation was observed both in fresh and memory runs using 20 and 30 mole% CO<sub>2</sub>. Memory refers to the case when the sample used in the experiment has previously experienced the hydrate formation. It is the general observation that memory sample has lower induction time compare to fresh sample [40]. It is observed that SDS and amino acids do not eliminate memory effects during the growth phase, as shown by lower induction time recorded in memory runs.

Induction time is also influenced by the driving force ( $P_{inj}-P_{eq}$ ), such that, as driving force increases, induction time decreases. The driving force can be influenced by incoming CO<sub>2</sub>/N<sub>2</sub> feed gas pressure ( $P_{inj}$ ) or hydrate equilibrium pressure ( $P_{eq}$ ) dependent on CO<sub>2</sub> concentration in the CO<sub>2</sub>/N<sub>2</sub> feed gas. As CO<sub>2</sub> concentration increase from 20% to 30%, the driving force increase due to the lowering of  $P_{eq}$ , which decreases the induction time. Thus, promotion increases both amino acids and SDS solutions.

For the SDS solution, it is observed that as the concentration increases from 500 ppm to 3000 ppm, induction time decreases. In our previous publication using the same setup and same experiment parameters (rocking rate, rocking angle, and solution volume), a similar trend is observed for methane hydrate formation in the presence of SDS [25]. Thus, it is likely that liquid phase properties in the presence of promoters are essential in understanding the induction time variation.

These experimental results confirm that the performance of amino acids correlated with their hydrophobicity. Hydrophilic acids L-histidine and L-arginine tend to delay hydrate formation, and no rapid formation is observed. This is also confirmed by visual inspection of L-histidine and L-arginine based hydrate. This is likely because of the perturbation of liquid water structure caused by the charged side branch of hydrophilic amino acids, which reduced the hydrogen bonding capacity of water molecules. This is also confirmed by previous studies by Sa et al. [21].

Hydrophobic amino acids have a lower induction time compared with water and hydrophilic amino acids. Memory run in case of 20 mole% CO<sub>2</sub>, suggests that SDS and hydrophobic amino acids, display strong memory effect reflected by lower induction time during memory run. When CO<sub>2</sub> mol% increase from 20% to 30%, it is observed that the rise in CO<sub>2</sub> mole% leads to a decrease in induction time, which suggests promotion capability improves both for SDS and Amino acids. However, no induction time is observed for hydrophilic amino acids. Highlight the induction time in the presence of SDS. Linga et al. [13] had reported the induction time for 16.9% CO<sub>2</sub> mixture at 0.6 °C, and 11.0 MPa, 10 MPa as 14/19 min (in two trials) and 10.3/16.3 min (in two trials), respectively in pure water system and concluded that decrease in pressure reduces the induction time for CO<sub>2</sub>/N<sub>2</sub> system. This is due to the fact that higher pressure, more N<sub>2</sub> go into the hydrate reflected by an increase in N<sub>2</sub> uptake in the hydrate. N<sub>2</sub> has higher hydrate formation pressure, thus influencing the induction time of the CO<sub>2</sub>/N<sub>2</sub> system. We did not observe any sharp growth phase of the CO<sub>2</sub>/N<sub>2</sub> system at 12 MPa using rocking cells. Thus, it can be said that there is a range of pressures at a given temperature (273 K–275 K) when flue gas hydrate can be formed in pure water cases. Their research also shows memory effect reflected by lower induction time recorded by memory run. Partoon et al. [41] has discussed the pure CO<sub>2</sub> hydrate formation in the presence of SDS and concluded that an increase in concentration decreases the induction time from 100 ppm to 1500 ppm at 35 bar and 273.65 K. The decrease in induction time in the presence of SDS is due to an increase in CO<sub>2</sub> solubility in water and lowering the surface tension [42].

**Table 6.** Effect of CO<sub>2</sub>% on induction time measurement. Induction time is measured for SDS and amino acids during fresh and memory run. Initial pressure and temperature conditions are 120 bar and 1 °C.

		Induction Time (t <sub>0</sub> ) (in mins)			
		20% CO <sub>2</sub> /80% N <sub>2</sub>		30% CO <sub>2</sub> /70% N <sub>2</sub>	
		Fresh	Memory	Fresh	Memory
Concentration (ppm)					
Water	0	No hydrae	No hydrate	No hydrate	-
SDS	500	No hydrae	No hydrate	No hydrate	-
	1000	72.5	70	27	-
	2000	14	4.5	13.5	-
	3000	4	1	2	-
Amino Acid	L-valine	44	30	35	-
	L-methionine	4.5	2	2	-
	L-histidine	No hydrae	35	No hydrate	-
	L-arginine	No hydrae	35.5	No hydrate	-

Hydrophilic amino acids show a decrease in induction time due to an increase in driving force caused by lower  $P_{eq}$  at higher CO<sub>2</sub> mole%. For SDS, a significant decline in induction time is noted between 1000 ppm and 2000 ppm for both 20 and 30 mole% CO<sub>2</sub>. In the case of 20 and 30% CO<sub>2</sub>, induction time for L-methionine and SDS at 3000 ppm are comparable.

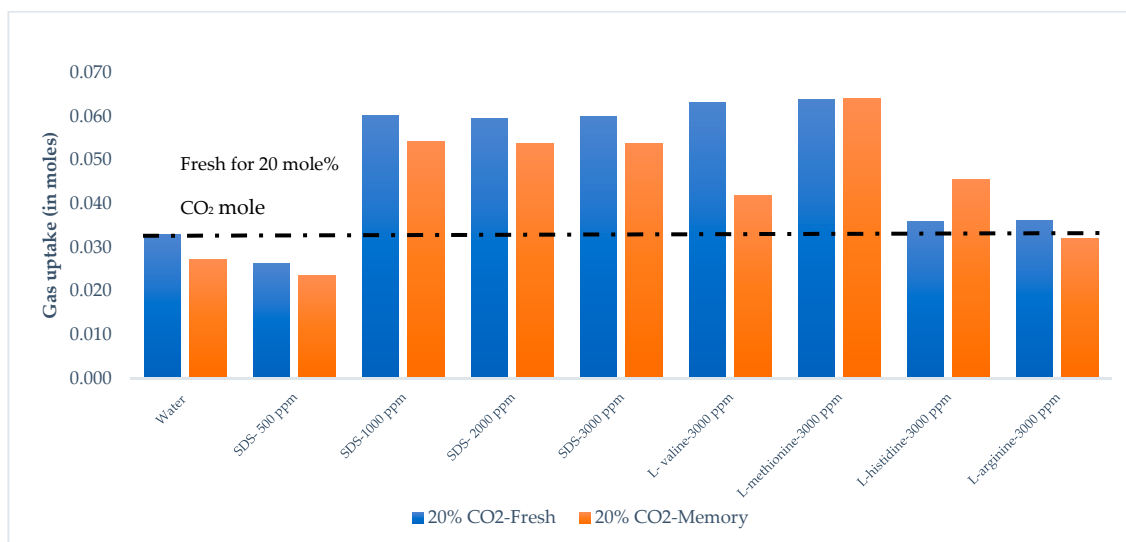
### 3.3. Gas Uptake and Growth Profile

Normalized gas uptake calculations are based on isothermal experiments at  $T = 1$  °C and  $P_{inj} = 120$  bar. Gas samples are collected for GC analysis at the end of fresh runs and combined fresh, and memory runs to study the gas composition of unused gas. Table 7 below summarizes the gas uptake calculations for fresh and memory runs at 20 mole% concentration CO<sub>2</sub> in CO<sub>2</sub>/N<sub>2</sub> mixture. Figure 5 graphically represents the results. For SDS, as the concentration increases, gas uptake increases; however, at 1000 ppm and above, an increase in gas uptake is marginal. Gas uptake for memory runs is lower than for fresh runs at all concentrations.

**Table 7.** Gas uptake calculations for 20% CO<sub>2</sub> in CO<sub>2</sub>/N<sub>2</sub> mixture for both fresh and memory runs. Starting conditions are  $P_i = 120$  bar and  $T_{op} = 1$  °C. Gas chromatography (GC) analysis is carried out after fresh and memory run to measure the moles of CO<sub>2</sub> and N<sub>2</sub> left in the gas phase after hydrate formed.  $P_{eq}$  for 20% CO<sub>2</sub> at 1 °C is calculated as 56.85 bar for pure water using CSMGem.

		20 % CO <sub>2</sub> /80 % N <sub>2</sub>						
		Fresh Run				Memory Run		
		$n_{Fluegas}$	$P_f$ (bar)	$P_i - P_f$ (bar)	$\Delta n_{fluegas}$ (mole/mole)	$P_f$ (bar)	$P_i - P_f$ (bar)	$\Delta n_{fluegas}$ (mole/mole)
$P_{inj}$	120							
$P_{eq}$ (20% CO <sub>2</sub> )	56.85							
$P_i - P_{eq}$	63.16							
Concentration (ppm)								
			110.31	9.69	0.03	110.47	9.53	0.03
Water	0	0.18	110.18	9.82	0.03	109.95	10.05	0.02
SDS	500	0.18	98.29	21.71	0.06	98.32	21.68	0.05
	1000	0.18	98.78	21.22	0.06	98.59	21.41	0.05
	2000	0.18	98.71	21.29	0.06	98.81	21.19	0.05
	3000	0.18	104.15	15.85	0.07	110.16	9.84	0.04
Amino acid	L-valine	0.18	102.94	17.06	0.07	99.75	20.25	0.06
	L-methionine	0.18	108.39	11.61	0.04	110.12	9.88	0.05
	L-histidine	0.18	109.06	10.94	0.03	110.63	9.37	0.03
	L-arginine	0.18	110.31	9.69	0.03	110.47	9.53	0.03

Gas uptake kinetics are controlled by heat and mass transfer; however, parameters affecting nucleation kinetics also play a role. It is clear from Figure 5 and Table 7 that surfactant and amino acids eliminate the memory effect for gas uptake calculations reflected by lower gas uptake value during the memory run.



**Figure 5.** Gas uptake (Fresh and Memory)-20% CO<sub>2</sub>: Gas uptake analysis for 20% CO<sub>2</sub> and effect of surfactant and different amino acids on the gas uptake for flue gas hydrate formation (20% CO<sub>2</sub>) measured during an isothermal experiment conducted at 120 bar and 1 °C recorded both for fresh and memory samples.

The gas uptake behavior of the surfactant for fresh runs shows that as the SDS concentration increases, gas uptake increases and then remains constant, not varying significantly. From the experiments, it is



also evident that gas uptake of SDS at higher concentrations is similar to hydrophobic amino acids. Gas uptakes for SDS at higher concentration and hydrophobic amino acids show distinctively higher value compare to water and hydrophilic amino acids. However, gas uptake drastically decreases for L-valine compared to L-methionine during memory runs. This is subject to further investigation.

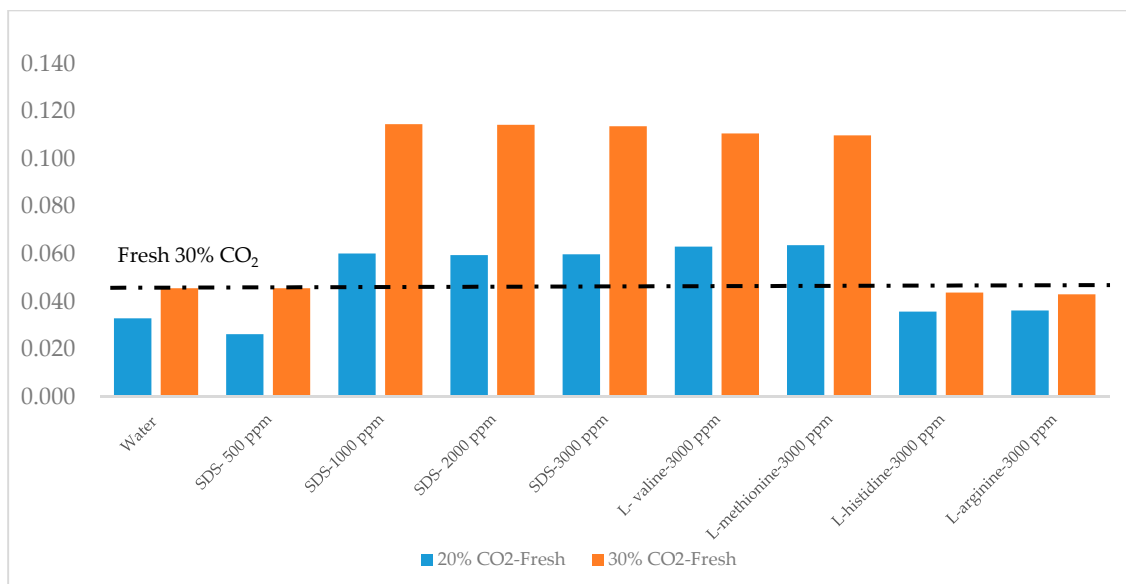
There is similarly found in nucleation temperature and growth kinetics, reflected in the case of amino acids. It suggests that hydrophilic amino acids with charged side chains inhibit CO<sub>2</sub> nucleation and growth kinetics due to the perturbation of liquid water molecules. It is known that kinetic hydrate inhibitors display different trends for nucleation kinetics and growth kinetics due to different surface phenomena [43]. However, in our research, it shows that the distinction between hydrophobic and hydrophilic behavior remains the same during nucleation and growth phases, thus confirm that hydrate formation capability of amino acids is independent of surface properties of the hydrate and connected with perturbation of the amino acids. Sa et al. have supported this conclusion previously in the case of hydrophilic amino acids with charged side chains [21].

Table 8 summarizes the gas uptake calculation for 20 and 30% CO<sub>2</sub> in flue gas for a fresh run. Figure 6 represents the change in gas uptake due to the change in CO<sub>2</sub> concentration in the CO<sub>2</sub>/N<sub>2</sub> mixture. From the results, it is clear that the increase in CO<sub>2</sub> concentration increases the gas uptake for all the solutions, including surfactant and amino acids. Final gas uptake is almost identical both for 20 and 30% CO<sub>2</sub> flue gas for L-histidine and L-arginine, which confirms the inhibition effect of hydrophilic amino acids and their impact on gas uptake. The performance of hydrophobic amino acids remains similar to SDS at a higher percentage, irrespective of CO<sub>2</sub> concentration in the flue gas. SDS under a large driving force and with a concentration between 1000 ppm to 3000 ppm is not able to show improvement in gas uptake. This could be due to the high solubility of SDS at large driving force ( $P_{inj}-P_{eq}$ ) and gas hydrate growth towards bulk liquid interfering with the hydrate growth by the capillary mechanisms [32].

**Table 8.** Gas uptake calculations for 20% CO<sub>2</sub>/80% N<sub>2</sub> and 30% CO<sub>2</sub>/70% N<sub>2</sub> for fresh runs. Starting conditions are  $P_i = 120$  bar and  $T_{op} = 1$  °C. GC analysis is carried out after a fresh run to measure the moles of CO<sub>2</sub> and N<sub>2</sub> left in the gas phase after hydrate formed.  $P_{eq}$  for 20% CO<sub>2</sub>/80% N<sub>2</sub> at 1 °C is calculated as 56.85 bar for pure water using CSMGem.

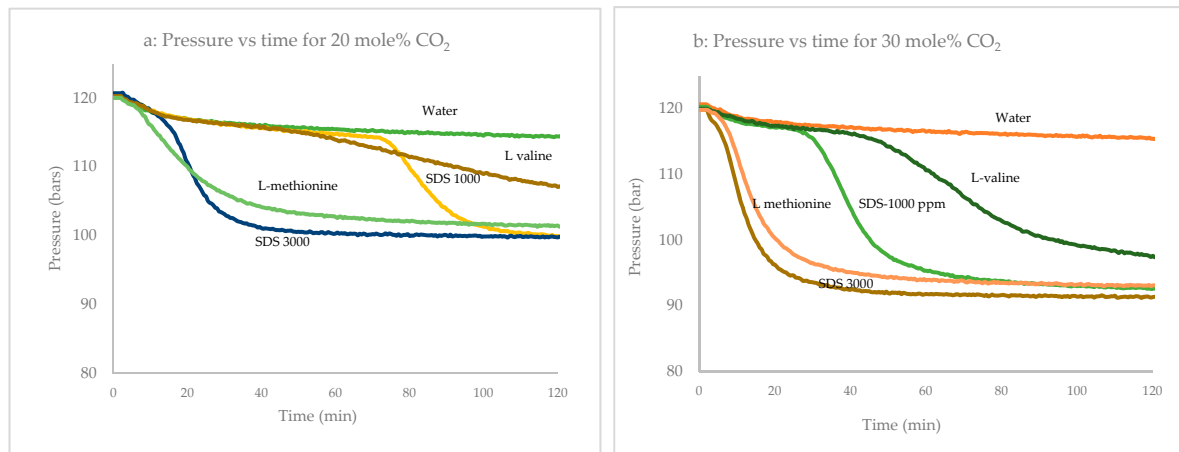
		20% CO <sub>2</sub> /80% N <sub>2</sub> (Fresh)				30% CO <sub>2</sub> /70% N <sub>2</sub> (Fresh)		
		$n_{Fluegas}$	$P_f$ (bar)	$P_i-P_f$ (bar)	$\Delta n_{fluegas}$ (mole/mole)	$P_f$ (bar)	$P_i-P_f$ (bar)	$\Delta n_{fluegas}$ (mole/mole)
$P_{inj}$ (bar)	120							
$P_{eq}$ (20% CO <sub>2</sub> )	56.84							
$P_{eq}$ (30% CO <sub>2</sub> )	41.45							
$P_i-P_{eq}$ (20% CO <sub>2</sub> )	63.16							
$P_i-P_{eq}$ (30% CO <sub>2</sub> )	78.55							
Concentration (ppm)								
Water	0	0.18	110.31	9.69	0.03	110.50	9.50	0.05
SDS	500	0.19	110.18	9.82	0.03	110.78	9.22	0.05
	1000	0.19	98.29	21.71	0.06	90.40	29.60	0.11
	2000	0.19	98.78	21.22	0.06	90.68	29.32	0.11
	3000	0.19	98.71	21.29	0.06	90.31	29.69	0.11
Amino Acid	L-valine	0.18	104.15	15.85	0.07	92.52	27.48	0.11
	L-methionine	0.18	102.94	17.06	0.07	91.74	28.26	0.11
	L-histidine	0.18	108.39	11.61	0.04	111.96	8.04	0.04
	L-arginine	0.18	109.06	10.94	0.03	110.68	9.32	0.04

Increasing the CO<sub>2</sub> concentration from 20% to 30% does not improve the gas uptake values of hydrophilic amino acids, and, for 30 mole% CO<sub>2</sub>, normalized gas uptake remains below or equal to the pure water case.



**Figure 6.** Gas uptake for fresh run (20% CO<sub>2</sub> and 30% CO<sub>2</sub>). Gas uptake during fresh runs for 20 and 30 mole% CO<sub>2</sub> with SDS and amino acid.

To understand the growth profile difference between the surfactant and hydrophobic amino acids, pressure variation is plotted against time for SDS and amino acids for 120 min, as shown in Figure 7. Hydrate kinetics after nucleation is a function of heat and mass transfer. The large contact area between liquid and gas and quick heat dissipation is necessary to accelerate hydrate growth. At this stage, extensive hydrate formation and crystal agglomeration at the gas/water interface occurs. We have seen the rapid formation in the case of surfactant and hydrophobic amino acids but did not observe a similar rapid formation in the case of hydrophilic amino acids. Figure 7 confirms that L-methionine shows pressure variation that is very similar to SDS at 3000 ppm concentration. As the CO<sub>2</sub> concentration increases, deviation in growth profile decreases. It is also visible through Figure 7 that within 120 min of the start of the experiment, the pressure is stabilized, and hydrate formation is saturated. L-valine has a lower hydrate formation rate compared to L-methionine, due to higher hydrophobicity of uncharged side chains in L-valine. Roosta et al. [28] have suggested that the CO<sub>2</sub> hydrate formation rate decreases as the hydrophobicity of the uncharged side chain increases.



**Figure 7.** Pressure vs. time variation surfactant and amino acid at isothermal temp 1 °C. (a) Pressure vs. time curve surfactant (SDS 3000 ppm and SDS 1000 ppm) and amino acid (L-valine and L-methionine) for 20 mole% CO<sub>2</sub> at pressure 120 bar and 1 °C. (b) Pressure vs. time curve surfactant (SDS 3000 ppm and SDS 1000 ppm) and amino acid (L-valine and L-methionine) for 30 mole% CO<sub>2</sub> at pressure 120 bar and 1 °C.

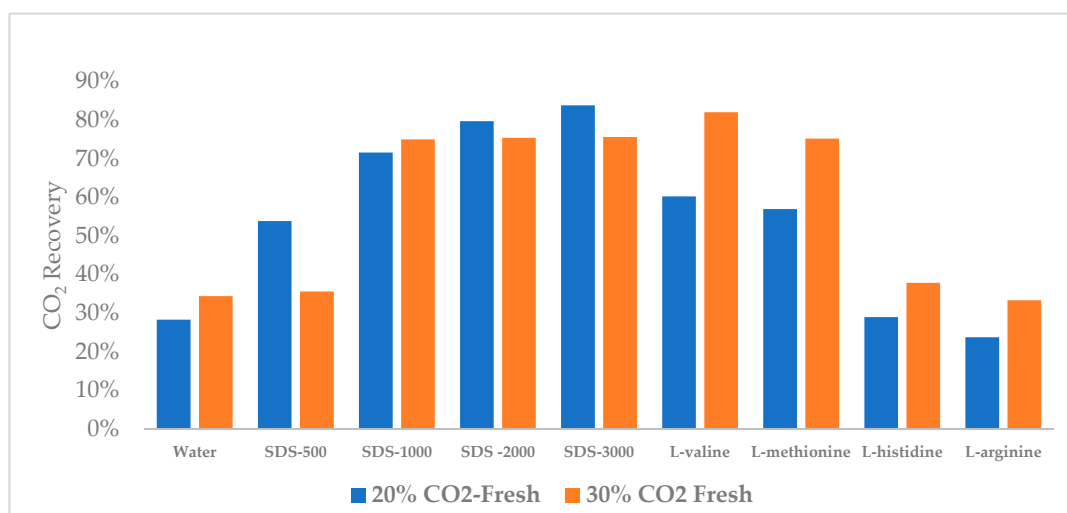
### 3.4. CO<sub>2</sub> Recovery

CO<sub>2</sub> recovery from flue gas in the form of gas hydrate can be quantified in terms of CO<sub>2</sub> recovery, which is calculated and presented in Table 9 and Figure 8. Gas uptake and CO<sub>2</sub> recovery rates are correlated, which suggests that amino acids assist the CO<sub>2</sub> molecule in occupying the gas hydrate cages. It is also visible from the CO<sub>2</sub> recovery that an increase in gas uptake results from an increase in CO<sub>2</sub> recovery captured via hydrate formation. It is evident from the table that the increase in CO<sub>2</sub> concentration leads to a decrease in CO<sub>2</sub> recovery for SDS. This could be due to no effect of SDS on CO<sub>2</sub> hydrate formation at the large driving force present due to high SDS concentration as well as higher CO<sub>2</sub> concentration in flue gas [32]. On the contrary, an increase in CO<sub>2</sub> concentration in CO<sub>2</sub>/N<sub>2</sub> mixture lead to an increase in CO<sub>2</sub> recovery hydrophobic and hydrophilic amino acids. Thus referring to the advantage of using amino acids over surfactant SDS for higher CO<sub>2</sub> concentration gas stream such as CO<sub>2</sub>/H<sub>2</sub>.

By comparing the gas uptake and CO<sub>2</sub> recovery analysis, due to the increase in CO<sub>2</sub> concentration in flue gas, it is likely that an increase in gas uptake is contributed by the increase in CO<sub>2</sub> recovery; however, overall CO<sub>2</sub> recovery % reduces. Kumar et al. [15] have also suggested in the presence of surfactant, CO<sub>2</sub> recovery, and separation not necessarily improve. In another study, Kumar et al. [7] have suggested that an increase in driving force leads to a decrease in CO<sub>2</sub> recovery.

**Table 9.** CO<sub>2</sub> recovery rate calculated in the presence of different solution including SDS and selected amino acids.

CO <sub>2</sub> 20%/80% N <sub>2</sub> (Fresh)					CO <sub>2</sub> 30%/70% N <sub>2</sub> (Fresh)		
	Concentration (ppm)	Moles of CO <sub>2</sub> Injected	Mole of CO <sub>2</sub> Remain in Gas	CO <sub>2</sub> Recovery	Moles of CO <sub>2</sub> Injected	A Mole of CO <sub>2</sub> Remain in Gas	CO <sub>2</sub> Recovery
Water	0	0.04	0.03	28%	0.06	0.04	34%
SDS	500	0.04	0.02	54%	0.06	0.04	36%
	1000	0.03	0.01	72%	0.06	0.01	75%
	2000	0.03	0.01	80%	0.06	0.01	75%
	3000	0.03	0.01	84%	0.06	0.01	76%
Amino Acid	L-valine	0.04	0.01	60%	0.06	0.01	82%
	L-methionine	0.04	0.02	57%	0.06	0.01	75%
	L-histidine	0.04	0.03	29%	0.06	0.04	38%
	L-arginine	0.04	0.03	24%	0.06	0.04	33%



**Figure 8.** Effect on CO<sub>2</sub> recovery % due to the change in CO<sub>2</sub> concentration in the Flue gas. Change in CO<sub>2</sub> recovery % for surfactant and selected amino acids due to change in CO<sub>2</sub> mole % in the feed gas, measured during an isothermal experiment conducted at 120 bar, 1 °C.

### 3.5. Review of the Role of Amino Acids during CO<sub>2</sub> Hydrate Formation

In this study, amino acids have been selected based on the difference in their structure, polarity, and the difference in hydropathy indices. A literature review concerning the role of selected amino acids in CO<sub>2</sub> hydrate formation is summarized in Table 10. To understand the role of an amino acid as a hydrate promoter or inhibitor, an important consideration is to understand the underlying mechanisms. It is proposed that amino acids do not interact with hydrate surface, reflected by similar nucleation and growth kinetics during flue gas hydrate formation in this study. The role of amino acids during the hydrate formation is explained by their ability to perturb liquid water structure. Sa et al. [21] have observed the perturbation of water structure in the presence of amino acids experimentally, and they studied the correlation of perturbation with CO<sub>2</sub> hydrate nucleation temperature and gas uptake. With the help of Raman spectroscopic experiments, they have confirmed that strongly hydrophobic amino acids tend to strengthen the water structure while hydrophilic amino acids tend to cause perturbations in the water structure, weakening the hydrogen bonds of water molecules. This observation is also supported by another study where the extent of perturbation is found to be strongly correlated with hydrophobicity [39]. It has been previously found that water structures around charged side chain become less ice-like [44], and the hydrogen-bonding network around hydrophilic amino acids is disrupted, while around hydrophobic alkyl chains, it is strengthened.

Our experimental results also suggest that L-methionine is a more effective hydrate promoter and shows rapid hydrate growth compared to L-valine during the flue gas hydrate formation. The difference between L-valine and L-methionine is the hydrophobicity based on the hydropathy index. Both amino acids are nonpolar in nature; however, L-valine has higher hydrophobicity compare to L-methionine. Roosta et al. [28] have demonstrated that the inhibition effect of uncharged amino acid tends to increase with the increasing length of side chains; therefore, L-methionine is expected to be better CO<sub>2</sub> hydrate promoter compared to L-valine due to lower hydrophobicity of the side chain. Cai et al. [45] have tested CO<sub>2</sub> hydrate kinetics in the presence of L-methionine and found CO<sub>2</sub> hydrate formation kinetics and hydrate storage capacity to be dramatically improved. The effect of different concentrations on the formation rate and storage capacity was studied, and it was found that 0.2 wt% of L-methionine gave the optimal performance. A key driver behind L-methionine promotion is due to optimal hydrophobic chain length and higher synergy between hydrophobic and carboxyl groups.

**Table 10.** Literature review of selected amino acids for CO<sub>2</sub> hydrate formation.

#	Amino Acid	Gas Tested	Test Type	Concentration	Pressure & Temperature	Reference	Remarks
1.1	L-valine	CO <sub>2</sub>	THI	0.1–0.5 mole %	273.05–281.45 K, 14.1–35.2 bar	[18]	Phase equilibrium conditions reported, The shift in CO <sub>2</sub> hydrate phase equilibrium to higher pressure and lower temperature is recorded L-valine solubility equal to 0.884 at 298.15 K and 0.763 at 273.15 K. Amino acid solubility in water affects the inhibition capability. THI function by lowering the water activity by associating with water molecules via hydrogen bonding. Increase in inhibiting effect increase with the increase in the size of the alkyl chain
1.2	L-valine	CO <sub>2</sub>	KHI	0.1 mole %	36 bar, 284.05 K Isothermal 273.45 K	[19]	Inhibition is two types, absorption-based and perturbation inhibition Most of the amino acids are zwitterions during CO <sub>2</sub> and interact with water due to strong electrostatic interaction generated by an electric charge on these molecules. The water structure around the electric charge becomes less ice-like. Less significant inhibition impacts due to the longer hydrophobic chain, thus increase hydrate promotion.
1.3	L-valine	CO <sub>2</sub>	KHP	0.5 wt%	33 bar, 298 K	[23]	The ability to inhibit the hydrate crystals decrease at higher mole fraction and at higher hydrophathy index, reduction of THI capability observed. At higher concentrations, it is found that amino acids tend to crystallize. L-valine show 20% higher gas uptake than pure CO <sub>2</sub> system under stirred conditions.
2.1	L-methionine	CO <sub>2</sub>	KHP + XRD	0.5 wt%	53 bar, 275 K	[23,47]	Gas uptake 20% more than pure CO <sub>2</sub> system. 90% of gas consumption in less than 45 min L-valine gas consumption takes double the time of the methionine system. L-methionine show highest gas uptake Near identical gas uptake in the stirred and unstirred system for L-methionine
2.2	L-methionine	CO <sub>2</sub>	KHP	0.02–1 wt%	33 bar, 273.2 K	[45]	The presence of salt decreases the hydrate formation capability. Impurities in CO <sub>2</sub> could have an impact on the CO <sub>2</sub> hydrate formation. Significantly promote hydrate formation uptake without stirring Induction time equal to 15 min for 90% if gas consumption The amphiphilic molecule acts as a dispersant to prevent hydrate particle agglomeration, and the overall formation mechanism consists of hydrate film development at gas-liquid interface, nucleation at gas-liquid interface, growth of hydrate crystal as porous structure, and capillary action to promote growth.
3.1	L-histidine	CO <sub>2</sub>	KHI+ Polarised Raman Spectroscopy	0.1 mol %	35 bar, 283.15 K	[21]	Experimental studies were carried out to study the perturbation of water in the presence of water. L-histidine display high hydrate inhibition effect caused by the perturbation of water structure. Perturbation is directly correlated with hydrophobicity. L-histidine show higher inhibition effect than glycine. The inhibition effect of the charged side chain is more than an uncharged side chain with a polar or non-polar group. Inhibition could be due to oxygen atom on carbonyl group forming a hydrogen bond with water or charged

Table 10. Cont.

#	Amino Acid	Gas Tested	Test Type	Concentration	Pressure & Temperature	Reference	Remarks
3.1	L-histidine	CO <sub>2</sub>	KHI+ Polarised Raman Spectroscopy	0.1 mol %	35 bar, 283.15 K	[21]	Experimental studies were carried out to study the perturbation of water in the presence of water. L-histidine display high hydrate inhibition effect caused by the perturbation of water structure. Perturbation is directly correlated with hydrophobicity.
3.2	L-histidine	CO <sub>2</sub>	KHI + Spectroscopy	0.5–2 wt%	30 bar, 275.15 K	[28]	L-histidine show higher inhibition effect than glycine. The inhibition effect of the charged side chain is more than an uncharged side chain with a polar or non-polar group. Inhibition could be due to oxygen atom on carbonyl group forming a hydrogen bond with water or charged side chain forming van der Waal interaction/electrostatic with crystal surface and therefore interrupting nucleation/disrupt growth. The chemical affinity model is used to quantify the kinetic parameter. L-histidine based CO <sub>2</sub> hydrate formation rate decrease with increasing concentration. Concentration range dependent on the amino acids. CO <sub>2</sub> hydrate formation rate decrease with increasing hydrophobicity of un charge side chain in amino acids.
4.1	L-arginine	CO <sub>2</sub>	THI	10 wt%	25.3–40 bar, 278.32–281.57 K	[46]	COSMO RS is used; the heat of dissociation is calculated to be 60.28 KJ/mole. Inhibit CO <sub>2</sub> hydrate formation, and inhibition effect is independent of the guest molecule and due to intermolecular forces. A higher alkyl chain in L-arginine leads to lower inhibition effect compare to glycine.

L-histidine has shown an inhibition effect during flue gas hydrate formation. The inhibition effect can be explained due to the presence of a high net charge on L-histidine (1.0005 at pH = 3.77, 2851.5 K, and 30 bar) [28]. Due to this net charge on the side chain, L-histidine interacts strongly with water molecules at the hydrate surface. It is also possible that water molecules around charged L-histidine or charged side chains are less likely to form a hydrate. Therefore, the nucleation and growth of CO<sub>2</sub> hydrate are decreased [28].

Bavoh et al. [46] have studied L-arginine as an inhibitor for CO<sub>2</sub> hydrate formation and modeled the CO<sub>2</sub> hydrate phase boundaries. It was observed that the phase diagram shifted to higher pressure and lower temperature in the presence of L-arginine, and it is classified as a thermodynamic inhibitor. By calculating CO<sub>2</sub> hydrate dissociation enthalpy, they have concluded that the amino acid does not participate in cage occupation or structure during the hydrate formation.

#### 4. Conclusions

Experiments were conducted in two temperature scheme to compare the hydrate formation capabilities of four selected amino acids with water and surfactant SDS during flue gas hydrate formation. Hydrophobic amino acids perform better than pure water and hydrophilic amino acids. This difference is due to their different perturbation effects on the water structure around the side chain. As CO<sub>2</sub> concentration is increased, the hydrate promotion effect in all solutions improved. SDS performance at lower concentrations (500–1000 ppm) was observed to be weak and similar to water; however, at higher concentrations, the performance is improved reflected by lower induction



time. At higher concentration (from 1000 ppm–3000 ppm), increase in CO<sub>2</sub> concentration in flue gas, do not increase gas uptake for SDS. At 3000 ppm concentration, L-methionine showed the best promotion ability among hydrophobic amino acids and very similar to SDS (1000–3000ppm) in terms of nucleation and growth kinetics. All tested solutions show a strong memory effect during induction time measurements, but they eliminated the memory effect for gas uptake and nucleation kinetics. The results highlight the potential of hydrophobic amino acids as environmentally friendly alternatives to SDS for CO<sub>2</sub> capture, as amino acids are green in nature, being cheap, non-toxic, and biodegradable. Pressure requirements to form flue gas hydrate are very high to be adopted at commercial scale; therefore, hydrophobic amino acids with thermodynamic promoters could provide the more eco-friendly solution for CO<sub>2</sub> capture by forming flue gas hydrates at moderate operational requirements and faster formation kinetics. The results also highlight the role of L-histidine and L-arginine as an inhibitor and could be of interest as a potential hydrate inhibitor in the oil and gas industry. The application of the rocking cell provides standardization during the experimental investigation. Overall, the results would help in understanding the role of amino acids and surfactants for flue gas hydrate formation to increase efficiency and reduce the costs.

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