


## Article

# The Carbon-Coated $\text{ZnCo}_2\text{O}_4$ Nanowire Arrays Pyrolyzed from PVA for Enhancing Lithium Storage Capacity

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**Abstract:** In this paper,  $\text{ZnCo}_2\text{O}_4$  nanowire arrays with a uniform carbon coating were introduced when polyvinyl alcohol (PVA) served as the carbon source. The coating process was completed by a facile bath method in PVA aqueous solution and subsequent pyrolyzation. The PVA-derived carbon-coated  $\text{ZnCo}_2\text{O}_4$  nanowire array composites can be used directly as the binder-free and self-supported anode materials for lithium-ion batteries. In the carbon-coated  $\text{ZnCo}_2\text{O}_4$  composites, the carbon layer carbonized from PVA can accelerate the electron transfer and accommodate the volume swing during the cycling process. The lithium storage properties of the carbon-coated  $\text{ZnCo}_2\text{O}_4$  composites are investigated. It is believed that the novel carbon-coating method is universal and can be applied to other nanoarray materials.

**Keywords:** polyvinyl alcohol; carbon coating; nanoarray composites; lithium-ion batteries

## 1. Introduction

Nowadays, a variety of electric vehicles and portable electronic devices (mobile phones, laptops, etc.) have been widely applied and developed rapidly in the world. Lithium-ion batteries as requisite energy storage devices for them receive more demand than ever before [1–4]. The commercial graphite anode has low theoretical capacity of  $372 \text{ mA}\cdot\text{h}\cdot\text{g}^{-1}$  and awkward rate performance. Hence, it is imperative to study and find promising alternate materials. In recent studies, transition metal oxides (TMOs) have attracted growing interest owing to their large theoretical capacities ( $>600 \text{ mA}\cdot\text{h}\cdot\text{g}^{-1}$ ) [5–8].  $\text{ZnCo}_2\text{O}_4$ , as one of the cobalt-based materials, stands out among a variety of anode materials for its high reversible capacities ( $976 \text{ mA}\cdot\text{h}\cdot\text{g}^{-1}$ ) and environmental benignity [9–14]. Nevertheless, poor electric conductivity and cycling stability still need improvement to realize its commercialization [15,16].

Recent research indicates that the free-standing nanoarray construction of electrode materials can effectively shorten the transport path of electrons and ions, alleviate the volume swings caused by the repeated ion insertion/extraction, as well as avoid the use of additional binder and conductive agent. These nanoarray electrodes displayed outstanding cycling stability and rate performance [17–23]. Wang et al. reported the fabrication of  $\text{NiCo}_2\text{O}_4$  nanowire arrays grown on carbon cloth as a binder-free electrode for both sodium and lithium-ion storage. Subsequently, an amorphous carbon layer has been deposited onto the surface of the  $\text{NiCo}_2\text{O}_4$  nanowires through the physical vapor deposition technique, further improving the conductivity and structural stability of the materials [24]. Wang and his co-workers at Tianjin University reported free-standing electrode of  $\text{Fe}_2\text{O}_3$  nanorods grown on the carbon fiber. They then coated  $\text{Fe}_2\text{O}_3$  with carbon by a hydrothermal method in an aqueous solution

containing glucose heated at 200 °C for 4 h [25]. Hence, the carbon coating can significantly enhance the electrochemical performances of active materials. However, the existing methods of coating on the nanoarrays are always complicated, either rely on the expensive instrument (e.g., CVD, PECVD, or ALD), or require high reaction temperature [26–30].

In this paper, we present a novel and facile in situ bath method in polyvinyl alcohol (PVA) aqueous solution to fabricate carbon-coated  $\text{ZnCo}_2\text{O}_4$  nanoarrays. The composites not only benefit the advantages of nanoarray architecture but also the introduction of a uniform carbon layer can greatly strengthen the electroconductivity and structural stability, leading to better electrochemical performance. Our work of the novel carbon-coating method has versatility and can be used to optimize the electrochemical performance of other electrode materials.

## 2. Experimental

### 2.1. Synthesis of Carbon-Coated $\text{ZnCo}_2\text{O}_4$ Nanowire Arrays

The  $\text{ZnCo}_2\text{O}_4$  nanowire arrays directly grown on the Ni foam were obtained according to the literature reported by our group [31]. Firstly, a mixed solution of 2 mM  $\text{Zn}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ , 4 mM  $\text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ , 4 mM  $\text{NH}_4\text{F}$ , and 10 mM  $\text{CO}(\text{NH}_2)_2$  was transferred into a Teflon-lined stainless steel autoclave. Then, a piece of Ni foam was placed into the solution with its top side covered by a PTFE (Poly tetra fluoroethylene) tape. The autoclave was maintained at 120 °C for 6 h and then annealed at 400 °C for 2 h in air. Subsequently, 0.3% wt. concentrations of PVA (PVA-105, Aladdin, Shanghai, China) solutions were prepared and kept ultrasonication for about half an hour to remove the bubbles. The as-prepared  $\text{ZnCo}_2\text{O}_4$  nanowire arrays were immersed into the PVA solutions for about 10 min and then freeze-dried for 24 h. Finally, the carbon-coated arrays were obtained by heat treatment at 500 °C for 2 h in an argon flow.

### 2.2. Materials Characterization

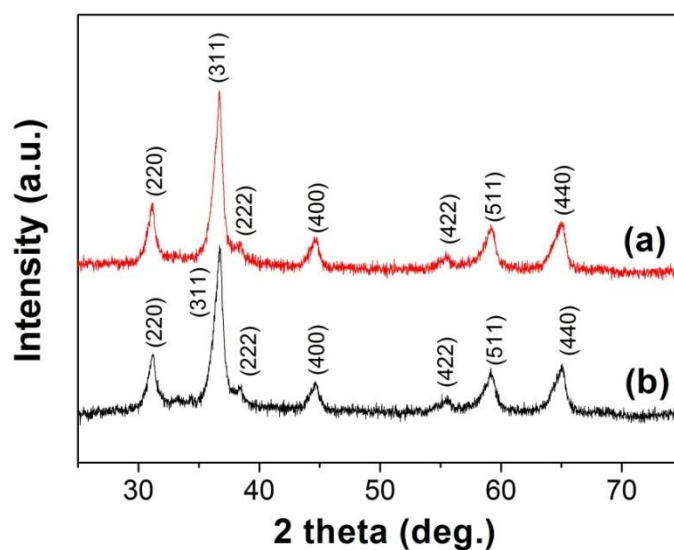
The X-ray diffraction (XRD) measurements were operated on a Rigaku D/max 2550 PC diffractometer using  $\text{Cu K}\alpha$  radiation ( $\lambda = 1.5418 \text{ \AA}$ ). The morphology and microstructure of samples were characterized by scanning electron microscopy (SEM, Hitachi, Tokyo, Japan, SU8010) and transmission electron microscopy (TEM, FEI, Hillsboro, OR, USA, F20). Thermogravimetric analysis (TGA) was performed in air with a temperature range of 25 °C to 800 °C to measure the carbon content in samples. Raman spectroscopic measurement was done using a NTEGRA Spectra AFM Raman Confocal SNOM instrument.

### 2.3. Electrochemical Measurement

The electrochemical tests were conducted using 2032 coin-type cells. These cells were assembled in an Ar-filled glove box ( $\text{O}_2 \leq 0.1 \text{ ppm}$ ,  $\text{H}_2\text{O} \leq 0.1 \text{ ppm}$ ), with the prepared samples directly as working electrode, lithium foil as counter/reference electrode. The electrolyte was a 1 M solution of  $\text{LiPF}_6$  in ethylene carbonate (EC) and Dimethyl carbonate (DMC) (1:1 by volume). Galvanostatic charge-discharge tests of cells were taken on a LAND test system in the voltage range of 0.01 and 2.5 V (vs.  $\text{Li/Li}^+$ ) at room temperature.

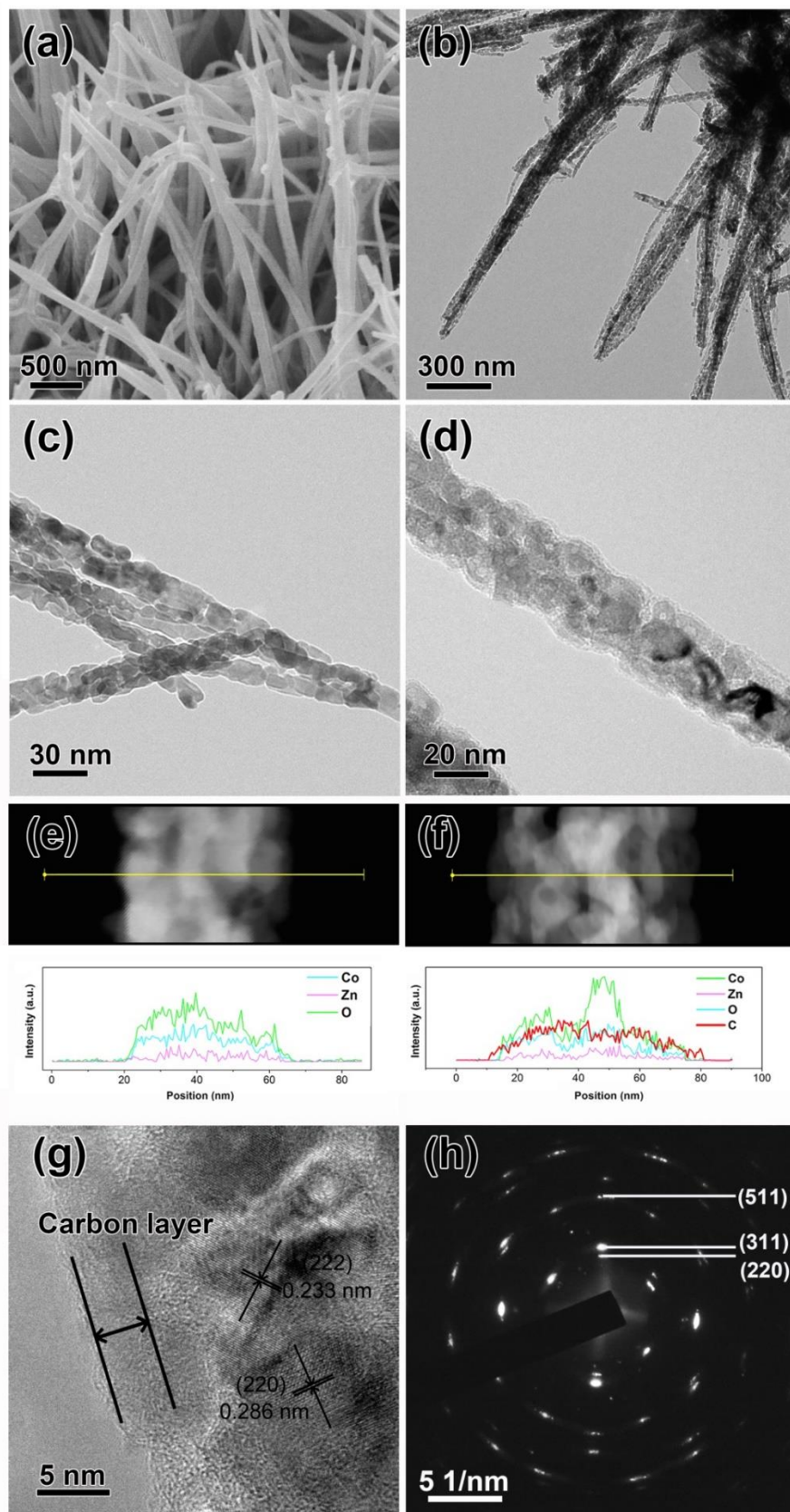
## 3. Results and Discussion

The XRD patterns of the  $\text{ZnCo}_2\text{O}_4$  and carbon-coated  $\text{ZnCo}_2\text{O}_4$  samples have been shown in Figure 1a and Figure 1b, respectively. Before the XRD test, in order to separate the nanomaterials from the nickel foam, the samples were wildly ultrasonicated for 30 min. It can be seen that both of the samples display well-defined peaks at 31.2, 36.8, 38.5, 44.7, 55.6, 59.3, and 65.1°, which can be indexed to the (220), (311), (222), (400), (422), (511) and (440) planes of spinel  $\text{ZnCo}_2\text{O}_4$  (PDF No. 23–1390). Moreover, no obvious peaks related to the carbon have been found in the XRD patterns in Figure 1b, which can be ascribed to the amorphous nature of the carbon layer.



**Figure 1.** XRD patterns of (a)  $\text{ZnCo}_2\text{O}_4$  and (b) carbon-coated  $\text{ZnCo}_2\text{O}_4$  samples.

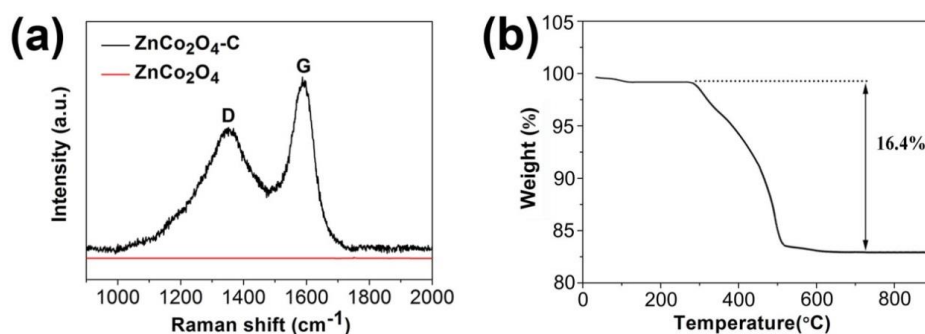
Figure 2 exhibits the characterizations of the morphology, structure, and composition of the carbon-coated  $\text{ZnCo}_2\text{O}_4$  nanostructures. Figure 2a shows the SEM image of the carbon-coated  $\text{ZnCo}_2\text{O}_4$  nanowires directly grown on the nickel foam. The nanocomposites still maintain the nanowire array structure after the immersion and thermal treatment process compared with the original  $\text{ZnCo}_2\text{O}_4$ . The specific characterizations of the morphology of bare  $\text{ZnCo}_2\text{O}_4$  nanowire arrays can be referred to in our pre-published literature [31]. Figure 2b displays the TEM image of the carbon-coated  $\text{ZnCo}_2\text{O}_4$  nanowires, in which more structural details can be seen. The  $\text{ZnCo}_2\text{O}_4$  nanowires are polycrystalline and consist of numerous nanoparticles. However, in neither SEM nor TEM images, can the carbon coating layer be obviously observed. Figure 2d gives a partial enlarged TEM image of a single carbon-coated  $\text{ZnCo}_2\text{O}_4$  nanowire. In Figure 2d, one can clearly see that the nanowire has been completely and uniformly surrounded by an amorphous layer. In comparison to the carbon-coated samples, a TEM image of the bare  $\text{ZnCo}_2\text{O}_4$  nanowire was illustrated in Figure 2c. Without the amorphous layer, the polycrystalline nanoparticles look more distinct. The compositional line profiles of the uncoated and coated nanowires are revealed in Figure 2e,f. In Figure 2f, there are significant signals of the Energy Dispersive X-Ray Spectroscopy (EDX) for Zn, Co, O, and C elements, which are expected to stem from  $\text{ZnCo}_2\text{O}_4$  and carbon layer, respectively. The corresponding line-scan EDX patterns further confirm the presence and location of the carbon layer. The HRTEM (Figure 2g) of the carbon-coated  $\text{ZnCo}_2\text{O}_4$  further indicates that the thickness of the coating layer is about 3 nm. The completeness, continuity, and suitable thickness of a carbon coating layer on the surface of the electrode materials is of great importance for the electrochemical reactions. If the coating layer is too thick, it may lead to negative effects for the ions to pass through; while too thin coating layers may not effectively maintain the micro-structure of the active materials. It has been reported that the thickness of about 2–3 nm of the carbon coating layer may be appropriate for repetitive insertion/extraction of ions [32–34]. In Figure 2g, there are distinct lattice fringes with lattice spacings of about 0.233 and 0.286 nm, corresponding to (222) and (220) planes of  $\text{ZnCo}_2\text{O}_4$ , respectively. The selected area electron diffraction (SAED) pattern of the carbon-coated  $\text{ZnCo}_2\text{O}_4$  in Figure 2h is comprised of sets of concentric rings with occasional bright spots which indicate the nano-crystalline nature of the composites. The diffraction rings can be readily assigned to the crystal planes of the spinel  $\text{ZnCo}_2\text{O}_4$  phase. These outcomes suggest that the carbon-coated  $\text{ZnCo}_2\text{O}_4$  core-shell nanostructures have been successfully manufactured after the immersion in PVA solution and subsequent carbonization process.



**Figure 2.** (a) A SEM image, (b) a TEM image of carbon-coated  $\text{ZnCo}_2\text{O}_4$  nanowires, (c,d) partially enlarged TEM images of uncoated and carbon-coated  $\text{ZnCo}_2\text{O}_4$  nanowires, compositional line profiles across (e) the uncoated and (f) carbon-coated  $\text{ZnCo}_2\text{O}_4$  nanowires probed by EDX spectroscopy, (g) HRTEM image, (h) SAED patterns of the carbon-coated  $\text{ZnCo}_2\text{O}_4$  nanowires.

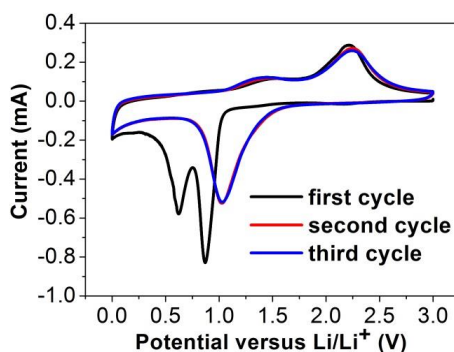
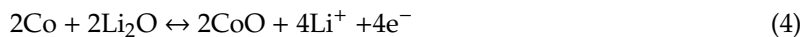
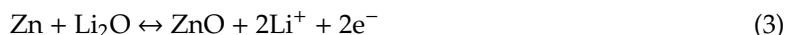
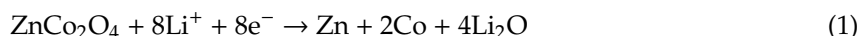


The presence of carbon in the composites can be further affirmed by the Raman spectrum (Figure 3a) and TGA analysis (Figure 3b). In Figure 3a, there are two distinct characteristic peaks for the carbon-coated samples. The peaks located at  $\sim 1340\text{ cm}^{-1}$  are D band, representing the disordered structures or structural defects in carbon materials; while the peaks at  $\sim 1590\text{ cm}^{-1}$  are G band, attributed to the degree of graphitization. Meanwhile, due to the relatively low signal intensity, the typical Raman spectrum of the uncoated samples is almost a straight line. The results of the Raman spectrum confirm the presence of carbon and partial graphitization of the carbon-coated  $\text{ZnCo}_2\text{O}_4$  samples. TGA was carried out to measure the carbon content in the composites in the air atmosphere to  $800^\circ\text{C}$ . In Figure 3b, the weight loss from  $200$  to  $600^\circ\text{C}$  is due to the combustion reaction of the carbon component in the composites. The carbon content of the carbon-coated  $\text{ZnCo}_2\text{O}_4$  composites is about 16.4% by weight. The above characterizations confirm the successful synthesis of  $\text{ZnCo}_2\text{O}_4\text{-C}$  core-shell nanocomposites.



**Figure 3.** (a) Typical Raman spectrum of the uncoated and carbon-coated  $\text{ZnCo}_2\text{O}_4$ , (b) TGA analysis of the carbon-coated  $\text{ZnCo}_2\text{O}_4$  nanocomposites.

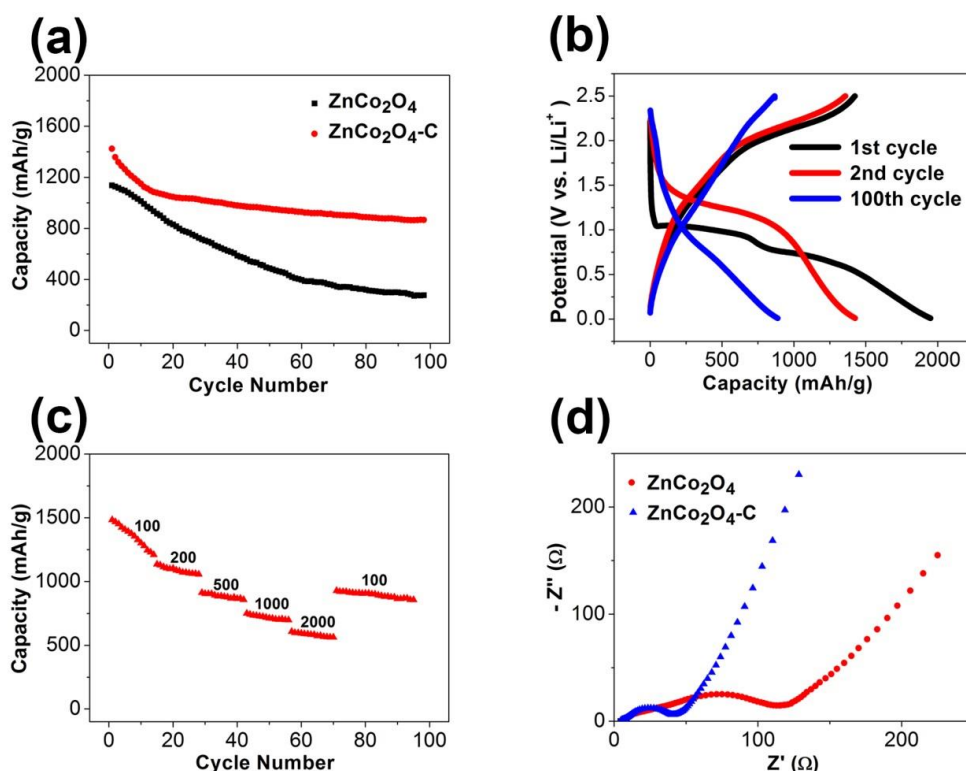
The one-dimensional nanowire structure, the orderly aligned arrays, and the uniform carbon coating make the composite a preeminent anode material for LIBs. Herein, the carbon-coated  $\text{ZnCo}_2\text{O}_4$  nanowire arrays grown on the Ni foam are directly evaluated as the anode. Figure 4 reveals the first three cyclic voltammogram (CV) curves of the carbon-coated  $\text{ZnCo}_2\text{O}_4$  nanowire electrodes in the potential range of  $0.01\text{--}3.0\text{ V}$  at a scan rate of  $0.1\text{ mV}\cdot\text{s}^{-1}$ . The electrochemical reactions of  $\text{ZnCo}_2\text{O}_4$  with Li can be clarified as follows [35,36]:



**Figure 4.** The first three CV curves of carbon-coated  $\text{ZnCo}_2\text{O}_4$  array electrodes between  $0.0\text{ V}$  to  $3.0\text{ V}$  versus  $\text{Li/Li}^+$  at a scan rate of  $0.1\text{ mV}\cdot\text{s}^{-1}$ .

It can be clearly seen that the curve of the first cathodic scan is different from the subsequent ones. In the first circle, the two reduction peaks located at 0.87 V and 0.62 V are corresponded to the reduction process of  $\text{ZnCo}_2\text{O}_4$  to Zn and Co accompanying the formation of the SEI layer. During the first anodic scan, two broad peaks corresponding to the oxidation of metallic Zn and Co are detected at 1.4 V and 2.2 V. In the following cycles, the major reduction peak shifts to a higher potential of 1.03 V, suggesting that the polarization of the electrode material is decreased after the activation process in the initial circle. Similar reports of shifts to a higher potential in the subsequent circles have been reported [37,38]. The shape and position of anodic peaks keep stable in the following circles.

The investigation about the effects of the carbon coating on the electrochemical performance has been carried out. Figure 5a illustrates the cycling performance of the carbon-coated  $\text{ZnCo}_2\text{O}_4$  in comparison with the pristine  $\text{ZnCo}_2\text{O}_4$  electrode for LIBs under the same test conditions. As can be seen, the carbon-coated  $\text{ZnCo}_2\text{O}_4$  electrodes display higher reversible capacities, compared with the untreated  $\text{ZnCo}_2\text{O}_4$ . The bare  $\text{ZnCo}_2\text{O}_4$  electrode goes through a rapid capacity fading, and remain the capacity of less than  $300 \text{ mA}\cdot\text{h}\cdot\text{g}^{-1}$  after 100 loops. On the other hand, the reversible capacities of the carbon-coated  $\text{ZnCo}_2\text{O}_4$  electrodes after 100 cycles were  $886.4 \text{ mA}\cdot\text{h}\cdot\text{g}^{-1}$ . The right amount of the carbon coating plays a significant role in the electrochemical performance of the anode materials [29–31]. If the coating layer is excessively thin, it probably cannot afford to hold the integrity of the materials during repeated ion insertion/extraction. However, coating that is too thick may block the traversing of the  $\text{Li}^+$  ions. The uniform and moderate carbon coating is able to hold back the aggregation and pulverization of the  $\text{ZnCo}_2\text{O}_4$  electrode, as well as accelerates the ion/electron transfer, which may be responsible for the enhanced cycling performance [39,40]. In addition, it is meaningful to make a comparison between our materials with those in the previously published literature, as illustrated in Table 1.



**Figure 5.** (a) The comparison of cycling performance of carbon-coated  $\text{ZnCo}_2\text{O}_4$  pyrolyzed from PVA solutions with pristine  $\text{ZnCo}_2\text{O}_4$  electrodes; (b) charge-discharge curves for lithium-ion batteries; (c) rate performance of the carbon-coated  $\text{ZnCo}_2\text{O}_4$  pyrolyzed from PVA solution; (d) Nyquist plots of the carbon-coated and pristine  $\text{ZnCo}_2\text{O}_4$  electrodes from 100 kHz to 0.01 Hz.

**Table 1.** Electrochemical performance of our carbon-coated ZnCo<sub>2</sub>O<sub>4</sub> compared with those previously published Co-based anodes for LIBs.

Electrode Materials	Initial Discharge/ Charge Capacity (mA·h·g <sup>-1</sup> )	Capacity after Cycling (mA·h·g <sup>-1</sup> )	Reference
porous ZnCo <sub>2</sub> O <sub>4</sub>	1332/979	721, 80th@100 mA g <sup>-1</sup>	[10]
ZnCo <sub>2</sub> O <sub>4</sub> /Co <sub>3</sub> O <sub>4</sub> on CC	750/736	481.9, 100th@300 mA g <sup>-1</sup>	[1]
Carbon fiber/ZnCo <sub>2</sub> O <sub>4</sub>	927.2/613.5	787.2, 150th@100 mA g <sup>-1</sup>	[2]
CoO/hollow carbon sphere	-	584, 50th@100 mA g <sup>-1</sup>	[41]
Co@CoO	1262/956	800, 50th@50 mA g <sup>-1</sup>	[42]
CoO@C	1564/-	~800, 70th@100 mA g <sup>-1</sup>	[43]
CoO-NiO-C	1125/729	562, 60th@100 mA g <sup>-1</sup>	[44]
ZnCo <sub>2</sub> O <sub>4</sub> @TiO <sub>2</sub>	1598/1343	827, 90th@100 mA g <sup>-1</sup>	[45]
C/ZnCo <sub>2</sub> O <sub>4</sub> @CNT	1947.1/763.1	800.6, 100th@100 mA g <sup>-1</sup>	[46]
Co-ZnO@C	940/862	538, 50th@100 mA g <sup>-1</sup>	[47]
ZnCo <sub>2</sub> O <sub>4</sub> -C	1951.4/1424.5	886.4, 100th @200 mA g <sup>-1</sup>	this work

Figure 5b displays the 1st, 2nd, and 100th discharge-charge curves of the carbon-coated ZnCo<sub>2</sub>O<sub>4</sub> anode material at a current density of 200 mA g<sup>-1</sup>. The obvious plateau at about 1.0 V and 0.7 V were observed in the first discharge curve, which is also consistent with the CV results. The initial discharge and charge capacities reached 1951.4 mA·h·g<sup>-1</sup> and 1424.5 mA·h·g<sup>-1</sup>, respectively. The capacity loss primarily dates from the formation of SEI and irreversible reactions of the electrode materials.

The carbon-coated ZnCo<sub>2</sub>O<sub>4</sub> anode material also demonstrated a remarkable rate performance in Figure 5c. When the current densities reached 100, 200, 500, 1000, and 2000 mA g<sup>-1</sup>, the reversible capacity reached 1210.0, 1056.4, 858.3, 698.5, and 562.8 mA·h·g<sup>-1</sup>, respectively. When the current densities return to 100 mA g<sup>-1</sup>, the reversible capacity can be recovered to 857.1 mA·h·g<sup>-1</sup>, which is ~70.8% of the initial reversible capacity.

Electrochemical impedance spectroscopy (EIS) was processed to evaluate the charge transfer properties of the electrode materials after 5 cycles in Figure 5d. The flat semicircle in the high-frequency region is ascribed to the charge transfer resistance (R<sub>ct</sub>) for the anode materials after cycling [48–50]. The Nyquist plots uncover that the diameters of the semicircles of the carbon-coated ZnCo<sub>2</sub>O<sub>4</sub> electrodes are significantly smaller than that of the bare ZnCo<sub>2</sub>O<sub>4</sub> electrode. From the above, the EIS results show that the coating of a layer of carbon is conducive to the reduction of the charge transfer resistance, further improving the electronic conductivity and electrochemical performance. The interface of the electrode materials for the lithium-ion batteries is one of the key issues for its long-term development, and deeper research in the future is essential [51].

#### 4. Conclusions

In summary, we have proposed an in-situ bath method to uniformly coat ZnCo<sub>2</sub>O<sub>4</sub> nanowire arrays with a carbon shell. The polymer of PVA has been chosen as the carbon source and the influence of PVA-derived carbon-coating layer on the electrochemical performance of ZnCo<sub>2</sub>O<sub>4</sub> has been investigated. The carbon-coated ZnCo<sub>2</sub>O<sub>4</sub> nanowire array composites can be used directly as an anode material for LIBs. The ZnCo<sub>2</sub>O<sub>4</sub> nanowires directly grown on the metal substrate can enhance the electron and ion transport and avoid the additional use of binder and conductive agent. The carbon layer is able to improve the conductivity of the electrode materials and restrain the pulverization, eventually resulting in a splendid electrochemical performance.

**Author Contributions:** Conceptualization, W.Z. and J.C.; methodology, Z.S., Y.Q.; writing—original draft preparation, W.Z.; writing—review and editing, J.C. All authors have read and agreed to the published version of the manuscript.

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**Conflicts of Interest:** The authors declare no conflict of interest.

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