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# Green Synthesis of a Functionalized Zirconium-Based Metal-Organic Framework for Water and Ethanol Adsorption

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**Abstract:** Aqueous synthesis of metal–organic frameworks (MOFs) at room temperature offers many advantages such as reduction in the generation of toxic byproducts and operation costs, as well as increased safety in the material’s production. Functional group-bearing MOFs have received growing attention compared to nonfunctionalized analogues due to enhanced adsorption properties of the former in many cases. Here, we report an aqueous solution-based synthesis of a robust zirconium MOF, UiO-66-NO<sub>2</sub>, at room temperature. We evaluated the phase purity, porosity, thermal stability, particle morphology and size of the resulting material. High uptake, as well as near complete recyclability of water and ethanol vapor isotherms at room temperature supports the potential of UiO-66-NO<sub>2</sub> as a solid adsorbent in adsorption-based cooling applications or water harvesting systems.

**Keywords:** zirconium-based MOFs; water adsorption; ethanol adsorption; porous materials; adsorption heat pump

## 1. Introduction

Metal–organic frameworks (MOFs) [1–5] are a promising class of crystalline porous materials with a wide range of applications including but not limited to heterogeneous catalysis [6–10], enzyme and nanoparticle encapsulation [11–14], water capture [15–19], gas storage, and separation [20–25]. Importantly, physical and chemical properties of MOFs can be fine-tuned with the help of reticular chemistry, where pre-selected molecular building blocks are combined to yield pre-designed frameworks for targeted applications [26–30].

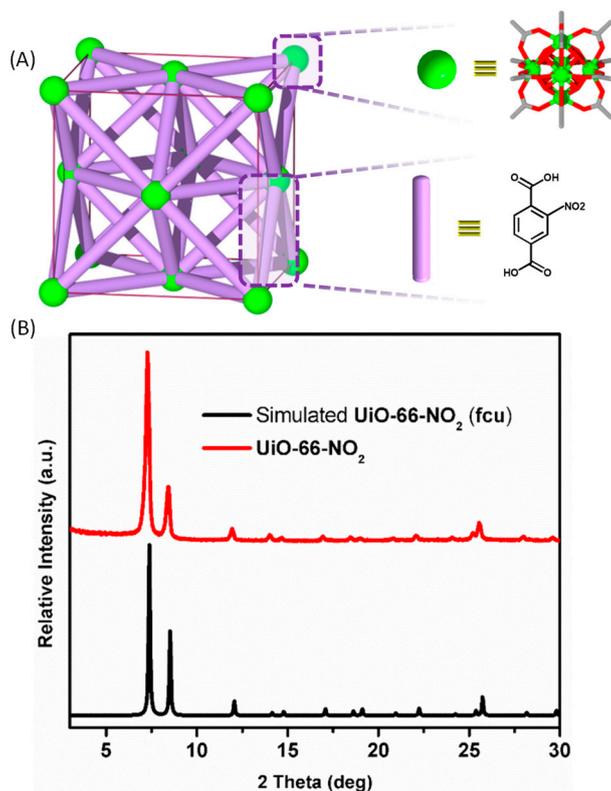
While gas adsorption is the most widely studied application of MOFs, vapor adsorption in MOFs has recently garnered some attention especially with the increasing number of water-stable MOFs reported [15–19,31–33]. MOFs with a high capacity for water uptake at room temperature have also gained increasing attention. By tuning features such as the topology, pore size, pore volume, and chemical functionality of MOFs, the relative pressure where the water uptake reaches maximum can be controlled. This is essential for many applications such as natural gas dehydration [34], water capture from air [16,35], adsorption-based cooling systems [15,19,36], and indoor humidity cooling systems [18]. Ethanol can be also used as refrigerant for adsorption-driven cooling application due to its lower freezing point compared to water [37]. Nevertheless, all of these adsorption-based applications such as adsorption beds [21,38], mixed-matrix membranes [39,40], or devices for capturing water from

air [35,41], require large amounts of adsorbents for operation. Consequently, developing green and scalable methods for synthesis of MOFs is of interest for the aforementioned applications.

We have recently reported a promising strategy for the scalable synthesis of UiO-66 (UiO = University of Oslo) analogues (i.e., UiO-66-(COOH)<sub>2</sub> and UiO-66-F<sub>4</sub>) for toxic chemical removal [42]. Herein, we extended this method to the synthesis of a highly stable Zr MOF, UiO-66-NO<sub>2</sub> [43–45]. Various characterization techniques were implemented to confirm phase purity, morphology, thermal stability, and porosity of the resulting MOF. Moreover, we studied water and ethanol vapor adsorption of UiO-66-NO<sub>2</sub>, which showed high uptake as well as good recyclability.

## 2. Results and Discussion

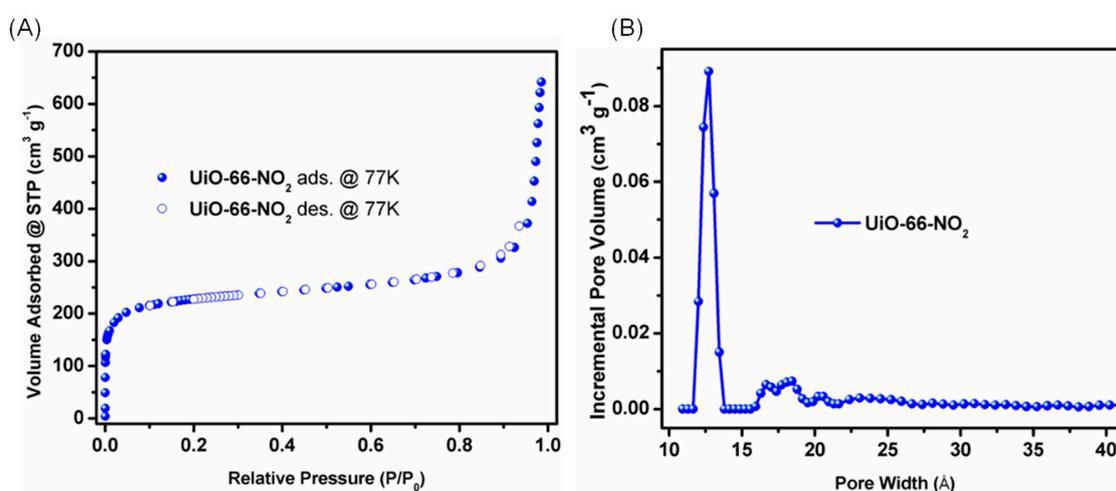
UiO-66 is constructed from 12-connected hexanuclear zirconium oxo-clusters and terephthalate linkers with face-centered cubic (fcu) topology [46]. Due to the exceptional chemical and thermal stability, UiO-66 has become one of the most studied MOFs since its invention in 2008 [43,46–49]. UiO-66 can be easily functionalized with different functional groups on account of its isoreticular chemistry [5], however, a vast majority of these studies involve *N,N*-dimethylformamide (DMF), a toxic organic solvent, and elevated temperatures [5]. We recently reported an inexpensive, scalable, and environmentally benign procedure to synthesize UiO-66 derivatives (UiO-66-(COOH)<sub>2</sub> and UiO-66-F<sub>4</sub>) in water at room temperature [42]. Building on the previous successful examples, we extended our strategy to synthesize another UiO-66 derivative, UiO-66-NO<sub>2</sub> (Figure 1A). Among those functionalized UiO-66 analogues, UiO-66-NO<sub>2</sub> is believed to be one of the most chemically stable Zr MOFs which makes it a potential candidate for broader applications [50,51].



**Figure 1.** (A) Structural representation of UiO-66-NO<sub>2</sub> constructed from 12-c Zr nodes and linear 2-nitroterephthalic acid (H<sub>2</sub>BDC-NO<sub>2</sub>) ligand. (B) Powder X-ray diffraction (PXRD) pattern of the as-synthesized sample matched well with the simulated UiO-66-NO<sub>2</sub>.

As illustrated in Figure 1B, our aqueous synthetic method at room temperature resulted in phase-pure UiO-66-NO<sub>2</sub> confirmed by the powder X-ray diffraction (PXRD) pattern. The crystal size of UiO-66-NO<sub>2</sub> was about 100 nm as shown in the scanning electron microscopy (SEM) image (Figure S1).

N<sub>2</sub> adsorption–desorption isotherms at 77 K revealed the experimental total pore volume at  $P/P_0 = 0.8$  was about 0.43 cm<sup>3</sup>·g<sup>-1</sup>, and the apparent Brunauer–Emmett–Teller (BET) area of UiO-66-NO<sub>2</sub> was calculated to be 840 m<sup>2</sup>·g<sup>-1</sup> based on the relative pressure range from  $P/P_0 = 0.004$  to  $P/P_0 = 0.0294$  (Figure 2A), in line with the surface areas obtained from solvothermal synthesis [52]. The pore-size distribution based on non-local density functional theory (NLDFT) model with the cylindrical kernel showed two major pores. The main peak centered at about 1.3 nm was assigned to octahedral cages of ideal UiO-66 structure, and the smaller shoulder near 1.8 nm was assigned to the presence of non-ordered missing cluster defects (Figure 2B) [53]. The sharp increase in the N<sub>2</sub> isotherm near saturation pressure was due to the condensation of the N<sub>2</sub> gas at interparticle voids, which is common in MOFs with nanosized crystallites [54].



**Figure 2.** (A) N<sub>2</sub> sorption isotherms of UiO-66-NO<sub>2</sub> at 77 K and (B) pore-size distribution based on a non-local density functional theory (NLDFT) model.

The thermogravimetric analyses (TGA) plot of UiO-66-NO<sub>2</sub> under continuous air flow indicated that the framework was stable up to about 300 °C (Figure S2). The number of the linkers per Zr<sub>6</sub> node inside UiO-66-NO<sub>2</sub> was estimated to be about 4.3 according to the TGA plot, assuming the full combustion of MOFs under the air flow. This indicated the defective nature of this UiO-66-NO<sub>2</sub> synthesized in water at room temperature, which was consistent with the pore size distribution, further indicating the presence of defects. Diffuse reflectance infrared Fourier transform spectroscopy (DRIFTS) spectra was collected on this material where a sharp band at 3667 cm<sup>-1</sup> was assigned to  $\mu_3$ -OH stretches on the hexanuclear zirconium oxo-clusters (Figure S3) which was red-shifted compared to its parent UiO-66 [55] and consistent with solvothermally synthesized UiO-66-NO<sub>2</sub> [52].

Encouraged by the moderate surface area and high stability of the UiO-66-NO<sub>2</sub>, we further studied the ethanol and water vapor sorption, which is important for practical applications such as adsorption-based heat pump and vapor capture from air, respectively. Ethanol adsorption of UiO-66-NO<sub>2</sub> at 298 K showed an uptake of about 112 cm<sup>3</sup>·g<sup>-1</sup> at  $P/P_0 = 0.3$  and about 186 cm<sup>3</sup>·g<sup>-1</sup> at  $P/P_0 = 0.9$ , respectively (Figure 3). Moreover, it retained its ethanol sorption performance for the second cycle after the simple activation under dynamic vacuum.

Water adsorption isotherm at room temperature is important to gauge MOFs for their potential for water vapor sorption related applications or to assess their moisture stability [19,35]. Thus, we performed the water vapor sorption at 298 K for UiO-66-NO<sub>2</sub>, which exhibited an S-shape adsorption isotherm (Figure 4). The uptake of water vapor at  $P/P_0 = 0.9$  was about 459 cm<sup>3</sup>·g<sup>-1</sup> (or 0.37 g·g<sup>-1</sup>), and this value was slightly smaller than the reported value [17] of parent UiO-66 (533 cm<sup>3</sup>·g<sup>-1</sup> or

$0.43 \text{ g}\cdot\text{g}^{-1}$ ), due to the reduction of pore volume after the introduction of nitro-groups. On the other hand, the water uptake of UiO-66-NO<sub>2</sub> at  $P/P_0 = 0.3$  was about  $177 \text{ cm}^3\cdot\text{g}^{-1}$  (or  $0.14 \text{ g}\cdot\text{g}^{-1}$ ) which was higher than that of UiO-66 ( $108 \text{ cm}^3\cdot\text{g}^{-1}$  or  $0.09 \text{ g}\cdot\text{g}^{-1}$ ), indicating the more hydrophilic nature of this material in comparison with the parent UiO-66. The second run of the same batch material also verified the good cycling ability and therefore the stability of UiO-66-NO<sub>2</sub>. Note, the UiO-66-NO<sub>2</sub> synthesized in aqueous solution at room temperature showed similar water uptake to that previously reported ( $459 \text{ cm}^3\cdot\text{g}^{-1}$  (or  $20.5 \text{ mmol}\cdot\text{g}^{-1}$ ) in this work, in comparison to  $533 \text{ cm}^3\cdot\text{g}^{-1}$  (or  $22 \text{ mmol}\cdot\text{g}^{-1}$ ) at  $P/P_0 = 0.9$  by Walton and coworkers [44]), further confirming the practicality of this synthesis method of UiO-66-NO<sub>2</sub> compared to that of solvothermal based methods.

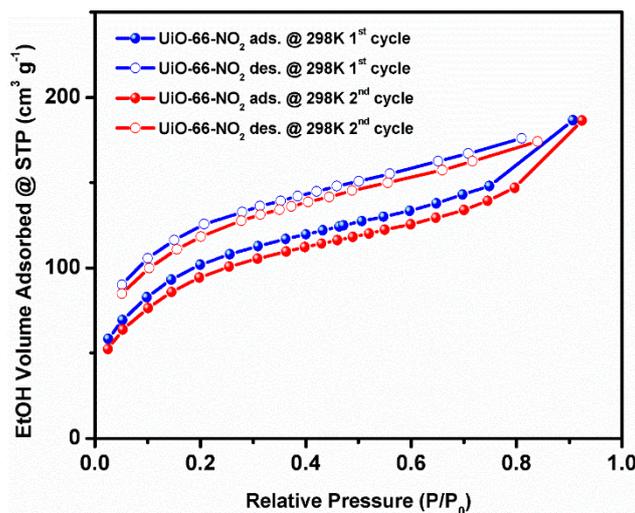


Figure 3. Ethanol sorption isotherms of UiO-66-NO<sub>2</sub> at 298 K (two cycles).

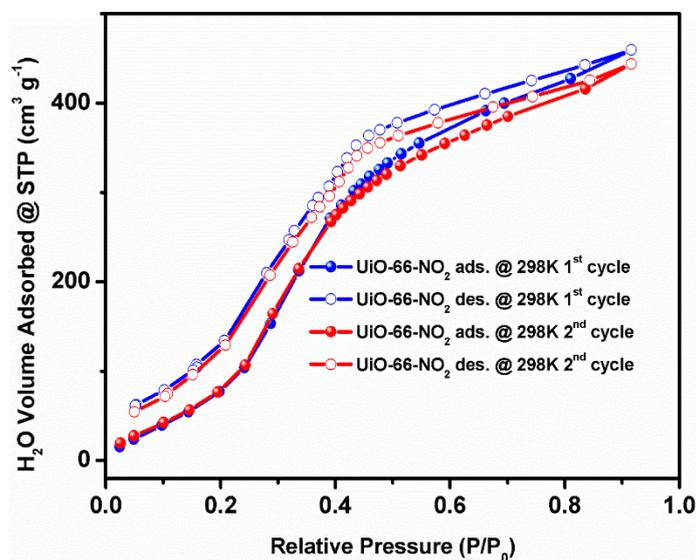


Figure 4. Water sorption isotherms of UiO-66-NO<sub>2</sub> at 298 K (two cycles).

### 3. Materials and Methods

All reagents were obtained from commercial sources and used without further purification, unless otherwise noted. Zirconium (IV) oxynitrate hydrate (99%) was purchased from Sigma-Aldrich (St. Louis, MO, USA).

### 3.1. X-ray Diffraction Analyses

Powder X-ray diffraction (PXRD) of MOFs were measured at room temperature on a STOE-STADIMP powder diffractometer (STOE & Cie GmbH, Darmstadt, Germany) equipped with an asymmetric curved Germanium monochromator (Cu K $\alpha$ 1 radiation,  $\lambda = 1.54056 \text{ \AA}$ ) and one-dimensional silicon strip detector (MYTHEN2 1K from DECTRIS, Baden, Switzerland). The line focused Cu X-ray tube was operated at 40 kV and 40 mA. The activated powder was sandwiched between two Kapton foils and measured in transmission geometry in a rotating holder. Intensity data from 1 to 30 degrees two theta were collected over a period of 6 min. The instrument was calibrated against a NIST Silicon standard (640d) prior to the measurement.

### 3.2. N<sub>2</sub> Sorption Measurements

N<sub>2</sub> adsorption and desorption isotherms on activated materials were measured on a Tristar (Micromeritics, Norcross, GA, USA) instrument at 77 K. In general, about 30–50 mg of sample was used in each measurement after activated at 120 °C for 24 h.

### 3.3. Thermogravimetric Analyses (TGA)

TGA was performed on a TGA/DCS 1 system (Mettler-Toledo AG, Schwerzenbach, Switzerland), which runs on a PC with STARe software. Samples were heated from 30 to 600 °C at a rate of 10 °C/min under air with flow rate 20 mL/min.

### 3.4. Diffuse Reflectance for Infrared Fourier Transform Spectroscopy (DRIFTS)

DRIFTS spectra were recorded on a Nicolet 6700 FTIR spectrometer (Thermo Nicolet Corp., Madison, WI, USA) equipped with an MCT detector. The detector was cooled with liquid N<sub>2</sub> and the spectra were collected under Ar atmosphere. KBr was utilized as a background spectrum.

### 3.5. Scanning Electron Micrographs (SEM)

Scanning electron micrographs (SEM) images were taken using a Hitachi SU8030 (Hitachi High Technologies Corporation, Tokyo, Japan) at the EPIC facility (NUANCE Center-Northwestern University). Samples were activated and coated with OsO<sub>4</sub> to ~9 nm thickness in a Denton Desk III TSC Sputter Coater before imaging.

### 3.6. Water and Ethanol Vapor Sorption Measurements

Water and ethanol isotherms were measured on a 3Flex (Micromeritics), and the water uptake in g·g<sup>-1</sup> unit is calculated as [(adsorbed amount of water)/(amount of adsorbent)]. Prior to the adsorption measurements, water or ethanol (analyte) was flash frozen under liquid nitrogen and then evacuated under dynamic vacuum at least 3 times to remove any gases in the water reservoir until the P<sub>0</sub> of the analyte gas matches the suggested P<sub>0</sub> at room temperature. The measurement temperature was controlled with an ISO Controller (Micromeritics).

### 3.7. Synthesis of UiO-66-NO<sub>2</sub>

1.5 mL of water was added to 20 mg (~0.094 mmol) of 2-nitroterephthalic acid. The mixture was then placed in the oven at 100 °C for 5 min. The color of the solution became pale yellow. Separately, 48 mg (~0.208 mmol) of zirconyl nitrate hydrate (ZrO(NO<sub>3</sub>)<sub>2</sub>·xH<sub>2</sub>O) was dissolved by adding 1.0 mL of water followed by sonication. The two solutions were combined, followed by an addition of 300  $\mu$ L (~5.24 mmol) of acetic acid. The mixture was stirred for 72 h at room temperature. The crystalline powders were collected by centrifugation (7000 rpm for 30 min). As-synthesized sample was washed 3 times with DI water and then sequentially washed 2 times with ethanol, and finally immersed in acetone for 2 days, during which time the acetone was replaced 2 times per day. The product was

obtained after initial drying via leaving exposed to ambient atmosphere overnight and activated at 120 °C for 24 h; yield: 25 mg.

#### 4. Conclusions

In conclusion, we extended the room temperature, aqueous methodology of Zr MOFs to access a nitro-functionalized UiO-66 derivative, UiO-66-NO<sub>2</sub>. Importantly, the UiO-66-NO<sub>2</sub> obtained here, using more environmentally benign methods, showed comparable BET area to MOFs obtained from solvothermal synthesis. Moreover, water and ethanol vapor sorption isotherms at room temperature indicated high uptake of both vapors, which suggests potential of UiO-66-NO<sub>2</sub> for adsorption-based cooling applications or water harvesting systems.

**Supplementary Materials:** The following are available online at <http://www.mdpi.com/2304-6740/7/5/56/s1>, Figures S1–S3: SEM image, TGA curve and DRIFTS spectra of UiO-66-NO<sub>2</sub>.

**Author Contributions:** Conceptualization, Z.C., T.I. and O.K.F.; methodology, Z.C.; software, Z.C.; validation, T.I. and O.K.F.; formal analysis, X.W., Z.C.; investigation, Z.C. and X.W.; resources, O.K.F.; data curation, Z.C. and X.W.; writing—original draft preparation, Z.C.; writing—review and editing, X.W., T.I. and O.K.F.; visualization, Z.C. and X.W.; supervision, T.I. and O.K.F.; project administration, Z.C., T.I. and O.K.F.; funding acquisition, O.K.F.

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**Conflicts of Interest:** The authors declare no conflict of interest.

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