

**Malonaldehyde-like systems:BeF<sub>2</sub> clusters. A subtle balance between hydrogen bonds, beryllium bonds and resonance.**

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## **Supporting Information**

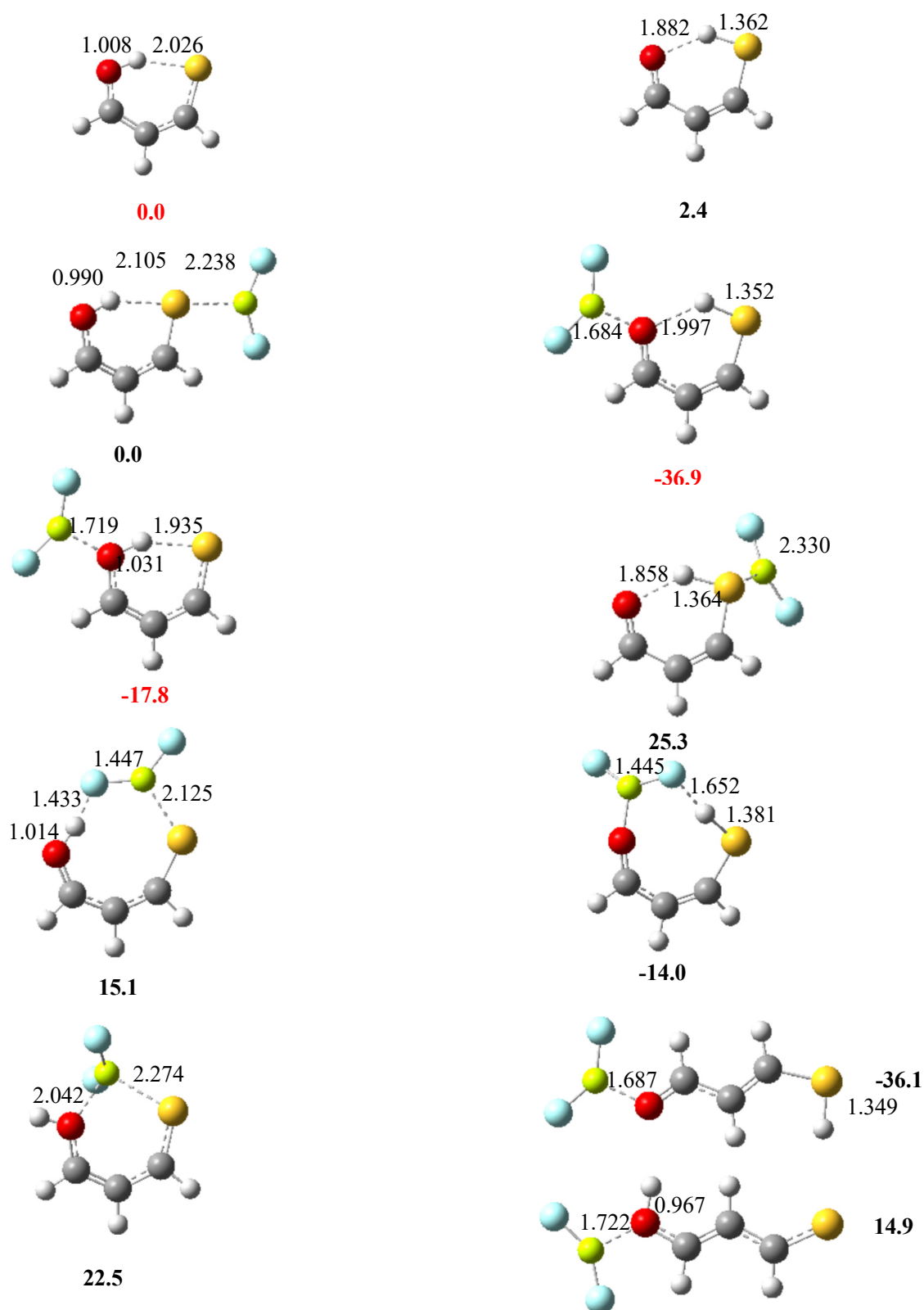
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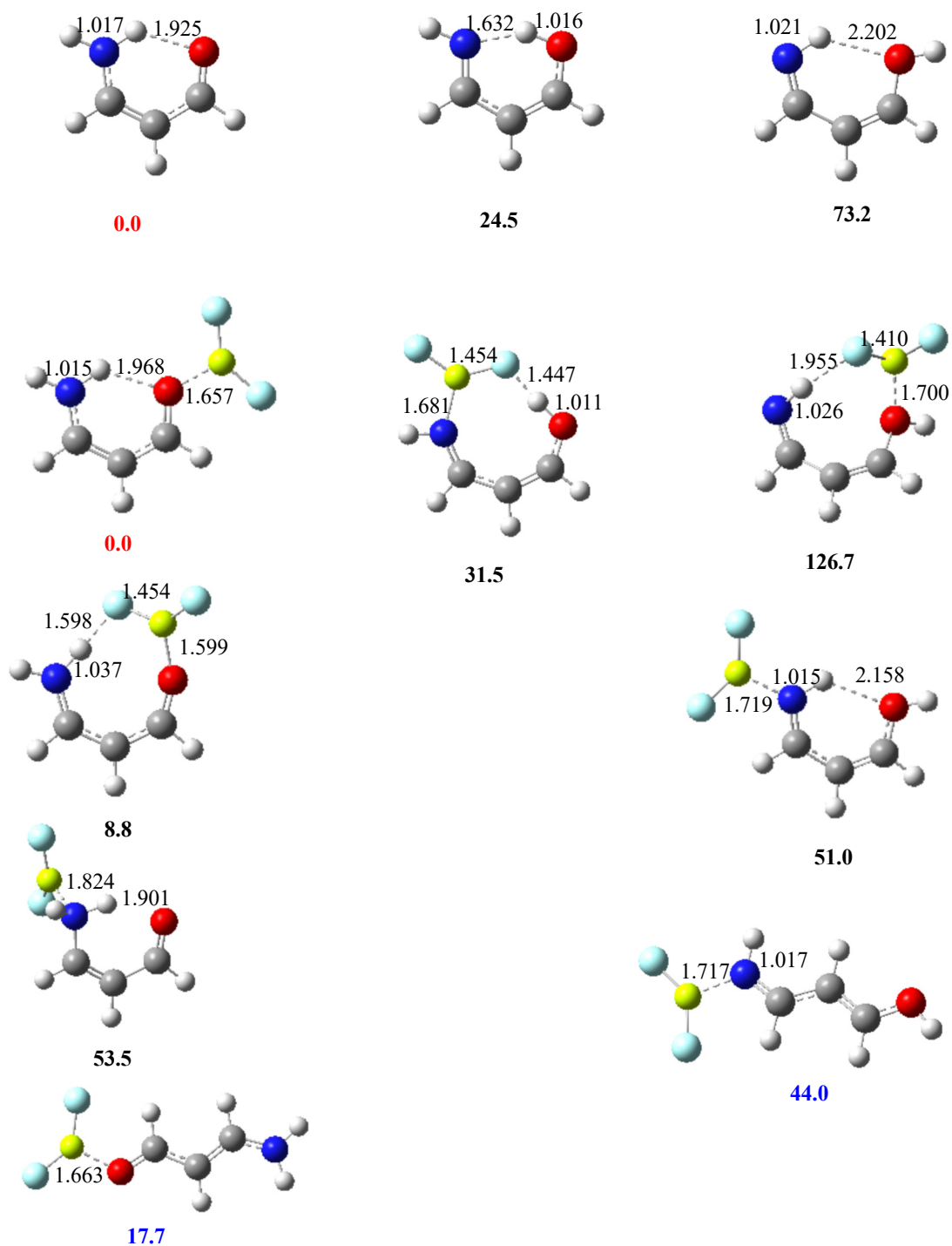
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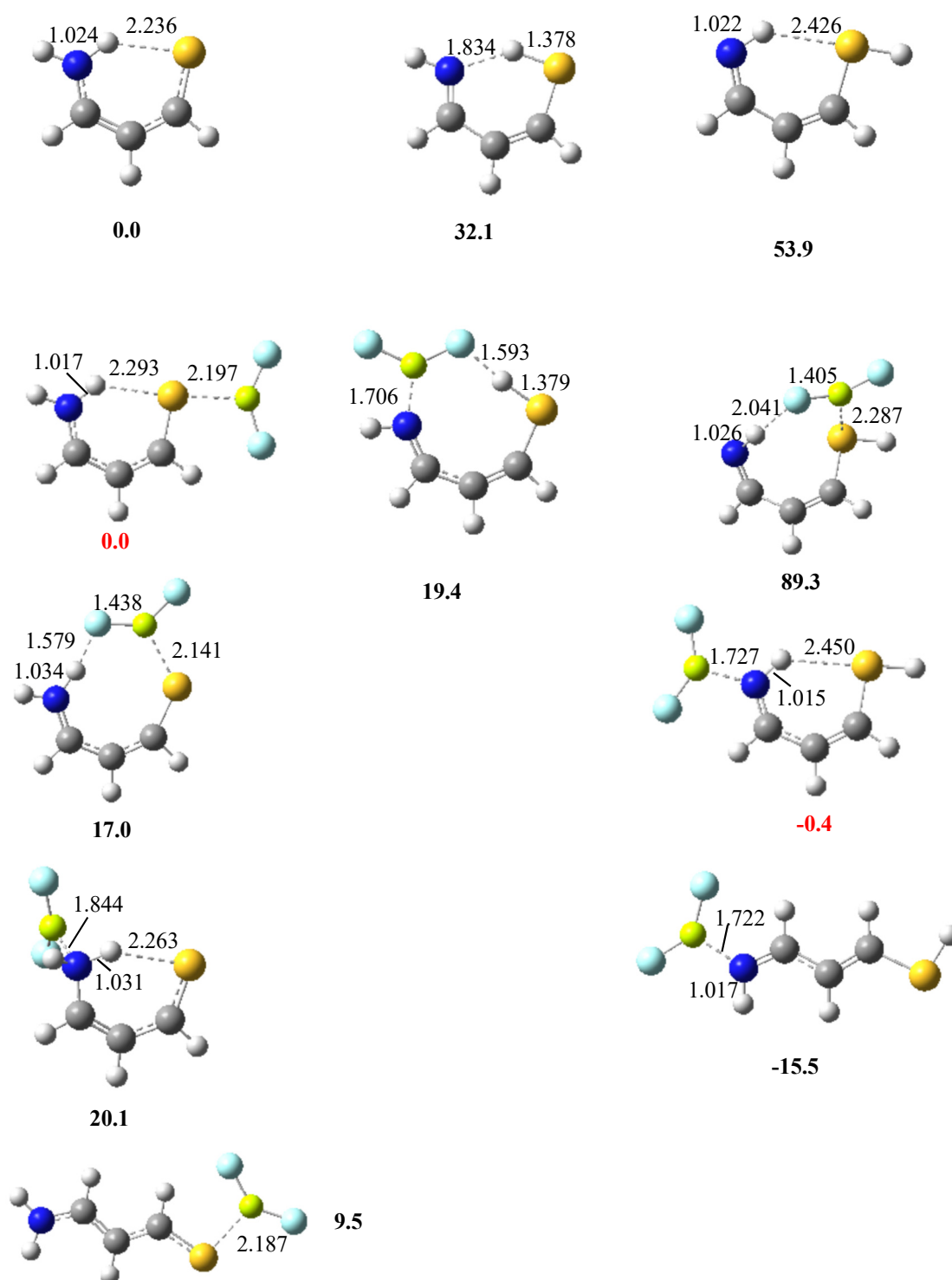
**Figure S4.** Molecular graphs of the complexes in which the BeF<sub>2</sub> molecule interacts with the imino forms of the complexes in which X = O, S. Electron densities at the bond critical points are in a.u.



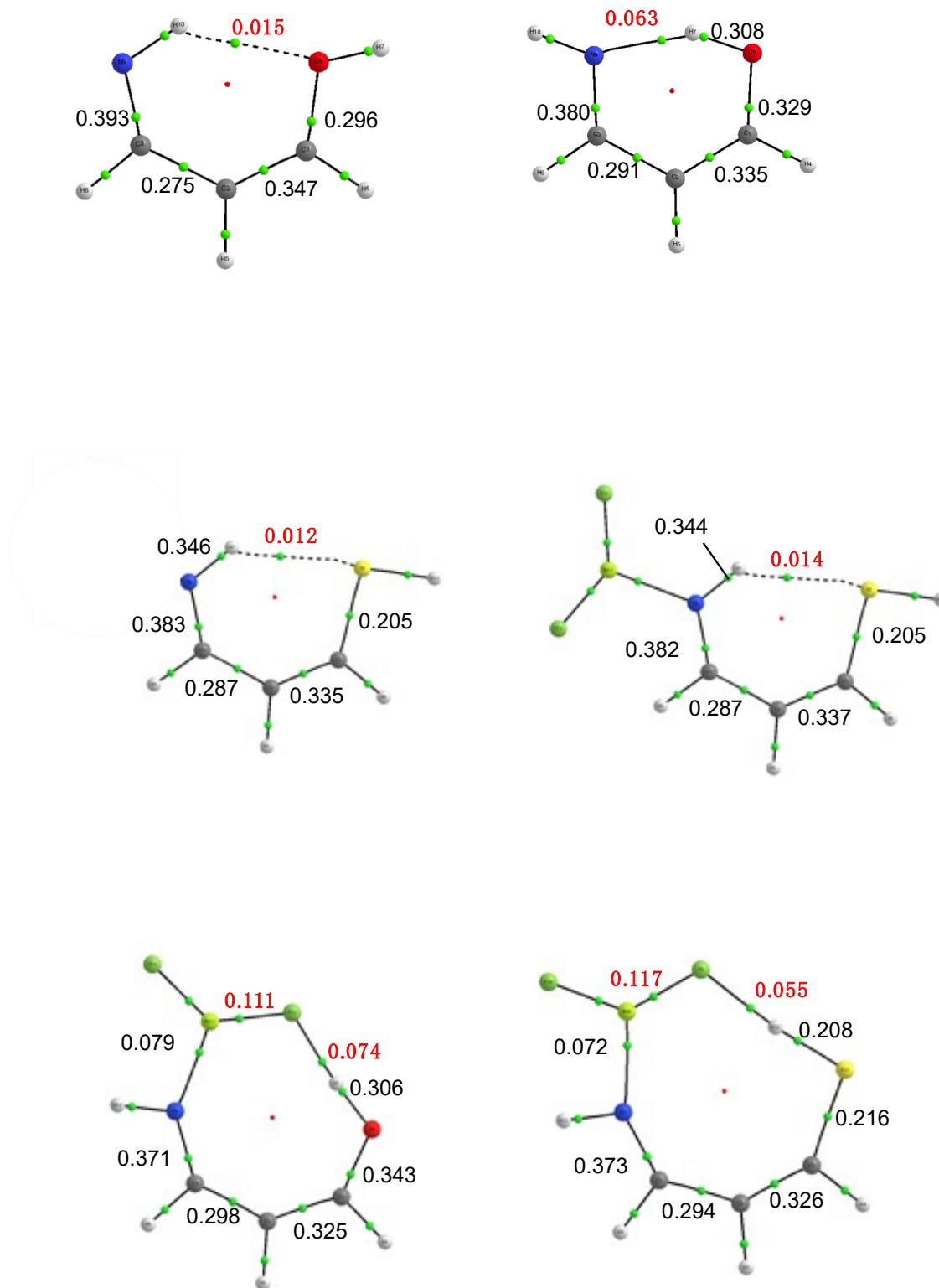
**Figure S1.** Optimized structures of the complexes between malonaldehyde-like (X = O, Y = S) with BeF<sub>2</sub>, showing the relative enthalpies (in kJ·mol<sup>-1</sup>) with respect to the global minimum (in red). Bond lengths are in Å. Atom color code: C, grey; H: white; O: red; S: mustard yellow; Be: lemon yellow; F: light blue.



**Figure S2.** Optimized structures of the complexes between malonaldehyde-like (X = O, Y = NH) with BeF<sub>2</sub>, showing the relative enthalpies (in kJ·mol<sup>-1</sup>) with respect to the global minimum (in red). Bond lengths are in Å. Same conventions as in Figure S1.



**Figure S3.** Optimized structures of the complexes between malonaldehyde-like (X = S, Y = NH) with BeF<sub>2</sub>, showing the relative enthalpies (in kJ·mol<sup>-1</sup>) with respect to the global minimum (in red). Bond lengths are in Å. Same conventions as in Figure S1.



**Figure S4.** Molecular graphs of the malonaldehyde-like complexes in which X = O, S and Y = NH with BeF<sub>2</sub> molecule. Electron densities at the bond critical points in a.u.